

# The Maintenance of a Lifelong Bibliographic Database

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with Emphasis on Thermodynamics Applied to  
Phase & Reaction Equilibria and Continuum  
Physics; and the Mathematical Description,  
Computer Programming and Calculation of  
Property Changes and Physical Conditions in  
said Systems

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## Foreword

This database has matured over three decades of research related to all kinds of problems related to thermodynamics, and especially to the mathematical description of thermodynamic equilibrium problems. No attempt has been made to create a complete database for any of the projects, but as it stands the database has a number of links to older literature and lesser-known assets, which may prove valuable to the next generation of researchers. Among several industrial projects and academic research topics covered by the database, I would like to mention:

- (1) Experimental data
  - (a)  $\text{NaF} - \text{AlF}_3 - \text{Na}_3\text{AlF}_6$
  - (b)  $\text{HCl} - \text{H}_2\text{O}$
  - (c)  $\text{HNO}_3 - \text{H}_2\text{O}$
  - (d)  $\text{HCOOH} - \text{H}_2\text{O}$
- (2) State description
  - (a) Equations of state
  - (b) Excess energy models
  - (c) Electrolyte theory
  - (d) Continuous thermodynamics
  - (e) Thermodynamic shocks
- (3) Thermodynamic theory
  - (a) Phase equilibrium
  - (b) Phase stability
  - (c) Critical points
  - (d) Chemical equilibrium
- (4) Applications
  - (a) Exergy analysis
  - (b) Finite time thermodynamics
  - (c) Irreversible thermodynamics
  - (d) Differential scanning calorimetry

About two-thirds of the articles referenced in the database, and a good number of the books, are available as scanned documents or downloaded PDF files. These files cannot (currently) be released for public use without violating copyright, but most universities are equipped with good online services, so this should not cause a serious problem for academic research staff.

## CHAPTER 1

# Cross-sectional Topics

### My Private Library (thwlib)

All kinds of books and articles dealing with mechanics, technical hobbies and the history of technology and science. However, a good part of the literature is hard science, including thermodynamics, physics, mathematics and chemistry.

**Related keywords:** thwlib.

#### Bibliography.

- [1] **Jacob Aall.** Om jernmalmleier og jerntilvirkning i Norge. In *Facsimilia scientia et technica Norvegica*, volume 12. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 97. Opptrykk av: *Om Jernmalmleier og Jerntilvirkning i Norge, Et fragmentarisk Forsøg*. Det Skandinaviske Litteraturselskaps Skrifter.
- [2] **Jacob Aall.** Om bjergværksvæsenet fornemmeligen med hensyn til Norges jernværker. In *Facsimilia scientia et technica Norvegica*, volume 45. NTH-trykk, Trondheim, 1981. Faksimile eksemplar 1. Opptrykk av: *Nutid og Fortid. Et Hæfteskift af Jacob Aall, Andet Hæfte, Arendal 1833*.
- [3] **Ivar Aavatsmark.** *Mathematische Einführung in die Thermodynamik der Gemische*. Akademie Verlag GmbH, Berlin, 1995.
- [4] **Michael M. Abbott** and Hendrick C. van Ness. *Theory and Problems of Thermodynamics: SI (metric) Edition*. Schaum's Outline Series. McGraw-Hill Book Company, Inc., New York, 1976.
- [5] **N. H. Abel.** Démonstration de l'impossibilité de la résolution générale des équations du cinquième degré. In *Facsimilia scientia et technica Norvegica*, volume 40. NTH-trykk, Trondheim, 1976. Faksimile eksemplar 302. Opptrykk av: *Mémoire sur les équations algébriques où on démontre l'impossibilité de la résolution générale des équations du cinquième degré par N. H. Abel, Christiania 1824*.
- [6] **Knut Alfsen** and Erik Alfsen. *Matematikk for gymnasene: funksjonslære 1*. H. Aschehough & Co. (W. Nygaard), Kristiania (Oslo), Norge, 1974.
- [7] **A. Almar-Næss.** *Metalliske Materialer: Struktur og egenskaper*. Tapir Forlag, Trondheim, Norge, 1981.
- [8] **Marcelo Alonso** and Edward J. Finn. *Physics*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1981.
- [9] **Fr. Chr. Holb. Arentz.** Undersøgning om hvorledes man paa korteste maade kan opløse saadanne æqvationer, som indeholde flere eller mange ubekjendte størrelser tillige. In *Facsimilia scientia et technica Norvegica*, volume 2. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 334. Opptrykk av: *Nye Samling af det Kongelige Norske Videnskabers Selskabs Skrifter, Andet bind, Kiøbenhavn 1788* side 251–286.
- [10] **Rutherford Aris.** *Vectors, Tensors and the Basic Equations of Fluid Mechanics*. Dover Publications, Inc., New York, 1989. Reprint of 1962 edition published by Prentice-Hall, Inc., Englewood Cliffs, New Jersey.
- [11] **D. Chr. Asbjørnsen.** Torv og torvdrift. In *Facsimilia scientia et technica Norvegica*, volume 43. NTH-trykk, Trondheim, 1979. Faksimile eksemplar 430. Publisert: *Christiania, 1868*.
- [12] **Isaac Asimov.** Motion, sound, and heat. In *Understanding Physics. 3 Volumes in 1*, volume I. Barnes & Noble, Inc., New York, 1966.
- [13] **Isaac Asimov.** Light, magnetism, and electricity. In *Understanding Physics. 3 Volumes in 1*, volume II. Barnes & Noble, Inc., New York, 1966.
- [14] **Isaac Asimov.** The electron, proton, and neutron. In *Understanding Physics. 3 Volumes in 1*, volume III. Barnes & Noble, Inc., New York, 1966.
- [15] **Gianni Astarita.** *Thermodynamics. An Advanced Textbook for Chemical Engineers*. Plenum Press, New York, 1989.
- [16] **Walter L. Badger** and Warren L. McCabe. *Elements of Chemical Engineering*. Chemical Engineering Series. McGraw-Hill Book Company, Inc., New York, second edition, 1936.
- [17] **Gordon M. Barrow.** *Physical Chemistry*. McGraw-Hill Book Company, Inc., New York, fourth edition, 1979.

- [18] Arthur Beiser, editor. *The World of Physics*. McGraw-Hill Book Company, Inc., New York, 1960.
- [19] Adrian Bejan. *Advanced Engineering Thermodynamics*. John Wiley & Sons, New York, 1988.
- [20] W. A. Bentley and W. J. Humphreys. *Snow Crystals*. Dover Publications, Inc., New York, 1962. Reprint of 1931 edition published by McGraw-Hill Book Company, Inc.
- [21] Gerhard Berge. *Vektor og tensoranalyse*. Matematisk institutt, Universitetet i Bergen, 1988.
- [22] Anthon Beuther. Bergkordning des löblichen neuen bergkwergrs auff dem Golmsbergk im königreich Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 13. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 207. Publisert: *Zwickau, 1540*.
- [23] Henrich Frantzen Blichfeld. Kort efterretning om bergverket i Sundhordlehn udi Bergens Stift i Norge, til oplysning om sandhed og befordring af rigets nytte, samt for at forebygge saavel de nærværende som tilkommende verkets eieres skade. In *Facsimilia scientia et technica Norvegica*, volume 34. NTH-trykk, Trondheim, 1970. Faksimile eksemplar 216. Publisert: *Kjøbenhavn, 1771*.
- [24] Carl Friedrich Böbert. Über das modumer blaufarbenwerk in norwegen. In *Facsimilia scientia et technica Norvegica*, volume 42. NTH-trykk, Trondheim, 1978. Faksimile eksemplar 279. Faksimile *Archiv für Mineralogie, Geognosie, Bergbau und Hüttenkunde, Band 21, Berlin 1847*.
- [25] Ildfast stein. Håndbok utgitt av Borgestad Fabrikker A/S.
- [26] David Eberhard Bradt. Om de norske glasværker i Aggershuus Stift i august 1781. In *Facsimilia scientia et technica Norvegica*, volume 6. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 124. Opptrykk av: *Minerva, I, 1790* side 208–237.
- [27] David Eberhard Bradt. Kort beskrivelse over det kongelige Modumske blaaifarveverk i Buskeruds amt 1781. In *Facsimilia scientia et technica Norvegica*, volume 22. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 119. Opptrykk av: *Topographisk Journal for Kongeriket Norge, Hefte 29, Kjøbenhavn 1802* side 145–179.
- [28] David Eberhard Bradt. Om guldverket i Edsvold i Norge. In *Facsimilia scientia et technica Norvegica*, volume 27. NTH-trykk, Trondheim, 1967. Faksimile eksemplar 107. Opptrykk av: *Minerva* september 1790.
- [29] David Eberhard Bradt. Beskrivelse over Walløe-saltverk, beliggende i Jarlsberg grevskap, en halv mil fra den ældgamle kjøbstad Tønsberg. In *Facsimilia scientia et technica Norvegica*, volume 28. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 147. Opptrykk av: *Minerva* mars 1790.
- [30] Gyles Brandreth. *The Great Book of Optical Illusions*. Sterling Publishing Co., Inc., New York, 1985. Earlier published as *The Big Book of Optical Illusions*, 1979.
- [31] Sven Brenner. *Materiallära*. Number 5 in Karlebo-serien. Maskinaktiebolaget Karlebo, Stockholm, Sweden, eighth edition, 1973.
- [32] Egil Brendsdal. *Computation of Phase Equilibrium in Fluid Mixtures*. Dr.ing. thesis 1999:7, Norwegian Institute of Technology, 1999.
- [33] J. N. Brönsted. *Lærebog i fysisk kemi*. Levin & Munksgaard, København, Danmark, 1936.
- [34] Georg Brochmann. *Jern*. J. W. Cappelens Forlag, Oslo, Norge, 1939.
- [35] M. Th. Brünnich. Forsøg til mineralogie for norge. In *Facsimilia scientia et technica Norvegica*, volume 41. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 814. Opptrykk av: *Et Pris-Skrift Belønnet formedelst Hans Kongel. Høyheds Arve-Prinsens Gavnildhed af Det Kongl. Norske Videnskabers Selskab og paa dets Bekostning udgivet til alminnelig Brug, Trondhiem, 1777*.
- [36] H. A. Buchdahl. *Seventeen Simple Lectures on General Relativity Theory*. John Wiley & Sons, New York, 1981.
- [37] Vincent Stoltenberg Bull. Viidenskabernes fordeele i den selskabelige omgang. In *Facsimilia scientia et technica Norvegica*, volume 36. NTH-trykk, Trondheim, 1972. Faksimile eksemplar 65. Publisert: *Christiania, 1791*.
- [38] Andreas Bull. Oekonomiske tanker om fabrikvæsenet og raae produkters forarbeidelse i landet. In *Facsimilia scientia et technica Norvegica*, volume 39. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 153. Publisert: *Kjøbenhavn, 1786*.
- [39] Herbert Callen. *Thermodynamics. An Introduction to the Physical Theories of Equilibrium Thermodynamics and Irreversible Thermodynamics*. John Wiley & Sons, New York, 1960.
- [40] Herbert Callen. *Thermodynamics and an Introduction to Thermostatistics*. John Wiley & Sons, New York, second edition, 1985.
- [41] Ashley H. Carter. *Classical and Statistical Thermodynamics*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 2001.
- [42] C. Carathéodory. Investigations into the foundations of thermodynamics. In Joseph Kestin, editor, *The Second Law of Thermodynamics*, number 5 in Benchmark papers on Energy, chapter 12. Dowden, Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976. This article was translated expressly for this Benchmark volume by J. Kestin, Brown University, from “Untersuchungen über die Grundlagen der Thermodynamik,” in *Math. Ann.* (Berlin), 67, 355–386 (1909).

- [43] **F. Albert Cotton** and Geoffrey Wilkinson. *Advanced Inorganic Chemistry*. John Wiley & Sons, New York, fourth edition, 1980.
- [44] **J. D. Cox**, D. D. Wagman, and V. A. Medvedev. *CODATA Key Values for Thermodynamics*. CODATA Series on Thermodynamic Properties. Hemisphere Publishing Corporation, New York, 1989. Final Report of the CODATA Task Group on Key Values for Thermodynamics.
- [45] **Day, Jr., R. A.** and A. L. Underwood. *Quantitative Analysis*. Prentice-Hall Chemistry Series. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1967.
- [46] **Robert T. DeHoff**. *Thermodynamics in Materials Science*. McGraw-Hill Series in Materials Science and Engineering. McGraw-Hill Book Company, Inc., New York, 1993.
- [47] **Kenneth Denbigh**. *The Principles of Chemical Equilibrium. With Applications in Chemistry and Chemical Engineering*. Cambridge University Press, Cambridge, London, fourth edition, 1981.
- [48] **Dyno Industrier**. *Sprengstoffer—Sprengningsteknikk*. Dyno Industrier A. S., Oslo, 1976.
- [49] **Ken Easterling**. *Tomorrow's Materials*. The Institute of Metals, London, 1988.
- [50] **A. Einstein**. *Über die spezielle und die allgemeine Relativitätstheorie*. Friedr. Vieweg & Sohn, Braunschweig, achte auflage edition, 1920.
- [51] **Sven Englundh**. Gamla kemisk tekniska recept. del I, 1980.
- [52] **Sven Englundh**. Kemisk tekniska recept. del II, 1980.
- [53] **W. G. Ernst**. Earth materials. In A. Lee McAlester, editor, *Foundations of Earth Science Series*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 1969.
- [54] **Michael Peterson Escholt**. Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10A. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 78. Opptrykk av: *Geologia Norvegica. Or, a brief instructive remembrancer, Concerning that very great and spacious earthquake, Which hapned almost quite through the South parts of Norway: Upon the 24<sup>th</sup>. day of April, in the year 1657. In the Danish Tongue, by Michael Peterson Escholt. And Englished by Daniel Collins*.
- [55] **Michael Peterson Escholt**. Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 290. Opptrykk av: *Geologia Norvegica. Eller: En kort Undervisning om det viit-begrebne Jordskjelff Som her udi Norge skeede mesten ofuer alt Syndenfields den 24. Aprilis udi nærværende Aar 1657*.
- [56] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh. 1790* side 387–449 og *Indhold, pl. I–II* side 450.
- [57] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. I* side 450.
- [58] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. II* side 450.
- [59] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh.*, volume D3, pages 387–449, *Indhold, pl. I–II*, 1790. Et priisskrift, som vandt det Kongelige Landhuusholdnings-Selskabs 2den Gulmedalje i Aaret 1782.
- [60] **Thorbjørn Faarlund**, Nils Sjøberg, and Thorleif Øisang, editors. *Molekylet*. Emnebiblioteket. Gyldendal Norsk Forlag, Oslo, Norge, 1970. Norsk oversettelse av *Building the Molecule* utgitt i 1969 av Marshall Cavendish Book Ltd. London.
- [61] **Enrico Fermi**. *Thermodynamics*. Dover Publications, Inc., New York, 1956. Reprint of 1937 edition published by Prentice-Hall Company.
- [62] **Donald D. Fitts**. *Nonequilibrium Thermodynamics*. McGraw-Hill Series in Advanced Chemistry. McGraw-Hill Book Company, Inc., New York, 1962.
- [63] **Ira Flatow**. *They All Laughed. From Light Bulbs to Lasers: The Fascinating Stories Behind the Great Inventions That Have Changed Our Lives*. HarperCollins Publishers, Inc., New York, 1993.
- [64] **Hans-Joachim Flechtner**. *Die Welt in der Retorte*. Deutscher Verlag, Berlin, 1938.
- [65] **Martin Richard Flor**. Bidrag til kundskab om naturvidenskabens fremskridt i Norge, et indbydelsesskrift til den offentlige examen i Christiania Kathedralskole i september 1813. In *Facsimilia scientia et technica Norvegica*, volume 11. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 500.
- [66] **Katrine Seip Førland**. *Kjemisk likevekt*. Tapir Forlag, Trondheim, Norge, 1978.
- [67] **Alfred W. Francis**. Liquid–liquid equilibrium. In Herbert M. Schoen and John J. Mcketta, Jr., editors, *Interscience Library of Chemical Engineering and Processing*, volume 3. John Wiley & Sons, New York, 1963.
- [68] **George Gamow**. *Gravitasjon*. Cappelens realbøker. J. W. Cappelens Forlag, Oslo, Norge, 1966. Oversatt av J.Randers. Bokens originaltittel: *Gravity*. Utgitt av Educational Services, New York.

- [69] **Martin Gardner**. *Fads and Fallacies in the Name of Science*. Dover Publications, Inc., New York, 1957. Reprint of 1952 edition published by G. P. Putnam's Sons under the title: In the Name of Science.
- [70] **Robert P. H. Gasser** and W. Graham Richards. *An Introduction to Statistical Thermodynamics*. World Scientific Publishing Co. Pte. Ltd., Singapore, second edition, 1995.
- [71] **Christie J. Geankoplis**. *Transport Processes and Unit Operations*. Allyn and Bacon, Inc., Boston, Massachusetts, 1978.
- [72] **J. W. Gibbs**. On the equilibrium of heterogeneous substances. In Joseph Kestin, editor, *The Second Law of Thermodynamics*, number 5 in Benchmark papers on Energy, chapter 11. Dowden, Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976. Reprinted from pp.354-372 of The Scientific Papers of J. W. Gibbs, Vol I: Thermodynamics, Dover Publications, Inc., New York, 1961 462 pp.
- [73] **R. P. Gillespie**. Partial differentiation. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, 1951.
- [74] **Robert Gilmore**. *Alice in Quantumland*. Springer-Verlag, New York, 1995.
- [75] **Samuel Glasstone**. *Theoretical Chemistry*. D. van Nostrand Company, Inc., New York, 1944.
- [76] **Samuel Glasstone**. *Thermodynamics for Chemists*. D. van Nostrand Company, Inc., New York, 1947.
- [77] **Arthur Good**. *Tom Tits experiment. Vetenskapliga förströelser*. Teknisk Litteraturtjänst, Stockholm, Sverige, 1987. Faksimil av första svenska utgåvan, utgiven av Alf Samuelssons Förlag, Stockholm 1898.
- [78] **Jens Gram**. Tanker om aarsagen til de norske jern-verkers nærværende tilstand og nogle herunder blandede anmerkninger samt forslag af hielpe-midler til deres opkomst. In *Facsimilia scientia et technica Norvegica*, volume 25. NTH-trykk, Trondheim, 1967. Faksimile eksemplar 59. Publisert: *Christiania, 1774*.
- [79] **Michael D. Greenberg**. *Foundations of Applied Mathematics*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 1978.
- [80] **E. A. Guggenheim** and J. E. Prue. *Physicochemical Calculations*. Series in Physics. North-Holland Publishing Company, Amsterdam, 1956.
- [81] **E. A. Guggenheim**. *Thermodynamics. An Advanced Treatment for Chemists and Physicists*. Series in Physics. North-Holland Publishing Company, Amsterdam, third edition, 1957.
- [82] **E. A. Guggenheim**. *Thermodynamics. An Advanced Treatment for Chemists and Physicists*. North-Holland Publishing Company, Amsterdam, sixth edition, 1977.
- [83] **C. M. Guldberg** and P. Waage. Studier over affiniteten. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapsselskapets forhandlinger* for 1864.
- [84] **C. M. Guldberg** and P. Waage. Om den chemiske affinitet. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapsselskapets forhandlinger* for 1879, nr. 4.
- [85] **D. ter Haar** and H. Wergeland. *Elements of Thermodynamics*. Addison-Wesley Series in Advanced Physics. Addison-Wesley Publishing Company, Reading, Massachusetts, 1966.
- [86] **Herm. Haeder**. *Dampfmaschinen. I. Band: Berechnung und Details*. Herm. Haeder, Duisburg am Rhein, siebente edition, 1903.
- [87] **Herm. Haeder**. *Dampfmaschinen. II. Band: Zeichnungen und Bilder*. Herm. Haeder, Duisburg am Rhein, 1903.
- [88] **Gustav Hallin**. *Handbok för stålbehandlare*. Josef Bergendahls Boktryckeri, Gøteborg, 1937.
- [89] **Staffan Hansson**. *Teknik-Historia. En historia om teknisk kunnande och dess betydelse för individ och samhälle från äldsta tid fram till 1900-talet*. Studentlitteratur, Lund, Sweden, 1990.
- [90] **Nora Hartsfield** and Gerhard Ringel. *Pearls in Graph Theory: A Comprehensive Introduction*. Dover Publications, Inc., New York, 2003. First Published by Academic Press in 1990.
- [91] **Tore Haug-Warberg**. *Den termodynamiske arbeidsboken: Boken for de to hundre hjem*. Kolofon forlag, Oslo, 2006.
- [92] **Tore Haug-Warberg**. *Sprengstoffer*. Eget forlag, 1979.
- [93] **Tore Haug-Warberg**. *Computation of Thermodynamic Equilibria*. Dr.ing. thesis 1988:42, Norwegian Institute of Technology, 1988.
- [94] **von Haxthausen, Ferdinand Christian**. Kort efterretning og beregning om Sølv-Verkets og grubernes drift, som ogsaa bergstaden Kongsberg i Norge. In *Facsimilia scientia et technica Norvegica*, volume 30. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 275. Publisert: *Kjøbenhavn, 1776*.
- [95] **Robert A. Heidemann**, Ayodeji A. Jeje, and Farhang Mohtadi. *An Introduction to the Properties of Fluids and Solids*. The University of Calgary Press, 1987.
- [96] **Christian Ernst Heltzen**. Indbydelse, plan og convention til et norskt skierper og berbygnings gewerckskab. In *Facsimilia scientia et technica Norvegica*, volume 9. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 001. Opptrykk av: *Indbydelse, Plan og Convention til et Norskt Skierper og Berbygnings Gewerckskab, Christiania 1782*.

- [97] Harold C. Helgeson, Joan M. Delany, H. Wayne Nesbitt, and Dennis K. Bird. Summary and critique of the thermodynamic properties of rock-forming minerals. *Am. J. Sci.*, 278A:1–229, 1978.
- [98] Kåre Hellan. *Emner fra fasthetslæren*. Tapir Forlag, Trondheim, Norge, 1981.
- [99] Kåre Hellan. *Fluidmekanikk*. Tapir Forlag, Trondheim, Norge, 1982.
- [100] Per Chr. Hemmer. *Statistisk mekanikk*. Tapir Forlag, Trondheim, Norge, 1970.
- [101] Gregorius Henckel. Et par ord til mine landsmænd om og fra den gamle hyttemand Gregorius Henckel, pensioneret hytteskriver ved Kongsberg Sølvværk. In *Facsimilia scientia et technica Norvegica*, volume 5. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 049. Publisert: *Drammen, 1830*.
- [102] Peder Harboe Hertzberg. Underretning for bønder i Norge om den meget nyttige jord-frugt potatoes. In *Facsimilia scientia et technica Norvegica*, volume 3. NTH-trykk, Trondheim, 1961. Opptrykk av: *Samling av Adskillige Andres Skrifter, Mag. Hans Mossins Bogtrykkerie, Andet Oplag, Bergen 1778*.
- [103] Terrell L. Hill. Thermodynamics of small systems. *J. Chem. Phys.*, 36(12):3182–3197, 1962.
- [104] Richard L. Hills. *Power from Steam: A History of the Stationary Steam Engine*. Cambridge University Press, Cambridge, London, 1989.
- [105] Terrell L. Hill. *Thermodynamics of Small Systems. Part I*. Dover Publications, Inc., New York, 1994. Two volumes bound as one. Reprint of 1963 edition published by W. A. Benjamin, Inc.
- [106] Terrell L. Hill. *Thermodynamics of Small Systems. Part II*. Dover Publications, Inc., New York, 1994. Two volumes bound as one. Reprint of 1964 edition published by W. A. Benjamin, Inc.
- [107] Peder Hiort. Historisk beretning om Røraas Kobberværk, fra dets første udfindelse og anlæg 1646 intil aar 1679. In *Facsimilia scientia et technica Norvegica*, volume 29. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 40. Opptrykk av: *Journal for Politik, Natur- og Menneske-Kundskab, for Julii 1819*.
- [108] G. Holtmark. *Lærebok i fysikk*. H. Aschehough & Co. (W. Nygaard), Kristiania (Oslo), Norge, 1922.
- [109] Hans Holtan. *Kjemisk termodynamikk*. Yrkesopplæringsrådet for håndverk og industri, Oslo, Norge, 1971.
- [110] Arnljot Høyland. Sannsynlighetslære. In *Sannsynlighetsregning og statistisk metodelære*, del I. Tapir Forlag, Trondheim, Norge, second edition, 1976.
- [111] Arnljot Høyland. Statistisk metodelære. In *Sannsynlighetsregning og statistisk metodelære*, del II. Tapir Forlag, Trondheim, Norge, second edition, 1977.
- [112] James M. Hyslop. Infinite series. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, fifth edition, 1959.
- [113] Thomas L. Jacobs, William E. Truce, and G. Ross Robertson. *Laboratory Practice of Organic Chemistry*. The Macmillan Press Ltd., London, fifth edition, 1974.
- [114] George J. Janz, Albert W. Davison, Henry S. van Klooster, and Walter H. Bauer. *Laboratory Manual of Physical Chemistry*. John Wiley & Sons, New York, fourth edition, 1956.
- [115] G. B. D. Johnson. Nogle ord om snedrev, snefog og snefonner, med fremstilling om maaden, hvorpaa antages at sammes skadelige virkning paa veie, paa ager- og engeland, etc., kan mangsted enten forebygges eller svækkes. In *Facsimilia scientia et technica Norvegica*, volume 31. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 184. Publisert: *Christiania, 1852*.
- [116] Claes Johnson. *Numerical Solution of Partial Differential Equations by the Finite Element Method*. Studentlitteratur, Lund, Sweden, 1987.
- [117] Thomas W. Judson. *Abstract Algebra*. The Prindle, Weber & Schmidt Series in Advanced Mathematics. PWS Publishing Company, Boston, Massachusetts, 1994.
- [118] Sprengningslære. Del 1. KA skolen.
- [119] Sprengningslære. Del 2. KA skolen.
- [120] Fjellsprengningsarbeid. Universitetsforlaget.
- [121] KanEnergi AS. Nye fornybare energikilder. Technical report, Norges forskningsråd i samarbeid med Norges vassdrags- og energidirektorat (NVE), Oslo, Norge, 2001. Revidert utgave 2001.
- [122] KanEnergi AS. Nye fornybare energikilder. Technical report, Norges forskningsråd i samarbeid med Norges vassdrags- og energiverk, Oslo, Norge, 1996.
- [123] Joseph Kestin, editor. *The Second Law of Thermodynamics*. Number 5 in Benchmark papers on Energy. Dowden, Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976.
- [124] A. I. Khinchin. *Mathematical Foundations of Statistical Mechanics*. Dover Publications, Inc., New York, 1949. Translated from the Russian by G. Gamow.
- [125] M. B. King. Phase equilibrium in mixtures. In P. V. Danckwerts, editor, *International Series of Monographs in Chemical Engineering*, volume 9. Pergamon Press, Oxford, London, 1969.
- [126] Emil Kirschbaum. *Distillation and Rectification*. Chemical Publishing Co., Inc., New York, 1948. Translated by M. Wulfinghoff.
- [127] Charles Kittel. *Introduction to Solid State Physics*. John Wiley & Sons, New York, 1953.
- [128] Knut Jostein Knutsen. *Formler og data i fysikk*. Tapir Forlag, Trondheim, Norge, 7 edition, 1994.

- [129] Svein A. Knudsen. *Norske oppfinnelser*. Universitetsforlaget, Oslo, Norge, 1996.
- [130] A. S. Kompaneyets. *Theoretical Physics*. Dover Publications, Inc., New York, 1962. Reprint of 1961 edition published by the Foreign Languages Publishing House, Moscow. Translated from the Russian by George Yankovsky.
- [131] Erwin Kreyszig. *Advanced Engineering Mathematics*. John Wiley & Sons, New York, fourth edition, 1979.
- [132] Norbert Adolph Lange and Gordon M. Forker. *Handbook of Chemistry*. McGraw-Hill Book Company, Inc., New York, tenth edition, 1961.
- [133] D. F. Lawden. *Principles of Thermodynamics and Statistical Mechanics*. John Wiley & Sons, New York, 1987.
- [134] Octave Levenspiel. *Chemical Reaction Engineering*. John Wiley & Sons, New York, second edition, 1972.
- [135] Lewis, Gilbert Newton and Merle Randall. *Thermodynamics and the Free Energy of Chemical Substances*. McGraw-Hill Book Company, Inc., New York, first edition, 1923.
- [136] Gilbert Newton Lewis and Merle Randall. *Thermodynamics*. McGraw-Hill Series in Advanced Chemistry. McGraw-Hill Book Company, Inc., New York, second edition, 1961. Revised by Kenneth S. Pitzer and Leo Brewer.
- [137] Kam W. Li. *Applied Thermodynamics: Availability Method and Energy Conversion*. Combustion: An International Series. Taylor & Francis, Washington, D. C., 1996.
- [138] Kenneth Libbrecht. *The Little Book of Snowflakes*. Voyageur Press, 2004.
- [139] Karl Heinrich Lieser. *Einführung in die Kernchemie*, volume 1 of *Kernchemie in Einzeldarstellungen*. Verlag Chemie GmbH, Weinheim, Germany, 1969.
- [140] Lindsay. *Lindsay's Chemical Cross Reference*. Lindsay Publications Inc., Bradley, Illinois, 1989.
- [141] David G. Luenberger. *Linear and Nonlinear Programming*. Addison-Wesley Publishing Company, Reading, Massachusetts, second edition, 1989.
- [142] Leif Lundby. *Forbrenningsmotorer. I. Grunnlag*. Universitetsforlaget, Oslo, Norge, 2 edition, 1975.
- [143] Einar Wang Lund. *Kvalitativ uorganisk analyse*. Universitetsforlaget, Oslo, Norge, 1976.
- [144] Tambs Lyche, R. Funksjoner av én fri variabel. In *Lærebok i Matematisk analyse*, del I. Gyldendal Norsk Forlag, Oslo, Norge, 1940.
- [145] Tambs Lyche, R. Funksjoner av flere fri variable. In *Lærebok i Matematisk analyse*, del II. Gyldendal Norsk Forlag, Oslo, Norge, 1941.
- [146] Tambs Lyche, R. Differensiallikninger. In *Lærebok i Matematisk analyse*, del III. Gyldendal Norsk Forlag, Oslo, Norge, 1941.
- [147] Aksel Lydersen. *Kjemiteknikk*. Tapir Forlag, Trondheim, Norge, 1972.
- [148] Aksel Lydersen. Fasthetsberegninger, mar 1978.
- [149] William Francis Magie. *A Source Book in Physics*. Source Books in the History of the Sciences. McGraw-Hill Book Company, Inc., New York, 1935.
- [150] Stanislaw Malanowski and Andrzej Anderko. *Modelling Phase Equilibria. Thermodynamic Background and Practical Tools*. Wiley Series in Chemical Engineering. John Wiley & Sons, New York, 1992.
- [151] Henry Margenau and George Moseley Murphy. *The Mathematics of Chemistry and Physics*. D. van Nostrand Company, Inc., New York, 1951.
- [152] Marvin Marcus and Henryk Minc. *A Survey of Matrix Theory and Matrix Inequalities*. Allyn and Bacon, Inc., Boston, Massachusetts, 1964.
- [153] Jon Mathews and R. L. Walker. *Mathematical Methods of Physics*. The Benjamin/Cummings Publishing Company, Menlo Park, California, second edition, 1970.
- [154] **Fundamentals of Mass Determination**. Handbook published by Mettler – Toledo AG.
- [155] **Glossary of Weighing Terms**. Handbook published by Mettler – Toledo AG.
- [156] Karl Meyer. *Meyers Vareleksikon*. Gyldendalske Boghandel, Kristiania, Norge, 2 edition, 1907.
- [157] Karl Meyer. *Meyers Vareleksikon*. Gyldendalske Boghandel, Kristiania, Norge, 3 edition, 1918.
- [158] M. Minoux. *Mathematical Programming. Theory and Algorithms*. John Wiley & Sons, New York, 1986. Reprint of 1983 edition published by Bordas Dunod Gauthier, Villars, Paris 1983. Original title *Programmation Mathématique: Théorie et Algorithmes*. Translated from the French by Steven Vajda.
- [159] Michael Modell and Robert C. Reid. *Thermodynamics and Its Applications*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1983.
- [160] Envolv Møller. Kongeriget norges herligheder, angaaende berg-værker osv. og anden videre landets frugtbarhed frem for forige tider. Til den ende sammenskreven, at enhver eftertænksom undersaat deraf kand bevæges til at prise og take den rige og runde GUD for sine utallige velgjerninger.

- In *Facsimilia scientia et technica Norvegica*, volume 18. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 213. Publisert: *Trondhjem, trøkt Aar 1758*.
- [161] [Christian Gran Molberg](#). Taksigelses tale paa Hendes Majestæts Enke-Dronningens Julianæ Mariæ Høye fødsels-dag den 4. sept. 1777 i anledning af det nye og lykkelige fuldførte vand-spring i Trondheim. In *Facsimilia scientia et technica Norvegica*, volume 24. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 27. Opptrykk av: *Trondhjems Adresse Contoirs Efterretninger No. 37, 1777 og Plakat af 30. Sept. 1777*.
- [162] [Nannestad, \(Imprimatur\), F.](#) Trende bergs-bønner, som bruges ved Røraas Kaaber-Verk paa de samtlige gruber, den første om morgenen, naar arbejderne skal anfare til arbejdet, den anden om eftermiddagen, naar de igjen anfarer, den tredje om aftenen, naar arbejderne kommer af gruben igjen. In *Facsimilia scientia et technica Norvegica*, volume 8. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 041. Publisert: *Trondhiem, 1751*.
- [163] [H. C. van Ness](#). *Understanding Thermodynamics*. Dover Publications, Inc., New York, 1983. Reprint of 1969 edition published by McGraw-Hill Book Company, Inc.
- [164] [Darrell Kirk Nordstrom](#) and James L. Munoz. *Geochemical Thermodynamics*. Blackwell Scientific Publications, Oxford, London, second edition, 1994.
- [165] [Øgrim](#), Ormestad, and Lunde. Elektrisitet, atomfysikk. In *Rom – stoff – tid*, bind 3. J. W. Cappelens Forlag, Oslo, Norge, 1970.
- [166] [Øgrim](#), Ormestad, and Lunde. Mekanikk. In *Rom – stoff – tid*, bind 1. J. W. Cappelens Forlag, Oslo, Norge, 1971.
- [167] [Øgrim](#), Ormestad, and Lunde. Varme, bølger, lys. In *Rom – stoff – tid*, bind 2. J. W. Cappelens Forlag, Oslo, Norge, 1972.
- [168] [Eric Olsen](#). Kort underretning om hellebrudd, hellernes behandling og oeconomiske nytte, samt om huul-veyters anlæggelse med brake eller eener. In *Facsimilia scientia et technica Norvegica*, volume 35. NTH-trykk, Trondheim, 1971. Faksimile eksemplar 527. Publisert: *Bergen, 1779*.
- [169] [E. Palm](#). *Hydromekanikk*. Tapir Forlag, Trondheim, Norge, 1977.
- [170] [David A. Palmer](#). *Handbook of Applied Thermodynamics*. CRC Press, Inc., Boca Raton, Florida, 1987.
- [171] [Robert Pehrson](#). *Boltzmann Temperature*. Kolofon AS, www.kolofon.com, 2004.
- [172] [E. G. Phillips](#). Functions of a complex variable. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, eighth edition, 1957.
- [173] [O. Phil](#). Anviisning til, paa en ny og forbedret maade, at anlægge og drive teglverk. In *Facsimilia scientia et technica Norvegica*, volume 37. NTH-trykk, Trondheim, 1973. Faksimile eksemplar 136. Publisert: *Kjøbenhavn, 1802*.
- [174] [A. B. Pippard](#). *Elements of Classical Thermodynamics for Advanced Students of Physics*. Cambridge University Press, Cambridge, London, 1960.
- [175] [John M. Prausnitz](#), Ruediger N. Lichtenthaler, and Edmundo Gomes de Azevedo. *Molecular Thermodynamics of Fluid-Phase Equilibria*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1986.
- [176] [William H. Press](#), Saul A. Teukolsky, William T. Vetterling, and Brian P. Flannery. *Numerical Recipes in C. The Art of Scientific Computing*. Cambridge University Press, Cambridge, London, second edition, 1992.
- [177] [I. Prigogine](#), A. Bellemans, and V. Mathot. *The Molecular Theory of Solutions*. Series in Physics. North-Holland Publishing Company, Amsterdam, 1957.
- [178] [I. Prigogine](#). *Introduction to Thermodynamics of Irreversible Processes*. John Wiley & Sons, New York, second edition, 1961.
- [179] [J. E. Prue](#). Ionic equilibria. In R. A. Robinson, editor, *Topic 15. Equilibrium Properties of Electrolyte Solutions*, volume 3 of *The International Encyclopedia of Physical Chemistry and Chemical Sciences*. Pergamon Press, Oxford, London, 1966.
- [180] [Lennart Råde](#) and Bertil Westergren. *BETA Mathematics Handbook*. Studentlitteratur, Lund, Sweden, second edition, 1990.
- [181] [Otto Redlich](#). *Thermodynamics: Fundamentals. Applications*. Elsevier, Amsterdam, Holland, 1976.
- [182] [\[Christian Ditlev\] Reventlow](#), Moltke, [Morten] Wormskiold, [Mads] Fridsch, and M. von Essen. Berganordning for kongeriget Norge. In *Facsimilia scientia et technica Norvegica*, volume 19. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 213. Publisert: *Fredriksberg, den 7 de September 1812*.
- [183] [\[Christian Ditlev\] Reventlow](#), [Carsten] Anker, and [J. F. A.] Schiffmann. Foundation for det Kongelige Norske Berg-Seminarium paa bergstaden Kongsberg. In *Facsimilia scientia et technica Norvegica*, volume 23. NTH-trykk, Trondheim, 1966. Publisert: *Christiansborg Slot, den 3die May 1786*.
- [184] [John D. Roberts](#) and Marjorie C. Caserio. *Basic Principles of Organic Chemistry*. W. A. Benjamin, Inc., Menlo Park, California, second edition, 1977.

- [185] Peter A. Rock. *Chemical Thermodynamics*. University Science Books, Mill Valley, California, 1983.
- [186] Werner W. Rogosinski. Volume and integral. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, 1952.
- [187] E. C. Rollason. *Metallurgy for Engineers*. Edward Arnold (Publishers) Ltd., London, second edition, 1955.
- [188] Terkel Rosenqvist. *Principles of Extractive Metallurgy*. McGraw-Hill Series in Materials Science and Engineering. McGraw-Hill Book Company, Inc., New York, 1974.
- [189] Jacob Rosted. Fysisk og oekonomisk beskrivelse over alumverket ved Opslo. In *Facsimilia scientia et technica Norvegica*, volume 44. NTH-trykk, Trondheim, 1980. Faksimile eksemplar 231. Opptrykk av: *Topographisk Journal for Kongeriket Norge, Hefte 3, Christiania 1793* side 1–74.
- [190] P. M. Røwde. *Nye råstoffer: et streiftog til verdenshusholdningens nye eventyrkilder*. P. M. Bye & Co. AS, Oslo, Norge, 1940.
- [191] J. S. Rowlinson. *Liquids and Liquid Mixtures*. Butterworths Scientific Publications. Butterworths Publications Ltd., London, 1959.
- [192] D. E. Rutherford. Vector methods. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, eighth edition, 1954.
- [193] Jens Rynning. Tanker over tangbrændingens indflydelse paa fiskerierne og agerdyrkingen, fremsat til det oplyste publikums overveielse. In *Facsimilia scientia et technica Norvegica*, volume 7. NTH-trykk, Trondheim, 1962. Faksimile eksemplar ?? (manko per 14-03-23). Opptrykk av: *Magasin for Næringsstanden, Trondhiem 1803*. Tildelt *Det Kongelige Norske Videnskabers Selskabs pris* på 200 Rdlr.
- [194] J. Sæchnann. Veiledning til dreiel- og teppen-vævning for begyndere. In *Facsimilia scientia et technica Norvegica*, volume 14. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 86. Publisert: *Levanger, 1849*.
- [195] Aamund Salvesson. *Lærebok i kjemi*. Yrkesopplæringsrådet for håndverk og industri, Oslo, Norge, 1956.
- [196] Jacob Henric Schou. Berg-ordinance i Norge. In *Facsimilia scientia et technica Norvegica*, volume 17. NTH-trykk, Trondheim, 1964. Faksimile eksemplar ?? (manko per 14-03-23). Opptrykk av: Jacob Henric Schou. *Chronologisk Register over de Kongelige Forordninger og aabne Breve samt andre trykte Anordninger som fra Aar 1670 af ere udkomne, tilligemed et nøiagtigt Udtog af de endnu giældende, for saavidt samme i Almindelighed angaae Undersaatterne i Danmark og Norge, forsynet med et alphabetisk Register. I Deel. Som indeholder K. Christian V Frr. fra 1670 til 1699 samt nogle før Hans Tid udkomne Anordninger. Anden Udgave. Kiøbenhavn, 1795* side 368–376. Publisert: *Hafniae die 23. Junij Anno 1683*.
- [197] Heinrich Schmidhuber. Bericht über das Kobaltwerk Snarum in Norwegen nach denen in monat juni 1846 daselbst gesammelten unterlagen bearbeitet von Heinrich Schmidhuber, k. s. berggeschwornem. In *Facsimilia scientia et technica Norvegica*, volume 38. NTH-trykk, Trondheim, 1974. Faksimile eksemplar 508. Publisert: *Leipzig, 1847*.
- [198] Erwin Schrödinger. *Statistical Thermodynamics*. Dover Publications, Inc., New York, 1989. Reprint of 1952 edition (2nd) published by Cambridge University Press with the following subtitle: *A Course of Seminar Lectures Delivered in January – March 1944, At the School of Theoretical Physics, Dublin Institute For Advanced Studies*.
- [199] Metaller. Katalog No. I utgitt av A/S Schreiner & Co.
- [200] Spesialstål og slitegods. Katalog No. IV utgitt av A/S Schreiner & Co.
- [201] Franz Seufert. *Versuche an Dampfmaschinen, Dampfkesseln, Dampfturbinen und Dieselmashinen*. Julius Springer, Berlin, fünfte edition, 1919.
- [202] Ascher H. Shapiro. *Shape and Flow. The Fluid Dynamics of Drag*. Number 20 in The Science Study Series. Heinemann Educational Books Ltd., London, 1981.
- [203] Morris H. Shamos. *Great Experiments in Physics. Firsthand Accounts from Galileo to Einstein*. Dover Publications, Inc., New York, 1987. Reprint of 1959 edition published by Holt, Rinehart and Winston, New York.
- [204] Volker Siepmann. *Process Modelling on a Canonical Basis*. Dr.ing. thesis 2006:71, Norwegian Institute of Technology, Jul 2006.
- [205] Michell J. Sienko and Robert A. Plane. *Chemistry*. McGraw-Hill Book Company, Inc., New York, fifth edition, 1976.
- [206] Lars Gunnar Sillén, Paul W. Lange, and C. Olof Gabrielson. *Fysikalisk-Kemiska Räkneuppgifter*. Almqvist & Wiksells Akademiska handböcker. Almqvist & Wiksells, Uppsala, Sverige, 1951.
- [207] Lars Monsen Sindberg. En berg-wiise, over det ypperlige og dyrebare kongsberg sølv-verck/ udi norge/ med grubernes og bergs-arbeydernes navne, under den tone: Guds godhed ville vi priise. In *Facsimilia scientia et technica Norvegica*, volume 26. NTH-trykk, Trondheim, 1967. Eksemplar nummer 258.

- [208] Nils Skogen. *Fysikk utan svarte boksar*. Fysisk institutt, Universitetet i Oslo, 1993.
- [209] J. M. Smith and H. C. van Ness. *Introduction to Chemical Engineering Thermodynamics*. McGraw-Hill Book Company, Inc., New York, third edition, 1975.
- [210] Alf Søreide. *Metallene i handverk og industri: Arbeidskunnskap for elever i framhaldsskolen, realskolen, yrkes- og bedriftsskoler*. J. W. Cappelens Forlag, Oslo, Norge, 1949.
- [211] K. S. Spiegel. *Principles of Energetics*. Springer-Verlag, New York, 1983. Based on: *Applications de la thermodynamique du non-équilibre* by P. Chartier, M. Cross and K. S. Spiegel.
- [212] Euan Squires. *The Mystery of the Quantum World*. Institute of Physics Publishing, Bristol, UK, second edition, 1994.
- [213] H. Eugene Stanley. *Introduction to Phase Transitions and Critical Phenomena*. International Series on Monographs on Physics. Oxford University Press, London, 1971.
- [214] Roger Y. Stanier, Michael Doudoroff, and Edward A. Adelberg. *General Microbiology*. Macmillans Student Editions. The Macmillan Press Ltd., London, third edition, 1972.
- [215] Johan Stivimoen. Geopatogene strålingssoner: jordstrålingspåvirkninger fra vannårer og lignende, aug 1988.
- [216] Hans Strøm. Om norske fabriker og om fabriker i Aggershuus-Stift i Norge som opprinnelig ble utgitt i samleren, et ugeskrivt første og andet bind, henholdsvis side 321–327, 113–136. In *Facsimilia scientia et technica Norvegica*, volume 32. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 109. Publisert: *Kiøbenhavn, 1787–1788*.
- [217] Gilbert Strang. *Linear Algebra and Its Applications*. Academic Press, Inc., Orlando, Florida, second edition, 1980.
- [218] George P. Sutton. *Rocket Propulsion Elements: An Introduction to the Engineering of Rockets*. John Wiley & Sons, New York, 1949.
- [219] Richard A. Swalin. *Thermodynamics of Solids*. Wiley Series on the Science and Technology of Materials. John Wiley & Sons, New York, 1962.
- [220] Knut Sydsæter, Atle Seierstad, and Arne Strøm. *Matematisk analyse. Bind II*. Universitetsforlaget, Oslo, Norge, third edition, 1996.
- [221] Jefferson W. Tester and Michael Modell. *Thermodynamics and Its Applications*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall PTR, Upper Saddle River, New Jersey, third edition, 1997.
- [222] Th. Tharaldsen. *Eksplisivstoffer*. Dreyers Forlag, Oslo, Norge, 1950.
- [223] Frederik Thaarup. Det norske kompanies octrojer og privilegier m.v. In *Facsimilia scientia et technica Norvegica*, volume 16. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 5. Opptrykk av: *Det norske Kompanies Octrojer og Privilegier m.v., Magazin for Danmarks og Norges topographiske, oekonomiske og statistiske Beskrivelse, Andet Bind, Kiøbenhavn, 1802–1803* side 78–117.
- [224] Peter Thorstensøn. Peter Thorstensøns tale holdt ved innvielsen av det Kongelige Norske Berg-Seminarium den 27. mai 1786. In *Facsimilia scientia et technica Norvegica*, volume 4. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 430. Opptrykk av: *Minerva, et Maanedsskift, Anden Aargangs Tredie Bind for Martii Maaned 1787* side 308–346.
- [225] Thomas, Jr., George B. *Calculus and Analytical Geometry*. Addison-Wesley Series in Mathematics. Addison-Wesley Publishing Company, Reading, Massachusetts, 1972.
- [226] Laszlo Tisza. *Generalized Thermodynamics*. The M. I. T. Press, Cambridge, 1966.
- [227] Gaston Tissandier. *Videnskabelige adspredelser*. Alb. Cammermeyer, Kristiania (Oslo), Norge, 1882. Oversatt af kommandørkaptein H. J. Müller.
- [228] Annanias Tveit. Vandampdiffusjonstabell for papp- og trefiberplater. Technical Report 9, Norges Byggeforskningsinstitutt, Oslo, Norge, 1954.
- [229] Annanias Tveit. Fukt og fukttransport i porøse materialer. Technical Report 39, Norges Byggeforskningsinstitutt, Oslo, Norge, 1964. En litteraturstudie.
- [230] H. J. V. Tyrrell. *Diffusion and Heat Flow in Liquids*. Butterworths Publications Ltd., London, 1961.
- [231] Gheorghe Vasaru, Günther Müller, Günther Reinhold, and Traian Fodor. The thermal diffusion column. Theory and practice with particular emphasis on isotope separation. In E. Krell, editor, *Physikalisch-chemische Trenn- und Messmethoden*, band 13. VEB Deutscher Verlag der Wissenschaften, Berlin, 1969.
- [232] Marianne Videm. *Cyclic Deformation of Pure Aluminum*. Dr.ing. thesis 1992:36, Norwegian Institute of Technology, 1992.
- [233] H. U. von Vogel. *Chemiker-Kalender*. Springer-Verlag, New York, 1974.
- [234] Johan Friederich Christian Voss. Det frivillige slaverie, eller kort begrep om Sølvverket Kongsberg nærværende tilstand. In *Facsimilia scientia et technica Norvegica*, volume 33. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 330. Publisert: *Kiøbenhavn, 1771*.
- [235] K. Walther and M. Röttinger. *Technische Wärmelehre*. Sammlung Göschen. G. J. Göschen'sche Verlagshandlung, Leipzig, 1905.

- [236] [Jearl Walker](#). *The Flying Circus of Physics with Answers*. John Wiley & Sons, New York, 1977.
- [237] [Robert Weast](#), editor. *CRC Handbook of Chemistry and Physics*. CRC Press, Inc., Boca Raton, Florida, 58th edition, 1978.
- [238] [Steven Weinberg](#). *De første tre minutter. Moderne forskning om universets opprinnelse*. EFI Forlag, Oslo, Norge, 1984. Oversatt av A. Quale. Bokens originale tittel: *The Thirst Three Minutes*.
- [239] [Ralph R. Wenner](#). *Thermochemical Calculations*. McGraw-Hill Book Company, Inc., New York, 1941.
- [240] [Thorstein Wereide](#). *Statistical Theory of Energy and Matter*. Gyldendalske Boghandel, Kristiania, 1915.
- [241] [Thorstein Wereide](#). *Relativitetsprincippet eller tidrummets struktur*. Gyldendalske Boghandel, Kristiania, 1922.
- [242] [Söfren Wesing](#). En kort underretning om brand-redskabets brug og anbringelse i nøds-tilfælde samt hvad ildebrande, som er overgaaet Trondhiems bye fra ao. 1328 og til 1740. In *Facsimilia scientia et technica Norvegica*, volume 20. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 5. Publisert: *Trondhjem, 1740*.
- [243] [Söfren Wesing](#). Kurtze, doch warhafftige beschreibung der silber und anderer bergwercke in Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 21. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 98. Publisert: 2. utgave fra 1685.
- [244] [J. H. Wolfenden](#). *Numerical Problems in Advanced Physical Chemistry*. Oxford University Press, London, 1938.
- [245] [M. M. Woolfson](#). *An Introduction to X-ray Crystallography*. Cambridge University Press, Cambridge, London, 1979.
- [246] [F. van Zegeer](#) and S. H. Storey. *The Computation of Chemical Equilibria*. Cambridge University Press, Cambridge, London, 1970.
- [247] [Mark W. Zemansky](#). *Heat and Thermodynamics*. McGraw-Hill Book Company, Inc., New York, second edition, 1943.
- [248] [Mark W. Zemansky](#). *Temperatures Very Low and Very High*. Dover Publications, Inc., New York, 1981. Reprint of 1964 edition published by D. van Nostrand Company, Inc.
- [249] [Steven S. Zumdahl](#). *Chemical Principles*. D. C. Heath and Company, Lexington, Massachusetts, second edition, 1995.

### Biographies (biography)

The life and work of scientists from the past. I am particularly keen to obtain biographical material, but time has not permitted me to collect the literature for more than about two-thirds of the biographies (yet).

#### Related keywords: biography.

##### Bibliography.

- [1] [Samia Al-Shayban](#). Aristotle biography: Immortal philosopher of antiquity. English Department, King Saud University, Riyadh City. Kingdom of Saudi Arabia, 2008.
- [2] [Samuel K. Allison](#). Enrico Fermi: 1901 – 1954. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1957.
- [3] [E. Arunan](#). van der Waals. *Resonance*, pages 584–587, jul 2010.
- [4] [Karl Egil Aubert](#). Niels Henrik Abel. *Normat*, 27(4):129–140, 1979.
- [5] [Olof Beckman](#). Anders Celsius, 300 år. Föredrag av Hr Olof Beckman vid Kungl. Vetenskaps-socitetens högtidssammanträde den 10 november 2001”. Årsbok / Kungl. Vetenskaps-societeten i Uppsala (2002) sid. 20–31., 2001.
- [6] [George W. Benthien](#). James Clerk Maxwell: His life and his faith, aug 2009.
- [7] [Klaus Beneke](#). Hermann Ludwig Ferdinand von Helmholtz und zur geschichte der russischen studentinnen und studenten in heidelberg im letzten jahrhundert. *Mitteilungen der Kolloid-Gesellschaft*, pages 106–150, 1999.
- [8] [Sean Bird](#). James Clerk Maxwell: 1831 – 1879. Covenant Christian High School for History of Mathematics at IUPUI (Indiana University — Purdue University Indianapolis), apr 2002.
- [9] [Jacqueline Boniface](#). Leopold Kronecker’s conception of the foundations of mathematics. *Philosophia Scientiæ*, CS 5:143–156, 2005.
- [10] [C. Brezinski](#). The life and work of André Cholesky. *Numer. Algor.*, 43:279–288, 2006.
- [11] [Georg Bredig](#). Kopp, Hermann. *Allgemeine Deutsche Biographie*, 55:820–826, 1910.
- [12] [E. Bretscher](#) and J. D. Cockcroft. Enrico Fermi: 1901 – 1954. *Biographical Memoirs of Fellows of the Royal Society*, 1:69–78, nov 1955.
- [13] [Viggo Brun](#). Det gjenfunne manuskript til Abels parisavhandling. *Normat*, 7:91–97, 1953.

- [14] [H. B. G. Casimir](#). Hugo Martin Tetrode. NRC Handelsblad — Wetenschap en Onderwijs bijlage, jan 1984.
- [15] [R. Chéret](#). The life and work of Pierre-Henri Hugoniot. *Shock Waves (Historical Archive)*, 2(1):1–4, mar 1992. Springer-Verlag.
- [16] [M. Duží](#). Kurt Gödel. Metamathematical results on formally undecidable propositions: Completeness vs. incompleteness. *Organon F*, XII(4):447–474, 2005.
- [17] [W. Dyck](#), S. Gundelfinger, J. Lüroth, and M. Noether. *Ludwig Otto Hesse's gesammelte Werke*, pages 711–721. Der Mathematisch-Physikalische Classe der Königlich Bayerische Akademie der Wissenschaften, München, 1897.
- [18] [Freeman Dyson](#). A meeting with Enrico Fermi. *Nature*, 27:297, jan 2004.
- [19] [Herman Erlichson](#). Sadi Carnot, “Founder of the second law of thermodynamics”. *Eur. J. Phys.*, 20:183–192, 1999.
- [20] [F. M. Feldhaus](#). Otto, Nikolaus. *Allgemeine Deutsche Biographie*, 52:734–735, 1906.
- [21] [James Ford](#). Thomas Newcomen: Inventor and baptist minister, 1663–1729. In Geo. P. Gould, editor, *Transactions of the Baptist Historical Society 2.2*. Biblical Studies, oct 1911.
- [22] [Jacob Goldberg](#). Robert Boyle: The man who changed the history of science & the history of science that changed the man. *Dartmouth Undergraduate Journal of Science*, 27(4):129–140, 2007.
- [23] [I. Grattan-Guinness](#). Daniel Bernoulli and the varieties of mechanics in the 18th century. *NAW*, 5/1(3):242–249, sep 2000. Johann Bernoulli lecture, Groningen, May 9, 2000.
- [24] [Ulrich Grigull](#). Fahrenheit, a pioneer of exact thermometry. *Heat Transfer*, 1:9–18, 1966. The Proceedings of the 8th International Heat Transfer Conference, San Francisco, 1966.
- [25] [Charles S. Hastings](#). Josiah Willard Gibbs: 1839–1903. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1909.
- [26] [Robert A. Hatch](#). Isaac Newton biography—Newton’s life, career, work. In *Encyclopedia Americana*, volume 20. Grolier Incorporated, Danbury, CT, 1998.
- [27] [André Hauser](#). Auf den spuren Leonhard Eulers. *Geschichte der Mathematik*, Pädagogische Hochschule Aargau, Institut Sekundarstufe 1, 2008.
- [28] [Walter Heinz](#). Kopp, Hermann. *Neue Deutsche Biographie*, 12:567–568, 1979.
- [29] [Joel H. Hildebrand](#). Gilbert Newton Lewis: 1875–1946. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1958.
- [30] [Walter Höflechner](#). Ludwig Boltzmann—Persönlichkeit—Karriere—Bedeutung. Vortrag gehalten in Wien am 24. November 2006 zu Boltzmanns 100. Todestag, 2006.
- [31] [Don A. Howard](#). Albert Einstein as a philosopher of science. *Physics Today*, pages 34–40, dec 2005.
- [32] [A. T. Humphrey](#). Lord Rayleigh—the last of the great victorian polymaths. *GEC Review*, 7(3):167–179, 1992.
- [33] [ICE \[unknown\]](#). Obituary. Rudolf Julius Emanuel Clausius, 1822–1888. *Minutes of the Proceedings of the Institution of Civil Engineers*, 96:307–316, 1889. Part 2.
- [34] [Michel Janssen](#). Albert Einstein: His biography in a nutshell. Hsci/Phys 1905 Spring 2003, Einstein for Everyone, 2003.
- [35] [John Ericsson](#): 1803–1889. John Ericsson Sällskapet, 2011.
- [36] [Edwin C. Kemble](#) and Francis Birch. Percy Williams Bridgman: 1882–1961. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1970.
- [37] [Alexander Kipnis](#). Sackur, Otto. *Neue Deutsche Biographie*, 22:344, 2005.
- [38] [Stephen C. Kleene](#). Kurt Gödel: 1906–1978. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1987.
- [39] [Carl Krafft](#). Niels Henrik Abel. *Naturen: Et illustreret Maanedsskrift for populær Naturvidenskab*, 8(1):1–4, Jan. 1884.
- [40] [V. Kumaran](#). Josiah Willard Gibbs. *Resonance*, pages 4–11, jul 2007.
- [41] [Pierre Lemay](#) and Ralph E. Oesper. Pierre Louis Dulong, his life and work. *Chymia*, 1:171–190, 1948.
- [42] [Carl Lira](#). Biography of James Watt. Prepared as a supplement to: Introductory Chemical Engineering Thermodynamics, 2001.
- [43] [James Prescott Joule \(1818–1889\)](#): Converting work into heat. The Museum of Science and Industry in Manchester, 2004.
- [44] [Melissa Moore](#). Pierre-Simon Laplace—biography paper. MATH 4010, University of Colorado, Denver, 2008.
- [45] [Walther Nernst](#). *Rudolf Clausius*. Ludwig Röhrscheid, Bonn, 1922.
- [46] [N. V. E. Nordenmark](#). Anders Celsius. In *Lychnos-bibliotek*, volume 1. Lärdomshistoriska samfundet, 1936.
- [47] [J. J. O’Connor](#) and E. F. Robertson. Carl Gustav Jacob Jacobi. MacTutor History of Mathematics archive, University of St Andrews, jan 2000.

- [48] [J. J. O'Connor](#) and E. F. Robertson. Evangelista Torricelli. MacTutor History of Mathematics archive, University of St Andrews, nov 2002.
- [49] [J. J. O'Connor](#) and E. F. Robertson. William Thomson (Lord Kelvin). MacTutor History of Mathematics archive, University of St Andrews, oct 2003.
- [50] [J. J. O'Connor](#) and E. F. Robertson. William John Macquorn Rankine. MacTutor History of Mathematics archive, University of St Andrews, 2005.
- [51] [J. J. O'Connor](#) and E. F. Robertson. Denis Papin. MacTutor History of Mathematics archive, University of St Andrews, dec 1996.
- [52] [J. J. O'Connor](#) and E. F. Robertson. Pierre-Simon Laplace. MacTutor History of Mathematics archive, University of St Andrews, jan 1999.
- [53] [J. J. O'Connor](#) and E. F. Robertson. Adrien-Marie Legendre. MacTutor History of Mathematics archive, University of St Andrews, jan 1999.
- [54] [Evangelos Papadopoulos](#). Heron of Alexandria (ca. 10–85 AD). Department of Mechanical Engineering, National Technical University of Athens, Greece, 2011.
- [55] [Kenneth S. Pitzer](#). Gilbert N. Lewis and the thermodynamics of strong electrolytes. *J. Chem. Educ.*, 61(2):104–107, 1984.
- [56] [W. O. Price](#) and Sheldon Krinsky. On thought experiments: Ernst Mach. *Philosophical Forum*, pages 449–457, jul 1973. Translated from “Über Gedankenexperimente”, *Zeitschrift für den physikalischen und chemischen Unterricht*, 10, 1–5 (1897).
- [57] [Renuka Ravindran](#), C. R. Pranesachar, and D. P. Patil. Joseph Louis Lagrange (1736–1813). *Resonance*, pages 2–4, apr 2006.
- [58] [Jan Sebestik](#). The dawn of cognitive science: Early European contributors. In Liliana Albertazzi, editor, *Synthese Library: Studies in Epistemology, Logic, Methodology, And Philosophy of Science*, volume 295, chapter Ernst Mach’s Evolutionary Theory of Representation. Springer-Science, 2001.
- [59] [R. M. Sillitto](#). Clerk Maxwell’s life and work. Talk given at JCMB (James Clerk Maxwell Building, University of Edinburgh), for a party of visiting Dutch schoolteachers, 1993.
- [60] [Wolfgang Steinicke](#). Wolfgang Pauli—leben und werk. Seminar Physik- und Chemiegeschichte IV, Universität Hamburg Fachbereich 11: Mathematik, May 2004.
- [61] [David J. Thomas](#) and Judith M. Smith. Joseph Raphson, F. R. S. *Notes Rec. R. Soc. Lond.*, 44(5):151–167, 1990.
- [62] [Encyclopedia of World Biography on Otto Von Guericke](#). Thomson Gale, a part of the Thomson Corporation, 2005.
- [63] [Constantine Tsonopoulos](#). Otto Redlich, 1896–1978: In memory and appreciation. *Fluid Phase Equilib.*, 12(1/2):1–9, 1983.
- [64] [John Archibald Wheeler](#). Albert Einstein: 1879–1955. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1980.
- [65] [Carl von Linde](#). Who’s Who, 2011.
- [66] [Richard Williams](#). François-Marie Raoult and Raoult’s law: May 23, 1887. *APS News*, 20(5):2, May 2015. Column *This Month in Physics History*.
- [67] [Jaime Wisniak](#). Benoit Paul Emile Clapeyron: A short bibliographical sketch. *Chem. Educator*, 5:83–87, 2000.
- [68] [Jaime Wisniak](#). Pierre Maurice Marie Duhem: A polemical scientist. *Chem. Educator*, 5:156–161, 2000.
- [69] [Jaime Wisniak](#). William Henry: His achievements and his law. *Chem. Educator*, 6:62–68, 2001.

### Historical Notes (histor(y,ical))

Books and articles of historical interest, also including a few biographies and discussions of the impact of science (and named scientists) on modern society.

**Related keywords:** [history](#) | [historical](#).

#### Bibliography.

- [1] [Jacob Aall](#). Om jernmalmleier og jerntilvirkning i Norge. In *Facsimilia scientia et technica Norvegica*, volume 12. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 97. Opptrykk av: *Om Jernmalmleier og Jerntilvirkning i Norge, Et fragmentarisk Forsøg*. Det Skandinaviske Litteraturselskaps Skrifter.
- [2] [Jacob Aall](#). Om bjergværksvæsenet fornemmeligen med hensyn til norges jernværker. In *Facsimilia scientia et technica Norvegica*, volume 45. NTH-trykk, Trondheim, 1981. Faksimile eksemplar 1. Opptrykk av: *Nutid og Fortid. Et Hæfteskrift af Jacob Aall, Andet Hæfte, Arendal 1833*.
- [3] [Stuart S. Antman](#). The tragicomical history of thermodynamics 1822–1854. By C. J. Truesdell. Springer-Verlag, New York, 1980. xii + 372 pp. *Am. Math. Monthly*, 90(5):343–346, 1983.

- [4] **Fr. Chr. Holb. Arentz**. Undersøgning om hvorledes man paa korteste maade kan opløse saadanne æqvationer, som indeholde flere eller mange ubekjendte størrelser tillige. In *Facsimilia scientia et technica Norvegica*, volume 2. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 334. Opptrykk av: *Nye Samling af det Kongelige Norske Videnskabers Selskabs Skrifter, Andet bind, Kiøbenhavn 1788* side 251–286.
- [5] **D. Chr. Asbjørnsen**. Torv og torvdrift. In *Facsimilia scientia et technica Norvegica*, volume 43. NTH-trykk, Trondheim, 1979. Faksimile eksemplar 430. Publisert: *Christiania, 1868*.
- [6] **G. Astarita**. Historical and philosophical background of thermodynamics. *Ind. Eng. Chem.*, 16(1):138–143, 1977.
- [7] **Rubin Battino**, Laurence E. Strong, and Scott E. Wood. A brief history of thermodynamics notion. *J. Chem. Educ.*, 74(3):304–305, Mar. 1997.
- [8] **Baumgartner, Friherre von, Andr.** Om varmens mekaniske ækvivalent og dens betydning i naturvidenskaben. *Polyteknisk Tidsskrift*, 3(20):305–316, 31. Okt. 1856. Oversatt til norsk fra tysk artikkel i *Dinglers Polyt. Journal*.
- [9] **Arthur Beiser**, editor. *The World of Physics*. McGraw-Hill Book Company, Inc., New York, 1960.
- [10] **Anthon Beuther**. Bergkordnung des löblichen neuen bergkwerchs auff dem Golmsbergk im königreich Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 13. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 207. Publisert: *Zwickau, 1540*.
- [11] **Henrich Frantzen Blichfeld**. Kort efterretning om bergverket i Sundhordlehn udi Bergens Stift i Norge, til oplysning om sandhed og befordring af rigets nytte, samt for at forebygge saavel de nærværende som tilkommende verkets eieres skade. In *Facsimilia scientia et technica Norvegica*, volume 34. NTH-trykk, Trondheim, 1970. Faksimile eksemplar 216. Publisert: *Kiøbenhavn, 1771*.
- [12] **Carl Friedrich Böbert**. Über das modumer blaufarbenwerk in norwegen. In *Facsimilia scientia et technica Norvegica*, volume 42. NTH-trykk, Trondheim, 1978. Faksimile eksemplar 279. Faksimile *Archiv für Mineralogie, Geognosie, Bergbau und Hüttenkunde, Band 21, Berlin 1847*.
- [13] **David Eberhard Bradt**. Om de norske glasværker i Aggershuus Stift i august 1781. In *Facsimilia scientia et technica Norvegica*, volume 6. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 124. Opptrykk av: *Minerva, I, 1790* side 208–237.
- [14] **David Eberhard Bradt**. Kort beskrivelse over det kongelige Modumske blaaifarveverk i Buskeruds amt 1781. In *Facsimilia scientia et technica Norvegica*, volume 22. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 119. Opptrykk av: *Topographisk Journal for Kongeriket Norge, Hefte 29, Kiøbenhavn 1802* side 145–179.
- [15] **Georg Brochmann**. *Jern*. J. W. Cappelens Forlag, Oslo, Norge, 1939.
- [16] **M. Th. Brünnich**. Forsøg til mineralogie for norge. In *Facsimilia scientia et technica Norvegica*, volume 41. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 814. Opptrykk av: *Et Pris-Skrift Belønnet formedelst Hans Kongel. Høyheds Arve-Prinsens Gavnildhed af Det Kongl. Norske Videnskabers Selskab og paa dets Bekostning udgivet til alminnelig Brug, Trondhiem, 1777*.
- [17] **Vincent Stoltenberg Bull**. Viidenskabernes fordeele i den selskabelige omgang. In *Facsimilia scientia et technica Norvegica*, volume 36. NTH-trykk, Trondheim, 1972. Faksimile eksemplar 65. Publisert: *Christiania, 1791*.
- [18] **Andreas Bull**. Oekonomiske tanker om fabrikvæsenet og raae produkters forarbeidelse i landet. In *Facsimilia scientia et technica Norvegica*, volume 39. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 153. Publisert: *Kiøbenhavn, 1786*.
- [19] **J. C. Butcher** and G. Wanner. Runge-kutta methods: Some historical notes. *Applied Numerical Mathematics*, 22:113–151, 1996.
- [20] **W. C. Chapin**. Chemical research problems. *J. Chem. Educ.*, 2(5):391–403, May 1925.
- [21] **K. C. Cheng**. Historical development of the theory of heat and thermodynamics: Review and some observations. *Heat Transfer Eng.*, 13(3):19–37, 1992.
- [22] **R. Clausius**. *The Mechanical Theory of Heat, with Its Applications to the Steam-Engine and to the Physical Properties of Bodies*. John Van Voorst, London, 1867. English translation of *Mechanische Warmetheorie* directed by John Tyndall.
- [23] **Freeman J. Dyson**. Willard Gibbs and the teaching of science. In *Proceedings of The Gibbs Symposium*, pages 269–276. Yale University, May 15–17 1989.
- [24] **Michael Peterson Escholt**. Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10A. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 78. Opptrykk av: *Geologia Norvegica. Or, a brief instructive remembrancer, Concerning that very great and spacious earthquake, Which hapned almost quite through the South parts of Norway: Upon the 24<sup>th</sup>. day of April, in the year 1657. In the Danish Tongue, by Michael Peterson Escholt. And Englished by Daniel Collins*.
- [25] **Michael Peterson Escholt**. Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 290. Opptrykk av: *Geologia Norvegica. Eller:*

- En kort Undervisning om det viit-begrebne Jordskjelff Som her udi Norge skeede mesten ofuer alt Syndenfields den 24. Aprilis udi nærværende Aar 1657.*
- [26] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh. 1790* side 387–449 og *Indhold, pl. I–II* side 450.
- [27] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. I* side 450.
- [28] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. II* side 450.
- [29] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh.*, volume D3, pages 387–449, *Indhold, pl. I–II*, 1790. Et priisskrift, som vandt det Kongelige Landhuusholdnings-Selskabs 2den Guldmedalje i Aaret 1782.
- [30] **Michael E. Fisher**. Phases and phase diagrams: Gibbs's legacy today. In *Proceedings of The Gibbs Symposium*, pages 39–72. Yale University, May 15–17 1989.
- [31] **Ira Flatow**. *They All Laughed. From Light Bulbs to Lasers: The Fascinating Stories Behind the Great Inventions That Have Changed Our Lives*. HarperCollins Publishers, Inc., New York, 1993.
- [32] **Hans-Joachim Flechtner**. *Die Welt in der Retorte*. Deutscher Verlag, Berlin, 1938.
- [33] **Martin Richard Flor**. Bidrag til kundskab om naturvidenskabens fremskridt i Norge, et indbydelsesskrift til den offentlige examen i Christiania Kathedralskole i september 1813. In *Facsimilia scientia et technica Norvegica*, volume 11. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 500.
- [34] **James Ford**. Thomas Newcomen: Inventor and baptist minister, 1663–1729. In Geo. P. Gould, editor, *Transactions of the Baptist Historical Society 2.2*. Biblical Studies, oct 1911.
- [35] **Martin Gardner**. *Fads and Fallacies in the Name of Science*. Dover Publications, Inc., New York, 1957. Reprint of 1952 edition published by G. P. Putnam's Sons under the title: *In the Name of Science*.
- [36] **Jacob Goldberg**. Robert Boyle: The man who changed the history of science & the history of science that changed the man. *Dartmouth Undergraduate Journal of Science*, 27(4):129–140, 2007.
- [37] **Jens Gram**. Tanker om aarsagen til de norske jern-verkers nærværende tilstand og nogle herunder blandede anmerkninger samt forslag af hielpe-midler til deres opkomst. In *Facsimilia scientia et technica Norvegica*, volume 25. NTH-trykk, Trondheim, 1967. Faksimile eksemplar 59. Publisert: *Christiania, 1774*.
- [38] **C. M. Guldberg** and P. Waage. Studier over affiniteten. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapselskapets forhandlinger for 1864*.
- [39] **C. M. Guldberg** and P. Waage. Om den chemiske affinitet. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapselskapets forhandlinger for 1879*, nr. 4.
- [40] **Staffan Hansson**. *Teknik-Historia. En historia om teknisk kunnande och dess betydelse för individ och samhälle från äldsta tid fram till 1900-talet*. Studentlitteratur, Lund, Sweden, 1990.
- [41] **von Haxthausen, Ferdinand Christian**. Kort efterretning og beregning om Sølv-Verkets og grubernes drift, som ogsaa bergstaden Kongsberg i Norge. In *Facsimilia scientia et technica Norvegica*, volume 30. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 275. Publisert: *Kjøbenhavn, 1776*.
- [42] **Christian Ernst Heltzen**. Indbydelse, plan og convention til et norskt skierper og berbygnings gewerckskab. In *Facsimilia scientia et technica Norvegica*, volume 9. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 001. Opptrykk av: *Indbydelse, Plan og Convention til et Norskt Skierper og Berbygnings Gewerckskab, Christiania 1782*.
- [43] **Gregorius Henckel**. Et par ord til mine landsmænd om og fra den gamle hyttemand Gregorius Henckel, pensioneret hytteskriver ved Kongsberg Sølvværk. In *Facsimilia scientia et technica Norvegica*, volume 5. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 049. Publisert: *Drammen, 1830*.
- [44] **J. Hertz**. Josiah Willard Gibbs and teaching thermodynamics of materials (history). *J. Phase Equilib.*, 13(5):450–458, 1992.
- [45] **Richard L. Hills**. *Power from Steam: A History of the Stationary Steam Engine*. Cambridge University Press, Cambridge, London, 1989.
- [46] **Peder Hiort**. Historisk beretning om Røraas Kobberværk, fra dets første udfindelse og anlæg 1646 intil aar 1679. In *Facsimilia scientia et technica Norvegica*, volume 29. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 40. Opptrykk av: *Journal for Politik, Natur- og Menneske-Kundskab, for Julii 1819*.

- [47] G. B. D. Johnson. Nogle ord om snedrev, snefog og snefonner, med fremstilling om maaden, hvorpaa antages at sammes skadelige virkning paa veie, paa ager- og engeland, etc., kan mangested enten forebygges eller svækkes. In *Facsimilia scientia et technica Norvegica*, volume 31. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 184. Publisert: *Christiania, 1852*.
- [48] M. J. Klein. The physics of J. Willard Gibbs in his time. In *Proceedings of The Gibbs Symposium*, pages 1–21. Yale University, May 15–17 1989.
- [49] Svein A. Knudsen. *Norske oppfinnelser*. Universitetsforlaget, Oslo, Norge, 1996.
- [50] Hermann Kopp. Investigations of the specific heat of solid bodies. *Philos. Trans. R. Soc. London*, 155:71–202, 1865.
- [51] Éric Lévénéz. Languages, Apr. 5 2006.
- [52] Maurice W. Lindauer. The evolution of the concept of chemical equilibrium from 1775 to 1923. *J. Chem. Educ.*, 39(8):384–390, aug 1962.
- [53] William Francis Magie. *A Source Book in Physics*. Source Books in the History of the Sciences. McGraw-Hill Book Company, Inc., New York, 1935.
- [54] Envold Møller. Kongeriget Norges herligheder, angaaende berg-værker osv. og anden videre landets frugtbarhed frem for forige tider. Til den ende sammenskreven, at enhver eftertænsom undersøgt deraf kand bevæges til at prise og take den rige og runde GUD for sine utallige velgjerninger. In *Facsimilia scientia et technica Norvegica*, volume 18. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 213. Publisert: *Trondhjem, trøkt Aar 1758*.
- [55] Christian Gran Molberg. Taksigelses tale paa Hendes Majestæts Enke-Dronningens Julianæ Mariæ Høye fødsels-dag den 4. sept. 1777 i anledning af det nye og lykkelige fuldførte vand-spring i Trondhiem. In *Facsimilia scientia et technica Norvegica*, volume 24. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 27. Opptrykk av: *Trondhjems Adresse Contoires Efterretninger No. 37, 1777 og Plakat af 30. Sept. 1777*.
- [56] Nannestad, (Imprimatur), F. Trende bergs-bønner, som bruges ved Røraas Kaaber-Verk paa de samtlige gruber, den første om morgenen, naar arbejderne skal anfare til arbejdet, den anden om eftermiddagen, naar de igjen anfarer, den tredje om aftenen, naar arbejderne kommer af gruben igjen. In *Facsimilia scientia et technica Norvegica*, volume 8. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 041. Publisert: *Trondhiem, 1751*.
- [57] Thomas Oikonomou and Gokhan B. Bagci. Clausius versus Sackur – Tetrode entropies, 2013. Accepted for publication in *Studies in History and Philosophy of Modern Physics Part B*.
- [58] Eric Olsen. Kort underretning om hellebrudd, hellernes behandling og oeconomicke nytte, samt om huul-veyters anlæggelse med brake eller eener. In *Facsimilia scientia et technica Norvegica*, volume 35. NTH-trykk, Trondheim, 1971. Faksimile eksemplar 527. Publisert: *Bergen, 1779*.
- [59] O. Phil. Anviisning til, paa en ny og forbedret maade, at anlægge og drive teglverk. In *Facsimilia scientia et technica Norvegica*, volume 37. NTH-trykk, Trondheim, 1973. Faksimile eksemplar 136. Publisert: *Kjøbenhavn, 1802*.
- [60] de Solla Price, Derek J. *Little Science, Big Sciences*. Columbia University Press, New York, 1963.
- [61] Quilez-Pardo, Juan and Joan Josep Solaz-Portolés. Students' and teachers' misapplication of le chatelier's principle: Implications for the teaching of chemical equilibrium. *Journal of Research in Science Teaching*, 32(9):939–957, 1995.
- [62] Otto Redlich. The so-called zeroth law of thermodynamics. *J. Chem. Educ.*, 47:740–741, 1970.
- [63] Otto Redlich. Science and mathematics. *J. Chem. Educ.*, 49:222–225, 1972.
- [64] [Christian Ditlev] Reventlow, Moltke, [Morten] Wormskiold, [Mads] Fridsch, and M. von Essen. Berganordning for kongeriget Norge. In *Facsimilia scientia et technica Norvegica*, volume 19. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 213. Publisert: *Fredriksberg, den 7 de September 1812*.
- [65] [Christian Ditlev] Reventlow, [Carsten] Anker, and [J. F. A.] Schiffmann. Foundation for det Kongelige Norske Berg-Seminarium paa bergstaden Kongsberg. In *Facsimilia scientia et technica Norvegica*, volume 23. NTH-trykk, Trondheim, 1966. Publisert: *Christiansborg Slot, den 3die May 1786*.
- [66] Jens Rynning. Tanker over tangbrændingens indflydelse paa fiskerierne og agerdyrkningen, fremsat til det oplyste publikums overveelse. In *Facsimilia scientia et technica Norvegica*, volume 7. NTH-trykk, Trondheim, 1962. Faksimile eksemplar ?? (manko per 14-03-23). Opptrykk av: *Magasin for Næringsstanden, Trondhiem 1803*. Tildelt *Det Kongelige Norske Videnskabers Selskabs pris på 200 Rdlr*.
- [67] J. Sæchmann. Veiledning til dreiel- og teppen-vævning for begyndere. In *Facsimilia scientia et technica Norvegica*, volume 14. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 86. Publisert: *Levanger, 1849*.
- [68] Jacob Henric Schou. Berg-ordinance i Norge. In *Facsimilia scientia et technica Norvegica*, volume 17. NTH-trykk, Trondheim, 1964. Faksimile eksemplar ?? (manko per 14-03-23). Opptrykk av: Jacob Henric Schou. *Chronologisk Register over de Kongelige Forordninger og aabne Breve samt andre trykte Anordninger som fra Aar 1670 af ere udkomne, tilligemed et nøiagtigt Udtog af de endnu giældende,*

- for saavidt samme i Almindelighed angaae Undersaatterne i Danmark og Norge, forsynet med et alphabetisk Register. I Deel. Som indeholder K. Christian V Frr. fra 1670 til 1699 samt nogle før Hans Tid udkomne Anordninger. Anden Udgaave. Kiøbenhavn, 1795 side 368–376. Publisert: *Hafniæ die 23. Junij Anno 1683*.
- [69] **Heinrich Schmidhuber**. Bericht über das Kobaltwerk Snarum in Norwegen nach denen in monat juni 1846 daselbst gesammelten unterlagen bearbeitet von Heinrich Schmidhuber, k. s. berggeschwornem. In *Facsimilia scientia et technica Norvegica*, volume 38. NTH-trykk, Trondheim, 1974. Faksimile eksemplar 508. Publisert: *Leipzig, 1847*.
- [70] **Lars Monsen Sindberg**. En berg-wiise, over det ypperlige og dyrebare kongsberg sølv-verck/ udi norge/ med grubernes og bergs-arbeydernes navne, under den tone: Guds godhed ville vi priise. In *Facsimilia scientia et technica Norvegica*, volume 26. NTH-trykk, Trondheim, 1967. Eksemplar nummer 258.
- [71] **Hans Strøm**. Om norske fabriker og om fabriker i Aggershuus-Stift i Norge som opprinnelig ble utgitt i samleren, et ugeskrift første og andet bind, henholdsvis side 321–327, 113–136. In *Facsimilia scientia et technica Norvegica*, volume 32. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 109. Publisert: *Kiøbenhavn, 1787–1788*.
- [72] **P. G. Tait**. *Sketch of Thermodynamics*. Edmonton and Douglas, Edinburgh, 1868.
- [73] **Frederik Thaarup**. Det norske kompagnies octrojer og privilegier m.v. In *Facsimilia scientia et technica Norvegica*, volume 16. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 5. Opptrykk av: *Det norske Kompagnies Octrojer og Privilegier m.v., Magazin for Danmarks og Norges topographiske, oekonomiske og statistiske Beskrivelse, Andet Bind, Kiøbenhavn, 1802–1803* side 78–117.
- [74] **Peter Thorstensøn**. Peter Thorstensøns tale holdt ved innvielsen av det Kongelige Norske Berg-Seminarium den 27. mai 1786. In *Facsimilia scientia et technica Norvegica*, volume 4. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 430. Opptrykk av: *Minerva, et Maanedsskift, Anden Aargang Tredie Bind for Martii Maaned 1787* side 308–346.
- [75] **R. H. Thurston**. *A History of the Growth of the Steam Engine*. D. Appleton & Company, 1878.
- [76] **Jos Uffink**. Bluff your way in the second law of thermodynamics. *Stud. Hist. Phil. Mod. Phys.*, 32(3):305–394, 2001.
- [77] **Johan Friederich Christian Voss**. Det frivillige slaverie, eller kort begrep om Sølvverket Kongsberg nærværende tilstand. In *Facsimilia scientia et technica Norvegica*, volume 33. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 330. Publisert: *Kiøbenhavn, 1771*.
- [78] **Johannes D. Van der Waals**. The equation of state for gases and liquids, Dec. 1910. Nobel Lecture.
- [79] **Söfren Wesing**. En kort underretning om brand-redskabets brug og anbringelse i nøds-tilfælde samt hvad ildebrande, som er overgaaet Trondhiems bye fra ao. 1328 og til 1740. In *Facsimilia scientia et technica Norvegica*, volume 20. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 5. Publisert: *Trondhjem, 1740*.
- [80] **Söfren Wesing**. Kurtze, doch warhafftige beschreibung der silber und anderer bergwercke in Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 21. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 98. Publisert: 2. utgave fra 1685.
- [81] **A. S. Wightman**. On the prescience of J. Willard Gibbs. In *Proceedings of The Gibbs Symposium*, pages 23–38. Yale University, May 15–17 1989.

### Reviews (review(s))

Relevant reviews on thermodynamics, physics, mathematics and chemistry. And of experimental techniques and experimental data. To a lesser extent, there are also articles dealing with the historical development of science.

**Related keywords:** review | reviews.

#### Bibliography.

- [1] **Tarek H. Ahmed**. Comparative study of eight equations of state for predicting hydrocarbon volumetric phase behavior. *SPE Reservoir Eng.*, 3(1):337–348, Feb 1988.
- [2] **Selby Angus**. Guide for the preparation of thermodynamic tables and correlations of the fluid state. *CODATA Bull.*, 51:1–43, Dec 1983. IUPAC Thermodynamic Tables Project Centre, Department of Chemical Engineering and Chemical Technology, Imperial College of Science and Technology, London.
- [3] **Ibrahim Ansara**, Ben Burton, Qing Chen, Mats Hillert, Armando Fernandez Guillermet, Suzana G. Fries, Hans Leo Lukas, Hans Jürgen Seifert, and W. Alan Oates. Applications of computational thermodynamics. Group 2: Models for composition dependence. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):19–40, 2000.
- [4] **J. de Swaan Arons**. Fascinating phenomena in the critical region of mixtures. *Fluid Phase Equilib.*, 52:319–336, 1989.

- [5] Christopher W. Bale and Gunnar Erikson. Metallurgical thermochemical databases—A review. *Can. Metall. Q.*, 29(2):105–132, 1990.
- [6] Rubin Battino and H. Lawrence Clever. The solubility of gases in liquids. *Chem. Rev. (Washington, D. C.)*, 66:395–463, 1966.
- [7] R. B. Berman, M. Engi, H. J. Greenwood, and T. H. Brown. Derivation of internally-consistent thermodynamic data by the technique of mathematical programming: A review with application to the system  $\text{MgO}-\text{SiO}_2-\text{H}_2\text{O}$ . *J. Petrol.*, 27(6):1331–1364, 1986.
- [8] S. Browne, J. Ziegler, and J. E. Shepherd. Numerical solution methods for shock and detonation jump conditions. Technical report, Aeronautics and Mechanical Engineering, California Institute of Technology, Pasadena, CA USA 91125, 2008. GALCIT Report FM2006.006 July 2004—Revised August 21, 2008.
- [9] Jerald A. Caton and Joe E. West. A review and thermodynamic analysis of an external combustion, reciprocating engine. In *1996 Spring Technical Conference of the ASME*, volume 3, pages 43–55, 1996.
- [10] A. Chapoy, A. H. Mohammadi, A. Chareton, B. Tohidi, and D. Richon. Measurement and modeling of gas solubility and literature review of the properties for the carbon dioxide–water system. *Ind. Eng. Chem. Res.*, 43:1794–1802, 2004.
- [11] K. C. Cheng. Historical development of the theory of heat and thermodynamics: Review and some observations. *Heat Transfer Eng.*, 13(3):19–37, 1992.
- [12] Peter Clark. Review: Elkana on Helmholtz and the conservation of energy. *The British Journal for the Philosophy of Science*, 27(2):165–176, jun 1976.
- [13] Henry Jermain Maude Creighton. How the nitrogen problem has been solved. *J. Franklin Inst.*, 187(6):705–735, 1919.
- [14] J. L. Daridon, H. Saint-Guirons, B. LaGourette, P. Xans, and C. Leibovici. A generalized process for phase equilibrium calculation with cubic equations of state. *Int. J. Thermophys.*, 14(4):1101–1108, 1993.
- [15] Gary W. Dilay and Robert A. Heidemann. Calculation of Joule–Thomson inversion curves from equations of state. *Ind. Eng. Chem. Fundam.*, 25(1):152–158, 1986.
- [16] Shalom Eliezer, Ajoy Ghatak, and Heinrich Hora. *Fundamentals of Equations of State*. World Scientific Publishing, Singapore, 2002.
- [17] H. T. Esendal. The selection of first programming language, 1993.
- [18] Alfred W. Francis. Liquid–liquid equilibrium. In Herbert M. Schoen and John J. Mcketta, Jr., editors, *Interscience Library of Chemical Engineering and Processing*, volume 3. John Wiley & Sons, New York, 1963.
- [19] Jorge F. Gabitto and Costas Tsouris. Physical properties of gas hydrates: A review. *J. Thermodynamics*, pages 1–12, 2010. ID 271291.
- [20] Rajeev Gautam and Warren D. Seider. Computation of phase and chemical equilibrium. Part I. Local and constrained minima in Gibbs free energy. *AIChE J.*, 25(6):991–999, Nov. 1979.
- [21] H. Gaye, J. Lehmann, P. Rocabois, and F. Ruby-Meyer. Slag modelling and industrial applications. *High Temp. Mater. Processes*, 20(3-4):285–292, 2001.
- [22] A. N. Gorban, G. S. Yablonskii, and V. I. Bykov. The path to equilibrium. *Int. Chem. Eng.*, 22(2):368–375, Apr 1982.
- [23] A. Harmens. Prediction of multicomponent low-temperature phase equilibria—A comparison of computation techniques. *Cryotech 73, Prod. Use Ind. Gases, Proc. Conf.*, 14:91–94, 1974.
- [24] A. Harmens. Phase equilibria from equation of state: Industrial application in cryogenics. In *Phase Equilibria and Fluid Properties in the Chemical Industry*, pages 379–388. European Federation of Chemical Engineers (EFCE), 1980. Proceedings Part II. Manuscripts of Invited Papers.
- [25] Tore Haug-Warberg. Computer science buzz-words, Sep. 2004.
- [26] R. W. Haywood. A critical review of the theorems of thermodynamic availability, with concise formulations. *J. Mech. Eng. Sci.*, 16(3):160–173, 1974.
- [27] R. A. Heidemann. Computation of high pressure phase equilibria. *Fluid Phase Equilib.*, 14:55–78, 1983.
- [28] Richard H. Heist and Honghai He. Review of vapor to liquid homogeneous nucleation experiments from 1968 to 1992. *J. Phys. Chem. Ref. Data*, 23(5):781–905, 1994.
- [29] Harold C. Helgeson, Joan M. Delany, H. Wayne Nesbitt, and Dennis K. Bird. Summary and critique of the thermodynamic properties of rock-forming minerals. *Am. J. Sci.*, 278A:1–229, 1978.
- [30] W. B. Holzapfel. Equations of state for solids under strong compression. *High Pressure Research*, 16:81–126, 1998.
- [31] Paul Hudak and Mark P. Jones. Haskell vs. Ada vs. C++ vs. Awk vs. . . . An experiment in software prototyping productivity, Jul. 1994.
- [32] Paul Johnson. Why Eiffel is better than C++ (for big projects), Mar. 1998.

- [33] Ursula R. Kattner, Gunnar Eriksson, Iris Hahn, Rainer Schmid-Fetzer, Bosse Sundman, Varghese Swamy, and Armin Kussmaul. Applications of computational thermodynamics. Group 4: Use of thermodynamic software in process modelling. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):55–94, 2000.
- [34] D. Kaustub and David Grimes. A comparison of object oriented scripting languages: Python and Ruby, Dec. 2001.
- [35] R. Krishna and J. A. Wesselingh. The Maxwell–Stefan approach to mass transfer. *Chem. Eng. Sci.*, 52(6):861–911, 1997.
- [36] Gerard D. C. Kuiken. *Thermodynamics of Irreversible Processes: Applications to Diffusion and Rheology*. John Wiley & Sons, New York, 1994.
- [37] K. Hemanth Kumar and Kenneth E. Starling. Comments on: “Cubic equations of state—Which?”. *Ind. Eng. Chem. Fundam.*, 19(1):128–129, 1980.
- [38] C. K. Kwok and D. R. Tilley. A review of some thermodynamic properties of the Van der Waals gas. *Phys. Educ.*, 14:422–425, 1979.
- [39] G. Lebon and P. Mathieu. Comparison of diverse theories of nonequilibrium thermodynamics. *Int. Chem. Eng.*, 23(4):651–662, Oct 1983.
- [40] Éric Lévénéz. Languages, Apr. 5 2006.
- [41] Maurice W. Lindauer. The evolution of the concept of chemical equilibrium from 1775 to 1923. *J. Chem. Educ.*, 39(8):384–390, aug 1962.
- [42] Stanislaw Malanowski and Andrzej Anderko. *Modelling Phase Equilibria. Thermodynamic Background and Practical Tools*. Wiley Series in Chemical Engineering. John Wiley & Sons, New York, 1992.
- [43] Joseph J. Martin. Cubic equations of state—Which? *Ind. Eng. Chem. Fundam.*, 18(2):81–97, 1979.
- [44] Joseph J. Martin. Comments on: “Cubic equations of state—Which?”. *Ind. Eng. Chem. Fundam.*, 19(1):130–131, 1980.
- [45] A. E. Mather. Phase equilibria and chemical reaction. *Fluid Phase Equilib.*, 30:83–100, 1986.
- [46] P. M. Mathias and M. S. Benson. Computational aspects of equations of state: Fact and fiction. *AIChE J.*, 32(12):2087–2090, Dec. 1986.
- [47] B. C. McGee, M. L. Hobbs, and M. R. Baer. Exponential 6 parameterization for the JCZ3-EOS. Technical report, Energetic and Multi-Phase Processes, Sandia National Laboratories, P. O. Box 5800 Albuquerque, New Mexico 87185-0834, jul 1998. SAND98-1191.
- [48] Georges A. Melhem, Riju Saini, and Bernhard M. Goodwin. A modified Peng–Robinson equation of state. *Fluid Phase Equilib.*, 47:189–237, 1989.
- [49] Bernd Mösl. A comparison of C++, FORTRAN 90 and Oberon-2 for scientific programming. Technical report, Arithmetica, May 1995.
- [50] Michael Neumann. A comparison between BETA, C++, Eiffel, Java, Object Pascal, Ruby and Smalltalk, Jul. 2000.
- [51] Darrell Kirk Nordstrom and James L. Munoz. *Geochemical Thermodynamics*. Blackwell Scientific Publications, Oxford, London, second edition, 1994.
- [52] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part I. Vapour–liquid equilibria. *Comput. Chem. Eng.*, 8(3/4):147–156, 1984.
- [53] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part II. Liquid–liquid and solid–liquid equilibria. *Comput. Chem. Eng.*, 8(3/4):157–162, 1984.
- [54] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part III. Multiphase equilibria. *Comput. Chem. Eng.*, 8(3/4):163–170, 1984.
- [55] Hasan Orbey and Stanley I. Sandler. A comparison of various cubic equation of state mixing rules for the simultaneous description of excess enthalpies and vapor–liquid equilibria. *Fluid Phase Equilib.*, 121:67–83, 1996.
- [56] David A. Palmer. *Handbook of Applied Thermodynamics*. CRC Press, Inc., Boca Raton, Florida, 1987.
- [57] V. L. Pan’kov and W. Ullmann. A comparative method for various approaches to the isothermal equation of state. *Pageoph*, 117:1001–1010, 1979.
- [58] Oliver Pfohl, Tim Giese, Ralf Dohrn, and Gerd Brunner. 1. Comparison of 12 equations of state with respect to gas-extraction processes: Reproduction of pure-component properties when enforcing the correct critical temperature and pressure. *Ind. Eng. Chem. Res.*, 37(8):2957–2965, 1998.
- [59] John M. Prausnitz, Ruediger N. Lichtenthaler, and Edmundo Gomes de Azevedo. *Molecular Thermodynamics of Fluid-Phase Equilibria*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1986.
- [60] Patrick Reany. Structured differentiation, 1987.
- [61] Patrick Reany. A structured differentiation for physicists (revised). *Arizona Journal of Natural Philosophy (AJNP)*, 4:4–20, jan 1992.
- [62] Patrick Reany. More structured differentiation. *Arizona Journal of Natural Philosophy (AJNP)*, 8:16–32, apr 1996.

- [63] Otto Redlich. The so-called zeroth law of thermodynamics. *J. Chem. Educ.*, 47:740–741, 1970.
- [64] Otto Redlich. Science and mathematics. *J. Chem. Educ.*, 49:222–225, 1972.
- [65] Syed S. H. Rizvi and Robert A. Heidemann. Vapor–liquid equilibria in the ammonia–water system. *J. Chem. Eng. Data*, 32(2):183–191, 1987.
- [66] J. S. Rowlinson. *Liquids and Liquid Mixtures*. Butterworths Scientific Publications. Butterworths Publications Ltd., London, 1959.
- [67] Paul H. Salim and Mark A. Trebble. A modified Trebble–Bishnoi equation of state: Thermodynamic consistency revisited. *Fluid Phase Equilib.*, 65:59–71, 1991.
- [68] G. Sarton. The discovery of the law of conservation of energy. *Isis*, 13:18–44, sep 1929. Facsimile reproductions of J. R. Mayer, J. P. Joule and Sadi Carnot.
- [69] M. Schacham, S. Macchietto, L. F. Stutzman, and P. Babcock. Equation oriented approach to process flowsheeting. *Comput. Chem. Eng.*, 6(2):79–85, 1982.
- [70] Heinz W. Schmidt and Stephen M. Omohundro. CLOS, Eiffel, and Sather: A comparison. Technical Report TR-91-047, International Computer Science Institute, Berkeley, California, Sep. 1991.
- [71] Warren D. Seider, Rajeev Gautam, and Charles W. White, III. Computation of phase and chemical equilibrium: A review. *ACS Symp. Ser.*, 124:115–134, 1980.
- [72] Jan V. Sengers and Anneke Levelt Sengers. The critical region. *Chem. Eng. News*, pages 104–118, Jun. 1968.
- [73] M. Shacham, S. Macchietto, L. F. Stutzman, and P. Babcock. Review. Equation oriented approach to process flowsheeting. *Comput. Chem. Eng.*, 6(2):79–95, 1982.
- [74] William R. Smith. The computation of chemical equilibria in complex systems. *Ind. Eng. Chem. Fundam.*, 19(1):1–10, 1980.
- [75] Buford D. Smith, Ol Muthu, Ashok Dewan, and Matthew Gierlach. Critical evaluation of vapor–liquid equilibrium, heat of mixing, and volume change of mixing data. General procedures. *J. Phys. Chem. Ref. Data*, 11(3):941–951, 1982.
- [76] Jannike Solsvik and Hugo A. Jakobsen. A review of the concepts for deriving the equations of change from the classical kinetic theory of gases: Single-component, multicomponent, and reactive gases. *European Journal of Mechanics B/Fluids*, 56:46–65, 2016.
- [77] G. R. Somayajulu, A. P. Kudchadker, and B. J. Zwolinski. Thermodynamics. *Annu. Rev. Phys. Chem.*, 16:213–244, 1965.
- [78] Philip J. Spencer, Tim J. Anderson, Tim G. Chart, André Costa e Silva, Bo Jansson, Beyong-Joo Lee, and Mikael Schalin. Applications of computational thermodynamics. Group 5: New applications of thermodynamic calculations. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):55–94, 2000.
- [79] Bo Sundman. Review of alloys modelling. *An. Fis., Ser. B*, 86:69–82, 1990.
- [80] P. G. Tait. *Sketch of Thermodynamics*. Edmonton and Douglas, Edinburgh, 1868.
- [81] Viktor Tekáč, Ivan Cibulka, and Robert Holub. *pvt* properties of liquids and liquid mixtures: A review of the experimental methods and the literature data. *Fluid Phase Equilib.*, 19:33–149, 1985.
- [82] Josef Templ. Oberon-2 vs. C++, Oct. 1994.
- [83] Jefferson W. Tester and Michael Modell. *Thermodynamics and Its Applications*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall PTR, Upper Saddle River, New Jersey, third edition, 1997.
- [84] Kaj Thomsen. Thermodynamics of electrolyte solutions. Department of Chemical Engineering, Technical University of Denmark, 2006.
- [85] Constantine Tsonopoulos and J. M. Prausnitz. Equations of state. A review for engineering applications. *Cryogenics*, 9(5):315–327, Oct. 1969.
- [86] Jos Uffink. Bluff your way in the second law of thermodynamics. *Stud. Hist. Phil. Mod. Phys.*, 32(3):305–394, 2001.
- [87] José O. Valderrama. The state of the cubic equations of state. *Ind. Eng. Chem. Res.*, 42(8):1603–1618, 2003.
- [88] Zhi-Chang Wang, Reinhard Lück, and Bruno Predel. New models for computing thermodynamic properties and phase diagrams of ternary systems. Part 1. Three-factor models. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(3):217–234, 1990.
- [89] Zhi-Chang Wang, Reinhard Lück, and Bruno Predel. New models for computing thermodynamic properties and phase diagrams of ternary systems Part 2. Multi-factor models. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(3):235–256, 1990.
- [90] Emmerich Wilhelm and Trevor Letcher, editors. *Volume Properties. Liquids, Solutions and Vapours*. The Royal Society of Chemistry, 2015.
- [91] Emmerich Wilhelm, Rubin Battino, and Robert J. Wilcock. Low-pressure solubility of gases in liquid water. *Chem. Rev. (Washington, D. C.)*, 77(2):219–262, 1977.

- [92] [F. van Zeggeren](#) and S. H. Storey. *The Computation of Chemical Equilibria*. Cambridge University Press, Cambridge, London, 1970.

### Miscellaneous (Misc)

The dreaded “misc” folder — stuff that doesn’t fit anywhere else. Every respectable database has a black hole, but that doesn’t mean it is an ideal solution. It is just needed.

**Related keywords:** `misc` | `miscellaneous`.

#### Bibliography.

- [1] [Samia Al-Shayban](#). Aristotle biography: Immortal philosopher of antiquity. English Department, King Saud University, Riyadh City. Kingdom of Saudi Arabia, 2008.
- [2] [D. Chr. Asbjørnsen](#). Torv og torvdrift. In *Facsimilia scientia et technica Norvegica*, volume 43. NTH-trykk, Trondheim, 1979. Faksimile eksemplar 430. Publisert: *Christiania, 1868*.
- [3] [W. A. Bentley](#) and W. J. Humphreys. *Snow Crystals*. Dover Publications, Inc., New York, 1962. Reprint of 1931 edition published by McGraw-Hill Book Company, Inc.
- [4] [David Eberhard Bradt](#). Om de norske glasværker i Aggershuus Stift i august 1781. In *Facsimilia scientia et technica Norvegica*, volume 6. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 124. Opptrykk av: *Minerva, I, 1790* side 208–237.
- [5] [David Eberhard Bradt](#). Om guldværket i Edsvold i Norge. In *Facsimilia scientia et technica Norvegica*, volume 27. NTH-trykk, Trondheim, 1967. Faksimile eksemplar 107. Opptrykk av: *Minerva* september 1790.
- [6] [David Eberhard Bradt](#). Beskrivelse over Walløe-saltverk, beliggende i Jarlsberg grevskap, en halv mil fra den ældgamle kjøbstad Tønsberg. In *Facsimilia scientia et technica Norvegica*, volume 28. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 147. Opptrykk av: *Minerva* mars 1790.
- [7] [Vincent Stoltenberg Bull](#). Viidenskabernes fordeele i den selskabelige omgang. In *Facsimilia scientia et technica Norvegica*, volume 36. NTH-trykk, Trondheim, 1972. Faksimile eksemplar 65. Publisert: *Christiania, 1791*.
- [8] [Andreas Bull](#). Oekonomiske tanker om fabrikkvæsenet og raae produkters forarbeidelse i landet. In *Facsimilia scientia et technica Norvegica*, volume 39. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 153. Publisert: *Kjøbenhavn, 1786*.
- [9] [John A. Eddy](#). The maunder minimum. *Science*, 192(4245):1189–1202, 18 Jun. 1976.
- [10] [F. M. Feldhaus](#). Otto, Nikolaus. *Allgemeine Deutsche Biographie*, 52:734–735, 1906.
- [11] [Martin Richard Flor](#). Bidrag til kundskab om naturvidenskabens fremskridt i Norge, et indbydelseskraft til den offentlige examen i Christiania Kathedralskole i september 1813. In *Facsimilia scientia et technica Norvegica*, volume 11. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 500.
- [12] [James Ford](#). Thomas Newcomen: Inventor and baptist minister, 1663–1729. In Geo. P. Gould, editor, *Transactions of the Baptist Historical Society 2.2*. Biblical Studies, oct 1911.
- [13] [Martin Gardner](#). *Fads and Fallacies in the Name of Science*. Dover Publications, Inc., New York, 1957. Reprint of 1952 edition published by G. P. Putnam’s Sons under the title: *In the Name of Science*.
- [14] [Staffan Hansson](#). *Teknik-Historia. En historia om teknisk kunnande och dess betydelse för individ och samhälle från äldsta tid fram till 1900-talet*. Studentlitteratur, Lund, Sweden, 1990.
- [15] [Odd Einar Haugen](#). Norrøn grammatikk i hovuddrag, 2009. Revidert utgåve, 16. juni 2009.
- [16] [Christian Ernst Heltzen](#). Indbydelse, plan og convention til et norskt skierper og berbygnings gewerckskab. In *Facsimilia scientia et technica Norvegica*, volume 9. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 001. Opptrykk av: *Indbydelse, Plan og Convention til et Norskt Skierper og Berbygnings Gewerckskab, Christiania 1782*.
- [17] [Peder Harboe Hertzberg](#). Underretning for bønder i Norge om den meget nyttige jord-frugt potatoes. In *Facsimilia scientia et technica Norvegica*, volume 3. NTH-trykk, Trondheim, 1961. Opptrykk av: *Samling av Adskillige Andres Skrifter, Mag. Hans Mossins Bogtrykkerie, Andet Oplag, Bergen 1778*.
- [18] [Teofil Holka](#). Heißluftmotor nach Manson-bauart. *Journal Dampf & Heißluft*, 3:32–37, 2014.
- [19] [G. B. D. Johnson](#). Nogle ord om snedrev, snefog og snefonner, med fremstilling om maaden, hvorpaa antages at sammes skadelige virkning paa veie, paa ager- og engeland, etc., kan mangested enten forebygges eller svækkes. In *Facsimilia scientia et technica Norvegica*, volume 31. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 184. Publisert: *Christiania, 1852*.
- [20] [John Ericsson](#): 1803–1889. John Ericsson Sällskapet, 2011.
- [21] [Sprenningslære](#). Del 1. KA skolen.
- [22] [Sprenningslære](#). Del 2. KA skolen.

- [23] Fjellsprengningsarbeid. Universitetsforlaget.
- [24] Svein A. Knudsen. *Norske oppfinnelser*. Universitetsforlaget, Oslo, Norge, 1996.
- [25] Carl Krafft. Niels Henrik Abel. *Naturen: Et illustreret Maanedsskrift for populær Naturvidenskab*, 8(1):1–4, Jan. 1884.
- [26] Kenneth Libbrecht. *The Little Book of Snowflakes*. Voyageur Press, 2004.
- [27] Carl Lira. Biography of James Watt. Prepared as a supplement to: Introductory Chemical Engineering Thermodynamics, 2001.
- [28] Envold Møller. Kongeriget Norges herligheder, angaaende berg-værker osv. og anden videre landets frugtbarhed frem for forige tider. Til den ende sammenskreven, at enhver eftertænksom undersaat deraf kand bevæges til at prise og take den rige og runde GUD for sine utallige velgjerninger. In *Facsimilia scientia et technica Norvegica*, volume 18. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 213. Publisert: *Trondhjem, trøkt Aar 1758*.
- [29] Christian Gran Molberg. Taksigelses tale paa Hendes Majestæts Enke-Dronningens Julianæ Mariæ Høye fødsels-dag den 4. sept. 1777 i anledning af det nye og lykkelige fuldførte vand-spring i Trondhiem. In *Facsimilia scientia et technica Norvegica*, volume 24. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 27. Optrykk av: *Trondhjems Adresse Contoires Efterretninger No. 37, 1777 og Plakat af 30. Sept. 1777*.
- [30] Nannestad, (Imprimatur), F. Trende bergs-bønner, som bruges ved Røraas Kaaber-Verk paa de samtlige gruber, den første om morgenen, naar arbejderne skal anfare til arbejdet, den anden om eftermiddagen, naar de igjen anfarer, den tredje om aftenen, naar arbejderne kommer af gruben igjen. In *Facsimilia scientia et technica Norvegica*, volume 8. NTH-trykk, Trondheim, 1962. Faksimile eksemplar 041. Publisert: *Trondhiem, 1751*.
- [31] J. J. O'Connor and E. F. Robertson. Evangelista Torricelli. MacTutor History of Mathematics archive, University of St Andrews, nov 2002.
- [32] J. J. O'Connor and E. F. Robertson. Denis Papin. MacTutor History of Mathematics archive, University of St Andrews, dec 1996.
- [33] Eric Olsen. Kort underretning om hellebrudd, hellernes behandling og oekonomiske nytte, samt om huul-veyters anleggelse med brake eller eener. In *Facsimilia scientia et technica Norvegica*, volume 35. NTH-trykk, Trondheim, 1971. Faksimile eksemplar 527. Publisert: *Bergen, 1779*.
- [34] Evangelos Papadopoulos. Heron of Alexandria (ca. 10–85 AD). Department of Mechanical Engineering, National Technical University of Athens, Greece, 2011.
- [35] O. Phil. Anviisning til, paa en ny og forbedret maade, at anlægge og drive teglverk. In *Facsimilia scientia et technica Norvegica*, volume 37. NTH-trykk, Trondheim, 1973. Faksimile eksemplar 136. Publisert: *Kjøbenhavn, 1802*.
- [36] de Solla Price, Derek J. *Little Science, Big Sciences*. Columbia University Press, New York, 1963.
- [37] [Christian Ditlev] Reventlow, [Carsten] Anker, and [J. F. A.] Schiffmann. Foundation for det Kongelige Norske Berg-Seminarium paa bergstaden Kongsberg. In *Facsimilia scientia et technica Norvegica*, volume 23. NTH-trykk, Trondheim, 1966. Publisert: *Christiansborg Slot, den 3die May 1786*.
- [38] Nils Kristian Rossing. *Pendeltegning*. Vitensenteret, Trondheim, 2003. Rev. 2.2 27.03.03.
- [39] Jacob Rosted. Physisk og oekonomisk beskrivelse over alumverket ved Opslo. In *Facsimilia scientia et technica Norvegica*, volume 44. NTH-trykk, Trondheim, 1980. Faksimile eksemplar 231. Optrykk av: *Topographisk Journal for Kongeriket Norge, Hefte 3, Christiania 1793* side 1–74.
- [40] Jens Rynning. Tanker over tangbrændingens indflydelse paa fiskerierne og agerdyrkingen, fremsat til det oplyste publikums overveielse. In *Facsimilia scientia et technica Norvegica*, volume 7. NTH-trykk, Trondheim, 1962. Faksimile eksemplar ?? (manko per 14-03-23). Optrykk av: *Magasin for Næringsstanden, Trondhiem 1803*. Tildelt *Det Kongelige Norske Videnskabers Selskabs* pris på 200 Rdlr.
- [41] J. Sæchnann. Veiledning til dreiel- og teppen-vævning for begyndere. In *Facsimilia scientia et technica Norvegica*, volume 14. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 86. Publisert: *Levanger, 1849*.
- [42] Jacob Henric Schou. Berg-ordinance i Norge. In *Facsimilia scientia et technica Norvegica*, volume 17. NTH-trykk, Trondheim, 1964. Faksimile eksemplar ?? (manko per 14-03-23). Optrykk av: Jacob Henric Schou. *Chronologisk Register over de Kongelige Forordninger og aabne Breve samt andre trykte Anordninger som fra Aar 1670 af ere udkomne, tilligemed et nøiagtigt Udtog af de endnu giældende, for saavidt samme i Almindelighed angaae Undersaatterne i Danmark og Norge, forsynet med et alphabetisk Register. I Deel. Som indeholder K. Christian V Frr. fra 1670 til 1699 samt nogle før Hans Tid udkomne Anordninger. Anden Udgaave. Kjøbenhavn, 1795* side 368–376. Publisert: *Hafniæ die 23. Junij Anno 1683*.
- [43] Jan Sebestik. The dawn of cognitive science: Early European contributors. In Liliana Albertazzi, editor, *Synthese Library: Studies in Epistemology, Logic, Methodology, And Philosophy of Science*, volume 295, chapter Ernst Mach's Evolutionary Theory of Representation. Springer-Science, 2001.

- [44] [Lars Monsen Sindberg](#). En berg-wiise, over det ypperlige og dyrebare kongsberg sølv-verck/ udi norge/ med grubernes og bergs-arbeydernes navne, under den tone: Guds godhed ville vi priise. In *Facsimilia scientia et technica Norvegica*, volume 26. NTH-trykk, Trondheim, 1967. Eksemplar nummer 258.
- [45] [Hans Strøm](#). Om norske fabriker og om fabriker i Aggershuus-Stift i Norge som opprinnelig ble utgitt i samleren, et ugeskrivt første og andet bind, henholdsvis side 321–327, 113–136. In *Facsimilia scientia et technica Norvegica*, volume 32. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 109. Publisert: *Kiøbenhavn, 1787–1788*.
- [46] [Åse Berit Strandskogen](#) and Rolf Strandskogen. *Preposisjoner og partikler i norsk*. Oris Forlag, 1985.
- [47] [Frederik Thaarup](#). Det norske kompanies octrojer og privilegier m.v. In *Facsimilia scientia et technica Norvegica*, volume 16. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 5. Opptrykk av: *Det norske Kompanies Octrojer og Privilegier m.v., Magazin for Danmarks og Norges topographiske, oekonomiske og statistiske Beskrivelse, Andet Bind, Kiøbenhavn, 1802–1803* side 78–117.
- [48] [Peter Thorstensøn](#). Peter Thorstensøns tale holdt ved innvielsen av det Kongelige Norske Berg-Seminarium den 27. mai 1786. In *Facsimilia scientia et technica Norvegica*, volume 4. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 430. Opptrykk av: *Minerva, et Maanedsskift, Anden Aargangs Tredie Bind for Martii Maaned 1787* side 308–346.
- [49] [Encyclopedia of World Biography on Otto Von Guericke](#). Thomson Gale, a part of the Thomson Corporation, 2005.
- [50] [Söfren Wesing](#). En kort underretning om brand-redskabets brug og anbringelse i nøds-tilfælde samt hvad ildebrande, som er overgaaet Trondhiems bye fra ao. 1328 og til 1740. In *Facsimilia scientia et technica Norvegica*, volume 20. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 5. Publisert: *Trondhjem, 1740*.
- [51] [Carl von Linde](#). Who's Who, 2011.

## CHAPTER 2

# Phase and Matter

### Binodal Curves (PhMatB)

The calculation of phase boundary curves, and not to mention the complete modelling of assessed phase diagrams, covers a vast literature. In most cases, they are purpose-built numerical procedures that require clever initializations and a lot of heuristics. The ultimate difficulty, however, occurs when the phase boundary has one, or at worst several, critical endpoints.

**Related keywords: phase & boundary | binodal & curve | phase & diagram | saturation.**

#### Bibliography.

- [1] Lee E. Baker and Kraemer D. Luks. Critical point and saturation pressure calculations for multipoint systems. *Soc. Pet. Eng. J. (1980-)*, 20(1):15–24, 1980.
- [2] Domingos Barbosa and Michael F. Doherty. The influence of equilibrium chemical reactions on vapor–liquid phase diagrams. *Chem. Eng. Sci.*, 43(3):529–540, 1988.
- [3] Maria A. Barrufet and Philip T. Eubank. Generalized saturation properties of pure fluids via cubic equations of state. *Chem. Eng. Educ.*, pages 168–175, 1989.
- [4] Clifton G. Bergeron and Subsash H. Risbud. *Introduction to Phase Equilibria in Ceramics*. American Ceramic Society, Inc., Columbus, Ohio, 1984.
- [5] Thomas H. Brown, Robert G. Berman, and Ernest H. Perkins. Ge0-Calc: Software package for calculation and display of pressure–temperature–composition phase diagrams using an IBM or compatible personal computer. *Comput. Geosci.*, 14(3):279–289, 1988.
- [6] Georgios N. Charos, Paulette Clancy, Keith Gubbins, and Chandrashekhar D. Naik. Three-dimensional *ptx* phase diagrams through interactive computer graphics. *Fluid Phase Equilib.*, 23:59–78, 1985.
- [7] Shuang-Lin Chen, Kuo-Chih Chou, and Y. A. Chang. On a new strategy for phase diagram calculation. 1. Basic principles. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(3):237–250, 1993.
- [8] Shuang-Lin Chen, Kuo-Chih Chou, and Y. Austin Chang. On a new strategy for phase calculation. 2. Binary systems. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(3):287–302, 1993.
- [9] Douglas A. Clibbens. *The Principles of the Phase Theory: Heterogeneous Equilibria between Salts and Their Aqueous Solutions*. Macmillan and Co., Limited, London, 1920.
- [10] J. A. D. Connolly and D. M. Kerrick. An algorithm and computer program for calculating composition phase diagrams. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 11(1):1–55, 1987.
- [11] P. Englezos and P. R. Bishnoi. Gibbs free energy analysis for the supersaturation limits of methane in liquid water and the hydrate–gas–liquid water phase behaviour. *Fluid Phase Equilib.*, 42:129–140, 1988.
- [12] E. Ulrich Franck. Fluids at high pressures and temperatures. *J. Chem. Thermodyn.*, 19:225–242, 1987.
- [13] D. Furman and Robert B. Griffiths. Global phase diagram for a Van der Waals model of a binary mixture. *Phys. Rev. A: At., Mol., Opt. Phys.*, 17(3):1139–1148, Mar. 1978.
- [14] J. Gaite, J. Margalef-Roig, and S. Miret-Artés. Analysis of a three-component model phase diagram by catastrophe theory. *Phys. Rev. B: Solid State*, 57(21):13527–13534, jun 1998.
- [15] A. M. Haseeb, M. Keskin, I. L. Pegg, and P. H. E. Meijer. A new method for tieline calculations in binary fluid systems. *Fluid Phase Equilib.*, 120:93–105, 1996.
- [16] P. H. van Konynenburg and R. L. Scott. Critical lines and phase equilibria in binary Van der Waals mixtures. *Philos. Trans. R. Soc. London, Ser. A*, 298(1442):495–540, 1980.
- [17] Erich Königsberger and Heniz Gamsjäger. Comment: Solid-solution aqueous-solution equilibria: Thermodynamic theory and representation. *Am. J. Sci.*, 292(3):199–214, 1992.

- [18] Laar, J. J., von. Über die dampftension von flüssigen gemischen, z. b. von brom und jod, beu annahme einer teilweisen (im grenzfall nicht- oder total-)dissoziierten verbindung. *Z. Phys. Chem., Stoechiom. Verwandtschaftsl.*, 47:129–145, 1904.
- [19] Monica H. Lamm and Carol K. Hall. Molecular simulation of complete phase diagrams for binary mixtures. *AIChE J.*, 47(7):1664–1675, Jul. 2001.
- [20] D. D. Lee, J. H. Choy, and J. K. Lee. Computer generation of binary and ternary phase diagrams via a convex hull method. *J. Phase Equilib.*, 13(4):365–372, 1992.
- [21] Lindsay Leveen, William H. Abraham, and Charles W. Schelin. Multicomponent vapor–liquid equilibrium: Direct integration of the Gibbs–Duhem equation. *Chem. Eng. Sci.*, 33:1191–1199, 1978.
- [22] P.-L. Lin, A. D. Pelton, C. W. Bale, and W. T. Thompson. An interactive computer program for calculating ternary phase diagrams. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 4(1):47–60, 1980.
- [23] Li Lin, P. Wollants, O. van der Biest, and L. Delaey. A simple method for the calculation of equilibrium phase diagrams. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 18(1):89–98, 1994.
- [24] Lars Lundegaard and Jørgen Møllerup. Calculation of phase diagrams of gas-hydrates. *Fluid Phase Equilib.*, 76:141–149, 1992.
- [25] Paul H. E. Meijer. Theory of coexisting states: Calculation of binodals. *Physica (Amsterdam)*, 237(1-2):31–44, 1997.
- [26] Michael L. Michelsen. Calculation of phase envelopes and critical points for multicomponent mixtures. *Fluid Phase Equilib.*, 4:1–10, 1980.
- [27] Michael L. Michelsen. Saturation point calculations. *Fluid Phase Equilib.*, 23:181–192, 1985.
- [28] Michael L. Michelsen. A simple method for calculation of approximate phase boundaries. *Fluid Phase Equilib.*, 98:1–11, 1994.
- [29] Long X. Nghiem, Yau-Kun Li, and Robert A. Heidemann. Application of the tangent plane criterion to saturation pressure and temperature computations. *Fluid Phase Equilib.*, 21:39–60, 1985.
- [30] Hideo Nishiumi. Pressure determination of vapor–liquid equilibrium using successive substitution method. *J. Chem. Eng. Jpn.*, 12(2):210–214, 1988.
- [31] Ding-Ju Peng. Accelerated successive substitution schemes for bubble-point and dew-point calculations. *Can. J. Chem. Eng.*, 69:978–985, 1991.
- [32] Fanya Moiseevna Perel'man. *Phase Diagrams of Multicomponent Systems. Geometric Methods*. Consultants Bureau Enterprises, Inc., New York, 1966. The original text was published by Nauka Press for the N. S. Kurnakov Institute of General and Inorganic Chemistry in Moscow 1965.
- [33] Marcel Pourbaix. A method of studying involved equilibria and its application to metallurgical processes. In *XXIst Congress of Pure and Applied Chemistry, London*, pages 1–19, Jul. 1947. Translation from French by J. O'M. Bockris.
- [34] John E. Ricci. *The Phase Rule and Heterogeneous Equilibrium*. Dover Publications, Inc., New York, 1966. Reprint of 1951 edition published by D. Van Nostrand Company.
- [35] Richard J. Sadus. *High Pressure Phase Behaviour of Multicomponent Fluid Mixtures*. Elsevier, Amsterdam, Holland, 1992.
- [36] Ketan D. Samant, David A. Berry, and Ka M. Ng. Representation of high-dimensional molecular solid–liquid phase diagrams. *AIChE J.*, 46(12):2435–2455, 2000.
- [37] Ketan D. Samant and Ka M. Ng. Representation of high-dimensional solid–liquid phase diagrams of ionic systems. *AIChE J.*, 47(4):861–879, 2001.
- [38] N. Saunders and A. P. Miodownik. *CALPHAD. Calculation of Phase Diagrams: A Comprehensive Guide*. Pergamon Materials Series. Pergamon Press, Oxford, London, 1998.
- [39] H. Segura and J. Wisniak. Calculation of pure saturation properties using cubic equations of state. *Comput. Chem. Eng.*, 21(12):1339–1347, 1997.
- [40] Levelt Sengers, Johanna. *How Fluids Unmix: Discoveries by the School of Van der Waals and Kamerlingh Onnes*. Koninklijke Nederlandse Akademie van Wetenschappen, Amsterdam, 2002.
- [41] J. M. Sorenson and G. R. van Hecke. A simple method for equal Gibbs energy analysis of phase boundaries. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 18(3):329–333, 1994.
- [42] Maciej Starzak and Mohamed Mathlouthi. Temperature dependence of water activity in aqueous solutions of sucrose. *Food Chemistry*, 96:346–370, 2006.
- [43] V. V. Sychev. *The Differential Equations of Thermodynamics*. MIR Publishers, Moscow, 1983. Translated from the Russian by Eugene Yankovsky.
- [44] Vicente Talanquer. Global phase diagram for reacting systems. *J. Chem. Phys.*, 96(7):5408–5421, Apr. 1992.
- [45] Max Teubner. Global constraint and the inverse problem for ternary phase diagrams. *J. Chem. Phys.*, 96(1):555–557, Jan 1992.

- [46] A. L. Udovskii. Computer modeling of phase-diagrams, thermodynamic properties, and structure of multicomponent systems. *Russian Metallurgy*, 2:132–153, 1990. Original publication: *Izvestiya Akademii Nauk SSSR Metally*, No.2, pp.136–157, 1990.
- [47] C. R. Wang, Z. B. Zhao, S. K. Xia, W. Q. Zhang, and R. Z. Zhu. The computer algorithm and program for the generation of phase stability diagrams. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(3):257–264, 1990.
- [48] Christiano Wibowo, Ketan D. Samant, and Ka M. Ng. High-dimensional solid–liquid phase diagrams involving compounds and polymorphs. *AIChE J.*, 48(10):2179–2192, Oct. 2002.
- [49] Christiano Wibowo and Ka M. Ng. Visualization of high-dimensional phase diagrams of molecular and ionic mixtures. *AIChE J.*, 48(5):991–1000, May 2002.
- [50] Christiano Wibowo and Ka M. Ng. Visualization of high-dimensional systems via geometric modeling with homogenous coordinates. *Ind. Eng. Chem. Res.*, 41(9):2213–2225, 2002.
- [51] Ivan Wichterle. High pressure vapour–liquid equilibrium. I. Phenomenological description. Part 1. *Fluid Phase Equilib.*, 1:161–172, 1977.
- [52] Ivan Wichterle. High pressure vapour–liquid equilibrium. II. Phenomenological description. Part 2. *Fluid Phase Equilib.*, 1:225–245, 1977/1978.
- [53] Ivan Wichterle. High pressure vapour–liquid equilibrium. III. Quantitative description. Part 1. *Fluid Phase Equilib.*, 1:305–316, 1977/1978.
- [54] Ivan Wichterle. High pressure vapour–liquid equilibrium. IV. Quantitative description. Part 2. *Fluid Phase Equilib.*, 2:49–78, 1978.
- [55] Ivan Wichterle. High pressure vapour–liquid equilibrium. V. Quantitative description. Part 3. *Fluid Phase Equilib.*, 2:143–159, 1978.
- [56] Ivan Wichterle. High pressure vapour–liquid equilibrium. VI. Experimental technique. *Fluid Phase Equilib.*, 2:293–295, 1979.
- [57] Hui-an Yin. Calculation and auto-plotting of the stable equilibrium  $p$ – $t$  phase diagram of an  $n$ -component system composed of  $n + 2$  phases. *Sci. China, Ser. B: Chem., Life Sci., Earth Sci.*, 15(10):1173–1180, Oct. 1992.

### Criticality (PhMatC)

Direct calculation of critical points.

**Related keywords:** critical | tricritical.

#### Bibliography.

- [1] M. A. Anisimov, S. B. Kiselev, J. V. Sengers, and S. Tang. Crossover approach to global critical phenomena in fluids. *Physica A (Amsterdam)*, 188:487–525, 1992.
- [2] J. de Swaan Arons. Fascinating phenomena in the critical region of mixtures. *Fluid Phase Equilib.*, 52:319–336, 1989.
- [3] James T. Bartis. Thermodynamic equations for tri- or third order critical points. *J. Chem. Phys.*, 59(10):5423–5430, Nov. 1973.
- [4] D. S. Billingsley and S. Lam. Critical point calculation with nonzero interaction parameters. *AIChE J.*, 32(8):1393–1396, Aug 1986.
- [5] W. P. H. de Boer. The Griffiths theory of tricritical phase behavior and its application to a model microemulsion system. *J. Chem. Phys.*, 91(11):7167–7172, Dec. 1989.
- [6] Z. Y. Chen, P. C. Albright, and J. V. Sengers. Crossover from singular critical to regular classical thermodynamic behavior of fluids. *Phys. Rev. A: At., Mol., Opt. Phys.*, 41(6):3161–3177, Mar. 1990.
- [7] Binny J. Cherayil. Critical phenomena in polymer solutions: Scaling of the free energy. *J. Chem. Phys.*, 72(1):9126–9133, Jun. 1993.
- [8] L. René Corrales and John C. Wheeler. Tetracritical and novel tricritical points in sulfur solutions: A Flory model for polymerization of rings and chains in a solvent. *J. Chem. Phys.*, 90(9):5030–5055, May 1989.
- [9] J. R. DiAndreth and M. E. Paulaitis. Multiphase behavior in ternary fluid mixtures: A case study of the isopropanol–water–CO<sub>2</sub> system at elevated pressures. *Chem. Eng. Sci.*, 44(5):1061–1069, 1989.
- [10] Enrique Fernandez-Fassnacht, Arthur G. Williamson, Alwarappa Sivaraman, Robert L. Scott, and Charles M. Knobler. Tricritical phenomena in quasi-binary mixtures. V. New measurements on ternary methane systems. *J. Chem. Phys.*, 86(7):4133–4137, Apr. 1987.
- [11] Michael E. Fisher and Yan Levin. Criticality in ionic fluids—Debye–Hückel theory, Bjerrum, and beyond. *Phys. Rev. Lett.*, 71(23):3826–3829, dec 1993.
- [12] M. Cynthia Goh, Jürgen Specovius, Robert L. Scott, and Charles M. Knobler. Tricritical phenomena in quasi-binary mixtures. IV. Ternary ethane systems. *J. Chem. Phys.*, 86(7):4120–4132, Apr. 1987.

- [13] M. Cynthia Goh, Robert L. Scott, and Charles M. Knobler. Tricritical phenomena in quasi-binary mixtures. VI. The binary system ethane + *n*-eicosane and some revised scaling parameters. *J. Chem. Phys.*, 89(4):2281–2285, Aug. 1988.
- [14] Sandra C. Greer. Liquid–liquid critical phenomena. *Accounts of Chemical Research*, 11:427–432, 1978.
- [15] Robert B. Griffiths and John C. Wheeler. Critical points in multicomponent systems. *Phys. Rev. A: At., Mol., Opt. Phys.*, 2(3):1047–1064, Sep. 1970.
- [16] Robert B. Griffiths. Thermodynamic model for tricritical points in ternary and quaternary fluid mixtures. *J. Chem. Phys.*, 60(1):195–206, Jan. 1974.
- [17] Robert A. Heidemann and Ahmed M. Khalil. The calculation of critical points. *AIChE J.*, 26(5):769–779, Sep. 1980.
- [18] Robert A. Heidemann. Critical points in reacting mixtures. *ACS Symp. Ser.*, pages 379–392, 1980.
- [19] Miron Kaufman and Robert B. Griffiths. Thermodynamic model for tricritical mixtures with application to ammonium sulfate + water + ethanol + benzene. *J. Chem. Phys.*, 76(3):1508–1524, Feb. 1982.
- [20] Webster B. Kay and Herbert A. Fisch. Phase relations of binary systems that form azeotropes: I. The ammonia–*n*-butane system. *AIChE J.*, 4(3):293–296, Sep. 1958.
- [21] Webster B. Kay and F. Morgan Warzel. Phase relations of binary systems that form azeotropes: II. The ammonia–isooctane system. *AIChE J.*, 4(3):296–299, Sep. 1958.
- [22] B. F. Kohse and R. A. Heidemann. Tricritical lines and multiphase equilibria in quaternary mixtures. *Fluid Phase Equilib.*, 75:11–22, 1992.
- [23] Bruce F. Kohse and Robert A. Heidemann. Computation of tricritical points in ternary systems. *AIChE J.*, 39(7):1242–1257, Jul. 1993.
- [24] I. R. Krichevski. Thermodynamic equations of a critical phase of higher order in a ternary system. *Dokl. Akad. Nauk*, 202(5):1125–1126, Feb 1972. Translated from Russian by Consultants Bureau, Plenum Publishing Corp., 227 West 17th St. New York.
- [25] Lang, Jr., J. C. and B. Widom. Equilibrium of three liquid phases and approach to the tricritical point in benzene–ethanol–water–ammonium sulfate mixtures. *Physica A (Amsterdam)*, 81:190–213, May 1975.
- [26] Z. Ludmer, R. Shinnar, and V. Yakhot. Solubility in binary mixtures at the immiscibility critical point. *AIChE J.*, 33(11):1776–1780, Nov. 1987.
- [27] Leo Lue and John M. Prausnitz. Thermodynamics of fluid mixtures near to and far from the critical region. *AIChE J.*, 44(6):1455–1466, Jun. 1998.
- [28] Leo Lue and John M. Prausnitz. Renormalization-group corrections to an approximate free-energy model for simple fluids near to and far from the critical region. *J. Chem. Phys.*, 108(13):5529–5536, Apr. 1998.
- [29] Ronald E. Marcotte and Luis C. Zepeda. Designs to demonstrate the critical state. *Chem. Eng. Educ.*, Winter:44–47, 1994.
- [30] Robert McGraw. The thermodynamic barrier to nucleation near a critical point. *J. Chem. Phys.*, 91(9):5655–5663, Nov. 1989.
- [31] Michael L. Michelsen and Robert A. Heidemann. Calculation of critical points from cubic two-constant equations of state. *AIChE J.*, 27(3):521–523, May 1981.
- [32] Michael L. Michelsen. Calculation of critical points and phase boundaries in the critical region. *Fluid Phase Equilib.*, 16:57–76, 1984.
- [33] Michael L. Michelsen and Robert A. Heidemann. Calculation of tri-critical points. *Fluid Phase Equilib.*, 39:53–74, 1988.
- [34] L. Mistura. Critical phases in multicomponent fluid mixtures. *J. Phys. A: Gen. Phys.*, 9:2139–2148, 1976.
- [35] F. Munoz and E. H. Chimowitz. Critical phenomena in mixtures. I. Thermodynamic theory for the binary critical azeotrope. *J. Chem. Phys.*, 99(7):5438–5449, Oct. 1993.
- [36] T. Narayanan and A. Kumar. On the decreasing exponent near a critical double point. *J. Chem. Phys.*, 98(2):1752–1753, Jan. 1993.
- [37] Juan J. de Pablo and John M. Prausnitz. Thermodynamics of liquid–liquid equilibria including the critical region. Transformation to non-classical coordinates using revised scaling. *Fluid Phase Equilib.*, 59:1–14, 1990.
- [38] C. L. Patton and K. D. Luks. Determination of partial molar volumes at critical end points. *Fluid Phase Equilib.*, 113:21–26, 1995.
- [39] Ian L. Pegg, Charles M. Knobler, and Robert L. Scott. Tricritical phenomena in quasibinary mixtures. VIII. Calculations from the van der Waals equation for binary mixtures. *J. Chem. Phys.*, 92(9):5442–5453, May 1990.
- [40] Ding-Yu Peng and Donald B. Robinson. A rigorous method for predicting the critical properties of multicomponent systems from an equation of state. *AIChE J.*, 23(2):137–144, Mar. 1977.

- [41] S. Peter and H. Wenzel. Calculations of phase equilibria at high pressures in the critical region. *Ber. Bunsen-Ges.*, 76(3/4):331–335, 1972.
- [42] Kenneth S. Pitzer, M. Conceicao P. de Lima, and Donald R. Schreiber. Critical point and phase separation for an ionic system. *J. Phys. Chem.*, 89(10):1854–1855, 1985.
- [43] Robert C. Reid and Bruce L. Beegle. Critical point criteria in Legendre transform notation. *AIChE J.*, 23(5):726–732, Sep. 1977.
- [44] Richard J. Sadus. Calculating critical transitions of fluid mixtures: Theory vs. experiment. *AIChE J.*, 40(8):1376–1403, Aug. 1994.
- [45] G. M. Schneider. Phase behavior and critical phenomena in fluid mixtures under pressure. *Ber. Bunsen-Ges.*, 76(3/4):325–331, 1972.
- [46] William Schotte. Solubilities near the solvent critical point. *AIChE J.*, 31(1):154–157, Jan. 1985.
- [47] R. L. Scott. Multicritical phenomena in fluid mixtures. *Proc. Symp. Thermophys. Prop.*, 1(8):397–404, 1981.
- [48] Robert L. Scott. Tricritical phenomena in quasi-binary mixtures. III. An extended classical treatment. *J. Chem. Phys.*, 86(7):4106–4119, Apr. 1987.
- [49] Jan V. Sengers and Anneke Levelt Sengers. The critical region. *Chem. Eng. News*, pages 104–118, Jun. 1968.
- [50] Levelt Sengers, J. M. H. and J. V. Sengers. Universality of critical behavior in gases. *Phys. Rev. A: At., Mol., Opt. Phys.*, 12(6):2622–2627, 1975.
- [51] A. S. Teja and H. W. Kropholler. Critical states of mixtures in which azeotropic behaviour persists in the critical region. *Chem. Eng. Sci.*, 30:435–436, 1975.
- [52] John C. Wheeler and Pierre Pfeuty. Critical points and tricritical points in liquid sulfur solutions. *J. Chem. Phys.*, 74(11):6415–6430, Jun. 1981.
- [53] K. G. Wilson. Critical phenomena in 3.99 dimensions. *Physica (Amsterdam)*, 73:119–128, 1974.

## Metastability (PhMatM)

Metastability in general and the limiting negative pressure of fluids in particular.

**Related keywords: metastable | spinodal | negative & pressure | superheating.**

### Bibliography.

- [1] V. G. Baidakov and T. A. Gurina.  $(p, \rho, t)$  of superheated liquid oxygen. *J. Chem. Thermodyn.*, 17(2):131–142, Feb. 1985.
- [2] Lyman J. Briggs. Limiting negative pressure of water. *J. Appl. Phys.*, 21:721–722, Jul. 1950.
- [3] Lyman J. Briggs. The limiting negative pressure of acetic acid, benzene, aniline, carbon tetrachloride and chloroform. *J. Chem. Phys.*, 19(7):970–972, Jul. 1951.
- [4] W. Döring. Die überhitzungsgrenze und Zerreißfestigkeit von Flüssigkeiten. *Z. Phys. Chem., Abt. B*, 36:371–386, 1937.
- [5] W. Döring. Berichtigung zu der Arbeit: Die überhitzungsgrenze und Zerreißfestigkeit von Flüssigkeiten. *Z. Phys. Chem., Abt. B*, 38:292–294, 1937.
- [6] G. V. Ermakov and V. P. Skripov. Experimental determination of the specific volumes of a superheated liquid. *High Temp. (Transl. of Teplofiz. Vys. Temp.)*, 6:86–92, 1968.
- [7] Alan T. J. Hayward. Negative pressure in liquids: Can it be harnessed to serve man? *Am. Sci.*, 59:434–443, Jul/Aug 1971.
- [8] Richard H. Heist and Honghai He. Review of vapor to liquid homogeneous nucleation experiments from 1968 to 1992. *J. Phys. Chem. Ref. Data*, 23(5):781–905, 1994.
- [9] Narayanan Menon. A simple demonstration of a metastable state. *Am. J. Phys.*, 67(12):1109–1110, dec 1999.
- [10] Julius Meyer. Zur Kenntnis des negativen Druckes in Flüssigkeiten. In W. Nernst, editor, *Abhandlungen der Deutschen Bunsen-Gesellschaft für angewandte physikalische Chemie*, band III. Verlag Chemie GmbH, Weinheim, Germany, 1911.
- [11] Jennie R. Patrick-Yeboah and Robert C. Reid. Superheat-limit temperatures of polar liquids. *Ind. Eng. Chem. Fundam.*, 20:315–317, 1981.
- [12] Robert C. Reid. Rapid phase transition from liquid to vapour. *Adv. Chem. Eng.*, 12:105–208, 1983.
- [13] V. P. Skripov and G. V. Ermakov. Pressure dependence of the limiting superheating of a liquid. *Russ. J. Phys. Chem. (Transl. of Zh. Fiz. Khim.)*, 38(2):208–213, Feb 1964.

## Phase Equilibrium (PhMatP)

Phase equilibrium calculation. Most of the literature is about 2-phase VLE and 3-phase VLLE formulations since these are important for the oil and gas industry, but there are also a good number of general algorithms.

**Related keywords: (multicomponent | multiphase) & equilibrium.**

### Bibliography.

- [1] L. Asselineau, G. Bogdanic, and J. Vidal. A versatile algorithm for calculating vapour–liquid equilibria. *Fluid Phase Equilib.*, 3:273–290, 1979.
- [2] Stefan Behme, Gabriele Sadowski, Yuhua Song, and Chau-Chyun Chen. Multicomponent flash algorithm for mixtures containing polydisperse polymers. *AIChE J.*, 49(1):258–268, 2003.
- [3] Andreas Bolz, Ulrich Deiters, Cor J. Peters, and Theo W. de Loos. Nomenclature for phase diagrams with particular reference to vapour–liquid and liquid–liquid equilibria. *Pure Appl. Chem.*, 70(11):2233–2257, 1998.
- [4] J. F. Boston and H. I. Britt. A radically different formulation and solution of the single-stage flash problem. *Comput. Chem. Eng.*, 2:109–122, 1978.
- [5] J. F. Boston. Inside-out algorithms for multicomponent separation process calculations. *ACS Symp. Ser.*, 124:135–151, 1980.
- [6] Egil Brendsdal. *Computation of Phase Equilibrium in Fluid Mixtures*. Dr.ing. thesis 1999:7, Norwegian Institute of Technology, 1999.
- [7] L. G. Bullard and L. T. Biegler. Iterated linear programming strategies for non-smooth simulation: A penalty based method for vapor–liquid equilibrium applications. *Comput. Chem. Eng.*, 17(1):95–109, 1993.
- [8] A. P. Büinz, R. Dohrn, and J. M. Prausnitz. Three-phase calculations for multicomponent systems. *Comput. Chem. Eng.*, 15(1):47–51, 1991.
- [9] M. Castier. Solutions of the isochoric–isoenergetic flash problem by direct entropy maximization. *Fluid Phase Equilib.*, 276:7–17, 2009.
- [10] P. Cavallotti, G. Celeri, B. Leonardis, and L. Gardini. Calculation of multicomponent multiphase equilibria. *Chem. Eng. Sci.*, 35:2297–2304, 1980.
- [11] Tanmoy Chakravarty, C. W. White, III, and W. D. Seider. Computation of phase equilibrium: Optimization with thermodynamic inconsistency. *AIChE J.*, 31(2):316–320, Feb. 1985.
- [12] H. H. Y. Chien. KB method in two phase flash calculations. *Comput. Chem. Eng.*, 8(1):61–64, 1984.
- [13] Henry Hung-yeh Chien. Formulations for three-phase flash calculations. *AIChE J.*, 40(6):957–965, Jun. 1994.
- [14] Wendy A. Cole and Stephen P. Goodwin. Flash calculations for gas hydrates: A rigorous approach. *Chem. Eng. Sci.*, 45(3):569–573, 1990.
- [15] Michael F. Doherty, David B. van Dongen, and James R. Haight. Response to comments on “Material stability of multicomponent mixtures and the multiplicity of solutions to phase-equilibrium equations. 1. Nonreacting mixtures”. *Ind. Eng. Chem. Fundam.*, 23(3):374–377, 1984.
- [16] David B. van Dongen, Michael F. Doherty, and James R. Haight. Material stability of multicomponent mixtures and the multiplicity of solutions to phase-equilibrium equations. 1. Nonreacting mixtures. *Ind. Eng. Chem. Fundam.*, 22(4):472–485, 1983.
- [17] E. Eckert and M. Kubíček. Multiple liquid–vapour equilibrium flash calculation by a global approach. *Comput. Chem. Eng.*, 17(9):879–884, 1993.
- [18] Michael E. Fisher. Phases and phase diagrams: Gibbs’s legacy today. In *Proceedings of The Gibbs Symposium*, pages 39–72. Yale University, May 15–17 1989.
- [19] R. L. Fournier and J. F. Boston. A quasi-Newton algorithm for solving multiphase equilibrium flash problems. *Chem. Eng. Commun.*, 8:305–326, 1981.
- [20] Lynne T. Fussell. A technique for calculating multiphase equilibria. *Soc. Pet. Eng. J.*, 19(4):203–210, Aug. 1979. Includes discussion with A. M. Rowe.
- [21] L. T. Fussell. Author’s reply. *Soc. Pet. Eng. J.*, 19(4):208–210, Aug. 1979.
- [22] P. W. Gallier, L. B. Evans, H. I. Britt, J. F. Boston, and P. K. Gupta. ASPEN: Advanced capabilities for modeling and simulation of industrial processes. *ACS Symp. Ser.*, 124:293–308, 1980.
- [23] Boyd George, L. P. Brown, C. H. Farmer, Paul Buthod, and F. S. Manning. Computation of multicomponent, multiphase equilibrium. *Ind. Eng. Chem. Process Des. Dev.*, 15(3):372–377, 1976.
- [24] G. Han and G. P. Rangaiah. A method for calculation of vapor–liquid and liquid–liquid equilibria. *Comput. Chem. Eng.*, 21(8):905–913, 1997.
- [25] Tore Haug-Warberg. Termodynamiske likevektsberegninger i CADAS, feb 1998.

- [26] Robert A. Heidemann. Three-phase equilibria using equations of state. *AIChE J.*, 20(5):847–855, Sep. 1974.
- [27] Robert A. Heidemann. Predict three-phase equilibria. *Hydrocarbon Process.*, pages 167–170, Nov. 1974.
- [28] R. A. Heidemann. Computation of high pressure phase equilibria. *Fluid Phase Equilib.*, 14:55–78, 1983.
- [29] E. M. Hendriks and A. R. D. van Bergen. Applications of a reduction method to phase equilibria calculations. *Fluid Phase Equilib.*, 74:17–34, 1992.
- [30] B. H. Jensen and A. Fredenslund. A simplified flash procedure for multicomponent mixtures containing hydrocarbons and one non-hydrocarbon using two-parameter cubic equations of state. *Ind. Eng. Chem. Res.*, 26:2129–2134, 1987.
- [31] X. Joulia, P. Maggioni, B. Koehret, H. Paradowski, and J. J. Bartuel. Hybrid method for phase equilibrium flash calculations. *Fluid Phase Equilib.*, 26:15–36, 1986.
- [32] Jeffrey P. Kingsley and Angelo Lucia. Equilibrium flash calculations when the number of phases present is unknown. In *11<sup>th</sup> IMACS World Congress on System Simulation and Scientific Computation, Oslo*, pages 249–252, 1985.
- [33] M. Kinoshita and T. Takamatsu. A powerful solution algorithm for single-stage flash problems. *Comput. Chem. Eng.*, 10(4):353–360, 1986.
- [34] Eric Kvaalen and Daniel Tondeur. Constraints on phase equilibrium equations. *Chem. Eng. Sci.*, 43(4):803–810, 1988.
- [35] Claude F. Leibovici and Jean Neoschil. A solution of Rachford – Rice equations for multiphase systems. *Fluid Phase Equilib.*, 112:217–221, 1995.
- [36] Yau-Kun Li and Long X. Nghiem. Phase equilibria of oil, gas and water/brine mixtures from a cubic equation of state and henry’s law. *Can. J. Chem. Eng.*, 64:486–496, Jun. 1986.
- [37] Angelo Lucia, Laxminarasimhan Padmanabhan, and S. Venkataraman. Multiphase equilibrium flash calculations. *Comput. Chem. Eng.*, 24:2557–2569, 2000.
- [38] Angelo Lucia, Peter A. DiMaggio, Meghan L. Bellows, and Leah M. Octavio. The phase behavior of *n*-alkane systems. *Comput. Chem. Eng.*, 29(11–12):2363–2379, oct 2005.
- [39] Angelo Lucia, D. C. Miller, and Ashwani Kumar. Thermodynamically consistent quasi-Newton formulae. *AIChE J.*, 31(8):1381–1388, Aug. 1985.
- [40] Angelo Lucia. Uniqueness of solutions to single-stage isobaric flash processes involving homogeneous mixtures. *AIChE J.*, 32(11):1761–1770, Nov 1986.
- [41] Angelo Lucia. Reply. *AIChE J.*, 34(5):878, May 1988.
- [42] P. Mathias. Comment on “A comparative study of numerical methods for calculating phase equilibria in fluid mixtures from an equation of state” by K. Fotouh and K. Shukla. *Chem. Eng. Sci.*, 52(12):2007, 1997.
- [43] C. Mauri. Unified procedure for solving multiphase – multicomponent vapor – liquid equilibrium calculations. *Ind. Eng. Chem. Process Des. Dev.*, 19:482–489, 1980.
- [44] Conor M. McDonald and Christodoulos A. Floudas. Global optimization and analysis for the Gibbs free energy function using the UNIFAC, Wilson and ASOG equations. *Ind. Eng. Chem. Res.*, 34(5):1674–1687, 1995.
- [45] Ken McKinnon and Marcel Mongeau. Global minimization of the Gibbs free energy. Technical Report 5, Univ. Edinburgh, Dept. Math. Stat., Univ. Edinburgh, Scotland, Mar. 1994.
- [46] R. K. Mehra, R. A. Heidemann, and K. Aziz. An accelerated successive substitution algorithm. *Can. J. Chem. Eng.*, 61:590–596, Aug. 1983.
- [47] Michael L. Michelsen. The isothermal flash problem. Part II. Phase-split calculation. *Fluid Phase Equilib.*, 9:21–40, 1982.
- [48] Michael L. Michelsen. The isothermal flash problem. Part I. Stability. *Fluid Phase Equilib.*, 9:1–19, 1982.
- [49] Michael L. Michelsen. Comments on “Material stability of multicomponent mixtures and the multiplicity of solutions to phase-equilibrium equations. 1. Nonreacting mixtures”. *Ind. Eng. Chem. Fundam.*, 23:373–374, 1984.
- [50] Michael L. Michelsen. Some aspects of multiphase calculations. *Fluid Phase Equilib.*, 30:15–29, 1986.
- [51] M. Michelsen. Simplified flash calculations for cubic equations of state. *Ind. Eng. Chem. Process Des. Dev.*, 25(1):184–188, 1986.
- [52] Michael L. Michelsen. Multiphase isenthalpic and isentropic flash algorithms. *Fluid Phase Equilib.*, 33:13–27, 1987.
- [53] M. L. Michelsen and R. A. Heidemann. To the editor. *AIChE J.*, 34(5):877–878, May 1988.
- [54] M. L. Michelsen. Calculation of multiphase equilibrium. *Comput. Chem. Eng.*, 18(7):545–550, 1994.
- [55] Michael L. Michelsen. State function based flash specifications. *Fluid Phase Equilib.*, 160(158):617–626, 1999.

- [56] Alexander Mitsos and Paul I. Barton. A dual extremum principle in thermodynamics. *AIChE J.*, 53(8):2131–2147, aug 2007.
- [57] P. A. Nelson. Rapid phase determination in multiple-phase flash calculations. *Comput. Chem. Eng.*, 11(6):581–591, 1987.
- [58] H. C. van Ness and M. M. Abbott. Vapor–liquid equilibrium. *AIChE J.*, 25(4):645–653, 1979.
- [59] Ivó Nezbeda, Jiří Kolafa, and William R. Smith. Global phase diagrams of binary mixtures: Systematic basis for describing types of phase equilibrium phenomena. *J. Chem. Soc., Faraday Trans.*, 93(17):3073–3080, 1997.
- [60] Long X. Nghiem and Yay-Kun Li. Computation of multiphase equilibrium phenomena with an equation of state. *Fluid Phase Equilib.*, 17:77–95, 1984.
- [61] Dan Vladimír Nichita, Susana Gomez, and Eduardo Luna. Multiphase equilibria calculation by direct minimization of Gibbs free energy with a global optimization method. *Comput. Chem. Eng.*, 26:1703–1724, 2002.
- [62] D. V. Nichita, E. Luna-Ortiz, and S. Gomez. Multiphase equilibria calculation by direct minimization of Gibbs free energy using the tunnelling global optimization method. *J. Can. Petr. Tech.*, 43(5):13–16, May 2004.
- [63] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part I. Vapour–liquid equilibria. *Comput. Chem. Eng.*, 8(3/4):147–156, 1984.
- [64] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part II. Liquid–liquid and solid–liquid equilibria. *Comput. Chem. Eng.*, 8(3/4):157–162, 1984.
- [65] M. O. Ohanomah and D. W. Thompson. Computation of multicomponent phase equilibria—Part III. Multiphase equilibria. *Comput. Chem. Eng.*, 8(3/4):163–170, 1984.
- [66] V. S. Parekh and P. M. Mathias. Efficient flash calculations for chemical process design—Extension of the Boston–Britt “inside-out” flash algorithm to extreme conditions and new flash types. *Comput. Chem. Eng.*, 22(10):1371–1380, 1998.
- [67] Frances E. Pereira, George Jackson, Amparo Galindo, and Claire S. Adjiman. A duality-based optimisation approach for the reliable solution of  $(p, t)$  phase equilibrium in volume-composition space. *Fluid Phase Equilib.*, 299:1–23, 2010.
- [68] Aaron V. Phoenix and Robert A. Heidemann. A non-ideal multiphase chemical equilibrium algorithm. *Fluid Phase Equilib.*, 150–151:255–265, 1998.
- [69] Rachford, Jr., H. H. and J. D. Rice. Procedure for use of electronic digital computers in calculating flash vaporization hydrocarbon equilibrium. *Petr. Trans. AIME*, 195(Technical Note 136):327–328, 1952.
- [70] J. M. Reneaume, M. Meyer, J. J. Letourneau, and X. Joulia. A global MINLP approach for phase equilibrium calculations. *Comput. Chem. Eng.*, 20:S303–S308, 1996.
- [71] Marinus P. W. Rijkers and Robert A. Heidemann. Convergence behavior of single-stage flash calculations. *ACS Symp. Ser.*, pages 476–493, 1986.
- [72] A. M. Rowe. Discussion. *Soc. Pet. Eng. J.*, 19(4):208, Aug. 1979.
- [73] M. E. Soares, A. G. Medina, C. McDermott, and N. Ashton. Three phase flash calculations using free energy minimisation. *Chem. Eng. Sci.*, 37:521–528, 1982.
- [74] Weixin Song and Maurice A. Larson. Phase equilibrium calculation by using large-scale optimization technique. *Chem. Eng. Sci.*, 46(10):2513–2523, 1991.
- [75] F. Tiscareño, A. Gómez, A. Jiménez, and R. Chávez. Multiplicity of the solutions of the flash equations. *Chem. Eng. Sci.*, 53(4):671–677, 1998.
- [76] R. J. Topliss, D. Dimitrelis, and J. M. Prausnitz. Computational aspects of a non-cubic equation of state for phase-equilibrium calculations. Effect of density-dependent mixing rules. *Comput. Chem. Eng.*, 12(5):483–489, 1988.
- [77] John A. Trangenstein. Customized minimization techniques for phase equilibrium computations in reservoir simulation. *Chem. Eng. Sci.*, 42(12):2847–2863, 1987.
- [78] Alessandro Vetere. Algorithm for fluid phase equilibria calculations by means of a multiconstant equation of state. *Chem. Eng. Sci.*, 39(12):1779–1784, 1984.
- [79] William A. Wakeham and Roumiana P. Stateva. Numerical solution of the isothermal, isobaric phase equilibrium problem. *Rev. Chem. Eng.*, 20(1/2):1–56, 2004.
- [80] Peng Wang and Erling H. Stenby. Non-iterative flash calculation algorithm in compositional reservoir simulation. *Fluid Phase Equilib.*, 94:93–108, 1994.
- [81] Curtis H. Whitson and Michael L. Michelsen. The negative flash. *Fluid Phase Equilib.*, 53:51–71, 1989.
- [82] E. Zhao and S. Saha. Applications of complex domain in vapor–liquid equilibrium calculations using a cubic equation of state. *Ind. Eng. Chem. Res.*, 37:1625–1633, 1998.

### Chemical Equilibrium (PhMatR)

Chemical equilibrium in homogeneous (gas) phases and combined chemical and phase equilibrium calculations. Most of the literature is about the homogeneous (ideal) gas case, but there are also a good number of articles about VSE formulations. Unfortunately, there are few articles on the general formulation of the problem.

**Related keywords:** simultaneous | (& phase & reaction) & equilibrium.

#### Bibliography.

- [1] R. G. Anthony and D. M. Himmelblau. Calculation of complex chemical equilibria by search techniques. *J. Phys. Chem.*, 67(5):1080–1083, 1963.
- [2] Rutherford Aris. Prolegomena to the rational analysis of systems of chemical reactions. *Arch. Ration. Mech. Anal.*, 19(2):81–99, 1965.
- [3] G. Astarita and R. Ocone. Heterogeneous chemical equilibria in multicomponent mixtures. *Chem. Eng. Sci.*, 44(10):2323–2331, 1989.
- [4] G. R. Atwood and W. W. Foster. Transformation of bounded variables in simplex optimization techniques. *Ind. Eng. Chem. Process Des. Dev.*, 12(4):485–486, 1973.
- [5] Domingos Barbosa and Michael F. Doherty. The simple distillation of homogeneous reactive mixtures. *Chem. Eng. Sci.*, 43(3):541–550, 1988.
- [6] A. A. Berezovskii and V. I. Rozenblyum. Thermodynamic equilibrium in complex systems. *Russ. J. Phys. Chem. (Transl. of Zh. Fiz. Khim.)*, 63(2):169–175, 1989.
- [7] P. H. Bjornbom. The relation between the reaction mechanism and the stoichiometric behavior of chemical reactions. *AIChE J.*, 23(3):285–288, 1977.
- [8] G. E. Blau, R. R. Klimpel, and E. C. Steiner. Equilibrium constant determination by nonlinear optimization. *Ind. Eng. Chem. Fundam.*, 9(3):334–339, 1970.
- [9] Milton Blander, Kenneth W. Ragland, Roger L. Cole, Joseph A. Libera, and Arthur Pelton. The inorganic chemistry of wood combustion for power production. *Biomass Bioenergy*, 8(1):29–38, 1995.
- [10] M. Bos and H. Q. J. Meershoek. A computer program for the calculation of equilibrium concentrations in complex systems. *Anal. Chim. Acta*, 61:185–199, 1972.
- [11] F. P. Boynton. Chemical equilibrium in multicomponent polyphase systems. *J. Chem. Phys.*, 32:1880–1881, 1960.
- [12] Brinkley, Jr., Stuart R. Calculation of the equilibrium composition of systems of many constituents. *J. Chem. Phys.*, 15(2):107–110, Feb. 1947.
- [13] Christian de Capitani and Thomas H. Brown. The computation of chemical equilibrium in complex systems containing non-ideal solutions. *Geochim. Cosmochim. Acta*, 51:2639–2652, 1987.
- [14] H. S. Caram and L. E. Scriven. Non-unique reaction equilibria in non-ideal systems. *Chem. Eng. Sci.*, 31:163–168, 1976.
- [15] Marcelo Castier, Peter Rasmussen, and Aage Fredenslund. Calculation of simultaneous chemical and phase equilibria in nonideal systems. *Chem. Eng. Sci.*, 44(2):237–248, 1989.
- [16] Richard J. Clasen. The solution of the chemical equilibrium programming problem with generalized Benders decomposition. *Operations Research*, 32(1):70–79, Jan/Feb 1984.
- [17] D. R. Cruise. Notes on the rapid computation of chemical equilibria. *J. Phys. Chem.*, 68(12):3797–3802, Dec 1964.
- [18] Florin-Emilian Danes and Dan Geană. The effect of changes in chemical potentials on the equilibrium of a gaseous reaction system, 1981.
- [19] J. H. Dluzniewski and S. B. Adler. Calculation of complex reaction and/or phase equilibria problems. *I. Chem. E. Symp. Ser.*, 4(35):21–26, 1972. Pages are numbered as 4:21–4:26.
- [20] J. H. Dluzniewski, S. B. Adler, H. Ozkardes, and H. E. Barner. Aid to correlation of complex equilibria. *Chem. Eng. Prog.*, 69(11):79–80, 1973.
- [21] Gunnar Eriksson. Thermodynamic studies of high temperature equilibria III. SOLGAS, a computer program for calculating the compositions and heat condition of an equilibrium mixture. *Acta Chem. Scand. (1947-1973)*, 25(7):2651–2658, 1971.
- [22] Gunnar Eriksson and Erik Rosén. Thermodynamic studies of high temperature equilibria. *Chem. Scr.*, 4:193–194, 1973.
- [23] Gunnar Eriksson. Thermodynamic studies of high temperature equilibria. XII. SOLGASMIX, a computer program for calculation of equilibrium compositions in multiphase systems. *Chem. Scr.*, 8:100–103, 1975.
- [24] Gunnar Eriksson and Klaus Hack. ChemSage—A computer program for the calculation of complex chemical equilibria. *Metall. Trans. B*, 21B:1013–1023, Dec. 1990.

- [25] T. P. Feenstra. A note on the calculation of concentration in the case of many simultaneous equilibria. *J. Chem. Educ.*, 56(2):104–105, 1979.
- [26] Ilie Fishtik, Ivan Gutman, and István Nagypál. Response reactions in chemical thermodynamics. *J. Chem. Soc., Faraday Trans.*, 92(19):3525–3532, 1996.
- [27] Katrine Seip Førland. *Kjemisk likevekt*. Tapir Forlag, Trondheim, Norge, 1978.
- [28] John C. Friedly. Extent of reaction in open systems with multiple heterogeneous reactions. *AIChE J.*, 37(5):687–693, May 1991.
- [29] John C. Friedly and Jacob Rubin. Solute transport with multiple equilibrium-controlled or kinetically controlled chemical reactions. *Water Resour. Res.*, 28(6):1935–1953, Jul 1992.
- [30] Rajeev Gautam and Warren D. Seider. Computation of phase and chemical equilibrium. Part I. Local and constrained minima in Gibbs free energy. *AIChE J.*, 25(6):991–999, Nov. 1979.
- [31] Rajeev Gautam and Warren D. Seider. Computation of phase and chemical equilibrium. Part II. Phase splitting. *AIChE J.*, 25(6):999–1006, Nov. 1979.
- [32] Rajeev Gautam and Warren D. Seider. Computation of phase and chemical equilibrium. Part III. Electrolytic solutions. *AIChE J.*, 25(6):1006–1015, Nov. 1979.
- [33] Rajeev Gautam and J. S. Wareck. Computation of physical and chemical equilibria—Alternate specifications. *Comput. Chem. Eng.*, 10(2):143–151, 1986.
- [34] Dan Geană. Calculul echilibrului chimic multicomponente în sisteme eterogene cu faze condensate pure. *Revista de Chimie*, 29(12):1171–1176, 1978.
- [35] D. Geană and F. Danes. The computation of multicomponent chemical equilibria in gases. *Seria Chimie-Metalurgie (Bucuresti)*, XXXX(3):31–40, 1978.
- [36] D. Geană. A min–max formulation of the problem of phase equilibrium in chemical reacting systems. *Rev. Roum. Chim.*, 29:395–405, 1984.
- [37] D. Geană. Transformări prin programare geometrică ale problemei echilibrului chimic și/sau de faze. I. Teorie. *Revista de Chimie*, 38(12):1080–1088, 1987.
- [38] H. Greiner. An efficient implementation of Newton’s method for complex nonideal chemical equilibria. *Comput. Chem. Eng.*, 15(2):115–123, 1991.
- [39] Jerry P. Greenberg and John H. Weare. Simultaneous multi-phase precipitation in the primal chemical equilibrium problem. *AIChE Symp. Ser.*, 90(298):51–63, 1994.
- [40] Charles E. Harvie, Jerry P. Greenberg, and John Weare. A chemical equilibrium algorithm for highly non-ideal multiphase systems: Free energy minimization. *Geochim. Cosmochim. Acta*, 51:1045–1057, 1987.
- [41] Tore Haug-Warberg. *Computation of Thermodynamic Equilibria*. Dr.ing. thesis 1988:42, Norwegian Institute of Technology, 1988.
- [42] O. Heuzé, Henri-Noël Presles, and P. Bauer. Computation of chemical equilibria. *J. Chem. Phys.*, 83(9):4734–4737, 1985.
- [43] D. Hildebrandt and D. Glasser. Predicting phase and chemical equilibrium using the convex hull of the Gibbs free energy. *The Chemical Engineering Journal and the Biochemical Engineering Journal*, 54(3):187–197, jul 1994.
- [44] H. P. Hutchison. Basis for calculating equilibrium gas composition on a digital computer. *Chem. Eng. Sci.*, 17:703, 1962.
- [45] F. Jalali and J. D. Seader. Homotopy continuation method in multi-phase multi-reaction equilibrium systems. *Comput. Chem. Eng.*, 23:1319–1331, 1999.
- [46] Bo Jansson. A general method for calculating phase equilibria under different types of conditions. Report series D, No 55, Division of Physical Metallurgy, Royal Institute of Technology, Stockholm, 1984.
- [47] Y. Jiang, G. R. Chapman, and W. R. Smith. On the geometry of chemical reaction and phase equilibria. *Fluid Phase Equilib.*, 118:77–102, 1996.
- [48] D. J. A. Jubb and R. W. Missen. The computer-derivation of thermodynamic equations part III. Chemical systems. *Can. J. Chem. Eng.*, 67:658–664, 1989.
- [49] H. J. Kandiner and S. R. Brinkley, Jr. Calculation of complex equilibrium relations. *Ind. Eng. Chem.*, 42(5):850–855, 1950.
- [50] Igor K. Karpov, Konstantin V. Chudnenko, and Dmitri A. Kulik. Modeling chemical mass transfer in geochemical processes: Thermodynamic relations, conditions of equilibria, and numerical algorithms. *Am. J. Sci.*, 297(8):767–806, Oct. 1997.
- [51] Laar, J. J., von. Zum begriff der unabhängigen bestandteile. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 47:228–230, 1904.
- [52] G. Lantagne, B. Marcos, and B. Cayrol. Computation of complex equilibria by nonlinear optimization. *Comput. Chem. Eng.*, 12(6):589–599, 1988.
- [53] Yeow Peng Lee, Gade Pandu Rangaiah, and Rein Luus. Phase and chemical equilibrium calculations by direct search optimization. *Comput. Chem. Eng.*, 23:1183–1191, 1999.

- [54] Howard B. Levine. Chemical equilibrium in complex mixtures. *J. Chem. Phys.*, 36(11):3049–3050, Jun 1962.
- [55] Michael Luckas, Klaus Lucas, and Hans Roth. Computation of phase and chemical equilibria in flue-gas/water systems. *AIChE J.*, 40(11):1892–1900, 1994.
- [56] Y. H. Ma and C. W. Shipman. On the computation of complex equilibria. *AIChE J.*, 18(2):299–304, 1972.
- [57] W. David Madeley and James M. Toguri. Computing chemical equilibrium compositions in multi-phase systems. *Ind. Eng. Chem. Fundam.*, 12(2):261–262, 1973.
- [58] A. E. Mather. Phase equilibria and chemical reaction. *Fluid Phase Equilib.*, 30:83–100, 1986.
- [59] Gerd Maurer. Phase equilibria in chemical reactive fluid mixtures. *Fluid Phase Equilib.*, 116:39–51, 1996.
- [60] Conor M. McDonald and Christodoulos A. Floudas. Global optimization for the phase and chemical equilibrium problem: Application to the NRTL equation. *Comput. Chem. Eng.*, 19(11):1111–1139, 1995.
- [61] Ken McKinnon and Marcel Mongeau. A generic global optimization algorithm for the chemical and phase equilibrium problem. Technical Report 24, Univ. Edinburgh, Dept. Math. Stat., Univ. Edinburgh, Scotland, Sep. 1994.
- [62] Ken McKinnon and Marcel Mongeau. A generic global optimization algorithm for the chemical and phase equilibrium problem. *J. Global Opt.*, 12:325–351, 1998.
- [63] K. Meintjes and A. P. Morgan. Performance of algorithms for calculating the equilibrium composition of a mixture of gases. *J. Comp. Phys.*, 60:219–234, 1985.
- [64] Michael L. Michelsen. Calculation of multiphase ideal solution chemical equilibrium. *Fluid Phase Equilib.*, 53:73–80, 1989.
- [65] Raul Mihail, Corneliu V. Radu, and Ion M. Belcea. Methods for complex reaction systems equilibrium computation. *Rev. Roum. Chim.*, 23(5):681–698, 1978.
- [66] Andrea K. Myers and Alan L. Myers. Numerical solution of chemical equilibria with simultaneous reactions. *J. Chem. Phys.*, 84(10):5787–5795, May 1986.
- [67] István Nagypál, Eufrozina Hoffmann, Ivan Gutman, and Ilie Fishtik. Response reactions: An interpretive concept in chemical thermodynamics. *Pure Appl. Chem.*, 70(3):583–590, 1998. Translated from *Fizika Goreniya i Vzryva*, Vol. 32, No. 1, pp. 66–70, January–February 1995.
- [68] G. W. Norval, M. J. Phillips, R. W. Missen, and W. R. Smith. Constrained chemical equilibrium and incompletely specified elemental abundance data. *Can. J. Chem. Eng.*, 69:1184–1192, 1991.
- [69] R. C. Oliver, S. E. Stephanou, and R. W. Baier. Calculating free-energy minimization. *Chem. Eng. (N. Y.)*, 68:121–128, Feb 1962.
- [70] H. G. Othmer. Nonuniqueness of equilibria in closed reacting systems. *Chem. Eng. Sci.*, 31:993–1003, 1976.
- [71] D. J. M. Park. Numerical methods for solving the chemical mass action equilibrium problem. *J. Chem. Phys.*, 65(8):3085–3091, 1976.
- [72] Robert L. Potter and Wouter Vanderkulk. Computation of the equilibrium composition of burnt gases. *J. Chem. Phys.*, 32(5):1304–1307, May 1960.
- [73] B. N. Raju and C. S. Krishnaswami. Free energy minimization method for calculating thermodynamic equilibrium composition of a chemical system. *Indian J. Tech.*, 4:99–100, 1966.
- [74] R. V. Sanderson and H. H. Y. Chien. Simultaneous chemical and phase equilibrium calculation. *Ind. Eng. Chem. Process Des. Dev.*, 12(1):81–85, 1973.
- [75] G. L. Schott. Computation of restricted equilibria by general methods. *J. Chem. Phys.*, 40:2065–2066, 1964.
- [76] D. R. Schneider and G. V. Reklaitis. On material balances for chemically reacting systems. *Chem. Eng. Sci.*, 30:243–247, 1975.
- [77] D. B. Scully. Calculation of the equilibrium compositions for multi-constituent systems. *Chem. Eng. Sci.*, 17:977–985, 1962.
- [78] Warren D. Seider, Rajeev Gautam, and Charles W. White, III. Computation of phase and chemical equilibrium: A review. *ACS Symp. Ser.*, 124:115–134, 1980.
- [79] Warren D. Seider and Soemantri Widagdo. Multiphase equilibria of reactive systems. *Fluid Phase Equilib.*, 123:283–303, 1996.
- [80] N. Z. Shapiro and L. S. Shapley. Mass action laws and the Gibbs free energy function. *J. Soc. Ind. Appl. Math.*, 13(2):353–375, Jun 1965.
- [81] K. T. Shih and W. E. Ibele. Generalized free energy method for equilibrium compositions in complex mixtures. *Int. J. Heat Mass Transfer*, 11:607–609, 1968.
- [82] W. R. Smith and R. W. Missen. Calculating complex chemical equilibria by an improved reaction-adjustment method. *Can. J. Chem. Eng.*, 46:269–272, Aug. 1968.

- [83] William R. Smith. Some remarks on the calculation of complex chemical equilibria by general methods. *Ind. Eng. Chem. Fundam.*, 15(3):227–229, 1976.
- [84] William R. Smith. The computation of chemical equilibria in complex systems. *Ind. Eng. Chem. Fundam.*, 19(1):1–10, 1980.
- [85] William R. Smith and Ronald W. Missen. *Chemical Reaction Equilibrium Analysis: Theory and Algorithms*. John Wiley & Sons, New York, 1982.
- [86] W. R. Smith and R. W. Missen. Strategies for solving the chemical equilibrium problem and an efficient microcomputer-based algorithm. *Can. J. Chem. Eng.*, 66:591–598, 1988.
- [87] William R. Smith. EQS, a general purpose PC-based chemical equilibrium program. *AIChE Symp. Ser.*, 90(298):88–93, 1994.
- [88] Roumiana P. Stateva and William A. Wakeham. Phase equilibrium calculations for chemical reacting systems. *Ind. Eng. Chem. Res.*, 36:5474–5482, 1997.
- [89] E. E. Stone. Complex chemical equilibria. Application of Newton–Raphson method to solve non-linear equations. *J. Chem. Educ.*, 43(5):241–244, May 1966.
- [90] S. H. Storey and F. Van Zeggeren. Solving complex chemical equilibria by a method of nested iterations. *Can. J. Chem. Eng.*, 45:323–326, 1967.
- [91] Robert B. Symonds and Mark H. Reed. Calculation of multicomponent chemical equilibria in gas–solid–liquid systems: Calculation methods, thermochemical data, and applications to studies of high-temperature volcanic gases with examples from Mount St. Helens. *Am. J. Sci.*, 293(8):758–864, 1993.
- [92] Christos Tsanas, Erling H. Stenby, and Wei Yan. Calculation of multiphase chemical equilibrium in electrolyte solutions with non-stoichiometric methods. *Fluid Phase Equilib.*, 482:81–98, 2019.
- [93] A. G. Turnbull. A general computer program for the calculation of chemical equilibria and heat balances. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 7(2):137–147, May 1983.
- [94] Sophie Ung and Michael F. Doherty. Vapor–liquid phase equilibrium in systems with multiple chemical reactions. *Chem. Eng. Sci.*, 50(1):23–48, 1995.
- [95] Sophie Ung and Michael F. Doherty. Theory of phase equilibria in multireaction systems. *Chem. Eng. Sci.*, 50(20):3201–3216, 1995.
- [96] D. S. Villars. A method of successive approximations for computing combustion equilibria on a high speed digital computer. *J. Phys. Chem.*, 63(4):521–525, 1959.
- [97] Petr Voňka and Jindřich Leitner. Calculation of chemical equilibria in heterogeneous multicomponent systems. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 19(1):25–36, 1995.
- [98] T. Westerlund, B. Skrifvars, and S. Karrila. On the uniqueness in pH calculations. *Chem. Eng. Sci.*, 40(6):973–976, 1985.
- [99] Tapio Westerlund. Stoichiometry and equilibrium calculations in chemical engineering. Dept. Chem. Eng., Åbo Akademi, 1985.
- [100] W. B. White, S. M. Johnson, and G. B. Dantzig. Chemical equilibrium in complex mixtures. *J. Chem. Phys.*, 28(5):751–755, May 1958.
- [101] W. B. White. Numerical determination of chemical equilibrium and the partitioning of free energy. *J. Chem. Phys.*, 46(11):4171–4175, Jun. 1967.
- [102] White, III, Charles W. and Warren D. Seider. Computation of phase and chemical equilibrium. Part IV: Approach to chemical equilibrium. *AIChE J.*, 27(3):466–471, May 1981.
- [103] W. B. White. Further simplification in the numerical determination of chemical equilibrium. *J. Chem. Phys.*, 84(11):6391–6393, Jun 1986.
- [104] Wen-De Xiao, Kai-Hong Zhu, Wei-Kang Yuan, and Henry Hung-Yeh Chien. An algorithm for simultaneous chemical and phase equilibrium calculation. *AIChE J.*, 35(11):1813–1820, Nov. 1989.
- [105] Fushan Yin. A simpler method for finding independent reactions. *Chem. Eng. Commun.*, 83:117–127, 1989.
- [106] F. van Zeggeren and S. H. Storey. *The Computation of Chemical Equilibria*. Cambridge University Press, Cambridge, London, 1970.
- [107] Frank J. Zeleznik and Sanford Gordon. Calculation of complex chemical equilibria. *Ind. Eng. Chem.*, 60(6):27–57, Jun. 1968. Applied Thermodynamics Symposium.
- [108] M. Zhao. Method of minimization of free energy in the calculation of complex chemical equilibria. *Int. Chem. Eng.*, 23(3):541–549, Jul 1983.

### Global (intrinsic) Stability (PhMatS)

Methods for the global minimization of equilibrium problems, typically using the Gibbs’ tangent plane criterion.

**Related keywords: stability | global & minimum | uniqueness | tangent & plane | phase & split.**

### Bibliography.

- [1] M. N. Ammar and H. Renon. The isothermal flash problem: New methods for phase split calculations. *AIChE J.*, 33(6):926–939, jun 1987.
- [2] Lee E. Baker, Alan C. Pierce, and Kraemer D. Luks. Gibbs energy analysis of phase equilibria. *Soc. Pet. Eng. J. (1980-)*, 22(5):731–742, Oct. 1982.
- [3] Jürgen Bausa and Wolfgang Marquardt. Quick and reliable phase stability test in VLLLE flash calculations by homotopy continuation. *Comput. Chem. Eng.*, 24:2447–2456, 2000.
- [4] Bruce L. Beegle, Michael Modell, and Robert C. Reid. Thermodynamic stability criterion for pure substances and mixtures. *AIChE J.*, 20(6):1200–1206, Nov. 1974.
- [5] John W. Cahn. Spinodal decomposition. *Trans. Metall. Soc. AIME*, 242:166–180, Feb. 1968.
- [6] Bernard D. Coleman. On the stability of equilibrium states of general fluids. *Arch. Ration. Mech. Anal.*, 36:1–32, 1970.
- [7] D. B. van Dongen and M. F. Doherty. Calculation and stability of multiple equilibrium points for nonideal mixtures. *Fluid Phase Equilib.*, 14:335–343, 1983.
- [8] Mark S. Ghiorso. Algorithms for the estimation of phase stability in heterogeneous thermodynamic systems. *Geochim. Cosmochim. Acta*, 58(24):5489–5501, 1994.
- [9] Anup K. Gupta, P. Raj Bishnoi, and Nicolas Kalogerakis. A method for the simultaneous phase equilibria and stability calculations for multiphase reacting and non-reacting systems. *Fluid Phase Equilib.*, 63:65–89, 1991.
- [10] Robert A. Heidemann. The criteria for thermodynamic stability. *AIChE J.*, 21(4):824–826, Jul. 1975.
- [11] R. A. Heidemann. Comments on the uniqueness in equilibrium calculations. *Chem. Eng. Sci.*, 42(11):2797–2798, 1987.
- [12] James Z. Hua, Joan F. Brennecke, and Mark A. Stadtherr. Reliable computation of phase stability using interval analysis: Cubic equation of state models. *Comput. Chem. Eng.*, 22(9):1207–1214, 1998.
- [13] Fouad M. Houry. Calculate the right density. *Hydrocarbon Process.*, pages 155–157, 1978.
- [14] S. B. Kiselev and I. G. Kostyukova. Spinodal and kinetic boundary of metastable region. *J. Chem. Phys.*, 98(8):6455–6464, Apr. 1993.
- [15] Conor M. McDonald and Christodoulos A. Floudas. Global optimization for the phase equilibrium problem using the NRTL equation. *I. Chem. E. Symp. Ser.*, 133:274–280, 1994.
- [16] Conor M. McDonald and Christodoulos A. Floudas. Global optimization for the phase stability problem. *AIChE J.*, 41(7):1798–1814, 1995.
- [17] Dan Vladimir Nichita, Cecilia de los Angeles Duran Valencia, and Susana Gomez. Volume-based thermodynamics global phase stability analysis. *Chem. Eng. Commun.*, 193:1194–1216, 2006.
- [18] G. V. Pasad and G. Venkatarathnam. A method for avoiding trivial roots in isothermal flash calculations using cubic equation of state. *Ind. Eng. Chem. Res.*, 38:3530–3534, 1999.
- [19] Ulf J. Plöcker and Helmut Knapp. Save time in computing density. *Hydrocarbon Process.*, pages 199–201, 1976.
- [20] Irene F. Radzysinski and Wallace B. Whiting. Fluid phase stability and equations of state. *Fluid Phase Equilib.*, 34:101–110, 1987.
- [21] David B. Shear. Stability and uniqueness of the equilibrium point in chemical reaction systems. *J. Chem. Phys.*, 48(9):4144–4147, May 1968.
- [22] Roumiana P. Stateva and Stefan G. Tsvetkov. A new method for thermodynamic stability analysis of multicomponent systems. *Hung. J. Ind. Chem.*, 19:179–188, 1991.
- [23] Amy C. Sun and Warren D. Seider. Homotopy-continuation method for stability analysis in the global minimization of the Gibbs free energy. *Fluid Phase Equilib.*, 103:213–249, 1995.
- [24] Y. S. Teh and G. P. Rangaiah. Tabu search for global optimization of continuous functions with application to phase equilibrium calculations. *Comput. Chem. Eng.*, 27:1665–1679, 2003.
- [25] Stephen R. Tessier, Joan F. Brennecke, and Mark A. Stadtherr. Reliable phase stability analysis for excess Gibbs energy modules. *Chem. Eng. Sci.*, 55:1785–1796, 2000.
- [26] Stanislaw K. Wasylkiewicz, Lakshmi N. Sridhar, Michael F. Doherty, and Michael F. Malone. Global stability analysis and calculations of liquid–liquid equilibrium in multicomponent mixtures. *Ind. Eng. Chem. Res.*, 35:1395–1408, 1996.
- [27] Tapio Westerlund and Henrik Saxén. On the uniqueness in equilibrium calculations. *Chem. Eng. Sci.*, 42(1):188–190, 1987.
- [28] Tapio Westerlund and Henrik Saxén. Author’s reply to comments by R. A. Heidemann. *Chem. Eng. Sci.*, 42(11):2798–2799, 1987.
- [29] Zhu Yushan and Xu Zhihong. A reliable method for liquid–liquid phase equilibrium calculation and global stability analysis. *Chem. Eng. Commun.*, 176:133–160, 1999.

## Transport Phenomena (PhMatT)

Methods and data for (mostly) diffusion theory.

**Related keywords: transport & phenomena | diffusion | maxwell & stefan.**

### Bibliography.

- [1] [Annis, B. K.](#), A. E. Humphreys, and E. A. Mason. Nonisothermal, nonstationary diffusion. *The Physics of Fluids*, 12(1):78–83, 1969.
- [2] [Richard J. Bearman](#). The thermo-osmosis of the rare gases through a rubber membrane. *J. Phys. Chem.*, 61:708–713, Jun 1957.
- [3] [Dieter Bothe](#). On the maxwell–stefan approach to multicomponent diffusion. *arXiv:1007.1775v1*, pages 1–13, Jul 2010.
- [4] [Stig Claesson](#) and Lars-Olof Sundelof. Diffusion libre au voisinage de la température critique de miscibilité. *J. Chim. Phys. Phys.-Chim. Biol.*, 54:914–919, 1957.
- [5] [Charles F. Curtiss](#) and Joseph O. Hirschfelder. Transport properties of multicomponent gas mixtures. *J. Chem. Phys.*, 17(6):550–555, Jun. 1949.
- [6] [E. L. Cussler](#). Cluster diffusion in liquids. *AIChE J.*, 26(1):43–51, Jan. 1980.
- [7] [J. L. Duda](#) and J. S. Vrentas. Analysis of free diffusion experiments in binary systems. *Ind. Eng. Chem.*, 4(3):301–308, Aug. 1965.
- [8] [J. L. Duda](#) and J. S. Vrentas. Numerical technique for solution of the diffusion equation in an infinite medium. *Ind. Eng. Chem.*, 5(1):69–74, Feb. 1966.
- [9] [J. B. Duncan](#) and H. L. Toor. An experimental study of three component gas diffusion. *AIChE J.*, 8(1):38–41, Mar. 1962.
- [10] [James F. Ely](#) and H. J. M. Hanley. Prediction of transport properties. 1. Viscosity of fluids and mixtures. *Ind. Eng. Chem. Fundam.*, 20(4):323–332, 1981.
- [11] [James F. Ely](#) and H. J. M. Hanley. Prediction of transport properties. 2. Thermal conductivity of pure fluids and mixtures. *Ind. Eng. Chem. Fundam.*, 22(1):90–97, 1983.
- [12] [Andrew R. Felmy](#) and John H. Weare. Calculation of multicomponent ionic diffusion from zero to high concentration: I. The system Na–K–Ca–Mg–SO<sub>4</sub>–H<sub>2</sub>O at 25°C. *Geochim. Cosmochim. Acta*, 55(1):113–131, 1991.
- [13] [R. Haase](#) and M. Siry. Diffusion im kritischen Entmischungsgebiet binärer flüssiger Systeme. *Z. Phys. Chem., Neue Folge*, 57:56–73, 1968.
- [14] [Hans Janssen](#). Thermal diffusion of water vapour in porous materials: Fact or fiction? *Int. J. Heat Mass Transfer*, 54(7):1548–1562, 2011.
- [15] [T. R. Marrero](#) and E. A. Mason. Gaseous diffusion coefficients. *J. Phys. Chem. Ref. Data*, 1(1):3–118, 1972.
- [16] [Monchick, L.](#), R. J. Munn, and E. A. Mason. Thermal diffusion in polyatomic gases: A generalized stefan—maxwell diffusion equation. *J. Chem. Phys.*, 45(8):3051–3058, 1966.
- [17] [A. S. Myerson](#) and Dennis Senol. Diffusion coefficients near the spinodal curve. *AIChE J.*, 30(6):1004–1006, Nov. 1984.
- [18] [J. A. Quinn](#), C. H. Lin, and J. L. Anderson. Measuring diffusion coefficients by Taylor’s method of hydrodynamic stability. *AIChE J.*, 32(12):2028–2033, Dec. 1986.
- [19] [Sebastian Rehfeldt](#) and Johann Stichlmair. Measurement and calculation of multicomponent diffusion coefficients in liquids. *Fluid Phase Equilib.*, 256(1):99–104, 2007. 16th Symposium on Thermophysical Properties.
- [20] [L. S. Sorell](#) and A. S. Myerson. Diffusivity of urea in concentrated, saturated and supersaturated solutions. *AIChE J.*, 28(5):772–779, Sep. 1982.
- [21] [Tung Tsang](#) and Carol A. Hammarstrom. Nonlinear diffusion in the solid phase. *Ind. Eng. Chem. Res.*, 26(4):855–857, 1987.
- [22] [Warren E. Stewart](#) and Richard Prober. Matrix calculation of multicomponent mass transfer in isothermal systems. *Ind. Eng. Chem.*, 3(3):224–234, Aug. 1964.
- [23] [Tutomu Takashi](#), Takahiro Ohrai, Hideo Yamamoto, and Jun Okamoto. Upper limit and heaving pressure derived by pore water pressure measurements of partially frozen soil. In *ISGF’80: The 2nd International Symposium on Ground Freezing*, pages 713–724, The Norwegian Institute of Technology, Jun 1980.
- [24] [Sterling A. Taylor](#) and Luigi Cavazza. The movement of soil moisture in response to temperature gradients. *Soil Sci. Soc. Am. Proc.*, 18(4):351–358, Oct 1954. Division I—soil physics.
- [25] [Ross Taylor](#) and R. Krishna. *Multicomponent Mass Transfer*. Wiley Series in Chemical Engineering. John Wiley & Sons, New York, 1993.
- [26] [H. L. Toor](#). Diffusion in three-component gas mixtures. *AIChE J.*, 3(2):198–207, Jun. 1957.

- [27] [H. L. Toor](#). Solution of the linearized equations of multicomponent mass transfer: I. *AIChE J.*, 10(4):448–455, Jul. 1964.
- [28] [H. L. Toor](#). Solution of the linearized equations of multicomponent mass transfer: II. Matrix methods. *AIChE J.*, 10(4):460–465, Jul. 1964.
- [29] [H. J. V. Tyrrell](#). *Diffusion and Heat Flow in Liquids*. Butterworths Publications Ltd., London, 1961.
- [30] [Wendt, R. P.](#), J. N. Mundy, Stanley Weissman, and E. A. Mason. Gaseous self-diffusion in a temperature gradient. *The Physics of Fluids*, 6(4):572–578, 1963.
- [31] [Walter G. Whitman](#). The two-film theory of gas absorption. It seems to explain satisfactorily the well-recognized differences of absorption rate for varying concentrations. *Int. J. Heat Mass Transfer*, 5:429–433, 1962.

## CHAPTER 3

# Thermodynamics

### Computer Algebra Systems (TheCAS)

The borderland between thermodynamics, computer programming and thermodynamics databases.

**Related keywords: thermodynamics & (computer & algebra | software | object & oriented | database).**

#### Bibliography.

- [1] [Marcelo Castier](#). Automatic implementation of thermodynamic models using computer algebra. *Comput. Chem. Eng.*, 23:1229–1245, 1999.
- [2] [Angeles Dominguez](#), José Tojo, and Marcelo Castier. Automatic implementation of thermodynamic models for reliable parameter estimation using computer algebra. *Comput. Chem. Eng.*, 26:1473–1479, 2002.
- [3] [Michael Frenkel](#), Robert D. Chirico, Vladimir V. Diky, Qian Dong, Svetlana Frenkel, Paul R. Franchois, Dale L. Embry, Thomas L. Teague, Kenneth N. Marsh, and Randolph C. Wilhoit. ThermoML—An XML-based approach for storage and exchange of experimental and critically evaluated thermo-physical and thermochemical property data. 1. Experimental data. *J. Chem. Eng. Data*, 48:2–13, 2003.
- [4] [L. Jourda](#), X. Joulia, and B. Kœhret. Introducing ATOM, the applied thermodynamics object-oriented model. *Comput. Chem. Eng.*, 20:S157–S164, 1996.
- [5] [Ursula R. Kattner](#), Gunnar Eriksson, Iris Hahn, Rainer Schmid-Fetzer, Bosse Sundman, Varghese Swamy, and Armin Kussmaul. Applications of computational thermodynamics. Group 4: Use of thermodynamic software in process modelling. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):55–94, 2000.
- [6] [A. Kilakos](#) and B. Kalitventzeff. A new implementation for analytical derivatives of thermodynamic properties and its beneficial application on dynamic simulation. *Comput. Chem. Eng.*, 17(5/6):441–450, May 1993.
- [7] [E. B. Rudnyi](#) and G. F. Voronin. Classes and objects of chemical thermodynamics in object-oriented programming. 1. A class of analytical functions of temperature and pressure. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 19(2):189–206, 1995.
- [8] [Stanley I. Sandler](#). Spreadsheets for thermodynamic instruction. *Chem. Eng. Educ.*, Winter:18–20, 1997.
- [9] [A. S. Silva](#) and M. Castier. Automatic differentiation and implementation of thermodynamic models using a computer algebra system. *Comput. Chem. Eng.*, 17:473–478, 1993.
- [10] [Philip J. Spencer](#), Tim J. Anderson, Tim G. Chart, André Costa e Silva, Bo Jansson, Beyong-Joo Lee, and Mikael Schalin. Applications of computational thermodynamics. Group 5: New applications of thermodynamic calculations. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):55–94, 2000.
- [11] [Bo Sundman](#), Bo Jansson, and Jan-Olof Andersson. The Thermo-Calc databank system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 9(2):153–190, 1985.
- [12] [Bo Sundman](#). Thermodynamic databanks, visions and facts. *Scand. J. Metall.*, 20:79–85, 1991.
- [13] [Régis Thiéry](#). A new object-oriented library for calculating analytically high-order multivariable derivatives and thermodynamic properties of fluids with equations of state. *Comput. Geosci.*, 22(7):810–815, 1996.

### Differential Scanning Calorimetry (TheDSC)

Calorimetry in general and Differential Scanning Calorimetry in particular.

**Related keywords: differential & scanning & calorimetry | dsc.**

**Bibliography.**

- [1] S. J. Ashcroft. The measurement of enthalpies of sublimation by thermogravimetry. *Thermochim. Acta*, 2:512–514, 1971.
- [2] Harvey E. Bair. *Glass Transition Measurements by DSC*, pages 50–74. American Society for Testing and Materials, Philadelphia, 1994. Ed. Rickey J. Seyler.
- [3] Barrall, II, Edward M. Precise determination of melting and boiling points by differential thermal analysis and differential scanning calorimetry. *Thermochim. Acta*, 5:377–389, 1973.
- [4] William P. Brennan, Bernard Miller, and John C. Whitwell. An improved method of analyzing curves in differential scanning calorimetry. *I&EC Fundam.*, 8(2):314–318, 1969.
- [5] W. P. Brennan, B. Miller, and J. C. Whitwell. Rate of change of reference temperature in DSC and DTA. *Thermochim. Acta*, 2:354–356, 1971.
- [6] Robert W. Carling. Heat capacities of  $\text{NaNO}_3$  and  $\text{KNO}_3$  from 350 to 800 K. *Thermochim. Acta*, 60:265–275, 1983.
- [7] Bruce Cassel and Michael P. DiVito. Use of DSC to obtain accurate thermodynamic and kinetic data. *Am. Lab. (Shelton, Conn.)*, 26(1):14+, Jan. 1994.
- [8] R. Courchinoux, N. B. Chanh, and Y. Haget. Use of the “shape factors” as an empirical method to determine the actual characteristic temperatures of binary phase diagrams by differential scanning calorimetry. *Thermochim. Acta*, 128:45–53, 1988.
- [9] Michael P. DiVito, R. B. Cassel, M. Margulies, and S. Goodkowsky. Dynamic differential scanning calorimetry. *Am. Lab. (Shelton, Conn.)*, 27(12):28+, Aug. 1995.
- [10] A. A. van Dooren and B. W. Müller. Influence of experimental variables of curves in differential scanning calorimetry. Part I. Study design and results of calibration checks. *Thermochim. Acta*, 49:151–161, 1981.
- [11] Robert E. Farritor and Luh C. Tao. An improved method of measurement of vaporization heat of volatile liquids with a differential scanning calorimeter. *Thermochim. Acta*, 1:297–304, 1970.
- [12] Joseph H. Flynn. Instrumental limitations upon the measurement of temperature and rate of energy production by differential scanning calorimetry. *J. Therm. Anal.*, 1:127–138, 1971.
- [13] F. E. Freeberg and T. G. Alleman. A sealed cell for use with a commercial differential scanning calorimeter. *Anal. Chem.*, 38(12):1806–1807, 1966.
- [14] Allan P. Gray. A simple generalized theory for the analysis of dynamic thermal measurement. In Roger S. Porter and Julian F. Johnson, editors, *Analytical Calorimetry*. Springer-Verlag, New York, 1968.
- [15] O. Greis, K. M. Bahamdan, and B. M. Uwais. The phase diagram of the system  $\text{NaNO}_3$ – $\text{KNO}_3$  studied by differential scanning calorimetry. *Thermochim. Acta*, 86:343–350, 1985.
- [16] W. Hemminger and G. Höhne. *Calorimetry—Fundamentals and Practice*. Verlag Chemie GmbH, Weinheim, Germany, 1984.
- [17] H. M. Heuvel and K. C. J. B. Lind. Computerized analysis and correction of differential scanning calorimetric data for effects due to thermal lag and heat capacity changes. *Anal. Chem.*, 42(9):1044–1048, 1970.
- [18] G. W. H. Höhne and J. E. K. Schawe. Dynamic behaviour of power compensated differential scanning calorimeters. Part 1. DSC as a linear system. *Thermochim. Acta*, 229:27–36, 1993.
- [19] K. Kostyrko and M. Skoczylas. Temperature standard reference materials for thermal analysis. *J. Therm. Anal.*, 38:2181–2188, 1992.
- [20] J. L. McNaughton and C. T. Mortimer. *Differential Scanning Calorimetry*. The Perkin-Elmer Corporation, Norwalk, Connecticut, 1975. Reprinted from: IRS, Physical Chemistry Series 2, 1975, Volume 10, Butterworths, London.
- [21] Moldover, M. R., W. L. Tew, and H. W. Yoon. Advances in thermometry. *Nature Physics*, 12:7–11, Jan 2016.
- [22] M. J. O’Neill. The analysis of a temperature-controlled scanning calorimeter. *Anal. Chem.*, 36(7):1238–1245, Jun. 1964.
- [23] M. J. O’Neill. Measurement of specific heat functions by differential scanning calorimetry. *Anal. Chem.*, 38(10):1331–1336, Sep. 1966.
- [24] E. Pella and M. Nebuloni. Temperature measurements with a differential calorimeter. *J. Therm. Anal.*, 3:229–246, 1971.
- [25] D. M. Price. Temperature calibration of differential scanning calorimeters. *J. Therm. Anal.*, 45:1285–1296, 1995.
- [26] J. E. K. Schawe, C. Schick, and G. W. H. Höhne. Dynamic behaviour of power compensated differential scanning calorimeters. Part 2. The signal flow. *Thermochim. Acta*, 229:37–52, 1993.
- [27] J. E. K. Schawe, G. W. H. Höhne, and C. Schick. Dynamic behaviour of power compensated differential scanning calorimeters. Part 3. The influence of material properties (an error evaluation). *Thermochim. Acta*, 244:33–48, 1994.

- [28] J. E. K. Schawe. A comparison of different evaluation methods in modulated temperature DSC. *Thermochim. Acta*, 260:1–16, 1995.
- [29] J. E. K. Schawe. Modulated temperature DSC measurements—The influence of the experimental conditions. *Thermochim. Acta*, 271:127–140, Jan. 1996.
- [30] C. Howard Shomate. A method for evaluating and correlating thermodynamic data. *J. Phys. Chem.*, 58(4):368–372, 1954.
- [31] I. Konkoly Thege. DSC studies of binary inorganic ammonium compound systems. *J. Therm. Anal.*, 27:275–286, 1983.
- [32] E. S. Watson, M. J. O’Neill, Joshua Justin, and Nathaniel Brenner. A differential scanning calorimeter for quantitative differential thermal analysis. *Anal. Chem.*, 36(7):1233–1238, Jun. 1964.
- [33] W. W. Wendland. A capillary tube holder for differential scanning calorimetry. *Anal. Chim. Acta*, 49:187–188, 1970.
- [34] H. G. Wiedemann. Applications of thermogravimetry for vapor pressure determination. *Thermochim. Acta*, 3:355–366, 1972.
- [35] Randolph C. Wilhoit. Recent developments in calorimetry. Part one—Introductory survey of calorimetry. *J. Chem. Educ.*, 44(7):A551–A588, Jul. 1967. Number XXXIII in the series Topics in Chemical Instrumentation.
- [36] B. Wunderlich. Trends in thermal analysis. *Anal. Chim. Acta*, 83:35–46, 1985.
- [37] B. Wunderlich. Development towards a single-run DSC for heat capacity measurements. *J. Therm. Anal.*, 32:1949–1955, 1987.
- [38] B. Wunderlich. Modeling the heat flow and heat capacity of modulated differential scanning calorimetry. *J. Therm. Anal.*, 48:207–224, 1997.

### Extended Thermodynamics (TheExt)

Thermodynamics of systems that are subject to external fields and to capillary forces etc.

**Related keywords: extended & thermodynamics | gravity | gravitational | electromagnetic | surface & energy.**

#### Bibliography.

- [1] N. L. Balazs and J. M. Dawson. On thermodynamic equilibrium in a gravitational field. *Physica (Amsterdam)*, 31:222–232, 1965.
- [2] P. W. Bridgman. On the application of thermodynamics to the thermo-electric circuit. *Proc. Natl. Acad. Sci. U. S. A.*, 15(10):765–768, oct 1929.
- [3] John P. Byrne. Rubber elasticity: A simple method for measurement of thermodynamic properties. *J. Chem. Educ.*, 71(6):531–533, Jun. 1994.
- [4] John W. Cahn and John E. Hilliard. Free energy of a nonuniform system. I. Interfacial free energy. *J. Chem. Phys.*, 28(2):258–267, feb 1958.
- [5] B. S. Carey, L. E. Scriven, and H. T. Davies. Semiempirical theory of surface tension of binary systems. *AIChE J.*, 26(5):705–711, sep 1980.
- [6] J. Casas-Vázquez, L. F. del Castillo, D. Jou, and M. Criado-Sancho. Legendre transform in the thermodynamics of flowing polymer solutions. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 63(057101):1–3, 2001.
- [7] P. M. W. Cornelisse, C. J. Peters, and J. de Swaan Arons. Simultaneous prediction of phase equilibria, interfacial tension and concentration profiles. *Mol. Phys.*, 80(4):941–955, Jun. 1993.
- [8] M. Folman and J. L. Shereshefsky. Liquid–vapour equilibrium in microscopic capillaries. II. Non-aqueous systems. *J. Phys. Chem.*, 59(7):607–610, Jul. 1955.
- [9] Kirill Glavatskiy. *Multicomponent Interfacial Transport as Described by the Square Gradient Model: Evaporation and Condensation*. PhD thesis, Norwegian University of Science and Technology, Sep 2009.
- [10] H. J. M. Hanley and Denis J. Evans. A thermodynamics for a system under shear. *J. Chem. Phys.*, 76(6):3225–3232, Mar. 1982.
- [11] David Jou and Ahmed Salhoumi. Legendre transforms in nonequilibrium thermodynamics: An illustration in electrical systems. *Physica A (Amsterdam)*, 283(3–4):163–167, May 2001.
- [12] F. O. Koenig. Note on thermodynamic equilibrium in the gravitational field. *J. Phys. Chem.*, 40(3):373–378, 1936.
- [13] Dilip Kondepudi and Ilya Prigogine. *Modern Thermodynamics. From Heat Engines to Dissipative Structures*. John Wiley & Sons, New York, 1998.

- [14] [Markku J. Lampinen](#), Jari Vuorisalo, and Ulla Pursiheimo. Mathematical analysis of phase rule for systems with electrostatic energy. *J. Chem. Phys.*, 95(11):8402–8409, Dec. 1991.
- [15] [Hong Lin](#), Yuan-Yuan Duan, and Qi Min. Gradient theory modeling of surface tension for pure fluids and binary mixtures. *Fluid Phase Equilib.*, 254:75–90, 2007.
- [16] [La Mer, Victor K.](#) and Ruth Gruen. A direct test of Kelvin’s equation connecting vapour pressure and radius of curvature. *Trans. Faraday Soc.*, 48:410–415, 1952.
- [17] [C. Miqueu](#), B. Mendiboure, A. Graciaa, and J. Lachaise. Modelling of the surface tension of pure components with the gradient theory of fluid interfaces: A simple and accurate expression for the influence parameters. *Fluid Phase Equilib.*, 207:225–246, 2003.
- [18] [C. Miqueu](#), B. Mendiboure, C. Graciaa, and J. Lachaise. Modelling of the surface tension of binary and ternary mixtures with the gradient theory of fluid interfaces. *Fluid Phase Equilib.*, 218:189–203, 2004.
- [19] [Christelle Miqueu](#), Bruno Mendiboure, Alain Graciaa, and Jean Lachaise. Modeling of the surface tension of multicomponent mixtures with the gradient theory of fluid interfaces. *Ind. Eng. Chem. Res.*, 44(9):3321–3329, 2005.
- [20] [Morris Muskat](#). Distribution of non-reacting fluids in the gravitational field. *Phys. Rev.*, 35:1384–1393, Jun. 1930.
- [21] [Yen V. Nguyen](#) and Colin R. Phillips. Calculations of surface excess. *Can. J. Chem. Eng.*, 52:541–542, Aug. 1974.
- [22] [George Oster](#), Alan Perelson, and Aharon Katchalsky. Network thermodynamics. *Nature*, 234:393–399, Dec. 1971.
- [23] [J. Pellicer](#), J. A. Manzanares, J. Zúñiga, and P. Utrillas. Thermodynamics of rubber elasticity. *J. Chem. Educ.*, 78(2):263–267, 2001.
- [24] [A. S. Perelson](#). Network thermodynamics. An overview. *Biophysical Journal*, 15(7):667–685, 1975. Thermodynamics of living systems: A special issue dedicated to Aharon Katchalsky.
- [25] [C. Price](#) and F. P. Wolf. Thermodynamics of rubber elasticity. *Proc. R. Soc. London, Ser. A*, 351(1666):331–350, Nov 1976.
- [26] [António J. Queimada](#), Christelle Miqueu, Isabel M. Marrucho, Georgios M. Kontogeorgis, and João A. P. Coutinho. Modeling vapor–liquid interfaces with the gradient theory in combination with the CPA equation of state. *Fluid Phase Equilib.*, 228–229:479–485, 2005.
- [27] [Sydney Ross](#) and Ian Douglas Morrison. *Colloidal Systems and Interfaces*. John Wiley & Sons, New York, 1988.
- [28] [B. H. Sage](#) and W. N. Lacey. Gravitational concentration gradients in static columns of hydrocarbon fluids. *Petroleum Transactions, AIME*, 132:120–131, Nov. 1939.
- [29] [A. M. Schulte](#). Compositional variations within a hydrocarbon column due to gravity. In *SPE-9235-MS*, SPE Annual Technical Conference and Exhibition, pages 1–10, Dallas, Texas, Sep. 21–24 1980.
- [30] [G. Sonnino](#). Thermodynamic field theory with applications. *Int. J. Quantum Chem.*, 98:191–221, 2004.
- [31] [You-Xiang Zuo](#) and Erling H. Stenby. Calculation of interfacial tensions with gradient theory. *Fluid Phase Equilib.*, 132:139–158, 1997.

### Classical Thermodynamics (TheCla)

The description of simple systems in  $T, p, N$  and  $T, V, N$  coordinates. Gibbsian thermodynamics mostly. There are also some articles and the not-so-simple description in generalized coordinates.

**Related keywords:** thermodynamics | classical | (simple & systems) | gibbsian.

#### Bibliography.

- [1] [Ivar Aavatsmark](#). *Mathematische Einführung in die Thermodynamik der Gemische*. Akademie Verlag GmbH, Berlin, 1995.
- [2] [Michael M. Abbott](#) and Hendrick C. van Ness. *Theory and Problems of Thermodynamics: SI (metric) Edition*. Schaum’s Outline Series. McGraw-Hill Book Company, Inc., New York, 1976.
- [3] [Sumiyoshi Abe](#), S. Martínez, F. Pennini, and A. Plastino. Nonextensive thermodynamic relations. *Physica A (Amsterdam)*, 281(2-3):126–130, Mar. 2001.
- [4] [Acree, Jr., W. E.](#) Comment on free energy and equilibrium. *J. Chem. Educ.*, 63(2):150, 1986.
- [5] [Stephen R. Addison](#). Homogeneous functions in thermodynamics. *Proceedings Arkansas Academy of Science*, 45:114–117, 1991. General Notes.
- [6] [Francesca Aicardi](#). On the classification of singularities in thermodynamics. *Physica D*, 158:175–196, 2001.

- [7] Robert A. Alberty. Use of Legendre transforms in chemical thermodynamics. *Pure Appl. Chem.*, 73(8):1349–1380, 2001. IUPAC Technical Report.
- [8] Robert A. Alberty. Legendre transforms in chemical thermodynamics. *Chem. Rev. (Washington, D. C.)*, 94(6):1457–1482, Sep/Oct 1994.
- [9] Fouad A. Aly and Lloyd L. Lee. Self-consistent equations for calculating the ideal gas heat capacity, enthalpy, and entropy. *Fluid Phase Equilib.*, 6:169–179, 1981.
- [10] Stanley W. Angrist and Loren G. Hepler. *Order and Chaos: Laws of*. Basic Books, New York, 1967.
- [11] Stuart S. Antman. The tragicomical history of thermodynamics 1822–1854. By C. J. Truesdell. Springer-Verlag, New York, 1980. xii + 372 pp. *Am. Math. Monthly*, 90(5):343–346, 1983.
- [12] E. Arunan. van der Waals. *Resonance*, pages 584–587, jul 2010.
- [13] G. Astarita and G. C. Sarti. Modern thermodynamics in chemical engineering and chemistry—I. *Chim. Ind. (Milan)*, 57(10):680–758, Oct. 1975.
- [14] G. Astarita and G. C. Sarti. Modern thermodynamics in chemical engineering and chemistry—II. *Chim. Ind. (Milan)*, 57(11):749–758, Oct. 1975.
- [15] Gianni Astarita. A note on the homogeneous and heterogeneous chemical equilibria. *Chem. Eng. Sci.*, 31:1224–1225, 1976.
- [16] G. Astarita. Historical and philosophical background of thermodynamics. *Ind. Eng. Chem.*, 16(1):138–143, 1977.
- [17] Gianni Astarita. *Thermodynamics. An Advanced Textbook for Chemical Engineers*. Plenum Press, New York, 1989.
- [18] Roger Balian. François massieu and the thermodynamic potentials. *C. R. Physique*, 18:526–530, 2017.
- [19] Scott D. Barnicki. How good are your data? Some thoughts on the measurement and interpretation of vapor–liquid equilibrium. *CEP*, 98(6):58–67, jun 2002.
- [20] Rubin Battino, Laurence E. Strong, and Scott E. Wood. A brief history of thermodynamics notion. *J. Chem. Educ.*, 74(3):304–305, Mar. 1997.
- [21] James A. Beattie and Irwin Oppenheim. *Principles of Thermodynamics*. Studies in Modern Thermodynamics 2. Elsevier, Amsterdam, Holland, 1979.
- [22] Bruce L. Beegle, Michael Modell, and Robert C. Reid. Legendre transforms and their applications in thermodynamics. *AIChE J.*, 20(6):1194–1200, Nov. 1974.
- [23] Henry A. Bent. A simplified algebraic method for obtaining thermodynamical formulas. *J. Chem. Phys.*, 21(8):1408–1409, 1953.
- [24] Klaus Beneke. Hermann Ludwig Ferdinand von Helmholtz und zur geschichte der russischen studentinnen und studenten in heidelberg im letzten jahrhundert. *Mitteilungen der Kolloid-Gesellschaft*, pages 106–150, 1999.
- [25] Pehr H. Björnbom. The independent reactions in calculations of complex equilibria. *Ind. Eng. Chem. Fundam.*, 14(2):102–106, 1975.
- [26] S. M. Blinder. Mathematical methods in elementary thermodynamics. *J. Chem. Educ.*, 43(2):85–92, Feb 1966.
- [27] Alan J. Brainard. The mathematical behavior of extensive and intensive properties of simple systems. *J. Chem. Educ.*, 46(2):104–107, Feb 1969.
- [28] D. E. Breen. To the editor. *J. Chem. Phys.*, 41(5):289, 1964.
- [29] P. W. Bridgman. A complete collection of thermodynamic formulas. *Phys. Rev.*, 3(4):273–281, 1914.
- [30] Brinkley, Jr., Stuart R. Note on the conditions of equilibrium for systems of many constituents. *J. Chem. Phys.*, 14(9):563–564, Sep 1946.
- [31] G. H. Bryan. *Thermodynamics: An Introductory Treatise Dealing Mainly with First Principles and Their Direct Applications*. B. G. Teubners Sammlung von Lehrbüchern auf dem Gebiete der mathematische Wissenschaften mit Einschluß ihrer Anwendungen. Band XXI. B. G. Teubner, Leipzig, 1907.
- [32] H. A. Buchdahl. *The Concepts of Classical Thermodynamics*. Cambridge University Press, Cambridge, London, 1966.
- [33] Joseph W. Bursik. An envelope characteristic of the inversion curve in the  $z, p_r$  plane. *Ind. Eng. Chem. Fundam.*, 10(4):644–646, 1971.
- [34] Herbert Callen. *Thermodynamics. An Introduction to the Physical Theories of Equilibrium Thermodynamics and Irreversible Thermodynamics*. John Wiley & Sons, New York, 1960.
- [35] Herbert Callen. *Thermodynamics and an Introduction to Thermostatistics*. John Wiley & Sons, New York, second edition, 1985.
- [36] C. Carathéodory. Untersuchungen über die grundlagen der thermodynamik. *Mathematische Annalen*, 67:355–386, 1909.
- [37] C. Carathéodory. Investigations into the foundations of thermodynamics. In Joseph Kestin, editor, *The Second Law of Thermodynamics*, number 5 in Benchmark papers on Energy, chapter 12. Dowden,

- Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976. This article was translated expressly for this Benchmark volume by J. Kestin, Brown University, from “Untersuchungen über die Grundlagen der Thermodynamik,” in *Math. Ann.* (Berlin), 67, 355–386 (1909).
- [38] B. Carroll and A. Lehrman. Relations between the derivatives of the thermodynamic functions. *J. Chem. Educ.*, 24(8):389–392, 1947.
- [39] Benjamin Carroll. On the use of Jacobians in thermodynamics. *J. Chem. Educ.*, 42(4):218–221, apr 1965.
- [40] S. Carnot. Réflexions sur la puissance motrice du feu et sur les machines propres à développer cette puissance. *Annales Scientifiques de L'É. N. S.*, 1:393–457, 1872.
- [41] John J. Carroll. What is Henry's law? *Chem. Eng. Prog.*, 87(9):48–52, Sep. 1991.
- [42] John J. Carroll. Henry's law. *J. Chem. Educ.*, 70(2):91–92, Feb. 1993.
- [43] Márcio J. E. de M. Cardoso and José de Medeiros. A thermodynamic framework for solutions based on the osmotic equilibrium concept—1. General formulation. *Pure Appl. Chem.*, 66(3):383–386, 1994.
- [44] J. J. Carroll. Henry's law revisited. *Chem. Eng. Prog.*, pages 49–56, 1999.
- [45] K. C. Cheng. Historical development of the theory of heat and thermodynamics: Review and some observations. *Heat Transfer Eng.*, 13(3):19–37, 1992.
- [46] V. A. Cimmelli and P. Rogolino. On the mathematical structure of thermodynamics with internal variables. *J. Non-Equilib. Thermodyn.*, 26(3):231–242, 2001.
- [47] R. Clausius. Über verschiedene für die anwendung bequeme formen der hauptgleichungen der mechanischen wärmetheorie. *Ann. Phys. Chem.*, CXXV(7):352–400, 1865.
- [48] R. Clausius. *The Mechanical Theory of Heat, with Its Applications to the Steam-Engine and to the Physical Properties of Bodies*. John Van Voorst, London, 1867. English translation of *Mechanische Wärmetheorie* directed by John Tyndall.
- [49] Peter Clark. Review: Elkana on Helmholtz and the conservation of energy. *The British Journal for the Philosophy of Science*, 27(2):165–176, jun 1976.
- [50] Manuk Colakyan and Richard Turton. Cramer's rule and partial thermodynamics properties: A revisit. *Ind. Eng. Chem. Res.*, 27(4):721–723, 1988.
- [51] Thomas W. Copeman and Fred P. Stein. On the derivation and application of solution thermodynamics for reactive components. *Fluid Phase Equilib.*, 9:149–165, 1982.
- [52] F. H. Crawford. Jacobian methods in thermodynamics. *Am. J. Phys.*, 17(1):1–5, 1949.
- [53] F. H. Crawford. On Jacobian methods in thermodynamics. *Am. J. Phys.*, 17(6):397, 1949.
- [54] F. H. Crawford. Maxwell's relations again. *Am. J. Phys.*, 17:450, 1949.
- [55] F. H. Crawford. On the use of curve differentials in thermodynamics. *Am. J. Phys.*, 19(5):284–289, 1951.
- [56] Robert T. DeHoff. *Thermodynamics in Materials Science*. McGraw-Hill Series in Materials Science and Engineering. McGraw-Hill Book Company, Inc., New York, 1993.
- [57] Pablo G. Debenedetti. Generalized Massieu–Planck functions: Geometric representation, extrema and uniqueness properties. *J. Chem. Phys.*, 85(4):2131–2139, Aug 1986.
- [58] Kenneth Denbigh. *The Principles of Chemical Equilibrium. With Applications in Chemistry and Chemical Engineering*. Cambridge University Press, Cambridge, London, fourth edition, 1981.
- [59] Gary W. Dilay and Robert A. Heidemann. Calculation of Joule–Thomson inversion curves from equations of state. *Ind. Eng. Chem. Fundam.*, 25(1):152–158, 1986.
- [60] J. Ernest Dunn and Roger L. Fosdick. The morphology and stability of material phases. *Arch. Ration. Mech. Anal.*, 74:1–99, 1980.
- [61] Freeman J. Dyson. Willard Gibbs and the teaching of science. In *Proceedings of The Gibbs Symposium*, pages 269–276. Yale University, May 15–17 1989.
- [62] Herman Erlichson. Sadi Carnot, “Founder of the second law of thermodynamics”. *Eur. J. Phys.*, 20:183–192, 1999.
- [63] Gentil A. Estévez, Kai Yang, and Basab B. Dasgupta. Thermodynamic partial derivatives and experimentally measurable quantities. *J. Chem. Educ.*, 66(11):890–892, Nov. 1989.
- [64] N. Farah and R. W. Missen. The computer-derivation of thermodynamic equations: Part I. First and second derivatives for simple systems. *Can. J. Chem. Eng.*, 64:154–157, Feb 1986.
- [65] N. Farah and R. W. Missen. The computer-derivation of thermodynamic equations: Part II. First and second derivatives for complex unrestricted systems. *Can. J. Chem. Eng.*, 65:137–141, Feb 1987.
- [66] D. V. Fenby and A. Chand. Enthalpy of the isotropic exchange. *Aust. J. Chem.*, 31:241–245, 1978.
- [67] Enrico Fermi. *Thermodynamics*. Dover Publications, Inc., New York, 1956. Reprint of 1937 edition published by Prentice-Hall Company.
- [68] Vladimir E. Fortov and Igor V. Lomonosov. Thermodynamics of extreme states of matter. *Pure Appl. Chem.*, 69(4):893–904, 1997.
- [69] Walter J. Gensler. Terminology reexamined. Physical versus chemical change. *J. Chem. Educ.*, 47(2):154–155, Feb 1970.

- [70] J. Willard Gibbs. Thermodynamics. In *The Scientific Papers of J. Willard Gibbs, Ph.D., LL.D. In Two Volumes*, volume I. Dover Publications, Inc., New York, 1961. Reprint of 1906 edition published by Longmans, Green and Company.
- [71] J. W. Gibbs. On the equilibrium of heterogeneous substances. In Joseph Kestin, editor, *The Second Law of Thermodynamics*, number 5 in Benchmark papers on Energy, chapter 11. Dowden, Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976. Reprinted from pp.354-372 of *The Scientific Papers of J. W. Gibbs, Vol I: Thermodynamics*, Dover Publications, Inc., New York, 1961 462 pp.
- [72] R. Giles. *Mathematical Foundations of Thermodynamics*. Pergamon Press, Oxford, London, 1964.
- [73] Robert Gilmore. Thermodynamic partial derivatives. *J. Chem. Phys.*, 75(12):5965–5966, Dec. 1981.
- [74] Robert Gilmore. Le Châtelier reciprocal relations. *J. Chem. Phys.*, 76(11):5551–5553, Jun. 1982.
- [75] Robert Gilmore. Higher thermodynamic partial derivatives. *J. Chem. Phys.*, 77(11):5853–5855, Dec. 1982.
- [76] Samuel Glasstone. *Thermodynamics for Chemists*. D. van Nostrand Company, Inc., New York, 1947.
- [77] David Glickenstein. The Legendre transform, jan 2000.
- [78] Raymond E. Goldstein and James S. Walker. Theory of multiple phase separations in binary mixtures: Phase diagrams, thermodynamic properties and comparisons with experiments. *J. Chem. Phys.*, 78(3):1492–1512, Feb 1983.
- [79] A. E. Green and P. M. Naghdi. A re-examination of the basic postulates of thermomechanics. *Proceedings: Mathematical and Physical Sciences*, 432(1885):171–194, Feb. 8 1991.
- [80] Ulrich Grigull. Zum ersten hauptsatz der thermodynamik. *Forsch. Ing.-Wes.*, 29(5):153–154, 1963.
- [81] D. H. E. Gross. Geometric foundation of thermo-statistics, phase transitions, second law of thermodynamics, but without thermodynamic limit. *J. Chem. Soc., Faraday Trans.*, 4:863–872, 2002.
- [82] J. Güémez, C. Fiolhais, and M. Fiolhais. Thermodynamics at work: The pressure derivative of specific heat. *Am. J. Phys.*, 67(12):1100–1108, dec 1999.
- [83] E. A. Guggenheim and J. E. Prue. *Physicochemical Calculations*. Series in Physics. North-Holland Publishing Company, Amsterdam, 1956.
- [84] E. A. Guggenheim. *Thermodynamics. An Advanced Treatment for Chemists and Physicists*. Series in Physics. North-Holland Publishing Company, Amsterdam, third edition, 1957.
- [85] E. A. Guggenheim. *Thermodynamics. An Advanced Treatment for Chemists and Physicists*. North-Holland Publishing Company, Amsterdam, sixth edition, 1977.
- [86] Albert G. Guy. Thermodynamics without entropy. *Appl. Phys. Commun.*, 9(4):305–327, 1989/90.
- [87] D. ter Haar and H. Wergeland. *Elements of Thermodynamics*. Addison-Wesley Series in Advanced Physics. Addison-Wesley Publishing Company, Reading, Massachusetts, 1966.
- [88] R. W. Hakala. A method for relating thermodynamic first derivatives. *J. Chem. Educ.*, 41(2):99–101, 1964.
- [89] Charles S. Hastings. Josiah Willard Gibbs: 1839–1903. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1909.
- [90] George N. Hatsopoulos and Joseph H. Keenan. *Principles of General Thermodynamics*. John Wiley & Sons, New York, 1965.
- [91] Tore Haug-Warberg. *Den termodynamiske arbeidsboken: Boken for de to hundre hjem*. Kolofon forlag, Oslo, 2006.
- [92] R. W. Haywood. *Equilibrium Thermodynamics*. Wiley Interscience. John Wiley & Sons, New York, 1980.
- [93] Robert A. Heidemann. Letter to the editor. *Chem. Eng. Sci.*, 31:1223–1224, 1976.
- [94] Robert A. Heidemann. Non-uniqueness in phase and reaction equilibrium computations. *Chem. Eng. Sci.*, 33:1517–1528, 1978.
- [95] K. F. Herzfeld and Maria Goeppert-Mayer. On the states of aggregation. *J. Chem. Phys.*, 2(1):38–45, jan 1934.
- [96] J. Hertz. Josiah Willard Gibbs and teaching thermodynamics of materials (history). *J. Phase Equilib.*, 13(5):450–458, 1992.
- [97] Terrell L. Hill. Thermodynamics of small systems. *J. Chem. Phys.*, 36(12):3182–3197, 1962.
- [98] Terrell L. Hill. *Thermodynamics of Small Systems. Part I*. Dover Publications, Inc., New York, 1994. Two volumes bound as one. Reprint of 1963 edition published by W. A. Benjamin, Inc.
- [99] Hans Holtan. *Kjemisk termodynamikk*. Yrkesopplæringsrådet for håndverk og industri, Oslo, Norge, 1971.
- [100] ICE [unknown]. Obituary. Rudolf Julius Emanuel Clausius, 1822–1888. *Minutes of the Proceedings of the Institution of Civil Engineers*, 96:307–316, 1889. Part 2.
- [101] Kenneth R. Jolls. Gibbs and the art of thermodynamics. In *Proceedings of The Gibbs Symposium*, pages 293–321. Yale University, May 15–17 1989.
- [102] Kenneth R. Jolls, Michael C. Schmitz, and Daniel C. Coy. Seeing is believing: A new look at an old subject. *Chem. Eng. J. (London)*, 30(497):42–46, May 1991.

- [103] [Kenneth R. Jolls](#) and Jeffrey L. Butterbaugh. Confirming thermodynamic stability. *Chem. Eng. Educ.*, pages 124–129, 1992.
- [104] [Kenneth R. Jolls](#) and Daniel C. Coy. Gibbs’s models visualized. *Phys. Today*, 45(3):96–98, Mar 1992.
- [105] [Vladimir V. Kechin](#). Melting curve equations at high pressure. *Phys. Rev. B: Solid State*, 65(5):052102, feb 2002.
- [106] [Joseph Kestin](#). *A Course in Thermodynamics*. Blaisdell Publishing Company, Waltham, Massachusetts, 1966.
- [107] [Joseph Kestin](#), editor. *The Second Law of Thermodynamics*. Number 5 in Benchmark papers on Energy. Dowden, Hutchinson & Ross, Inc., Stroudsburg, Pennsylvania, 1976.
- [108] [M. B. King](#). Phase equilibrium in mixtures. In P. V. Danckwerts, editor, *International Series of Monographs in Chemical Engineering*, volume 9. Pergamon Press, Oxford, London, 1969.
- [109] [S. J. Kline](#) and F. O. Koenig. The state principle—Some general aspects of the relationships among the properties of systems. *J. Appl. Mech.*, 24:29–34, mar 1957.
- [110] [Yoshikata Koga](#). Vapor pressures of dilute aqueous *t*-butyl alcohol: How dilute is the Henry’s law region? *J. Phys. Chem.*, 99:6231–6233, Jun. 1995.
- [111] [Hermann Kopp](#). Investigations of the specific heat of solid bodies. *Philos. Trans. R. Soc. London*, 155:71–202, 1865.
- [112] [Mordechai L. Kremer](#). Thermodynamic state functions of mixtures. *J. Chem. Educ.*, 42(12):649–650, Dec 1965.
- [113] [V. Kumaran](#). Josiah Willard Gibbs. *Resonance*, pages 4–11, jul 2007.
- [114] [S. K. Kumar](#) and R. C. Reid. Derivation of the relationships between partial derivatives of Legendre transforms. *AIChE J.*, 32(7):1224–1226, Jul 1986.
- [115] [C. K. Kwok](#) and D. R. Tilley. A review of some thermodynamic properties of the Van der Waals gas. *Phys. Educ.*, 14:422–425, 1979.
- [116] [Frank Lerman](#). A method for deriving expressions for the first partial derivatives of thermodynamic functions. *J. Chem. Phys.*, 5:792–794, oct 1937.
- [117] [Lewis, Gilbert Newton](#) and Merle Randall. *Thermodynamics and the Free Energy of Chemical Substances*. McGraw-Hill Book Company, Inc., New York, first edition, 1923.
- [118] [Gilbert Newton Lewis](#) and Merle Randall. *Thermodynamics*. McGraw-Hill Series in Advanced Chemistry. McGraw-Hill Book Company, Inc., New York, second edition, 1961. Revised by Kenneth S. Pitzer and Leo Brewer.
- [119] [Elliott H. Lieb](#) and Jakob Yngvason. A fresh look at entropy and the second law of thermodynamics. *Phys. Today*, 53(4):32–37, 2000.
- [120] [Elliot H. Lieb](#) and Jakob Yngvason. The physics and mathematics of the Second law of thermodynamics. *Phys. Rep.*, 310(1):1–96, 1999.
- [121] [Elliot H. Lieb](#) and Jakob Yngvason. Erratum—The physics and mathematics of the Second Law of thermodynamics. *Phys. Rep.*, 314:669, 1999.
- [122] [Maurice W. Lindauer](#). The evolution of the concept of chemical equilibrium from 1775 to 1923. *J. Chem. Educ.*, 39(8):384–390, aug 1962.
- [123] [Paolo Lubini](#) and Michele D’Anna. Adopting the chemical potential in the high school curriculum: Why not? *Chimica*, 72(1/2):32–35, 2018.
- [124] [A. Rey de Luna](#) and M. Zamora. Canonical formalism in equilibrium thermodynamics. *J. Chem. Phys.*, 84(8):4612–4618, Apr. 1986.
- [125] [A. Rey de Luna](#) and M. Zamora. Canonical formalism in equilibrium thermodynamics. II. Homogeneity, stability, and generalized Maxwell relations. *J. Chem. Phys.*, 84(8):4619–4624, Apr. 1986.
- [126] [C. H. P. Lupis](#). *Chemical Thermodynamics of Materials*. North-Holland Publishing Company, Amsterdam, 1983.
- [127] [Alan Macdonald](#). A new statement of the second law of thermodynamics. *Am. J. Phys.*, 63(12):1122–1127, dec 1995.
- [128] [A. J. Majumdar](#) and Rustum Roy. Test of the applicability of the Clapeyron relationship to a few cases of solid–solid transitions. *J. Inorg. Nucl. Chem.*, 27:1961–1973, 1965.
- [129] [Stanislaw Malanowski](#) and Andrzej Anderko. *Modelling Phase Equilibria. Thermodynamic Background and Practical Tools*. Wiley Series in Chemical Engineering. John Wiley & Sons, New York, 1992.
- [130] [Paul M. Mathias](#) and John P. O’Connell. The Gibbs–Helmholtz equation and the thermodynamic consistency of chemical absorption data. *Ind. Eng. Chem. Res.*, 51:5090–5097, 2012.
- [131] [José Luis de Medeiros](#) and Márcio José Estillac de Mello Cardoso. On the derivation of thermodynamic equilibrium criteria. *Fluid Phase Equilib.*, 136(1–2):1–13, Nov. 1997.
- [132] [Michael L. Michelsen](#) and Jørgen Møllerup. Partial derivatives of thermodynamic properties. *AIChE J.*, 32(8):1389–1392, Aug 1986.

- [133] L. Mistura. A symmetry property of Van der Waals like equations of state. *J. Chem. Phys.*, 78(2):1002–1003, Jan 1983.
- [134] Michael Modell and Robert C. Reid. *Thermodynamics and Its Applications*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1983.
- [135] Jørgen M. Møllerup and Michael L. Michelsen. Calculation of thermodynamic equilibrium properties. *Fluid Phase Equilib.*, 74:1–15, 1992.
- [136] K. Morgan. An expansion equation of state subroutine. *Comput. Phys. Comm.*, 5:64–68, 1973.
- [137] Eddie Morild. Termodynamikkens matematiske grunnlag. Teknisk rapport, Kjemisk institutt, Universitetet i Bergen, 1975.
- [138] R. Mrugala. Geometrical formulation of equilibrium phenomenological thermodynamics. *Rep. Math. Phys.*, 14(3):419–427, 1978.
- [139] James A. Mullins, James B. Rawlings, and Keith P. Johnston. Partial derivative quantities from phase equilibria relationships for multicomponent mixtures, 1992.
- [140] James A. Mullins, James B. Rawlings, and Keith P. Johnston. Partial derivative quantities from phase equilibria relationships for mixtures. *AIChE J.*, 39(8):1363–1369, Aug 1993.
- [141] C. E. Mungan. Legendre transforms for dummies, 2009.
- [142] Walther Nernst. *Rudolf Clausius*. Ludwig Röhrscheid, Bonn, 1922.
- [143] van Ness, Hendrick C. and Michael M. Abbott. *Classical Thermodynamics of Nonelectrolyte Solutions: With Applications to Phase Equilibria*. McGraw-Hill Chemical Engineering Series. McGraw-Hill Book Company, Inc., New York, 1982.
- [144] H. C. van Ness. *Understanding Thermodynamics*. Dover Publications, Inc., New York, 1983. Reprint of 1969 edition published by McGraw-Hill Book Company, Inc.
- [145] R. Nieto, M. C. González, and F. Herrero. Thermodynamics of mixtures: Functions of mixing and excess functions. *Am. J. Phys.*, 67(12):1096–1099, dec 1999.
- [146] Darrell Kirk Nordstrom and James L. Munoz. *Geochemical Thermodynamics*. Blackwell Scientific Publications, Oxford, London, second edition, 1994.
- [147] James D. Nulton and Peter Salamon. Geometry of the ideal gas. *Phys. Rev. A: At., Mol., Opt. Phys.*, 31(4):2520–2523, Apr. 1985.
- [148] J. J. O’Connor and E. F. Robertson. William John Macquorn Rankine. MacTutor History of Mathematics archive, University of St Andrews, 2005.
- [149] John P. O’Connell. Thermodynamics. A structure for teaching and learning about much of reality. *Chem. Eng. Educ.*, pages 96–101, 1993.
- [150] W. A. Oates. Ideal solutions. *J. Chem. Educ.*, 46(8):501–504, Aug. 1969.
- [151] Thomas Oikonomou and Gokhan B. Bagci. Clausius versus Sackur–Tetrode entropies, 2013. Accepted for publication in Studies in History and Philosophy of Modern Physics Part B.
- [152] Claudio Olivera-Fuentes. Multicomponent fugacity coefficients and residual properties from pressure-explicit equations of state. *Chem. Eng. Sci.*, 46(8):2019–2029, 1991.
- [153] James D. Olson. Thermodynamic consistency testing of *ptx*-data via the Gibbs–Helmholtz equation. *Fluid Phase Equilib.*, 14:383–392, 1983.
- [154] Giulio Ottonello. *Principles of Geochemistry*. Columbia University Press, New York, 1997.
- [155] David A. Palmer. *Handbook of Applied Thermodynamics*. CRC Press, Inc., Boca Raton, Florida, 1987.
- [156] Athanassios Z. Panagiotopoulos and Robert C. Reid. A unified treatment of departure functions, fugacities, and activity coefficients, Nov. 1985.
- [157] A. Z. Panagiotopoulos and R. C. Reid. On the relationship between pairwise fluctuations and thermodynamic derivatives. *J. Chem. Phys.*, 85(8):4650–4653, Oct. 1986.
- [158] Gavin D. Peckham and Ian J. McNaught. Heat and work are not “forms of energy”. *J. Chem. Educ.*, 70(2):103–104, feb 1993.
- [159] Gavin D. Peckham and Ian J. McNaught. Phase diagrams of one-component systems. What most textbooks don’t say, but should! *J. Chem. Educ.*, 70(7):560–561, Jul. 1993.
- [160] Gavin D. Peckham and Ian J. McNaught. First-law problem solving. *J. Chem. Educ.*, 70(8):625–626, Aug. 1993.
- [161] Willi Pentermann. Zur Problematik der Vorausberechnung von Verdampfungsgleichgewichten überkritischer Systeme mit Hilfe von Aktivitätskoeffizienten. *Chem.-Ing.-Tech.*, 56(9):703–705, 1984.
- [162] Randal L. Perry, John C. Telotte, and John P. O’Connell. Solution thermodynamics for “reactive” components. *Fluid Phase Equilib.*, 5:245–277, 1980/1981.
- [163] Pierre Perrot. *A to Z of Thermodynamics*. Oxford University Press, London, 1998. Original French edition: Dictionnaire de thermodynamique.
- [164] Mark A. Peterson. Analogy between thermodynamics and mechanics. *Am. J. Phys.*, 47(6):488–490, jun 1979.

- [165] J. M. Phillips. Mnemonic diagrams for thermodynamic systems. *J. Chem. Educ.*, 64(8):674–675, 1987.
- [166] F. P. Pike. Vapor–liquid equilibrium in water–nitric acid systems at 760mmHg. Technical Report CF-50-6-143, Oak Ridge National Laboratory (US), 1950.
- [167] R. C. Pinkerton. A Jacobian method for the rapid evaluation of thermodynamic derivatives, without the use of Tables. *J. Phys. Chem.*, 56:799–800, 1952.
- [168] A. B. Pippard. *Elements of Classical Thermodynamics for Advanced Students of Physics*. Cambridge University Press, Cambridge, London, 1960.
- [169] Max Planck. *Treatise on Thermodynamics*. Dover Publications, New York, 1945.
- [170] A. Plastino and A. R. Plastino. On the universality of thermodynamics Legendre transform structure. *Physica A (Amsterdam)*, 226:257–263, Feb. 1997.
- [171] L. Pogliani. The diagrammatic method, the Planck and the Massieu functions. *J. Chem. Educ.*, 78(5):680–681, 2001.
- [172] L. Pogliani. Graphs and thermodynamics. *J. Math. Chem*, 46:15–23, 2009.
- [173] Lionello Pogliani and Camillo La Mesa. The mnemonic diagram for thermodynamic relationships. *J. Chem. Educ.*, 69(10):808–809, 1992.
- [174] John M. Prausnitz, Rüdiger N. Lichtenthaler, and Edumdo Gomes de Azevedo. *Molecular Thermodynamics of Fluid-Phase Equilibria: Solutions Manual to Accompany*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, third edition, 2000.
- [175] John M. Prausnitz, Ruediger N. Lichtenthaler, and Edmundo Gomes de Azevedo. *Molecular Thermodynamics of Fluid-Phase Equilibria*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1986.
- [176] I. Prigogine, A. Bellemans, and V. Mathot. *The Molecular Theory of Solutions*. Series in Physics. North-Holland Publishing Company, Amsterdam, 1957.
- [177] Quílez-Pardo, Juan and Joan Josep Solaz-Portolés. Students’ and teachers’ misapplication of le chatelier’s principle: Implications for the teaching of chemical equilibrium. *Journal of Research in Science Teaching*, 32(9):939–957, 1995.
- [178] Norman F. Ramsey. Thermodynamics and statistical mechanics at negative absolute temperatures. *Phys. Rev.*, 103(1):20–28, Jul. 1956.
- [179] Y. K. Rao. Extended form of the Gibbs phase rule. *Chem. Eng. Educ.*, pages 40–43,46–49, 1985.
- [180] Kjelstrup Ratkje, Signe. Euler homogenous functions and their application to thermodynamics. Technical report, Div. Phys. Chem., Norwegian Techn. Univ., 1974.
- [181] Otto Redlich. Fundamental thermodynamics since Caratheodory. *Rev. Mod. Phys.*, 40(3):556–563, jul 1968.
- [182] Otto Redlich. Terminology reexamined. Intensive and extensive properties. *J. Chem. Educ.*, 47(2):154,156, Feb 1970.
- [183] Otto Redlich. The so-called zeroth law of thermodynamics. *J. Chem. Educ.*, 47:740–741, 1970.
- [184] Otto Redlich. Science and mathematics. *J. Chem. Educ.*, 49:222–225, 1972.
- [185] Otto Redlich. *Thermodynamics: Fundamentals. Applications*. Elsevier, Amsterdam, Holland, 1976.
- [186] Robert C. Reid, John M. Prausnitz, and Thomas K. Sherwood. *The Properties of Gases and Liquids*. McGraw-Hill Book Company, Inc., New York, 3 edition, 1977.
- [187] Robert C. Reid, John M. Prausnitz, and Bruce E. Poling. *The Properties of Gases and Liquids*. McGraw-Hill Book Company, Inc., New York, 4 edition, 1987.
- [188] M. Ricou. The laws of thermodynamics for non-cyclic processes. In J. Serrin, editor, *New Perspectives in Thermodynamics*, chapter 6. Springer-Verlag, New York, 1986.
- [189] Peter A. Rock. *Chemical Thermodynamics*. University Science Books, Mill Valley, California, 1983.
- [190] W. H. Rodebush. The third law of thermodynamics. *J. Chem. Phys.*, 2(10):668–670, oct 1934.
- [191] J. Rodriguez and A. J. Brainard. An improved mnemonic diagram for thermodynamic relationships. *J. Chem. Educ.*, 66(6):495–496, 1989.
- [192] D. H. Root, H. T. Haselton, Jr., and I-Ming Chou. An alternative method for the reduction of thermodynamic derivatives. *J. Chem. Educ.*, 71(4):303–306, Apr 1994.
- [193] E. B. Rudnyi and L. N. Sidorov. Once more on  $\delta g$  of chemical reaction. Department of Chemistry, Moscow State University, 1983.
- [194] Peter Salamon, Bjarne Andresen, Paul D. Gait, and R. Stephen Berry. The significance of Weinhold’s length. *J. Chem. Phys.*, 73(2):1001–1002, Jul. 1980.
- [195] P. Salamon, B. Andresen, P. D. Gait, and R. S. Berry. Erratum: The significance of Weinhold’s length. *J. Chem. Phys.*, 73(10):5407, Nov. 1980.
- [196] P. Salamon, J. Nulton, and E. Ihrig. On the relation between entropy and energy versions of thermodynamic length. *J. Chem. Phys.*, 80(1):436–437, Jan. 1984.
- [197] Michael R. Samuels. The maximum number of solid phases in a multiphase reacting system. *Ind. Eng. Chem. Fundam.*, 10(4):643–644, 1971.

- [198] Michael R. Samuels. Interpreting thermodynamic consistency: How bad is “bad”? *Ind. Eng. Chem. Fundam.*, 11(3):422–424, 1972.
- [199] Michael R. Samuels, Dean L. Ulrichson, and F. Dee Stevenson. Interpretation of overall area tests for thermodynamic consistency: The effect of random error. *AIChE J.*, 18(5):1004–1009, sep 1972.
- [200] G. Sarton. The discovery of the law of conservation of energy. *Isis*, 13:18–44, sep 1929. Facsimile reproductions of J. R. Mayer, J. P. Joule and Sadi Carnot.
- [201] George Scatchard. Equilibrium in non-electrolyte mixtures. *Chem. Rev. (Washington, D. C.)*, 44(1):7–35, 1949.
- [202] S. F. Sciamanna and J. M. Prausnitz. Thermodynamics of highly dilute solutions and the quest for ultrapurity. *AIChE J.*, 33(8):1315–1321, Aug. 1987.
- [203] J. Serrin, editor. *New Perspectives in Thermodynamics*. Springer-Verlag, New York, 1986.
- [204] A. N. Shaw. The derivation of thermodynamical relations for a simple system. *Philos. Trans. R. Soc. London*, 234(740):299–328, 1935.
- [205] David Siminovitch. Entropy revisited, gorilla and all. *Phys. Today*, 53(10):11, 2000.
- [206] N. K. Simha and K. Bhattacharya. Equilibrium conditions at corners and edges of an interface in a multiphase solid. *Mater. Sci. Eng., B*, 238(1):32–41, 1997.
- [207] J. M. Smith and H. C. van Ness. *Introduction to Chemical Engineering Thermodynamics*. McGraw-Hill Book Company, Inc., New York, third edition, 1975.
- [208] Jeffrey V. Smith, Ronald W. Missen, and William R. Smith. General optimality criteria for multiphase multireaction chemical equilibrium. *AIChE J.*, 39(4):707–710, Apr. 1993.
- [209] Joan Josep Solaz and Juan Quílez. Changes of extent of reaction in open chemical equilibria. *Chemistry Education: Research and Practice in Europe*, 2(3):303–312, 2001.
- [210] G. R. Somayajulu, A. P. Kudchadker, and B. J. Zwolinski. Thermodynamics. *Annu. Rev. Phys. Chem.*, 16:213–244, 1965.
- [211] Daniel F. Styer. A thermodynamic derivative means an experiment. *Am. J. Phys.*, 67(12):1094–1095, dec 1999.
- [212] Richard A. Swalin. *Thermodynamics of Solids*. Wiley Series on the Science and Technology of Materials. John Wiley & Sons, New York, 1962.
- [213] Józef Szarawara and Andrzej Gawdzik. Method of calculation of fugacity coefficients from cubic equations of state. *Chem. Eng. Sci.*, 44(7):1489–1494, 1989.
- [214] P. G. Tait. *Sketch of Thermodynamics*. Edmonton and Douglas, Edinburgh, 1868.
- [215] P. G. Tait. *Sketch of Thermodynamics*. David Douglas, Edinburgh, second edition, 1877.
- [216] P. G. Tait. Poincaré’s Thermodynamique, 1892.
- [217] Jefferson W. Tester and Michael Modell. *Thermodynamics and Its Applications*. Prentice Hall International Series in the Physical and Chemical Engineering Sciences. Prentice Hall PTR, Upper Saddle River, New Jersey, third edition, 1997.
- [218] Kaj Thomsen. Thermodynamics of electrolyte solutions. Department of Chemical Engineering, Technical University of Denmark, 2006.
- [219] James B. Thompson. Chemical reactions in crystals. *Am. Mineral.*, 54(3/4):341–375, 1969.
- [220] Laszlo Tisza. The thermodynamics of phase equilibrium. *Ann. Phys. (N. Y.)*, 13:1–92, 1961.
- [221] Laszlo Tisza. *Generalized Thermodynamics*. The M. I. T. Press, Cambridge, 1966.
- [222] Arthur Tobolsky. A systematic method of obtaining the relations between thermodynamic derivatives. *J. Chem. Phys.*, 10:644–645, oct 1942.
- [223] Myron Tribus. *Thermostatistics and Thermodynamics*. D. van Nostrand Company, Inc., New York, 1961.
- [224] C. Truesdell. *Rational Thermodynamics*. McGraw-Hill Series in Modern Applied Mathematics. McGraw-Hill Book Company, Inc., New York, 1969.
- [225] Clifford Truesdell. The tragicomedy of classical thermodynamics. In *Course held at the Department of Mechanics of Solids*, number 70 in Courses and Lectures, pages 1–41, Udine, Jul 1971. International Centre for Mechanical Sciences.
- [226] Constantine Tsouopoulos. Otto Redlich, 1896 – 1978: In memory and appreciation. *Fluid Phase Equilib.*, 12(1/2):1–9, 1983.
- [227] J. L. Tveekrem, R. H. Cohn, and S. C. Greer. The effect of critical fluctuations on chemical equilibrium. *J. Chem. Phys.*, 86(6):3602–3606, Mar. 1987.
- [228] Jos Uffink. Bluff your way in the second law of thermodynamics. *Stud. Hist. Phil. Mod. Phys.*, 32(3):305–394, 2001.
- [229] S. Velasco, F. L. Román, and J. A. White. On the Clausius–Clapeyron vapor pressure equation. *J. Chem. Educ.*, 86(1):106–111, 2009.
- [230] K. Walther and M. Röttinger. *Technische Wärmelehre*. Sammlung Göschen. G. J. Göschen’sche Verlagshandlung, Leipzig, 1905.

- [231] E. Weltin. Let a computer balance your chemical equations and determine the number of independent reactions. *J. Chem. Educ.*, 71(4):295–297, Apr. 1994. Computer Series, 160.
- [232] Peter Westh, Charles A. Haynes, and Yoshikata Koga. How dilute is the Henry’s law region? II. *J. Phys. Chem. B*, 102:4982–4987, Jun. 1998.
- [233] John C. Wheeler. Geometric constraints at triple points. *J. Chem. Phys.*, 61(11):4474–4489, Dec. 1974.
- [234] R. J. Wheaton. Treatment of variations of composition with depth in gas-condensate systems. *SPE Reservoir Eng.*, pages 239–244, May 1991.
- [235] J. C. Whitwell and S. R. Dartt. Independent reactions in the presence of isomers. *AIChE J.*, 19(6):1114–1120, Nov 1973.
- [236] B. Widom. Two ideas from Gibbs: The entropy inequality and the dividing surface. In *Proceedings of The Gibbs Symposium*, pages 73–87. Yale University, May 15–17 1989.
- [237] Emmerich Wilhelm and Trevor Letcher, editors. *Volume Properties. Liquids, Solutions and Vapours*. The Royal Society of Chemistry, 2015.
- [238] Emmerich Wilhelm and Rubin Battino. Thermodynamic functions of the solubilities of gases in liquids at 25°C. *Chem. Rev. (Washington, D. C.)*, 73(1):1–9, feb 1973.
- [239] R. C. Wilhoit. Ideal gas thermodynamic functions. *TRC Curr. Data News*, 3(2):1–4, 1975.
- [240] Jaime Wisniak. Benoit Paul Emile Clapeyron: A short bibliographical sketch. *Chem. Educator*, 5:83–87, 2000.
- [241] Jaime Wisniak. Frederick Thomas Trouton: The man, the rule, and the ratio. *Chem. Educ.*, 6:55–61, Oct. 2001.
- [242] R. H. Wood, P. T. Thompson, and T. H. Lilley. Improvements in the representation of the thermodynamic properties of fluids. In *Fluid and Fluid Mixtures*, pages 103–111, Teddington Middlesex, UK, 1979. IPC Science and Technology Press. Proceedings of the NPL Conference: Chemical Thermodynamic Data on Fluid and Fluid Mixtures: Their Estimation, Correlation and Use, 11–12 September 1978.
- [243] P. A. H. Wyatt. Vapour phase reactions and the Duhem–Margules equation (with some reference to nitric acid). *Trans. Faraday Soc.*, 50:352–357, 1954.
- [244] Takuya Yamano. On the robust thermodynamical structures against arbitrary entropy form and energy mean value. *The European Physical Journal B*, 18:103–106, 2000.
- [245] Mark W. Zemansky. *Heat and Thermodynamics*. McGraw-Hill Book Company, Inc., New York, second edition, 1943.
- [246] Mark W. Zemansky and Richard H. Dittman. *Heat and Thermodynamics: An Intermediate Textbook*. McGraw-Hill Book Company, Inc., New York, 1997.

## Continuous Thermodynamics (TheCon)

Non-discrete description of polydisperse systems, polymers, hydrocarbon fluids, etc.

**Related keywords: thermodynamics & continuous.**

### Bibliography.

- [1] Ronald L. Cotterman and John M. Prausnitz. Flash calculations for continuous or semicontinuous mixtures using an equation of state. *Ind. Eng. Chem.*, 24:434–443, 1985.
- [2] Ronald L. Cotterman, Rainer Bender, and John M. Prausnitz. Phase equilibria for mixtures containing very many components. Development and application of continuous thermodynamics for chemical process design. *Ind. Eng. Chem.*, 24:194–203, 1985.
- [3] T. P. J. Halpin and Nick Quirke. A new method of continuous thermodynamics applied in an equation of state. *SPEJ, Soc. Pet. Eng. J.*, pages 617–622, 1990.
- [4] Horst Kehlen, M. T. Rätzsch, and Joachim Bergmann. Continuous thermodynamics of multicomponent systems. *AIChE J.*, 31(7):1136–1148, Jul. 1985.
- [5] Horst Kehlen and Margit T. Rätzsch. Complex multicomponent distillation calculations by continuous thermodynamics. *Chem. Eng. Sci.*, 42(2):221–232, 1987.
- [6] Jia Lin Liu and David Shan Hill Wong. Rigorous implementation of continuous thermodynamics using orthonormal polynomials. *Fluid Phase Equilib.*, 129(1-2):113–127, 1997.
- [7] A. R. Plastino, H. G. Miller, and A. Plastino. General thermostatical formalism based on parameterized entropic measures. *Continuum Mech. Thermodyn.*, 16:269–277, 2004.
- [8] M. T. Rätzsch and H. Kehlen. *Kontinuierliche Thermodynamik von Vielstoffgemischen*. Number 13 in Sitzungsberichte der Akademie der Wissenschaften der DDR. Akademie Verlag GmbH, Berlin, 1982.
- [9] M. T. Rätzsch and H. Kehlen. Continuous thermodynamics of complex mixtures. *Fluid Phase Equilib.*, 14:225–234, 1983.

- [10] Shao-hwa Wang and Wallace B. Whiting. A group-contribution, continuous-thermodynamics framework for calculation of vapor–liquid equilibria. *Can. J. Chem. Eng.*, 65:651–661, Aug. 1987.
- [11] B. T. Willman and A. S. Teja. Continuous thermodynamics of phase equilibria using a multivariate distribution function and an equation of state. *AIChE J.*, 32(12):2067–2078, 1986.

### Finite Time Thermodynamics (TheFin)

Finite time thermodynamics. Analyses of machinery.

**Related keywords:** thermodynamics & (finite & time | optimum & work | machines).

#### Bibliography.

- [1] B. Agnew, A. Anderson, and T. H. Frost. Optimisation of a steady-flow Carnot cycle with external irreversibilities for maximum specific output. *Appl. Therm. Eng.*, 17(1):3–15, 1997.
- [2] B. Andresen, M. H. Rubin, and R. S. Berry. Availability for finite time processes. General theory and model. *J. Phys. Chem.*, 87:2704–2713, 1983.
- [3] Bjarne Andresen and J. M. Gordon. Constant thermodynamic speed for minimizing entropy production in thermodynamic processes and simulated annealing. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 50(6):4346–4351, Dec. 1994.
- [4] Bjarne Andresen. Finite-time thermodynamics and thermodynamic length. *Rev. Gen. Therm.*, 35:647–650, 1996.
- [5] F. Angulo-Brown. An ecological optimization criterion for finite-time heat engines. *J. Appl. Phys.*, 69(11):7465–7469, 1991.
- [6] F. Angulo-Brown and R. Páez-Hernández. Endoreversible thermal cycle with a nonlinear heat-transfer law. *J. Appl. Phys.*, 74(4):2216–2219, 1993.
- [7] F. Angulo-Brown, J. A. Rocha-Martinez, and T. D. Navarrete-Gonzalez. A non-endoreversible Otto cycle model: Improving power output and efficiency. *J. Phys. D: Appl. Phys.*, 29:80–83, 1996.
- [8] F. Angulo-Brown, L. A. Arias-Hernandez, and R. Paez-Hernandez. A general property of non-endoreversible thermal cycles. *J. Phys. D: Appl. Phys.*, 32:1415–1420, 1999.
- [9] L. A. Arias-Hernández and F. Angulo-Brown. A general property of endoreversible thermal engines. *J. Appl. Phys.*, 81(7):2973–2979, 1997.
- [10] Adrian Bejan. Engineering advances on finite-time thermodynamics. *Am. J. Phys.*, 62(1):11–12, Jan. 1994.
- [11] Adrian Bejan. Entropy generation minimization: The new thermodynamics of finite-size devices and finite time processes. *J. Appl. Phys.*, 79(3):1191–1218, Feb. 1996.
- [12] A. Bejan. The first NATO advanced study institute on thermodynamic optimization (Neptun, Romania, 1998). *Energy (Oxford)*, 24:753–759, 1999.
- [13] C. H. Blanchard. Coefficient of performance for finite-speed heat pump. *J. Appl. Phys.*, 51(5):2471–2472, May 1980.
- [14] G. R. Brown, S. Snow, P. Andresen, and P. Salamon. Finite-time thermodynamics of a porous plug. *Phys. Rev. A: At., Mol., Opt. Phys.*, 34:4370–4379, 1986.
- [15] Martin E. Carrera-Patiño. Adiabatic processes in monoatomic gases. *J. Chem. Phys.*, 89(4):2271–2277, Aug. 1988.
- [16] Lixuan Chen and Zijun Yan. The effect of heat-transfer law on performance of a two-heat-source endoreversible cycle. *J. Chem. Phys.*, 90(7):3740–3743, apr 1989.
- [17] L. G. Chen, F. R. Sun, C. Wu, and R. L. Kiang. Theoretical analysis of the performance of a regenerative closed Brayton cycle with internal irreversibilities. *Energy Convers. Manage.*, 38(9):871–877, 1997.
- [18] Lingen Chen, Fengrui Sun, and Chih Wu. Influence of heat transfer law on the performance of a Carnot engine. *Appl. Therm. Eng.*, 17(3):277–282, 1997.
- [19] Jincan Chen, Zijun Yan, Lixuan Chen, and Bjarne Andresen. Efficiency bound of a solar-driven Stirling heat engine system. *Int. J. Energy Res.*, 22:805–812, 1998.
- [20] Lingen Chen, Fengrui Sun, and Chih Wu. Optimal expansion of a heated working fluid with phenomenological heat transfer. *Energy Convers. Manage.*, 39(3):149–156, 1998.
- [21] Ching-Yang Cheng and Cha’o-Kuang Chen. Ecological optimization of an endoreversible Brayton cycle. *Energy Convers. Manage.*, 39(1/2):33–44, 1998.
- [22] L. G. Chen, C. Wu, and F. R. Sun. Finite time thermodynamic optimization or entropy generation minimization of energy systems. *J. Non-Equilib. Thermodyn.*, 24:327–359, 1999.
- [23] Lingen Chen, Fengrui Sun, and Chih Wu. Effect of heat transfer law on the performance of a generalized irreversible Carnot engine. *J. Phys. D: Appl. Phys.*, 32:99–105, 1999.

- [24] Ching-Yang Cheng and Cha'o-Kuang Chen. Ecological optimization of an irreversible Brayton heat engine. *J. Phys. D: Appl. Phys.*, 32(3):350–357, 1999.
- [25] Cummins, Jr., C. Lyle. *Internal Fire*. Society of Automotive Engineers, Inc., Warrendale, PA, 1989.
- [26] F. L. Curzon and B. Ahlhorn. Efficiency of a Carnot engine at maximum power output. *Am. J. Phys.*, 43:22–24, jan 1975.
- [27] Massimo Dentice D'Accadia, Maurizio Sasso, and Sergio Sibilio. Optimum performance of heat engine-driven heat pumps: A finite-time approach. *Energy Convers. Manage.*, 38(4):401–413, 1997.
- [28] L. B. Erbay and H. Yavuz. Analysis of the Stirling heat engine at maximum power conditions. *Energy (Oxford)*, 22(7):645–650, 1997.
- [29] L. B. Erbay and H. Yavuz. An analysis of an endoreversible heat engine with combined heat transfer. *J. Phys. D: Appl. Phys.*, 30(20):2841–2847, 1997.
- [30] L. B. Erbay and H. Yavuz. The maximum cooling density of a realistic Stirling refrigerator. *J. Phys. D: Appl. Phys.*, 31(3):291–293, 1998.
- [31] L. Berrin Erbay and Hasbi Yavuz. Analysis of an irreversible Ericsson engine with a realistic regenerator. *Appl. Energy*, 62(3):155–167, 1999.
- [32] J. M. Gordon. Generalized power versus efficiency characteristics of heat engines—The thermoelectric generator as an instructive illustration. *Am. J. Phys.*, 59(6):551–555, 1991.
- [33] J. M. Gordon. A response to Yan and Chen's "comment on "generalized power versus efficiency characteristics of heat engines: The thermoelectric generator as an instructive illustration". *Am. J. Phys.*, 61(4):381, 1993.
- [34] J. M. Gordon and V. N. Orlov. Performance characteristics of endoreversible chemical engines. *J. Appl. Phys.*, 74(9):5303–5309, 1993.
- [35] D. Gutowicz-Krusin, I. Procaccia, and J. Ross. On the efficiency of rate processes, power and efficiency of heat engines. *J. Chem. Phys.*, 69(9):3898–3906, nov 1978.
- [36] K. H. Hoffman, Stanley J. Watowich, and R. S. Berry. Optimal paths for thermodynamic systems. The ideal Diesel-cycle. *J. Appl. Phys.*, 58(6):2125–2134, sep 1985.
- [37] H. G. Ladas and O. M. Ibrahim. Finite-time view of the Stirling engine. *Energy (Oxford)*, 19(8):837–843, 1994.
- [38] Valentina A. Mironova, Anatolii M. Tsirlin, Vladimir A. Kazakov, and R. Stephen Berry. Finite-time thermodynamics: Exergy and optimization of time-constrained processes. *J. Appl. Phys.*, 76(2):629–636, Jul. 1994.
- [39] F. Moukalled, R. Y. Nuwayhid, and S. Fattal. A universal model for studying the performance of Carnot-like engines at maximum power conditions. *Int. J. Energy Res.*, 20:203–214, 1996.
- [40] Michael Mozurkewich and R. Stephen Berry. Optimal paths for thermodynamic systems. The ideal Otto-cycle. *J. Appl. Phys.*, 53(1):34–42, jan 1982.
- [41] Mary Jo Ondrechen, Bjarne Andresen, Michael Mozurkewich, and R. Stephen Berry. Maximum work from a finite reservoir by sequential Carnot cycles. *Am. J. Phys.*, 49(7):681–685, 1981.
- [42] Mary Jo Ondrechen, Morton H. Rubin, and Yehuda B. Band. The generalized Carnot cycle: A working fluid operating in finite time between finite heat sources and sinks. *J. Chem. Phys.*, 78(7):4721–4727, Apr. 1983.
- [43] C. A. Ordonez. Liquid nitrogen fueled, closed Brayton cycle cryogenic heat engine. *Energy Convers. Manage.*, 41(4):331–341, 2000.
- [44] V. N. Orlov and R. S. Berry. Estimation of minimal heat consumption for heat-driven separation processes via methods of finite-time thermodynamics. *J. Phys. Chem.*, 95:5624–5628, 1991.
- [45] V. N. Orlov and R. S. Berry. Power and efficiency limits for internal-combustion engines via methods of finite-time thermodynamics. *J. Appl. Phys.*, 74(7):4317–4322, 1993.
- [46] Georghe Popescu, Vsevolod Radcenco, Monica Costea, and Michel Feidt. Finite-time thermodynamics optimisation of an endo- and exo-irreversible Stirling motor. *Rev. Gen. Therm.*, 35(41):656–661, 1996.
- [47] J. M. M. Roco, S. Velasco, A. Medina, and A. C. Hernandez. Optimum performance of a regenerative Brayton thermal cycle. *J. Appl. Phys.*, 82(6):2735–2741, 1997.
- [48] Fredy Romm. On optimizing maximum efficiency of finite-time engines. *J. Appl. Phys.*, 74(9):5310–5313, Nov. 1993.
- [49] Morton H. Rubin. Optimal configuration of an irreversible heat engine with fixed compression ratio. *Phys. Rev. A: At., Mol., Opt. Phys.*, 22(4):1741–1752, oct 1980.
- [50] B. Sahin, A. Kodal, T. Yilmaz, and H. Yavuz. Maximum power density analysis of an irreversible Joule–Brayton engine. *J. Phys. D: Appl. Phys.*, 29(5):1162–1167, 1996.
- [51] Peter Salamon and Abraham Nitzan. Finite time optimizations of a Newton's law Carnot cycle. *J. Chem. Phys.*, 74(6):3546–3560, Mar. 1981.
- [52] Peter Salamon and R. Stephen Berry. Thermodynamic length and dissipated availability. *Phys. Rev. Lett.*, 51(13):1127–1130, Sep. 1983.

- [53] Stanislaw Sieniutycz, Peter Salamon, et al. Finite-time thermodynamics and thermoeconomics. In Stanislaw Sieniutycz and Peter Salamon, editors, *Advances in Thermodynamics*, volume 4. Taylor & Francis, Washington, D. C., 1990.
- [54] Stanislaw Sieniutycz and Michael R. von Spakovsky. Finite time generalization of thermal exergy. *Energy Convers. Manage.*, 39(14):1423–1447, 1998.
- [55] Stanislaw Sieniutycz. Generalized Carnot problem of maximum work in finite time via Hamilton–Jacobi–Bellman theory. *Energy Convers. Manage.*, 39(16–18):1735–1743, 1998.
- [56] Stanislaw Sieniutycz. Carnot problem of maximum work from a finite resource interacting with environment in a finite time. *Physica A (Amsterdam)*, 26:234–263, 1999.
- [57] Anatolii M. Tsirlin, Vladimir A. Kazakov, and R. Stephen Berry. Finite-time thermodynamics: Limiting performance of rectification and minimal entropy production in mass transfer. *J. Phys. Chem.*, 98(13):3330–3336, 1994.
- [58] J. V. C. Vargas and A. Bejan. Thermodynamic optimization of the match between two streams with phase change. *Energy (Oxford)*, 25:15–33, 2000.
- [59] S. Velasco, J. M. M. Roco, A. Medina, J. A. White, and A. C. Hernández. Optimization of heat engines including the saving of natural resources and the reduction of thermal pollution. *J. Phys. D: Appl. Phys.*, 33:355–359, 2000.
- [60] Alexis de Vos. Efficiency of some heat engines at maximum power conditions. *Am. J. Phys.*, 53(6):570–573, jun 1984.
- [61] A. de Vos. Endoreversible thermoeconomics. *Energy Convers. Manage.*, 36(1):1–5, 1995.
- [62] A. de Vos. Endoreversible economics. *Energy Convers. Manage.*, 38(4):311–317, 1997.
- [63] A. de Vos. Endoreversible thermodynamics versus economics. *Energy Convers. Manage.*, 40:1009–1019, 1999.
- [64] Chih Wu and William H. Schulden. Maximum obtainable specific power of high-temperature waste heat engines. *Heat Recovery Systems & CHP*, 15(1):13–17, 1995.
- [65] Chih Wu, Lingen Chen, and Fengrui Sun. Effect of heat transfer law on the finite-time exergoeconomic performance of heat engines. *Energy (Oxford)*, 21(12):1127–1134, 1996.
- [66] Chin Wu, Lingen Chen, and Fengrui Sun. Performance of a regenerative Brayton heat engine. *Energy (Oxford)*, 21(2):71–76, 1996.
- [67] Chih Wu, Lingen Chen, and Fengrui Sun. Effect of heat transfer law on finite-time exergoeconomic performance of Carnot heat pump. *Energy Convers. Manage.*, 39(7):579–588, 1998.
- [68] F. Wu, L. G. Chen, C. Wu, and F. G. Sun. Optimum performance of irreversible Stirling engine with imperfect regeneration. *Energy Convers. Manage.*, 39(8):727–732, 1998.

### Irreversible Thermodynamics (TheIrr)

Onsager irreversible thermodynamics and transport phenomena (not radiation).

**Related keywords:** thermodynamics & irreversible | transport | diffusion.

#### Bibliography.

- [1] K. G. Denbigh. Entropy creation in open reaction systems. *Trans. Faraday Soc.*, 48:389–395, 1952.
- [2] B. C. Eu and L. S. Garcia-Colin. Irreversible processes and temperature. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 54(3):2501–2512, 1996.
- [3] Donald D. Fitts. *Nonequilibrium Thermodynamics*. McGraw-Hill Series in Advanced Chemistry. McGraw-Hill Book Company, Inc., New York, 1962.
- [4] Tormod Førland and Signe Kjelstrup Ratkje. Irreversible thermodynamics of isotope flow. *Lab. Phys. Chem.*, The Norwegian Institute of Technology.
- [5] Soney C. George and Sabu Thomas. Transport phenomena through polymeric systems. *Prog. Polym. Sci.*, 26:985–1017, 2001.
- [6] A. N. Gorban, G. S. Yablonskii, and V. I. Bykov. The path to equilibrium. *Int. Chem. Eng.*, 22(2):368–375, Apr 1982.
- [7] S. R. de Groot and P. Mazur. *Non-Equilibrium Thermodynamics*. North-Holland Publishing Company, Amsterdam, 1969.
- [8] S. R. de Groot and P. Mazur. *Non-Equilibrium Thermodynamics*, chapter 2-5, pages 11–57. North-Holland Publishing Company, Amsterdam, 1969.
- [9] Sybren R. de Groot and Peter Mazur. *Non-Equilibrium Thermodynamics*. Dover Publications, Inc., New York, 1984. Reprint of 1962 edition published by North-Holland Publishing Company, Amsterdam.
- [10] Rolf Haase. *Thermodynamics of Irreversible Processes*. Dover Publications, Inc., New York, 1990. Reprint of 1969 edition published by Addison-Wesley Publishing Company, Reading, Massachusetts.

- [11] [Haslach, Jr., Henry W.](#) *Maximum Dissipation Non-Equilibrium Thermodynamics and its Geometric Structure*. Springer-Verlag, New York, 2011.
- [12] [Tamio Ikeshoji](#) and Signe Kjelstrup Ratkje. Thermoelectric power of a cell with complex formation. *J. Electrochem. Soc.*, 133(6):1107–1113, Jun 1986.
- [13] [Edwin T. Jaynes](#). The minimum entropy production principle (1980). In R. D. Rosenkrantz, editor, *E. T. Jaynes: Papers on Probability, Statistics and Statistical Physics*, chapter 14. D. Reidel Publishing Company, Dordrecht, Holland, 1983.
- [14] [R. Krishna](#). A unified theory of separation processes based on irreversible thermodynamics. In *Fourth BOC Priestly Conference*, pages 64–88, 1986.
- [15] [Gerard D. C. Kuiken](#). *Thermodynamics of Irreversible Processes: Applications to Diffusion and Rheology*. John Wiley & Sons, New York, 1994.
- [16] [G. Lebon](#) and P. Mathieu. Comparison of diverse theories of nonequilibrium thermodynamics. *Int. Chem. Eng.*, 23(4):651–662, Oct 1983.
- [17] [Kristian Lindgren](#) and Bengt Å. G. Månsson. Entropy production in a chaotic chemical system. *Z. Naturforsch., A: Phys., Phys. Chem., Kosmophys.*, 41:1111–1117, Apr 1986.
- [18] [Bengt Å. G. Månsson](#). Entropy production in oscillating chemical systems. *Z. Naturforsch., A: Phys., Phys. Chem., Kosmophys.*, 40:877–884, 1985.
- [19] [R. E. Morel](#) and George Fleck. A fourth law of thermodynamics. *Chemistry*, 15(4):305–310, 2006.
- [20] [Eddie Morild](#). Dissipative (energiforbrukende) strukturer. *Naturen*, 102(4):161–165, 1978.
- [21] [V. P. S. Nain](#) and John R. Ferron. Prediction of binary diffusion coefficient for polar gas mixtures. *Ind. Eng. Chem. Fundam.*, 11(3):420–421, May 1972.
- [22] [R. E. Nettleton](#). The Gibbs equation from maximum entropy. *J. Chem. Phys.*, 93(11):8247–8253, Dec. 1990.
- [23] [L. Onsager](#). Reciprocal relations in irreversible processes. I. *Phys. Rev.*, 37:405–426, Feb 1931.
- [24] [L. Onsager](#). Reciprocal relations in irreversible processes. II. *Phys. Rev.*, 38:2265–2279, Dec. 1931.
- [25] [I. Prigogine](#). *Introduction to Thermodynamics of Irreversible Processes*. John Wiley & Sons, New York, second edition, 1961.
- [26] [I. Prigogine](#). *Introduction to Thermodynamics of Irreversible Processes*. John Wiley & Sons, New York, third edition, 1967.
- [27] [J. G. Ramos](#), A. R. Vasconcellos, and R. Luzzi. Considerations on nonequilibrium entropy and temperature. *Brazilian Journal of Physics*, 30(3):617–646, sep 2000.
- [28] [Kjelstrup Ratkje, Signe](#). Thermodynamics of evolution. Technical report, Div. Phys. Chem., The Norwegian Institute of Technology, Oct 1974.
- [29] [Kjelstrup Ratkje, Signe](#). Temperature gradient snow metamorphosis. *Polar Research*, 3:141–143, 1985.
- [30] [Signe Kjelstrup Ratkje](#) and Jakob de Swaan Arons. Denbigh revisited: Reducing lost work in chemical processes. *Chem. Eng. Sci.*, 50(10):1551–1560, 1995.
- [31] [Torben Smith Sørensen](#) and Lars Eger. On thermodynamics, variational principles and stability in chemical reaction systems. Towards a rational thermodynamics. IV. *Acta Chem. Scand., Ser. A*, 39:775–797, 1985.
- [32] [Rod Swenson](#). The fourth law of thermodynamics or the law of maximum entropy production (LMEP). *Chemistry*, 18(5):333–339, 2009.
- [33] [Richard C. Tolman](#) and Paul C. Fine. On the irreversible production of entropy. *Rev. Mod. Phys.*, 20(1):51–77, 1948.
- [34] [D. Tondeur](#) and E. Kvaalen. Equipartition of entropy production. An optimality criterion for transfer and separation processes. *Ind. Eng. Chem. Res.*, 26:50–56, 1987.
- [35] [R. J. Tykodi](#). Thermodynamics of steady states: Is the entropy-production surface convex in the thermodynamic space of steady currents? *J. Chem. Phys.*, 80(4):1652–1655, Feb. 1984.

## Second Law Analyses (The2nd)

Second law analysis of chemical plants, energy availability studies, exergy efficiency, etc.

**Related keywords:** second & law | exergy & analysis.

### Bibliography.

- [1] [Joachim Ahrendts](#). Reference states. *Energy (Oxford)*, 5:667–677, 1980.
- [2] [Ichiro Aoki](#). Exergy analysis of network systems at steady state. *Ecol. Modell.*, 62:183–193, 1992.
- [3] [Vedat S. Arpacı](#) and Ahmet Selamet. Entropy production in flames. In *87-HT-55*, National Heat Transfer Conference, pages 1–6, Pittsburg, Pennsylvania, Aug. 9-12 1987. The American Society of Mechanical Engineers.
- [4] [Vedat S. Arpacı](#) and Ahmet Selamet. Entropy production in flames. *Combust. Flame*, 73:251–259, 1988.

- [5] Robert U. Ayres. The second law, the fourth law, recycling and limits to growth. *Ecol. Econ.*, 29:473–483, 1999.
- [6] Baumgartner, Friherre von, Andr. Om varmens mekaniske ækvivalent og dens betydning i naturvidenskaben. *Polyteknisk Tidsskrift*, 3(20):305–316, 31. Okt. 1856. Oversatt til norsk fra tysk artikkel i *Dinglers Polyt. Journal*.
- [7] Adrian Bejan. Graphic techniques for teaching engineering thermodynamics. *Thermal and Fluids Engineering News*, pages 26–28, 1977.
- [8] Adrian Bejan. General criterion for rating heat-exchanger performance. *Int. J. Heat Mass Transfer*, 21:655–658, 1978.
- [9] Adrian Bejan. Two thermodynamic optima in the design of sensible heat units for energy storage. *J. Heat Transfer*, 100:708–712, Nov 1978.
- [10] Adrian Bejan. Second law analysis in heat transfer. *Energy (Oxford)*, 5:721–732, 1980.
- [11] Adrian Bejan. *Entropy Generation through Heat and Fluid Flow*. John Wiley & Sons, New York, 1982.
- [12] Adrian Bejan. *Advanced Engineering Thermodynamics*. John Wiley & Sons, New York, 1988.
- [13] A. Bejan. Theory of heat transfer-irreversible power plants. *Int. J. Heat Mass Transfer*, 31(6):1211–1219, 1988.
- [14] A. Bejan, J. V. C. Vargas, and M. Sokolov. Optimal allocation of a heat-exchanger inventory in heat driven refrigerators. *Int. J. Heat Mass Transfer*, 38(16):2997–3004, 1995.
- [15] A. Bejan and M. R. Errera. Maximum power from a hot stream. *Int. J. Heat Mass Transfer*, 41(13):2025–2035, 1998.
- [16] Giacomo Bisio. On a general statement for efficiency. *Chem. Eng. Commun.*, 81:177–195, 1989.
- [17] Jerald A. Caton and Joe E. West. A review and thermodynamic analysis of an external combustion, reciprocating engine. In *1996 Spring Technical Conference of the ASME*, volume 3, pages 43–55, 1996.
- [18] Jincan Chen. The efficiency of an irreversible combined cycle at maximum specific power output. *J. Phys. D: Appl. Phys.*, 29:2818–2822, 1996.
- [19] Jincan Chen. Optimal performance analysis of irreversible cycles used as heat pumps and refrigerators. *J. Phys. D: Appl. Phys.*, 30:582–587, 1997.
- [20] K. G. Denbigh. The second-law efficiency of chemical processes. *Chem. Eng. Sci.*, 6(1):1–9, 1956.
- [21] W. R. Dunbar, N. Lior, and R. A. Gaggioli. The component equations of energy and exergy. *J. Energy Resour. Technol.*, 114:75–83, Mar. 1992.
- [22] H. G. Fortak. Local balance equations for atmospheric exergies and anergies. *Meteorol. Atmos. Phys.*, 67:169–180, 1998.
- [23] Richard A. Gaggioli. The concept of available energy. *Chem. Eng. Sci.*, 16:87–96, 1961.
- [24] R. A. Gaggioli. The concept of thermodynamic friction, thermal available energy, chemical available energy and thermal energy. *Chem. Eng. Sci.*, 17:523–530, 1962.
- [25] Richard A. Gaggioli and Peter J. Petit. Use the second law first. *CHEMTECH*, 7:496–506, Aug 1977.
- [26] R. A. Gaggioli, D. A. Sama, Qian Sanhong, and Y. M. El-Sayed. Integration of a new process into an existing site: A case study in the application of exergy analysis. *J. Eng. Gas Turbines Power*, 113:170–183, Apr. 1991.
- [27] Ernest S. Geskin. Second law analysis of fuel consumption in furnaces. *Energy (Oxford)*, 5:949–954, 1980.
- [28] A. J. G. G. Graveland and E. Gisolf. Exergy analysis: An efficient tool for process optimization and understanding. *Comput. Chem. Eng.*, 22:S545–S552, 1998. Supplement: European Symposium on Computer Aided Process Engineering—8.
- [29] M. A. Habib. Thermodynamic analysis of the performance of cogeneration plants. *Energy (Oxford)*, 17(5):485–491, 1992.
- [30] R. W. Haywood. A critical review of the theorems of thermodynamic availability, with concise formulations. *J. Mech. Eng. Sci.*, 16(3):160–173, 1974.
- [31] A. P. Hinderink, F. P. J. M. Kerkhof, A. B. K. Lie, J. de Swaan Arons, and H. J. van der Koo. Exergy analysis with a flowsheeting simulator—I. Theory: Calculating exergies of material streams. *Chem. Eng. Sci.*, 51(20):4693–4700, 1996.
- [32] A. P. Hinderink, F. P. J. M. Kerkhof, A. B. K. Lie, J. de Swaan Arons, and H. J. van der Koo. Exergy analysis with a flowsheeting simulator—II. Application: Synthesis gas production from natural gas. *Chem. Eng. Sci.*, 51(20):4701–4715, 1996.
- [33] J. Jeżowski. The pinch design method for tasks with multiple pinches. *Comput. Chem. Eng.*, 16(2):129–133, 1992.
- [34] Hongguang Jin and Masaru Ishida. Graphic presentation of exergy loss in mixing on an energy-utilization diagram. *J. Chem. Eng. Jpn.*, 25(6):755–757, 1992.
- [35] Joseph Kestin. Availability: The concept and associated terminology. *Energy (Oxford)*, 5:679–692, 1980.

- [36] G. E. Kniel, K. Delmarco, and J. G. Petrie. Life cycle assessment applied to process design: Environmental and economic analysis and optimization of a nitric acid plant. *Environmental Progress*, 15(4):221–228, 1996.
- [37] T. J. Kotas. Exergy method of thermal and chemical plant analysis. *Chem. Eng. Res. Des.*, 64:212–229, 1986.
- [38] Robert J. Krane. A second law analysis of a thermal energy storage system with Joulean heating of the storage element. In *85-WA/HT-19*, Winter Annual Meeting, pages 1–10, Miami Beach, Florida, Nov. 17–21 1985. The American Society of Mechanical Engineers.
- [39] L. N. Lauerhass and D. F. Rudd. On the thermodynamics of the chemical heat pump. *Chem. Eng. Sci.*, 36:803–807, 1981.
- [40] Kam W. Li. *Applied Thermodynamics: Availability Method and Energy Conversion*. Combustion: An International Series. Taylor & Francis, Washington, D. C., 1996.
- [41] F. di Liberto, G. Monroy, F. Peruggi, and P. Ruggiero. Comment on “thermodynamic exploration of an unconventional heat-engine: The circular cycle”. *J. Phys. D: Appl. Phys.*, 36:1222–1226, 2003.
- [42] B. Linnhoff and E. Hindmarsh. The pinch design method for heat exchanger networks. *Chem. Eng. Sci.*, 38(5):745–763, 1983.
- [43] B. Linnhoff. Pinch technology for the synthesis of optimal heat and power systems. *J. Energy Resour. Technol.*, 111:137–147, 1989.
- [44] John Malone and the Invention of Liquid-based Engines. Los Alamos Science (21), 1993.
- [45] Jussi Manninen and X. X. Zhu. Thermodynamic analysis and mathematical optimisation of power plants. *Comput. Chem. Eng.*, 22:537–544, 1998.
- [46] Thomas V. Marcella and Eric Sheldon. Thermodynamic exploration of an unconventional heat-engine: The circular cycle. *J. Phys. D: Appl. Phys.*, 33:2402–2406, 2000.
- [47] David R. Morris and Jan Szargut. Standard chemical exergy of some elements and compounds on the planet Earth. *Energy (Oxford)*, 11(8):733–755, 1986.
- [48] David R. Morris, Frank R. Steward, and Jan Szargut. Technological assessment of chemical metallurgical processes. *Can. Metall. Q.*, 33(4):289–295, 1994.
- [49] M. J. Moran and E. Sciubba. Exergy analysis: Principles and practice. *J. Eng. Gas Turbines Power*, 116(2):285–290, 1994.
- [50] O. Mullins and R. S. Berry. Minimization of entropy production in distillation. *J. Phys. Chem.*, 88:723–728, 1984.
- [51] W. Muschik. Formulations of the second law—Recent developments. *J. Phys. Chem. Solids*, 49(6):709–720, 1988.
- [52] Bertil Myr en. Entropibalanser och kemitekniska processers genomf rbarhet. *Kem. - Kemi*, 6(5):227–230, 1979.
- [53] Noel de Nevers and J. D. Seader. Mechanical lost work, thermodynamic lost work, and thermodynamic efficiencies of processes. *Lat. Am. J. Heat Mass Transfer*, 8:77–105, 1984.
- [54] R. Y. Nuwayhid, F. Moukalled, and N. Noueihed. On entropy generation in thermoelectric devices. *Energy Convers. Manage.*, 41:891–914, 2000.
- [55] Luis Rodriguez. Calculation of available-energy quantities. *ACS Symp. Ser.*, 122:39–60, 1980.
- [56] D. A. Sama. The use of the second law of thermodynamics in process design. *J. Energy Resour. Technol.*, 117:179–185, Sep. 1995.
- [57] D. A. Sama. Differences between second law analysis and pinch technology. *J. Energy Resour. Technol.*, 117(3):186–191, 1995.
- [58] Erik Sauar, Signe Kjelstrup, and Kristian M. Lien. Equipartition of forces—Extension to chemical reactors. *Comput. Chem. Eng.*, 21:29–34, 1997.
- [59] William James Sidis. *The Animate and the Inanimate*. Richard G. Badger, Boston, 1925.
- [60] Stanislaw Sieniutycz. Hamilton–Jacobi–Bellman theory of dissipative thermal availability. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 56(5):5051–5064, nov 1997.
- [61] S. Sieniutycz. Hamilton–Jacobi–Bellman analysis of irreversible thermal exergy. *Int. J. Heat Mass Transfer*, 41(1):183–195, 1998.
- [62] M. Sorin, A. Hammache, and O. Diallo. Exergy based approach for process synthesis. *Energy (Oxford)*, 25:105–129, 2000.
- [63] M. V. Sorin and V. M. Brodyansky. A method for thermodynamic optimization—I. Theory and application to an ammonia synthesis plant. *Energy (Oxford)*, 11:1019–1031, 1992.
- [64] M. V. Sorin and P. Le Goff. A method for the thermodynamic optimization—II. Application to an absorption heat pump and generalization of the method. *Energy (Oxford)*, 11:1033–1048, 1992.
- [65] M. Sorin and J. Paris. Combined exergy and pinch approach to process analysis. *Comput. Chem. Eng.*, 21:S23–S28, 1997. Supplement.
- [66] M. Sorin, J.-C. Bonhivers, and J. Paris. Exergy efficiency and conversion of chemical reactions. *Energy Convers. Manage.*, 39(16–18):1863–1868, 1998.

- [67] [M. Sorin](#) and [J. Paris](#). Integrated exergy load distribution method and pinch analysis. *Comput. Chem. Eng.*, 23:497–507, 1999.
- [68] [K. S. Spiegler](#). *Principles of Energetics*. Springer-Verlag, New York, 1983. Based on: *Applications de la thermodynamique du non-équilibre* by [P. Chartier](#), [M. Cross](#) and [K. S. Spiegler](#).
- [69] [M. V. Sussman](#). Steady-flow availability and the standard chemical availability. *Energy*, 5:793–802, 1980.
- [70] [Jan Szargut](#). International progress in second law analysis. *Energy (Oxford)*, 5:709–718, 1980.
- [71] [Jan Szargut](#). Chemical exergies of the elements. *Appl. Energy*, 32:269–286, 1989.
- [72] [André Thess](#). Was ist Entropie? Eine Antwort für Unzufriedene. *Forsch. Ingenieurwes.*, DOI 10.1007/s10010-007-0063-7, 2008.
- [73] [Göran Wall](#). Exergy flows in industrial processes. *Energy (Oxford)*, 13(2):197–208, 1988.
- [74] [L. C. Witte](#) and [N. Shamsundar](#). A thermodynamic efficiency concept for heat exchange devices. *J. Eng. Power*, 105:199–203, 1983.
- [75] [Larry C. Witte](#). Second law optimization of heat exchangers. In *87-HT-65*, National Heat Transfer Conference, pages 1–8, Pittsburg, Pennsylvania, Aug. 9-12 1987. The American Society of Mechanical Engineers.
- [76] [Roger J. Zemp](#), [Sergio H. B. de Faria](#), and [Maria de L. Oliveira Maia](#). Driving force distribution and exergy loss in the thermodynamic analysis of distillation columns. *Comput. Chem. Eng.*, 21(Supplement):S523–S528, 1997.
- [77] [Jizhong Zhou](#), [Shijun Ma](#), and [George W. Hinman](#). Ecological exergy analysis: A new method for ecological energetics research. *Ecol. Modell.*, 84:291–303, 1996.

## CHAPTER 4

# Thermodynamic Models

### Semi-Empirical Correlations (ModCor)

Semi-empirical models other than thermodynamic ones, general correlations, corresponding state principle, etc.

**Related keywords:** empirical | correlations | corresponding & state.

#### Bibliography.

- [1] [Nikolay N. Akinfiev](#) and Larryn W. Diamond. A simple predictive model of quartz solubility in water – salt – CO<sub>2</sub> systems at temperatures up to 1000°C and pressures up to 1000 MPa. *Geochim. Cosmochim. Acta*, 73:1597–1608, 2009.
- [2] [Robert A. Alberty](#). Thermodynamics of chemical reactions written in terms of homologous series. *J. Chem. Phys.*, 86(4):2243–2248, Feb 1987.
- [3] [F. A. Ashraf](#) and J. H. Vera. On the derivation of the equations for the estimation of the minimum solubility of “solids” in compressed gases. *Chem. Eng. Commun.*, 4:563–567, 1980.
- [4] [Valentín García Baonza](#), Mercedes Cáceres, and Javier Núñez. Universal compressibility behavior of dense phases. *Phys. Rev. B: Solid State*, 51(1):28–37, Jan. 1995.
- [5] [Eberhard Bender](#), Udo Klein, Wolfgang Ph. Schmitt, and John M. Prausnitz. Thermodynamics of gas solubility: Relation between equation-of-state and activity-coefficient models. *Fluid Phase Equilib.*, 15:241–255, 1984.
- [6] [Scott W. Cambell](#) and George Thodos. Prediction of saturated-liquid densities and critical volumes for polar and nonpolar substances. *J. Chem. Eng. Data*, 30(1):102–111, 1985.
- [7] [Scott W. Campbell](#) and George Thodos. Prediction of saturated-liquid densities and critical volumes for polar and nonpolar substances. *J. Chem. Eng. Data*, 30(1):102–111, 1985.
- [8] [Scott W. Campbell](#). A good initial estimate for pure-component vapor pressures in equation of state calculations. *Ind. Eng. Chem. Res.*, 27:1333–1335, 1988.
- [9] [Eldred H. Chimowitz](#), Thomas F. Anderson, Sandro Macchietto, and Leroy F. Stutzman. Local models for representing phase equilibria in multicomponent, nonideal vapor – liquid and liquid – liquid systems. 1. Thermodynamic approximation functions. *Ind. Eng. Chem. Process Des. Dev.*, 22:217–225, 1983.
- [10] [Eldred H. Chimowitz](#), Sandro Macchietto, Thomas F. Anderson, and Leroy F. Stutzman. Local models for representing phase equilibria in multicomponent, nonideal vapor – liquid and liquid – liquid systems. 2. Application to process design. *Ind. Eng. Chem. Process Des. Dev.*, 23:609–618, 1984.
- [11] [James S. Chickos](#), William E. Acree, Jr., and Joel F. Liebman. Estimating solid – liquid phase change enthalpies and entropies. *J. Phys. Chem. Ref. Data*, 28(6):1535–1673, 1999.
- [12] [N. Cohen](#). Revised group additivity values for enthalpies of formation (at 298 K) of carbon – hydrogen and carbon – hydrogen – oxygen compounds. *J. Phys. Chem. Ref. Data*, 25(6):1411–1481, 1996.
- [13] [Eugene S. Domalski](#) and Elizabeth D. Hearing. Estimation of the thermodynamic properties of hydrocarbons at 298.15 K. *J. Phys. Chem. Ref. Data*, 17(4):1637–1678, 1988.
- [14] [P. G. Glugla](#) and S. M. Sax. Vapor liquid equilibrium for salt-containing systems: A correlation of vapor pressure depression and a prediction of multicomponent systems. *AIChE J.*, 31(11):1911–1914, nov 1985.
- [15] [M. Hillestad](#), Carsten Sørli, T. F. Anderson, I. Olsen, and T. Hertzberg. On estimating the error of local thermodynamic models—A general approach. *Comput. Chem. Eng.*, 13(7):789–796, 1989.
- [16] [Mohamed W. M. Hisham](#) and Sidney W. Benson. The thermochemistry of inorganic solids. IV. Enthalpies of formation of compounds of the formula MX<sub>a</sub>Y<sub>b</sub>. *J. Phys. Chem. Ref. Data*, 16(3):467–470, 1987.
- [17] [Lee Hong-Yi](#) and Liu Guojie. A generalized equation of state for liquid density calculation. *Fluid Phase Equilib.*, 108:15–25, 1995.
- [18] [W. Alexander van Hook](#). Vapor pressures of the deuterated ethanes. *J. Chem. Phys.*, 44(1):234–251, Jan. 1966.

- [19] G. A. Iglesias-Silva, J. C. Holste, P. T. Eubank, K. N. Marsh, and K. R. Hall. A vapor pressure equation from extended asymptotic behavior. *AIChE J.*, 33(9):1550–1556, Sep 1987.
- [20] Jack K. Kung, Frank N. Nazario, Joseph Joffe, and Dimitrios Tasslos. Prediction of Henry’s constants in mixed solvents from binary data. *Chem. Eng. Commun.*, 23(1):170–175, 1984.
- [21] V. A. Kuz. Thermodynamic vapour pressure equation. Triple and critical point applications. Prediction of a linear logarithmic relation between surface tension and latent heat of evaporation. *Fluid Phase Equilib.*, 66:113–124, 1991.
- [22] Byung Ik Lee and Michael G. Kesler. A generalized thermodynamic correlation based on three-parameter corresponding states. *AIChE J.*, 21(3):510–527, May 1975.
- [23] Hong-Yi Lee and Guojie Liu. Prediction of thermal pressure coefficient by the law of corresponding state and group contribution. *Fluid Phase Equilib.*, 105:141–152, 1995.
- [24] Sandro Macchietto, Eldred H. Chimowitz, Thomas F. Anderson, and Leroy F. Stutzman. Local models for representing phase equilibria in multicomponent, nonideal vapor–liquid and liquid–liquid systems. 3. Parameter estimation and update. *Ind. Eng. Chem. Process Des. Dev.*, 25:674–682, 1986.
- [25] T. R. Marrero and E. A. Mason. Correlation and prediction of gaseous diffusion coefficients. *AIChE J.*, 19(3):498–503, May 1973.
- [26] Joseph J. Martin. Correlation of second virial coefficients using a modified cubic equation of state. *Ind. Eng. Chem. Fundam.*, 23(4):454–459, 1984.
- [27] Kh. Mejbri and A. Bellagi. Corresponding states correlation for the saturated vapor pressure of pure fluids. *Thermochim. Acta*, 436:140–149, 2005.
- [28] Ray A. Mentzer, Robert A. Greenkorn, and Kwang-Chu Chao. Principle of corresponding states and vapor–liquid equilibria of molecular fluids, and their mixtures with light gases. *Ind. Eng. Chem. Process Des. Dev.*, 20:240–252, 1981.
- [29] Vincenzo Piacente, Danilo Fontana, and Paolo Scardala. Enthalpies of vaporization of a homologous series of *n*-alkanes determined from vapor pressure measurements. *J. Chem. Eng. Data*, 39:231–237, 1994.
- [30] Kenneth Pitzer. Corresponding states for perfect liquids. *J. Chem. Phys.*, 7:583–590, Aug 1939.
- [31] D. H. L. Prasad. Edmister’s rule for the acentric factor. *Trans. Inst. Chem. Eng.*, 72:123–124, Jan 1994.
- [32] Harold G. Rackett. Equation of state for saturated liquids. *J. Chem. Eng. Data*, 15(4):514–517, 1970.
- [33] H. Reiss and S. W. Mayer. Law of corresponding states for fused salts. *J. Chem. Phys.*, 35(3):820–826, Sep 1961.
- [34] K. Roth, U. Wolf, and G. Wolf. A simple method for modelling and prediction of the specific heat and density of aqueous electrolyte solutions. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 21(4):475–481, 1997.
- [35] Robert L. Scott. Corresponding states treatment of nonelectrolyte solutions. *J. Chem. Phys.*, 25(2):193–205, Aug. 1956.
- [36] Buford D. Smith, Ol Muthu, Ashok Dewan, and Matthew Gierlach. Critical evaluation of vapor–liquid equilibrium, heat of mixing, and volume change of mixing data. General procedures. *J. Phys. Chem. Ref. Data*, 11(3):941–951, 1982.
- [37] Giorgio S. Soave. Estimation of the critical constants of heavy hydrocarbons for their treatment by the Soave–Redlich–Kwong equation of state. *Fluid Phase Equilib.*, 13:29–39, 1998.
- [38] Calvin F. Spencer and Ronald P. Danner. Improved equation for prediction of saturated liquid density. *J. Chem. Eng. Data*, 17(2):236–241, 1972.
- [39] Sverre Støren and Terje Hertzberg. Local thermodynamic models applied in dynamic process simulation: A simplified approach. *Chem. Eng. Res. Des.*, 72(A3):395–401, 1994.
- [40] H. Theliander and U. Grén. A simple algorithm for the estimation of the density of aqueous solutions containing two or more different salts. *Comput. Chem. Eng.*, 13(4/5):419–424, 1989.
- [41] Constantine Tsonopoulos. An empirical correlation of second virial coefficients. *AIChE J.*, 20(2):263–272, Mar. 1974.

### Electrolyte Models (ModEle)

Excess Gibbs energy models for electrolytes.

**Related keywords:** (activity & coefficients | excess & gibbs & energy | thermodynamic & properties) & electrolytes & (aqueous | mixed & solvent).

#### Bibliography.

- [1] K. Aasberg-Petersen, E. Stenby, and A. Fredenslund. Prediction of high-pressure gas solubilities in aqueous mixtures of electrolytes. *Ind. Eng. Chem. Res.*, 30(9):2080–2085, 1991.

- [2] Zareen Abbas, Magnus Gunnarsson, Elisabet Ahlberg, and Sture Nordholm. Corrected Debye–Hückel theory of salt solutions: Size asymmetry and effective diameters. *J. Phys. Chem. B*, 106:1403–1420, 2002.
- [3] V. Abovsky, Y. Liu, and S. Waranasiri. Representation of nonideality in concentrated electrolyte solutions using the electrolyte NRTL model with concentration-dependent parameters. *Fluid Phase Equilib.*, 150–151:277–286, 1998.
- [4] M. S. Ananth and Sunder Ramachandran. Self-consistent local composition model of electrolyte solutions. *AIChE J.*, 36(3):370–386, Mar. 1990.
- [5] A. Anderko, P. Wang, and M. Rafal. Electrolyte solutions: From thermodynamic and transport property models to the simulation of industrial processes. *Fluid Phase Equilib.*, 194–197:123–142, 2002.
- [6] Archer, Donald G. Effect of revisions of debye–hückel limiting law coefficients on the thermodynamic parameters for strong-electrolyte solutions. *J. Chem. Eng. Data*, 35(3):340–344, 1990.
- [7] G. Atkinson, A. Kumar, and B. L. Atkinson. Modeling the PVT properties of concentrated electrolytes in water. *AIChE J.*, 32(9):1561–1566, Sep. 1986.
- [8] Eftymios Balomenos, Dimitrios Panias, and Ioannis Paspaliaris. Modeling chemical equilibrium of electrolyte solutions. *Miner. Process. Extr. Metall. Rev.*, 27(1):1–60, 2006.
- [9] Francois-Xavier Ball, Henri Planche, Walter Fürst, and Henri Renon. Representation of deviation from ideality in concentrated aqueous solutions of electrolytes using a mean spherical approximation molecular model. *AIChE J.*, 31(8):1233–1240, aug 1985.
- [10] M. Bernardis, G. Carvoli, and P. Delogu.  $\text{NH}_3$ – $\text{CO}_2$ – $\text{H}_2\text{O}$  VLE calculation using an extended UNIQUAC equation. *AIChE J.*, 35(2):314–317, feb 1989.
- [11] Didler Beutier and Henri Renon. Representation of  $\text{NH}_3$ – $\text{H}_2\text{S}$ – $\text{H}_2\text{O}$ ,  $\text{NH}_3$ – $\text{CO}_2$ – $\text{H}_2\text{O}$  and  $\text{NH}_3$ – $\text{SO}_2$ – $\text{H}_2\text{O}$  vapor–liquid equilibria. *Ind. Eng. Chem. Process Des. Dev.*, 17(3):220–230, 1978.
- [12] Niels Bjerrum. Die verdünnungswärme einer ionenlösung in der theorie von Debye und Hückel. zugleich ein betrag zur theorie der wärmeeffekte in einem dielektrikum. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 119(10):145–160, 1926.
- [13] W. R. Bousfield. The determination of the ionisation of an aqueous solution. *Trans. Faraday Soc.*, 15:47–73, 1919.
- [14] Daniel J. Bradley and Kenneth S. Pitzer. Thermodynamics of electrolytes. 12. Dielectric properties of water and Debye–Hückel parameters to  $350^\circ\text{C}$  and 1 kbar. *J. Phys. Chem.*, 83(12):1599–1603, 1979.
- [15] Leroy A. Bromley. Thermodynamic properties of strong electrolytes in aqueous solutions. *AIChE J.*, 19(2):313–320, Mar. 1973.
- [16] Márcio J. E. de M. Cardoso and John P. O’Connell. Activity coefficients in mixed solvent electrolyte solutions. *Fluid Phase Equilib.*, 33:315–326, 1987.
- [17] Chen, Chau-Chyun and Yuhua Song. Generalized electrolyte-NRTL model for mixed-solvent electrolyte systems. *AIChE J.*, 50(8):1928–1941, 2004.
- [18] Chau-Chyun Chen, Herbert I. Britt, Joseph F. Boston, and Lawrence B. Evans. Extension and application of the Pitzer equation for vapor–liquid equilibrium of aqueous electrolyte systems with molecular solutes. *AIChE J.*, 25(5):820–879, Sep. 1979.
- [19] Chau-Chyun Chen, Herbert I. Britt, Joseph F. Boston, and Lawrence B. Evans. Two new activity coefficient models for the vapor–liquid equilibrium of electrolyte systems. *ACS Symp. Ser.*, 133:61–89, 1980.
- [20] Chau-Chyun Chen, H. I. Britt, J. F. Boston, and L. B. Evans. Local composition model for excess Gibbs energy of electrolyte systems. Part I: Single solvent, single completely dissociated electrolyte systems. *AIChE J.*, 28(4):588–596, 1982.
- [21] Chau-Chyun Chen and L. B. Evans. A local composition model for the excess Gibbs energy of aqueous electrolyte systems. *AIChE J.*, 32(3):444–454, 1986.
- [22] Simon L. Clegg and Kenneth S. Pitzer. Thermodynamics of multicomponent, miscible, ionic solutions: Generalized equations for symmetrical electrolytes. *J. Phys. Chem.*, 96:3513–3520, 1992.
- [23] Simon L. Clegg and Michael Whitfield. A chemical model of seawater including dissolved ammonia and the stoichiometric dissociation constant of ammonia in estuarine water and seawater from  $-2$  to  $40^\circ\text{C}$ . *Geochim. Cosmochim. Acta*, 59(12):2403–2421, 1995.
- [24] Simon L. Clegg, Peter Brimblecombe, and Anthony S. Wexler. Thermodynamic model of the system  $\text{H}^+$ – $\text{NH}_4^+$ – $\text{Na}^+$ – $\text{SO}_4^{2-}$ – $\text{NO}_3^-$ – $\text{Cl}^-$ – $\text{H}_2\text{O}$  at 298.15 K. *J. Phys. Chem.*, 102(12):2155–2171, 1998.
- [25] Simon L. Clegg, Peter Brimblecombe, and Anthony S. Wexler. Thermodynamic model of the system  $\text{H}^+$ – $\text{NH}_4^+$ – $\text{SO}_4^{2-}$ – $\text{NO}_3^-$ – $\text{H}_2\text{O}$  at tropospheric temperatures. *J. Phys. Chem.*, 102(12):2137–2154, 1998.
- [26] Horacio Roberto Corti. Prediction of activity coefficients in aqueous electrolyte mixtures using the mean spherical approximation. *J. Phys. Chem.*, 91(3):686–689, 1987.

- [27] Jose-Luis Cruz and Henri Renon. A new thermodynamic representation of binary electrolyte solutions nonideality in the whole range of concentrations. *AIChE J.*, 24(5):817–830, Sep. 1978.
- [28] Ioannis G. Economou, Cor J. Peters, and Jakob de Swaan Arons. Water-salt phase equilibria at elevated temperatures and pressures: Model development and mixture predictions. *J. Phys. Chem.*, 99:6182–6193, 1995.
- [29] T. J. Edwards, Gerd Maurer, John Newman, and J. M. Prausnitz. Vapor–liquid equilibria in multicomponent aqueous solutions of volatile weak electrolytes. *AIChE J.*, 24(6):966–976, 1978.
- [30] Andrew R. Felmy and John H. Weare. The prediction of borate mineral equilibria in natural waters: Application to Searles Lake, California. *Geochim. Cosmochim. Acta*, 50:2771–2783, 1986.
- [31] H. L. Friedman. Lewis–Randall to McMillian–Mayer conversion for the thermodynamic excess function of solutions. Part I. Partial free energy coefficients. *J. Solution Chem.*, 1(5):387–412, 1972.
- [32] H. L. Friedman. Lewis–Randall to McMillian–Mayer conversion for the thermodynamic excess function of solutions. Part II. Excess energy and volume. *J. Solution Chem.*, 1(5):413–417, 1972.
- [33] H. L. Friedman. Lewis–Randall to McMillian–Mayer conversion for the thermodynamic excess function of solutions. Part III. Common-ion mixtures of two electrolytes. *J. Solution Chem.*, 1(5):419–431, 1972.
- [34] Raymond M. Fuoss and Charles A. Kraus. On electrolytes in media of low dielectric constant. *J. Chem. Phys.*, 2(7):386–389, jul 1934.
- [35] Philipp Gross and Otto Halpern. On electrolytes in media of small dielectric constant. *J. Chem. Phys.*, 2(4):188–192, apr 1934.
- [36] Eduard Hála. Vapour–liquid equilibria in systems of electrolytic components. *I. Chem. E. Symp. Ser.*, 32(3):8–16, 1969.
- [37] A. H. Harvey, T. W. Copeman, and J. M. Prausnitz. Explicit approximation to the mean spherical approximation for electrolyte systems with unequal ion sizes. *J. Phys. Chem.*, 92:6432–6436, 1988.
- [38] T. Haug-Warberg and S. Skogestad. Prediction of VLE behavior in concentrated electrolyte solutions. *Fluid Phase Equilib.*, 13:341–350, 1983.
- [39] C. A. Haynes and J. Newman. On converting from McMillian–Mayer framework I. Single-solvent systems. *Fluid Phase Equilib.*, 145:255–268, 1998.
- [40] Erik Hessen. The electrolyte NRTL model. Department of Chemical Engineering, NTNU, Norway, sep 2008.
- [41] Raji Heyrovská. Physical electrochemistry of strong electrolytes based on partial dissociation and hydration. Quantitative interpretation of the thermodynamic properties of NaCl(aq) from 'zero to saturation'. *J. Electrochem. Soc.*, 143:1789–1793, Jun 1996.
- [42] Alan A. Humffray. Extension of Chen's 1982 theory to correlate densities of aqueous electrolytes. *AIChE J.*, 35(2):293–299, Feb. 1989.
- [43] Lars-Johan Jansson and Ian A. Furzer. A comparison of thermodynamic models for VLE data in electrolyte systems. *AIChE J.*, 35(6):1044–1048, Jun. 1989.
- [44] A. M. Kalinkin. Calculation of phase equilibria in the NaNO<sub>3</sub>–KNO<sub>3</sub>–HNO<sub>3</sub>–H<sub>2</sub>O system at 25 degrees C over the range from 0 to 20 m HNO<sub>3</sub>. *Russian J. Appl. Chem.*, 69(5):664–668, May 1996. Translated from Zhurnal Prikladnoi Khimii, 69(5), 1996, pp. 743–748.
- [45] Álvaro Pérez-Salado Kamps. Model for the Gibbs excess energy of mixed-solvent (chemical-reacting and gas-containing) electrolyte system. *Ind. Eng. Chem. Res.*, 44:201–225, 2005.
- [46] R. E. Kidder and H. E. deWitt. Application of a modified Debye–Hückel theory to fully ionized gases. *J. Nucl. Energy, Part C*, 2:218–223, 1961.
- [47] Hee-taik Kim and William J. Frederick, Jr. Evaluation of Pitzer ion interaction parameters of aqueous mixed electrolyte solutions at 25°C. 2. Ternary mixing parameters. *J. Chem. Eng. Data*, 33(3):278–283, 1988.
- [48] Roland Kjellander. Modified Debye–Hückel approximation with effective charges: An application of dressed ion theory for electrolyte solutions. *J. Phys. Chem.*, 99:10392–10407, 1995.
- [49] A. R. Kolker. Thermodynamic modelling of concentrated aqueous electrolyte and non-aqueous systems. *Fluid Phase Equilib.*, 69:155–169, 1991.
- [50] Alexander Kolker and Juan J. de Pablo. Thermodynamic modeling of concentrated aqueous electrolyte and nonelectrolyte solutions. *AIChE J.*, 41(6):1563–1571, 1995.
- [51] Alexander Kolker and Juan J. de Pablo. Thermodynamic modeling of concentrated multicomponent aqueous electrolyte and non-electrolyte solutions. *Chem. Eng. Sci.*, 50(12):1953–1959, 1995.
- [52] Kazuo Kondo and Charles A. Eckert. Nonideality of single and mixed electrolyte solutions up to moderately high concentrations: Theory based on Debye–Hückel radial distribution function. *Ind. Eng. Chem. Fundam.*, 22:283–292, 1983.
- [53] Jadwiga Krop. New approach to simplify the equation for the excess Gibbs free energy of aqueous solutions of electrolytes applied to the modelling of the NH<sub>3</sub>–CO<sub>2</sub>–H<sub>2</sub>O vapour and liquid equilibria. *Fluid Phase Equilib.*, 163:209–229, 1999.

- [54] Anil Kumar and V. S. Patwardhan. Aqueous solutions of single electrolytes: Thermodynamic properties at high temperature and concentration. *Chem. Eng. Sci.*, 47(15/16):4039–4047, 1992.
- [55] L. L. Lee. Thermodynamic consistency and reference scale conversion in multisolvent electrolyte solutions. *J. Mol. Liq.*, 87:129–147, 2000.
- [56] M. Conceição P. de Lima and Kenneth S. Pitzer. Thermodynamics of saturated aqueous solutions including mixtures of NaCl, KCl and CsCl. *J. Solution Chem.*, 12(3):171–185, 1983.
- [57] Yunda Liu, Allan H. Harvey, and John M. Prausnitz. Thermodynamics of concentrated electrolyte solutions. *Chem. Eng. Commun.*, 77:43–66, 1989.
- [58] Yunda Liu and Suphat Watanasiri. Representation of liquid–liquid equilibrium of mixed-solvent electrolyte systems using the extended electrolyte NRTL model. *Fluid Phase Equilib.*, 116:193–200, 1996.
- [59] Joseph R. Loehe and Marc D. Donohue. Recent advances in modelling thermodynamic properties of aqueous strong electrolyte systems. *AIChE J.*, 43(1):180–194, Jan. 1997.
- [60] Xiao-hua Lu and G. Maurer. Model for describing activity coefficients in mixed electrolyte aqueous solutions. *AIChE J.*, 39(9):1527–1538, 1993.
- [61] Xiaohua Lu, Luzheng Zhang, Yanru Wang, and Jun Shi. Simultaneous prediction of activity coefficients and enthalpy for aqueous electrolyte solutions at high temperatures. *Fluid Phase Equilib.*, 116:201–208, 1996.
- [62] Eugenia A. Macedo, Per Skovborg, and Peter Rasmussen. Calculation of phase equilibria for solutions of strong electrolytes in solvent-water mixtures. *Chem. Eng. Sci.*, 45(4):875–882, 1990.
- [63] H. P. Meissner and C. L. Kusik. Activity coefficients of strong electrolytes in multicomponent aqueous solutions. *AIChE J.*, 18(2):294–298, Mar. 1972.
- [64] Bill Mock, L. B. Evans, and Chau-Chyun Chen. Thermodynamic representation of phase equilibria of mixed-solvent electrolyte systems. *AIChE J.*, 32(10):1655–1664, 1986.
- [65] J. A. Myers and S. I. Sandler. An equation of state for electrolyte solutions covering wide ranges of temperature, pressure, and composition. *Ind. Eng. Chem. Res.*, 41:3282–3297, 2002.
- [66] Jason A. Myers, Stanley I. Sandler, and Robert H. Wood. An equation of state for electrolyte solutions covering wide ranges of temperature, pressure, and composition. *Ind. Eng. Chem. Res.*, 41(13):3282–3297, 2002.
- [67] Kenneth E. Newman. A Kirkwood–Buff theoretical approach to Debye–Hückel theory. *J. Chem. Soc., Faraday Trans. 1*, 85(3):485–492, 1989.
- [68] O’Connell, J. P. and Anthony E. DeGance. Thermodynamic properties of strong electrolyte solutions from correlation functions. *J. Solution Chem.*, 4(9):763–778, Sep 1975.
- [69] Bernhard A. Pailthorpe, D. John Mitchell, and Barry W. Ninham. Ion-solvent interactions and the activity coefficients of real electrolyte solutions. *J. Chem. Soc., Faraday Trans. 2*, 80(2):115–139, 1984.
- [70] N. Papaiconomou, J.-P. Simonin, O. Bernard, and W. Kunz. MSA-NRTL model for the description of the thermodynamic properties of electrolyte solutions. *J. Chem. Soc., Faraday Trans.*, 4:4435–4443, 2002.
- [71] N. Papaiconomou, J.-P. Simonin, O. Bernard, and W. Kunz. Description of vapor–liquid equilibria for CO<sub>2</sub> in electrolyte solutions using the mean spherical approximation. *J. Phys. Chem. B*, 107:5948–5957, 2003.
- [72] Nicolas Papaiconomou. *Thermodynamic Modelling of Industrial Relevant Electrolyte Solutions*. PhD thesis, Naturwissenschaftlichen Fakultät IV Chemie & Pharmazie der Universität Regensburg, 2003.
- [73] N. Papaiconomou, J.-P. Simonin, O. Bernard, and W. Kunz. New approaches to the calculation of thermodynamic properties of electrolyte solutions. *J. Mol. Liq.*, 113:5–8, 2004.
- [74] J. P. Passarello and W. Fürst. Representation of the equilibrium properties of the H<sub>2</sub>O–HNO<sub>3</sub>–N<sub>2</sub>O<sub>5</sub> systems using the MSA electrolyte model. *Fluid Phase Equilib.*, 116:177–184, 1996.
- [75] V. S. Patwardhan and Anil Kumar. A unified approach for prediction of aqueous mixed-electrolyte solutions. Part I: Vapor pressure and heat of vaporization. *AIChE J.*, 32(9):1419–1428, sep 1986.
- [76] V. S. Patwardhan and Anil Kumar. A unified approach for prediction of aqueous mixed-electrolyte solutions. Part II: Volume, thermal, and other properties. *AIChE J.*, 32(9):1429–1438, sep 1986.
- [77] Kenneth S. Pitzer. Thermodynamics of electrolytes. I. Theoretical basis and general equations. *J. Phys. Chem.*, 77(2):268–277, 1973.
- [78] Kenneth S. Pitzer and Guillermo Mayorga. Thermodynamics of electrolytes. II. Activity and osmotic coefficients for strong electrolytes with one or both ions univalent. *J. Phys. Chem.*, 77(19):2300–2308, 1973.
- [79] Kenneth S. Pitzer and Guillermo Mayorga. Thermodynamics of electrolytes. III. Activity and osmotic coefficients for 2–2 electrolytes. *J. Solution Chem.*, 3(7):539–546, 1974.
- [80] Kenneth S. Pitzer and Janice J. Kim. Thermodynamics of electrolytes. IV. Activity and osmotic coefficients for mixed electrolytes. *J. Am. Chem. Soc.*, 96(18):5701–5707, Sep. 1974.

- [81] [Kenneth S. Pitzer](#). Thermodynamics of electrolytes. V. Effects of higher-order electrostatic terms. *J. Solution Chem.*, 4(3):249–265, 1975.
- [82] [Kenneth S. Pitzer](#). Electrolyte theory—Improvements since Debye and Hückel. *Acc. Chem. Res.*, 10:371–377, 1977.
- [83] [Kenneth S. Pitzer](#). Electrolytes. From dilute solutions to fused salts. *J. Am. Chem. Soc.*, 102(9):2902–2906, 1980.
- [84] [Kenneth S. Pitzer](#). Thermodynamics of aqueous electrolytes at various temperatures, pressures, and compositions. *ACS Monogr.*, pages 451–466, 1980.
- [85] [Kenneth S. Pitzer](#). The treatment of ionic solutions over the entire miscibility range. *Ber. Bunsen-Ges.*, 85:952–959, 1981.
- [86] [Kenneth S. Pitzer](#). Thermodynamics of unsymmetrical electrolyte mixtures. Enthalpy and heat capacity. *J. Phys. Chem.*, 87(13):2360–2364, 1983.
- [87] [Kenneth S. Pitzer](#). Ionic fluids. *J. Phys. Chem.*, 88(13):2689–2697, 1984.
- [88] [Kenneth S. Pitzer](#). Gilbert N. Lewis and the thermodynamics of strong electrolytes. *J. Chem. Educ.*, 61(2):104–107, 1984.
- [89] [Kenneth S. Pitzer](#). Thermodynamic properties of ionic fluids over wide ranges of temperature. *Pure Appl. Chem.*, 59(1):1–6, 1987.
- [90] [J. Rennotte](#) and [B. Kalitventzeff](#). Contribution to the thermodynamic study of electrolyte solutions. *Comput. Chem. Eng.*, 12(5):461–467, 1988.
- [91] [J. Rennotte](#), [J. H. Massillon](#), and [B. Kalitventzeff](#). A new model for the simulation of the behaviour of electrolytic aqueous solutions. Comparison with three well-known previous models. *Comput. Chem. Eng.*, 13(4/5):411–417, 1989.
- [92] [Bo Sander](#), [Aage Fredenslund](#), and [Peter Rasmussen](#). Calculation of vapour–liquid equilibria in mixed Solvent/Salt systems using an extended UNIQUAC equation. *Chem. Eng. Sci.*, 41(5):1171–1183, 1986.
- [93] [Wolfram Schröer](#) and [Hermann Weingärtner](#). Structure and criticality of ionic fluids. *Pure Appl. Chem.*, 76(1):19–27, 2004.
- [94] [Leonard F. Silvester](#) and [Kenneth S. Pitzer](#). Thermodynamics of electrolytes. 8. High-temperature properties, including enthalpy and heat capacity, with application to sodium chloride. *J. Phys. Chem.*, 81(19):1822–1828, 1977.
- [95] [S. M. Sterner](#), [A. R. Felmy](#), and [K. S. Pitzer](#). Correlation of thermodynamic data for aqueous electrolyte solutions to very high ionic strength using INSIGHT: Vapor saturated water activity in the system  $\text{CaCl}_2 - \text{H}_2\text{O}$  to 250°C and solid saturation, Jun. 1997.
- [96] [R. H. Stokes](#) and [R. A. Robinson](#). Ionic hydration and activity in electrolyte solutions. *J. Am. Chem. Soc.*, 70:1870–1878, May 1948.
- [97] [S. Teitler](#) and [N. Ginsburg](#). Debye–Hückel theory and the concentration dependent dielectric constant. *J. Chem. Phys.*, 25(4):783–784, 1956.
- [98] [Kaj Thomsen](#), [Peter Rasmussen](#), and [Rafiqul Gani](#). Correlation and prediction of thermal properties and phase behaviour for a class of aqueous electrolyte systems. *Chem. Eng. Sci.*, 51(14):3675–3683, 1996.
- [99] [R. Triolo](#), [J. R. Grigera](#), and [L. Blum](#). Simple electrolytes in the mean spherical approximation. *J. Phys. Chem.*, 80(17):1858–1861, 1976.
- [100] [Peiming Wang](#), [Andrzej Anderko](#), and [Robert D. Young](#). A speciation-based model for mixed-solvent electrolyte systems. *Fluid Phase Equilib.*, 203(1–2):141–176, 2002.
- [101] [Meng Wang](#), [Maximilian B. Gorenssek](#), and [Chau-Chyun Chen](#). Thermodynamic representation of aqueous sodium nitrate and nitric acid solution with electrolyte NRTL model. *Fluid Phase Equilib.*, 407:105–116, 2016.
- [102] [Meng Wang](#), [Harnoor Kaur](#), and [Chau-Chyun Chen](#). Thermodynamic modeling of  $\text{HNO}_3 - \text{H}_2\text{SO}_4 - \text{H}_2\text{O}$  ternary system with symmetric electrolyte NRTL model. *AIChE J.*, 63(7):3110–3117, 2017.
- [103] [Hermann Weingärtner](#). Corresponding states for electrolyte solutions. *Pure Appl. Chem.*, 73(11):1733–1748, 2001.
- [104] [Dongqing Wei](#) and [Lesser Blum](#). Internal energy in the mean spherical approximation as compared to Debye–Hückel theory. *J. Phys. Chem.*, 91:4342–4343, 1987.
- [105] [Rong-Song Wu](#) and [L. L. Lee](#). Vapor–liquid equilibria of mixed-solvent electrolyte solutions: Ion-size effects based on the MSA theory. *Fluid Phase Equilib.*, 78:1–24, 1992.
- [106] [You-Xiang Zuo](#) and [Tian-Min Guo](#). An EOS for high-pressure aqueous electrolyte systems—Predicting the solubility of natural gas in formation water. *SPE Form. Eval.*, 7(2):181–184, Jun 1992.

## Equations of State (ModEOS)

Equation of state models in Helmholtz (fluids mostly) and Gibbs energy (solids mostly) representations, i.e.  $p = p(T, V, N)$  and  $V = V(T, p, N)$ .

**Related keywords:** pvt | eos | fundamental & (helmholtz & energy | gibbs & energy) | equation & of & state | surface.

### Bibliography.

- [1] Tarek H. Ahmed. Comparative study of eight equations of state for predicting hydrocarbon volumetric phase behavior. *SPE Reservoir Eng.*, 3(1):337–348, Feb 1988.
- [2] P. C. Albright, T. J. Edwards, Z. Y. Chen, and J. V. Sengers. A scaled fundamental equation for the thermodynamic properties of carbon dioxide in the critical region. *J. Chem. Phys.*, 87(3):1717–1725, Aug. 1987.
- [3] Orson L. Anderson. Equation for thermal expansivity in planetary interiors. *J. Geophys. Res.*, 72(14):3661–3668, Jul. 1967.
- [4] Selby Angus. Guide for the preparation of thermodynamic tables and correlations of the fluid state. *CODATA Bull.*, 51:1–43, Dec 1983. IUPAC Thermodynamic Tables Project Centre, Department of Chemical Engineering and Chemical Technology, Imperial College of Science and Technology, London.
- [5] E. M. Apfelbaum and V. S. Vorob'ev. Correspondence between thermodynamics of lattice models and real substances at the liquid–gas domain of the phase diagram. *J. Phys. Chem. B*, 114:9820–9826, 2010.
- [6] Gustavo A. Arteca, Francisco M. Fernández, and Eduardo A. Castro. A new method of analysis of critical-point singularities from power series expansions: Application to eigenvalue problems and virial series. *J. Chem. Phys.*, 85(11):6713–6719, Dec. 1986.
- [7] Valentín García Baonza, Mercedes Cáceres Alonso, and Javier Núñez Delgado. Universal behavior of compressed liquids. *J. Phys. Chem.*, 98(19):4955–4958, May 1994.
- [8] Von R. Becker. Eine Zustandsgleichung für Stickstoff bei großen Dichten. *Z. Phys.*, 4:393–409, 1921.
- [9] Felipe J. Blas and Lourdes F. Vega. Critical behavior and partial miscibility phenomena in binary mixtures of hydrocarbons by the statistical associating fluid theory. *J. Chem. Phys.*, 109(17):7405–7413, Nov. 1998.
- [10] Felipe J. Blas and Lourdes F. Vega. Prediction of binary and ternary diagrams using the statistical associating fluid theory (SAFT) equation state. *Ind. Eng. Chem. Res.*, 37(17):660–674, 1998.
- [11] A. Boushehri and E. A. Mason. Equation of state for compressed liquids and their mixtures from the cohesive energy density. *Int. J. Thermophys.*, 14(4):685–697, 1993.
- [12] Martin Braithwaite and Neil L. Allan. Thermodynamic representations for solid products in ideal detonation predictions. In *Proceedings Twelfth International Detonation Symposium, Aug 11–16 2002, San Diego, Ca*, pages 601–610. Office of Naval Research (ONR 333-05-02), 2002.
- [13] W. Byers Brown, Z. Feng, and M. Braithwaite. Williamsburg equation of state for modelling non-ideal detonation. *Journal de Physique IV*, 5(C4):209–214, May 1995. Colloque C4, supplément au Journal de Physique III.
- [14] P. Colonna and Paolo Silva. Dense gas thermodynamic properties of single and multicomponent fluids for fluid dynamics simulations. *J. Flu. Eng.*, 125:414–427, 2003.
- [15] J. L. Daridon, H. Saint-Guirons, B. LaGourette, P. Xans, and C. Leibovici. A generalized process for phase equilibrium calculation with cubic equations of state. *Int. J. Thermophys.*, 14(4):1101–1108, 1993.
- [16] N. Dass and M. Kumari. Derivation of some equations of state for solids. A new approach. *Phys. Stat. Sol. (b)*, 127(103):103–108, 1985.
- [17] Shalom Eliezer, Ajoy Ghatak, and Heinrich Hora. *Fundamentals of Equations of State*. World Scientific Publishing, Singapore, 2002.
- [18] J. Richard Elliot and Thomas E. Daubert. Revised procedures for phase equilibrium calculations with the Soave equation of state. *Ind. Eng. Chem. Process Des. Dev.*, 24(3):743–748, 1985.
- [19] Laurence E. Fried and W. Michael Howard. Explicit Gibbs free energy equation of state applied to the carbon phase diagram. *Phys. Rev. B: Solid State*, 61(13):8734–8743, apr 2000.
- [20] Kazuhiro Fuchizaki. Murnaghan's equation of state revisited. *Journal of the Physical Society of Japan*, 75(3):8734–8743, mar 2006.
- [21] L. Gerward. The bulk modulus and its pressure derivative for 18 metals. *J. Phys. Chem. Solids*, 46(8):925–927, 1985.
- [22] M. Góral. Cubic equation of state for calculation of phase equilibria in association systems. *Fluid Phase Equilib.*, 118:27–59, 1996.

- [23] R. Gosset, G. Heyen, and B. Kalitventzeff. An efficient algorithm to solve cubic equations of state. *Fluid Phase Equilib.*, 25:51–64, 1986.
- [24] Michael S. Graboski and Thomas E. Daubert. A modified Soave equation of state for phase equilibrium calculations. 1. Hydrocarbon systems. *Ind. Eng. Chem. Process Des. Dev.*, 17(4):443–448, 1978.
- [25] Michael S. Graboski and Thomas E. Daubert. A modified Soave equation of state for phase equilibrium calculations. 2. Systems containing CO<sub>2</sub>, H<sub>2</sub>S, N<sub>2</sub>, and CO. *Ind. Eng. Chem. Process Des. Dev.*, 17(4):448–454, 1978.
- [26] Joachim Gross and Gabriele Sadowski. Perturbed-chain SAFT: An equation of state based on a perturbation theory for chain molecules. *Ind. Eng. Chem. Res.*, 40:1244–1260, 2001.
- [27] A. Fernández Guillermet. Thermodynamic properties of the generalized Murnaghan equation of state of solids. *Int. J. Thermophys.*, 16(4):1009–1026, 1995.
- [28] Truls Gundersen. Numerical aspects of the implementation of cubic equations of state in flash calculation routines. *Comput. Chem. Eng.*, 6(3):245–255, 1982.
- [29] Tian-Min Guo, Lian-Gui Du, K. S. Pedersen, and Aage Fredenslund. Application of the Du–Guo and SRK equations of state to predict the phase behavior of Chinese reservoir fluids. *SPE Reservoir Eng.*, pages 379–388, Aug 1991.
- [30] Kenneth R. Hall, Gustavo A. Iglesias-Silva, and G. Ali Mansoori. Quadratic mixing rules for equations of state. Origins and relationships to the virial expansion. *Fluid Phase Equilib.*, 91:67–76, 1993.
- [31] Harvey, A. H. and P. H. Huang. First-principles calculation of the air–water second virial coefficient. *Int. J. Thermophys.*, 28(2):556–565, Apr 2007.
- [32] A. Harmens. Prediction of multicomponent low-temperature phase equilibria—A comparison of computation techniques. *Cryotech 73, Prod. Use Ind. Gases, Proc. Conf.*, 14:91–94, 1974.
- [33] A. Harmens. Phase equilibria from equation of state: Industrial application in cryogenics. In *Phase Equilibria and Fluid Properties in the Chemical Industry*, pages 379–388. European Federation of Chemical Engineers (EFCE), 1980. Proceedings Part II. Manuscripts of Invited Papers.
- [34] Allan H. Harvey and John M. Prausnitz. Thermodynamics of high-pressure aqueous systems containing gases and salts. *AIChE J.*, 35(4):635–643, Apr. 1989.
- [35] Holland, T. J. B. and R. Powell. An improved and extended internally consistent thermodynamic dataset for phases of petrological interest, involving a new equation of state for solids. *Journal of Metamorphic Geology*, 29(3):333–383, 2011.
- [36] W. B. Holzapfel. Equations of state for solids under strong compression. *High Pressure Research*, 16:81–126, 1998.
- [37] Reed A. Howald. Use of the Murnaghan–Hildebrand equation of state in teaching thermodynamics. *Chem. Educ.*, 3(2):1–18, 1998.
- [38] Marie-Jose Huron, Guy-Noel Dufour, and Jean Vidal. Vapour–liquid equilibrium and critical locus curve calculations with the Soave equation for hydrocarbon systems with carbon dioxide and hydrogen sulphide. *Fluid Phase Equilib.*, 1:247–265, 1977/78.
- [39] G. Ihm, Yuhua Song, and E. A. Mason. A new strong principle of corresponding states for nonpolar fluids. *J. Chem. Phys.*, 94(5):3839–3848, mar 1991.
- [40] M. Jaeschke, A. Benito, A. Fredheim, J.-M. Henault, M. Sangalli, P. V. Wesenbeeck, R. Klimeck, O. Kunz, R. Span, and W. Wagner. GERG project: Wide-range reference equation of state for natural gases. Technical Report TP18, GERG, VDI Verlag GmbH, Düsseldorf, Germany, 2003.
- [41] Jean-Noël Jaubert and Fabrice Mutelet. VLE predictions with the Peng–Robinson equation of state and temperature dependent  $k_{ij}$  calculated through a group contribution method. *Fluid Phase Equilib.*, 224:285–304, 2004.
- [42] Joseph Joffe. Vapor–liquid equilibria and densities with the Martin equation of state. *Ind. Eng. Chem. Process Des. Dev.*, 20(1):168–172, 1981.
- [43] James W. Johnson, Erich H. Oelkers, and Harold C. Helgeson. SUPCRT92: A software package for calculating the standard molal thermodynamic properties of minerals, gases, aqueous species, and reactions from 1 to 5000 bar and 0 to 1000°C. *Comput. Geosci.*, 18(7):899–947, 1992.
- [44] Stevan Jovanović and Ratimir Paunović. Generating appropriate density values from a cubic state equation to avoid false unit K values. Application to distillation problems. *Ind. Eng. Chem. Process Des. Dev.*, 23(4):801–805, 1984.
- [45] Thomas Kraska and Ulrich K. Deiters. Systematic investigation of the phase behavior in binary fluid mixtures. II. Calculations based on the Carnahan–Starling–Redlich–Kwong equation of state. *J. Chem. Phys.*, 96(1):539–547, Jan 1992.
- [46] K. Hemanth Kumar and Kenneth E. Starling. Comments on: “Cubic equations of state—Which?”. *Ind. Eng. Chem. Fundam.*, 19(1):128–129, 1980.
- [47] Munish Kumar. High pressure equation of state for solids. *Physica B: Condensed Matter*, 212(4):391–394, 1995.

- [48] O. Kunz, R. Klimeck, W. Wagner, and M. Jaeschke. The GERG-2004 wide-range equation of state for natural gases and other mixtures. Technical Report TM15, GERG, VDI Verlag GmbH, Düsseldorf, Germany, 2007.
- [49] E. Lang and H. Wenzel. Extension of a cubic equation of state to solids. *Fluid Phase Equilib.*, 51:101–117, 1989.
- [50] Akami S. Lawal. A consistent rule for selecting roots in cubic equations of state. *Ind. Eng. Chem. Res.*, 26(4):857–859, 1987.
- [51] C. F. Leibovici. Variant and invariant properties from cubic equations of state. *Fluid Phase Equilib.*, 84:1–8, 1993.
- [52] Eric W. Lemmon and Marcia L. Huber. Thermodynamic properties of *n*-dodecane. *Energy & Fuels*, 18(4):960–967, 2004.
- [53] Don S. Lemons and Carl M. Lund. Thermodynamics of high temperature, mie–grüneisen solids. *Am. J. Phys.*, 67(12):1105–1108, Dec 1999.
- [54] Hong Lin and Yuan-Yuan Duan. Empirical correction to the Peng–Robinson equation of state for the saturated region. *Fluid Phase Equilib.*, 233:194–203, 2005.
- [55] Hong Lin, Yuan-Yuan Duan, Tao Zhang, and Zhi-Min Huang. Volumetric property improvement for the Soave–Redlich–Kwong equation of state. *Ind. Eng. Chem. Res.*, 45:1829–1839, 2006.
- [56] Yen-liang Lin, Paul R. Bienkowski, Vinod M. Shah, and Hank D. Cochran. Extension of a generalized quartic equation of state to pure polar fluids. *AIChE J.*, 42(2):562–579, 1996.
- [57] L. Maftoon-Azad and A. Boushehri. An analytical equation of state for some saturated liquid metals. *Int. J. Thermophys.*, 25(3):726–731, 2004.
- [58] David E. Mainwaring, Richard J. Sadus, and Colin L. Young. Deiters’ equation of state and critical phenomena. *Chem. Eng. Sci.*, 43(3):459–466, Jun. 1988.
- [59] Joseph J. Martin. Cubic equations of state—Which? *Ind. Eng. Chem. Fundam.*, 18(2):81–97, 1979.
- [60] Joseph J. Martin. Comments on: “Cubic equations of state—Which?”. *Ind. Eng. Chem. Fundam.*, 19(1):130–131, 1980.
- [61] Abraham Marmur. Mathematical properties of equations of state. *Chem. Eng. Sci.*, 40(10):1881–1884, 1985.
- [62] Paul M. Mathias. A versatile phase equilibrium equation of state. *Ind. Eng. Chem. Process Des. Dev.*, 22(3):385–391, 1983.
- [63] Paul M. Mathias and Thomas W. Copeman. Extension of the Peng–Robinson equation of state to complex mixtures: Evaluation of the various forms of the local composition concept. *Fluid Phase Equilib.*, 13:91–108, 1983.
- [64] P. M. Mathias, J. F. Boston, and S. Watanasiri. Effective utilization of equations of state for thermodynamic properties in process simulation. *AIChE J.*, 30(2):182–186, Mar. 1984.
- [65] P. M. Mathias and M. S. Benson. Computational aspects of equations of state: Fact and fiction. *AIChE J.*, 32(12):2087–2090, Dec. 1986.
- [66] N. Mehdipour and A. Boushehri. An analytical equation of state for mercury. *Int. J. Thermophys.*, 18(5):1329–1334, 1997.
- [67] Georges A. Melhem, Riju Saini, and Bernhard M. Goodwin. A modified Peng–Robinson equation of state. *Fluid Phase Equilib.*, 47:189–237, 1989.
- [68] Michael L. Michelsen. Comments on “design of a combined mixing rule for the prediction of vapor–liquid equilibria using neural networks”. *Ind. Eng. Chem. Res.*, 39:240, 2000.
- [69] Michael L. Michelsen and Eric M. Hendriks. Physical properties from association models. *Fluid Phase Equilib.*, 180:165–174, 2001.
- [70] Michael L. Michelsen. A method for incorporating excess Gibbs energy models in equation of state. *Fluid Phase Equilib.*, 60:47–58, 1990.
- [71] Michael L. Michelsen. Matching equation of state mixing rules to activity coefficient model expressions. *Fluid Phase Equilib.*, 121:15–26, 1996.
- [72] Martin B. Mills, M. John Wills, and Vasant L. Bhirud. The calculation of density by the BWRS equation of state in process simulation contexts. *AIChE J.*, 26(6):902–910, Nov. 1980.
- [73] George D. Miron, Thomas Wagner, Dmitrii A. Kulik, and Barbara Lothenbach. An internally consistent thermodynamic dataset for aqueous species in the system Ca–Mg–Na–K–Al–Si–O–H–C–Cl to 800°C and 5 kbar. *Am. J. Sci.*, 317:755–806, Sep. 2017.
- [74] K. Morita, V. Sobolev, and M. Flad. Critical parameters and equation of state for heavy liquid metals. *Journal of Nuclear Materials*, 362:227–234, 2007.
- [75] J. M. Moysan, M. J. Huron, H. Paradowski, and J. Vidal. Prediction of the solubility of hydrogen in hydrocarbon solvents through cubic equations of state. *Chem. Eng. Sci.*, 38(7):1085–1092, 1983.
- [76] F. D. Murnaghan. The compressibility of media under extreme pressure. *Proc. Natl. Acad. Sci. U. S. A.*, 30:244–247, 1944.

- [77] Hideo Nishiumi, Tsutomu Arai, and Katsuhiko Takeuchi. Generalization of the binary interaction parameter of the Peng–Robinson equation of state by component family. *Fluid Phase Equilib.*, 42:43–62, 1988.
- [78] Hasan Orbey and Stanley I. Sandler. On the combination of equation of state and excess free energy models. *Fluid Phase Equilib.*, 111:53–70, 1995.
- [79] Hasan Orbey and Stanley I. Sandler. A comparison of various cubic equation of state mixing rules for the simultaneous description of excess enthalpies and vapor–liquid equilibria. *Fluid Phase Equilib.*, 121:67–83, 1996.
- [80] V. L. Pan'kov and W. Ullmann. A comparative method for various approaches to the isothermal equation of state. *Pageoph*, 117:1001–1010, 1979.
- [81] Navin C. Patel and Aryn S. Teja. A new cubic equation of state for fluids and fluid mixtures. *Chem. Eng. Sci.*, 37(3):463–473, 1982.
- [82] Ding-Yu Peng and Donald B. Robinson. A new two-constant equation of state. *Ind. Eng. Chem. Fundam.*, 15(1):59–64, 1976.
- [83] André Péneloux, Evelyne Rauzy, and Richard Fréze. A consistent correction for Redlich–Kwong–Soave volumes. *Fluid Phase Equilib.*, 8:7–23, 1982.
- [84] Oliver Pfohl, Tim Giese, Ralf Dohrn, and Gerd Brunner. 1. Comparison of 12 equations of state with respect to gas-extraction processes: Reproduction of pure-component properties when enforcing the correct critical temperature and pressure. *Ind. Eng. Chem. Res.*, 37(8):2957–2965, 1998.
- [85] Andrey V. Plyasunov. Thermodynamics of  $\text{Si}(\text{OH})_4$  in the vapor phase of water: Henry's and vapor–liquid distribution constants, fugacity and cross virial coefficients. *Geochim. Cosmochim. Acta*, 77:215–231, 2012.
- [86] Thomas G. Plymate, James H. Stout, and Mark E. Cavaleri. Pressure-volume-temperature behavior and heterogeneous equilibria of the non-quenchable body-centered tetragonal polymorph of metallic tin. *J. Phys. Chem. Solids*, 49(11):1339–1348, 1988.
- [87] Thomas G. Plymate and James H. Stout. A five-parameter temperature-corrected Murnaghan equation for  $P$ – $V$ – $T$  surfaces. *JGR, J. Geophys. Res.*, 94(B7):9477–9483, Jul. 1989.
- [88] Bruce E. Poling, Edward A. Grens, and John M. Prausnitz. Thermodynamic properties from a cubic equation of state: Avoiding trivial roots and spurious derivatives. *Ind. Eng. Chem. Process Des. Dev.*, 20(1):127–130, 1981.
- [89] P. Proust and J. H. Vera. The Stryjek–Vera modification of the Peng–Robinson equation of state. Parameters for other pure compounds of industrial interest. *Can. J. Chem. Eng.*, 67:170–173, 1989.
- [90] Willigert Raatschen, Allan H. Harvey, and John M. Prausnitz. Equation of state for solutions of electrolytes in mixed solvents. *Fluid Phase Equilib.*, 38:19–38, 1987.
- [91] Otto Redlich and J. N. S. Kwong. On the thermodynamics of solutions. V. An equation of state. Fugacities of gaseous solutions. *Chem. Rev. (Washington, D. C.)*, 44:233–244, 1949.
- [92] D. B. Robinson and D. Y. Peng. The characterization of the heptanes and heavier fractions for the GPA Peng–Robinson programs. Technical Report 28, University of Alberta, Edmonton, mar 1978. Project 756.
- [93] Paul H. Salim and Mark A. Trebble. A modified Trebble–Bishnoi equation of state: Thermodynamic consistency revisited. *Fluid Phase Equilib.*, 65:59–71, 1991.
- [94] Paul H. Salim and Mark A. Trebble. Modelling of solid phases in thermodynamic calculations via translation of a cubic equation of state at the triple point. *Fluid Phase Equilib.*, 93:75–99, 1994.
- [95] Stanley I. Sandler and Lawrence R. Dodd. Letter to the editor: On the problem of phase identification in a mixture. *Fluid Phase Equilib.*, 31:313–316, 1986.
- [96] J. Schwartzentruber, H. Renon, and S. Watanasiri. Development of a new cubic equation of state for phase equilibrium calculations. *Fluid Phase Equilib.*, 52:127–134, 1989.
- [97] J. Schwartzentruber and H. Renon. K-values for non-ideal systems: An easier way. *Chem. Eng. [Int. Ed.]*, 93(3):118–124, Mar. 1990.
- [98] Werner Schulze. A simple generalization of the binary temperature-dependent interaction parameters in the Soave–Redlich–Kwong equations of state for helium-mixtures. *Fluid Phase Equilib.*, 87:199–211, 1993.
- [99] N. Shamsundar and R. P. Reddy. An invariance property associated with  $p$ – $v$ – $t$  equations of state. *Ind. Eng. Chem. Res.*, 30:2172–2180, 1991.
- [100] Giorgio Soave. Equilibrium constants from a modified Redlich–Kwong equation of state. *Fluid Phase Equilib.*, 27:1197–1203, 1972.
- [101] G. Soave. Application of equations of state and the theory of group solutions to phase equilibrium prediction. *Fluid Phase Equilib.*, 87:23–35, 1993.
- [102] Nicolas von Solms, Michael L. Michelsen, and Georgios M. Kontogeorgis. Computational and physical performance of a modified PC-SAFT equation of state for highly asymmetric and associating mixtures. *Ind. Eng. Chem. Res.*, 42:1098–1105, 2003.

- [103] [Yuhua Song](#) and E. A. Mason. Statistical-mechanical theory of a new analytical equation of state. *J. Chem. Phys.*, 91(12):7840–7853, dec 1989.
- [104] [Yuhua Song](#) and E. A. Mason. Analytical equation of state for molecular fluids: Comparison with experimental data. *Phys. Rev. A: At., Mol., Opt. Phys.*, 42(8):4749–4755, oct 1990.
- [105] [Yuhua Song](#), B. Caswell, and E. A. Mason. Compressibility of liquids: Theoretical basis for a century of empiricism. *Int. J. Thermophys.*, 12(5):855–868, 1991.
- [106] [Yuhua Song](#), B. Caswell, and E. A. Mason. Equation of state of compressed liquids: Statistical-mechanical basis. *Fluid Phase Equilib.*, 88:25–34, 1993.
- [107] [R. Span](#) and W. Wagner. Equations of state for technical applications. I. Simultaneously optimized functional forms for nonpolar and polar fluids. *Int. J. Thermophys.*, 24(1):1–39, jan 2003.
- [108] [R. Span](#) and W. Wagner. Equations of state for technical applications. II. Results for nonpolar fluids. *Int. J. Thermophys.*, 24(1):41–109, jan 2003.
- [109] [R. Span](#) and W. Wagner. Equations of state for technical applications. III. Results for polar fluids. *Int. J. Thermophys.*, 24(1):111–162, jan 2003.
- [110] [R. Span](#), H.-J. Collmann, and W. Wagner. Simultaneous optimization as a method to establish generalized functional forms for empirical equations of state. *Int. J. Thermophys.*, 19(2):491–500, 1998.
- [111] [R. Stryjek](#) and J. H. Vera. PRSV: An improved Peng–Robinson equation of state for pure compounds and mixtures. *Can. J. Chem. Eng.*, 64:323–333, apr 1986.
- [112] [R. Stryjek](#) and J. H. Vera. PRSV: An improved Peng–Robinson equation of state with new mixing rules for strongly nonideal mixtures. *Can. J. Chem. Eng.*, 64:334–340, apr 1986.
- [113] [R. Stryjek](#) and J. H. Vera. PRSV2: A cubic equation of state for accurate vapor–liquid equilibria calculations. *Can. J. Chem. Eng.*, 64:820–826, oct 1986.
- [114] [F.-M. Tao](#) and E. A. Mason. Equation of state for mixtures of nonpolar fluids: Prediction from experimental constants of the components. *Int. J. Thermophys.*, 13(6):1053–1060, 1992.
- [115] [Mercedes Taravillo](#), Valentín Garcé Baonza, Javier Núñez, and Mercedes Cáceres. Application of a new equation of state for solids. *High Temp. - High Pressures*, 30:97–103, 1998.
- [116] [Constantine Tsonopoulos](#) and J. M. Prausnitz. Equations of state. A review for engineering applications. *Cryogenics*, 9(5):315–327, Oct. 1969.
- [117] [José O. Valderrama](#). The state of the cubic equations of state. *Ind. Eng. Chem. Res.*, 42(8):1603–1618, 2003.
- [118] [J. Vidal](#) and G. Bogdanić. Equations of state, mixing rules and phase equilibrium calculations. Part 2. Mixing rules and vapor–liquid equilibrium calculation. *Chem. Biochem. Eng. Q.*, 13(3):101–125, Jul. 1999.
- [119] [Visco, Jr., Donald P.](#), David A. Kofke, and Rajiv R. Singh. Thermal properties of hydrogen fluoride from EOS + Association model. *AIChE J.*, 43(9):2381–2384, Sep. 1997.
- [120] [Nikolaos Voros](#), Sofia Stamataki, and Dimitrios Tassios. Effect of translation on the prediction of saturated densities of binary mixtures with a Peng–Robinson-equation of state. *Fluid Phase Equilib.*, 96:51–63, 1994.
- [121] [Johannes D. Van der Waals](#). The equation of state for gases and liquids, Dec. 1910. Nobel Lecture.
- [122] [Uwe Walzer](#), Wolfgang Ullmann, and V. L. Pan'kov. Comparison of some equation-of-state theories by using experimental high-compression data. *Physics of the Earth and Planetary Interiors*, 18:1–12, 1979.
- [123] [P. T. Wedepohl](#). Accuracy of the Murnaghan equation compared with that of an exponential equation of state. *J. Phys. D: Appl. Phys.*, 10:43–48, 1977.
- [124] [H. Wenzel](#) and G. Schmidt. A modified van der Waals equation of state for the representation of phase equilibria between solids, liquids and gases. *Fluid Phase Equilib.*, 5:3–17, 1980.
- [125] [H. Wenzel](#), R. A. S. Moorwood, and M. Baumgärtner. Calculation of vapour–liquid equilibrium of associated systems by an equation of state. *Fluid Phase Equilib.*, 9:225–266, 1982.
- [126] [C. H. Whitson](#), F. da Silva, and I. Søreide. Simplified compositional formulation for modified black-oil simulators. *SPE*, pages 511–526, 1988. SPE 18315.
- [127] [R. A. Wilsak](#) and George Thodos. An equation of state: Its development from argon data and its application to other substances. *AIChE J.*, 31(5):729–739, May 1985.
- [128] [Gang Xu](#), Joan F. Brennecke, and Mark A. Stadtherr. Reliable computation of phase stability and equilibrium from the SAFT equation of state. *Ind. Eng. Chem. Res.*, 41(5):938–952, 2002.
- [129] [Zabaloy, Marcelo S.](#) and Juan H. Vera. Cubic equation of state for pure compound vapor pressures from the triple point to the critical point. *Ind. Eng. Chem. Res.*, 35(3):829–836, 1996.
- [130] [F. Zielke](#) and D. A. Lempe. Generalized calculation of phase equilibria by using cubic equations of state. *Fluid Phase Equilib.*, 141(1-2):63–85, 1997.

- [131] Kurt Zimmer, Yilun Zhang, Peng Lu, Yanyan Chen, Guanru Zhang, Mehmet Dalkilic, and Chen Zhu. SUPCRTBL: A revised and extended thermodynamic dataset and software package of SUPCRT92. *Comput. Geosci.*, 90:97–111, 2016.

### Parameter Regression (ModFit)

Parameter regression of physical models.

**Related keywords:** parameter & (fitting | estimation | regression) | maximum & likelihood | least & squares.

#### Bibliography.

- [1] M. M. Abbott and Hendrick C. van Ness. Vapor–liquid equilibrium: Part III. Data reduction with precise expressions for  $g^e$ . *AIChE J.*, 21(1):62–71, Jan. 1975.
- [2] M. M. Abbott, J. K. Floess, G. E. Walsh, Jr., and H. C. van Ness. Vapor–liquid equilibrium: Part IV. Reduction of  $p$ – $x$  data for ternary systems. *AIChE J.*, 21(1):72–76, Jan. 1975.
- [3] M. M. Abbott and H. C. van Ness. An extension of Barker’s method for reduction of VLE data. *Fluid Phase Equilib.*, 1:3–11, 1977.
- [4] Joachim Ahrendts and Hans Dieter Baehr. Die Anwendung nichtlinearer Regressionsverfahren bei der Aufstellung thermodynamischer Zustandsgleichungen. *Forsch. Ingenieurwes.*, 45(2):51–56, 1979.
- [5] T. F. Anderson, D. S. Abrams, and E. A. Grens II. Evaluation of parameters for nonlinear thermodynamic models. *AIChE J.*, 24(1):20–29, 1978.
- [6] I. Ashour and G. Aly. Effect of computation techniques for equation of state binary interaction parameters on the prediction of binary VLE data. *Comput. Chem. Eng.*, 20(1):79–91, 1996.
- [7] J. A. Barker. Determination of activity coefficients from total pressure measurements. *Aust. J. Chem.*, 6:207–210, 1953.
- [8] R. B. Berman, M. Engi, H. J. Greenwood, and T. H. Brown. Derivation of internally-consistent thermodynamic data by the technique of mathematical programming: A review with application to the system MgO–SiO<sub>2</sub>–H<sub>2</sub>O. *J. Petrol.*, 27(6):1331–1364, 1986.
- [9] K.-J. Böttcher. Parameter sensitivity of deterministic and stochastic search methods. *Lecture Notes in Econom. and Math. Syst.*, 458:249–278, 1998.
- [10] Stanley M. Byer, Richard E. Gibbs, and Hendrick C. van Ness. Vapor–liquid equilibrium: Part II. Correlations from  $p$ – $x$  data for 15 systems. *AIChE J.*, 19(2):245–251, Mar. 1973.
- [11] J. David Chase. The qualification of pure component physical property data. *Chem. Eng. Prog.*, 4:63–67, Apr. 1984.
- [12] S.-L. Chen, Y. Zuo, C. R. Kao, and Y. A. Chang. On the optimization of solution model parameter values of phases and the calculation of phase diagrams. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(1):47–56, 1993.
- [13] Vladimír Dohnal and Dana Fenclová. Verification of the accuracy of complete binary vapor–liquid equilibrium data. *Fluid Phase Equilib.*, 19:1–12, 1985.
- [14] T. A. Duever, S. E. Keeler, and P. M. Reilly. An application of the error-in-variables model—Parameter estimation from van Ness-type vapour–liquid equilibrium experiments. *Chem. Eng. Sci.*, 42(3):403–412, 1987.
- [15] David Garvin, Vivian B. Parker, Donald D. Wagman, and William H. Evans. A combined least sums and least squares approach to the solution of thermodynamic data networks. In *The Proceedings of the Fifth Biennial International CODATA Conference*, International CODATA Conference, pages 567–575, Boulder, Colorado, Jun/Jul 1976.
- [16] M. Góral. Error analysis in Barker’s method of vapor pressure isotherm data processing. *Z. Phys. Chem. (Leipzig)*, 258:1040–1044, 1977.
- [17] M. F. Guest, J. B. Pedley, and M. Horn. Analysis by computer of thermochemical data on boron compounds. *J. Chem. Thermodyn.*, 1:345–352, 1969.
- [18] Bo Jansson. Evaluation of parameters in thermochemical models using different types of experimental data simultaneously. Report series D, No 57, Division of Physical Metallurgy, Royal Institute of Technology, Stockholm, 1984.
- [19] S. Kemeny, J. Manczinger, S. Skjold-Jørgensen, and K. Toth. Reduction of thermodynamic data by means of the multiresponse maximum likelihood principle. *AIChE J.*, 28(1):20–30, 1982.
- [20] In-Won Kim, Michael J. Liebman, and Thomas F. Edgar. Robust error-in-variables estimation using nonlinear programming techniques. *AIChE J.*, 36(7):985–993, 1990.
- [21] G. Kolašnička and P. Oracz. Generalization of error analysis and problem of vapor non-ideality in Barker’s method of vapor pressure isotherm data processing. *Z. Phys. Chem. (Leipzig)*, 260:169–173, 1979.

- [22] Erich Königsberger and Heinz Gamsjäger. Analysis of phase diagrams employing Bayesian excess parameter estimation. *Monatsh. Chem.*, 121:119–127, 1990.
- [23] Erich Königsberger and Gunnar Eriksson. A new optimization routine for ChemSage. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 19(2):207–214, 1995.
- [24] H. L. Lukas, E. Th. Henig, and B. Zimmermann. Optimization of phase diagrams by a least squares method using simultaneously different types of data. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 1(3):225–236, 1977.
- [25] H. L. Lukas and S. G. Fries. Demonstration of the use of “BINGSS” with the Mg–Zn system as example. *J. Phase Equilib.*, 13(5):532–541, 1992.
- [26] Evelyne Neau and André Pénélox. Estimation of model parameters. Comparison of methods based on the maximum likelihood principle. *Fluid Phase Equilib.*, 6:1–19, 1981.
- [27] Hendrick C. van Ness, Stanley M. Byer, and Richard E. Gibbs. Vapor–liquid equilibrium: Part I. An appraisal of data reduction methods. *AIChE J.*, 19(2):238–244, Mar. 1973.
- [28] H. C. van Ness, Finn Pedersen, and Peter Rasmussen. Vapor–liquid equilibrium: Part V. Data reduction by maximum likelihood. *AIChE J.*, 24(6):1055–1063, nov 1978.
- [29] Vicki G. Niesen and Victor F. Yesavage. Application of a maximum likelihood method using implicit constraints to determine equation of state parameters from binary phase behavior data. *Fluid Phase Equilib.*, 50:249–266, 1989.
- [30] Hugo Patino-Leal and P. M. Reilly. Statistical estimation of parameters in vapor–liquid equilibrium. *AIChE J.*, 28(4):580–587, Jul. 1982.
- [31] André Pénélox, Richard Deyrieux, Évelyne Canals, and Évelyne Neau. The maximum likelihood test and the estimation of experimental inaccuracies. Application to the data reduction for liquid–vapor equilibrium. *J. Chim. Phys. Phys.-Chim. Biol.*, 73(7-8):706–716, 1976.
- [32] André Pénélox, Evelyne Neau, and Adolfo Gramajo. Variance analysis fifteen years ago and now. *Fluid Phase Equilib.*, 56:1–16, 1990.
- [33] Charles A. Plank, James D. Olson, Harold R. Null, Ol Muthu, and Bufford D. Smith. Reduction of total-pressure vapor–liquid equilibrium data. Common pitfalls encountered. *Fluid Phase Equilib.*, 6:39–59, 1981.
- [34] Henri Renon. Qualities of models for evaluation, representation and prediction of fluid phase equilibrium data. *Fluid Phase Equilib.*, 2:101–118, 1978.
- [35] Vladimír Rod and Vladislav Hančil. Iterative estimation of model parameters when measurements of all variables are subject to error. *Comput. Chem. Eng.*, 4:33–38, 1980.
- [36] R. G. Rubio, J. A. R. Renuncio, and M. Diaz Peña. Regression of vapor–liquid equilibrium data based on application of the maximum-likelihood principle. *Fluid Phase Equilib.*, 12:217–234, 1983.
- [37] Steen Skjold-Jørgensen. On statistical principles in reduction of thermodynamic data. *Fluid Phase Equilib.*, 14:273–288, 1983.
- [38] Dean L. Ulrichson and F. Dee Stevenson. Effects of experimental errors on thermodynamic consistency and on representation of vapor–liquid equilibrium data. *Ind. Eng. Chem. Fundam.*, 11(3):287–293, 1972.
- [39] Petr Voňka, Josef Novák, and Jaroslav Matouš. Application of the maximum likelihood method to the parameter evaluation in heterogeneous systems. *Collect. Czech. Chem. Commun.*, 54:2823–2839, 1989.

### Gas hydrates and clathrates (ModHyd)

Thermodynamic models for gas hydrates or so-called gas clathrates. Most of the models are derived from Van der Waals–Platteeuw theory.

**Related keywords:** gas & hydrate | clathrate.

#### Bibliography.

- [1] Dimitrios Anatassios Avlonitis. *Thermodynamics Of Gas Hydrate Equilibria*. PhD thesis, Heriot-Watt University, Feb 1992.
- [2] Dimitrios Avlonitis and Nikos Varotsis. Modelling gas hydrate thermodynamic behaviour: Theoretical basis and computational methods. *Fluid Phase Equilib.*, 123(1–2):107–130, 1996.
- [3] A. Djaavidnia, A. A. Izadpanah, M. V. Sefti, and F. Varaminian. The equilibrium data and thermodynamic modeling of hydrate formation for CO<sub>2</sub> and mixtures of CO<sub>2</sub> and CH<sub>4</sub> in the presence of methanol. *Pet. Sci. Technol.*, 31:2013–2021, 2013.
- [4] P. Englezos and P. R. Bishnoi. Prediction of gas hydrate formation conditions in aqueous electrolyte solutions. *AIChE J.*, 34(10):1718–1721, Oct. 1988.
- [5] Jorge F. Gabitto and Costas Tsouris. Physical properties of gas hydrates: A review. *J. Thermodynamics*, pages 1–12, 2010. ID 271291.

- [6] Jean-Michel Herri, Amina Bouchemoua-Benaissa, Matthias Kwaterski, Amara Fezoua, Yamina Ouabbas, and Ana Cameirão. Gas hydrate equilibria for CO<sub>2</sub>-N<sub>2</sub> and CO<sub>2</sub>-CH<sub>4</sub> gas mixtures—Experimental studies and thermodynamic modelling. *Fluid Phase Equilib.*, 301(2):171–190, 2011.
- [7] Ehsan Khosravani, Gholamreza Moradi, and Sami Sajjadifar. An accurate thermodynamic model to predict phase behavior of clathrate hydrates in the absence and presence of methanol based on the genetic algorithm. *J. Chem. Thermodyn.*, 57:286–294, 2013.
- [8] H. C. Kim, P. R. Bishnoi, R. A. Heidemann, and S. S. H. Rizvi. Kinetics of methane hydrate decomposition. *Chem. Eng. Sci.*, 42(7):1645–1653, 1987.
- [9] Jeffery B. Klauda and Stanley I. Sandler. Phase behavior of clathrate hydrates: A model for single and multiple gas component hydrates. *Chem. Eng. Sci.*, 58:27–41, 2003.
- [10] Lars Lundegaard and Jørgen Møllerup. The influence of gas phase fugacity and solubility on correlation of gas-hydrate formation pressure. *Fluid Phase Equilib.*, 70:199–213, 1991.
- [11] Amir H. Mohammadi and Dominique Richon. Thermodynamic model for predicting liquid water–hydrate equilibrium of the water–hydrocarbon system. *Ind. Eng. Chem. Res.*, 47:1346–1350, 2008.
- [12] Gholamreza Moradia and Ehsan Khosravani. Modeling of hydrate formation conditions for CH<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, C<sub>3</sub>H<sub>8</sub>, N<sub>2</sub>, CO<sub>2</sub> and their mixtures using the PRSV2 equation of state and obtaining the kihara potential parameters for these components. *Fluid Phase Equilib.*, 338:179–187, 2013.
- [13] Mohammad M. Shabani, Ole J. Nydal, and Roar Larsen. A proposed thermodynamic model for gas hydrate equilibrium in electrolyte solutions. *Heat Transfer Eng.*, 32(2):168–175, 2011.
- [14] Ioannis N. Tsimpanogiannis, Nikolaos I. Papadimitriou, and Athanassios K. Stubos. On the limitation of the Van der Waals–Platteeuw-based thermodynamic models for hydrates with multiple occupancy of cavities. *Mol. Phys.*, 110(11–12):1213–1221, 2012.
- [15] Waals, J. H. van der and J. C. Platteeuw. Clathrate solutions. In *Advances in Chemical Physics*. John Wiley & Sons, New York, 1958.
- [16] Bahman Zarenezhad, Mona Mottahedin, and Ali Haghghi Asl. Effect of CO<sub>2</sub> concentration on the performance of different thermodynamic models for prediction of CH<sub>4</sub> + CO<sub>2</sub> + H<sub>2</sub>O hydrate equilibrium conditions. *Korean J. Chem. Eng.*, 28(3):949–953, 2011.
- [17] Julian Youxiang Zuo, Dan Zhang, and Erling H. Stenby. A thermodynamic model for gas hydrates in the presence of salts and methanol. *Chem. Eng. Commun.*, 184(1):175–192, 2001.

### Lattice Structure Models (ModLat)

Excess Gibbs energy models for electrolytes (mainly), metallic, and intermetallic systems.

**Related keywords:** sublattice | quasichemical | cell & theory | lattice & structure | metallic & alloys | temkin.

#### Bibliography.

- [1] J.-O. Andersson, A. Fernandez Guillermet, M. Hillert, B. Jansson, and B. Sundman. A compound-energy model of ordering in a phase with sites of different coordination numbers. *Acta Metall.*, 34(3):437–445, 1986.
- [2] Ibrahim Ansara, Ben Burton, Qing Chen, Mats Hillert, Armando Fernandez Guillermet, Suzana G. Fries, Hans Leo Lukas, Hans Jürgen Seifert, and W. Alan Oates. Applications of computational thermodynamics. Group 2: Models for composition dependence. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 24(1):19–40, 2000.
- [3] Patrice Chartrand and Arthur D. Pelton. The modified quasi-chemical model: Part III. Two sublattices. *Metall. Mater. Trans. A*, 32A:1397–1407, Jun. 2001.
- [4] V. Daněk and T. Ličko. Thermodynamic model and physico-chemical properties of silicate melts. *Chem. Geol.*, 96:439–447, 1992.
- [5] Y. Dessureault and A. D. Pelton. Contribution to the quasichemical model of reciprocal molten salt solutions. *J. Chim. Phys. Phys.-Chim. Biol.*, 88:1811–1830, 1991.
- [6] H. Flood, T. Førland, and K. Grjøtheim. Über den Zusammenhang zwischen Konzentrationen und Aktivitäten in geschmolzenen Salzmischungen. *Z. Anorg. Allg. Chem.*, 276:289–315, 1954.
- [7] Tormod Førland and Kai Grjøtheim. Thermodynamics of slag-metal equilibrium. *Metall. Trans. B*, 8:645–650, Dec 1977.
- [8] Tormod Førland and Kai Grjøtheim. Application of the activity concept in the physical chemistry of slags. *Metall. Trans. B*, 9B:45–49, Mar 1978.
- [9] H. Gaye, J. Lehmann, P. Rocabois, and F. Ruby-Meyer. Slag modelling and industrial applications. *High Temp. Mater. Processes*, 20(3-4):285–292, 2001.

- [10] H. Gaye and J. Welfringer. Modelling of the thermodynamic properties of complex metallurgical slags. In *Second International Symposium on Metallurgical Slags and Fluxes*, pages 357–375, Warrendale, PA, 1984. TMS.
- [11] P. Heransymenko and G. E. Speight. Ionic theory of slag-metal equilibria. *J. Iron Steel Inst., London*, 166:169–183, Nov 1950.
- [12] Mats Hillert. The compound energy formalism. *J. Alloys Compd.*, 320:161–176, 2001.
- [13] M. Hillert and L.-I. Staffansson. The regular solution model for stoichiometric phases and ionic melts. *Acta Chem. Scand. (1947-1973)*, 24:3618–3626, 1970.
- [14] Mats Hillert and Magnus Jarl. A model for alloying effects in ferromagnetic metals. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 2(3):227–238, 1978.
- [15] Mats Hillert, B. Jansson, B. Sundman, and John Ågren. A two-sublattice model for molten solutions with different tendency for ionization. *Metall. Trans. A*, 16:261–266, 1985.
- [16] Mats Hillert, Bo Jansson, and Bo Sundman. Application of the compound-energy model to oxide systems. *Z. Metallkd.*, 79(2):81–87, 1988.
- [17] Ivan Horsák and Ivo Sláma. A lattice model of electrolytes for the whole concentration range. *Collect. Czech. Chem. Commun.*, 52:1672–1679, 1987.
- [18] L. A. Kleintjens and R. Koningsveld. Lattice-gas treatment of supercritical phase behavior in fluid mixtures. *J. Electrochem. Soc.*, 127(11):2352–2355, Nov. 1980.
- [19] L. A. Kleintjens and R. Koningsveld. Mean-field lattice-gas description of the system CO<sub>2</sub>/H<sub>2</sub>O. *Sep. Sci. Technol.*, 17(1):215–233, 1982.
- [20] L. A. Kleintjens. Mean-field lattice gas description of vapour–liquid and supercritical equilibria. *Fluid Phase Equilib.*, 10:183–190, 1983.
- [21] Martijn H. R. Lankhorst, Henny J. M. Bouwmeester, and Henk Verweij. Thermodynamics and transport of ionic and electronic defects in crystalline oxides. *J. Am. Ceram. Soc.*, 80(9):2175–2198, 1997.
- [22] H. Marouschek and S. Peter. Description of the solubility of salts in supercritical steam on the basis of a mean-field lattice-gas model. *Fluid Phase Equilib.*, 1985(21):61–75, 1985.
- [23] A. D. Pelton, S. A. Degterov, G. Eriksson, C. Robelin, and Y. Dessureault. The modified quasichemical model I—Binary solutions. *Metall. Mater. Trans. B*, 31B:651–659, aug 2000.
- [24] Arthur D. Pelton and Patrice Chartrand. The modified quasi-chemical model: Part II. Multicomponent solutions. *Metall. Mater. Trans. A*, 32A:1355–1360, Jun. 2001.
- [25] Arthur D. Pelton, Patrice Chartrand, and Gunnar Eriksson. The modified quasi-chemical model: Part IV. Two sublattice quadruplet approximation. *Metall. Mater. Trans. A*, 32A:1409–1416, Jun. 2001.
- [26] Bo Sundman and John Ågren. A regular solution model for phases with several components and sublattices, suitable for computer applications. *J. Phys. Chem. Solids*, 42:297–301, 1981.
- [27] Bo Sundman. Review of alloys modelling. *An. Fis., Ser. B*, 86:69–82, 1990.
- [28] M. Temkin. Mixtures of fused salts as ionic solutions. *Acta Physicochim. URSS*, 20(4):411–420, 1945.
- [29] N. J. Trappeniers, J. A. Schouten, and C. A. Ten Seldam. Gas-gas equilibrium and the two-component lattice-gas model. *Chem. Phys. Lett.*, 5(9):541–544, Jun. 1970.
- [30] B. J. Wood and J. Nicholls. The thermodynamic properties of reciprocal solid solutions. *Contrib. Mineral. Petrol.*, 66:389–400, 1978.

### Models for Molecular Fluids (ModMol)

Excess Gibbs energy models for non-electrolytes.

**Related keywords:** (nonpolar | molecular) & excess & gibbs & energy.

#### Bibliography.

- [1] Denis S. Abrams and John M. Prausnitz. Statistical thermodynamics of liquid mixtures: A new expression for the excess Gibbs energy of partly or completely miscible systems. *AIChE J.*, 21(1):116–128, 1975.
- [2] C. W. Bale and A. D. Pelton. Mathematical representation of thermodynamic properties in binary systems and solution of Gibbs–Duhem equation. *Metall. Trans.*, 5:2323–2337, Nov. 1974.
- [3] C. W. Bale and A. D. Pelton. Series representations of the thermodynamic properties of solutions. *Can. Metall. Q.*, 14(3):213–219, 1975.
- [4] Gary L. Bertrand, William E. Acree, Jr., and Thomas E. Burchfield. Thermochemical excess properties of multicomponent systems: Representation and estimation from binary mixing data. *J. Solution Chem.*, 12(5):327–346, 1983.
- [5] Cline Black. Phase equilibria in binary and multicomponent systems. *Ind. Eng. Chem.*, 50(3):403–412, Mar. 1958.
- [6] Cabezas, Jr., H. and J. P. O’Connell. Some uses and misuses of thermodynamic models for dilute liquid solutions. *Ind. Eng. Chem. Res.*, 32:2892–2904, 1993.

- [7] Harrison C. Carlson and Allan P. Colburn. Vapor–liquid equilibria of nonideal solutions. *Ind. Eng. Chem.*, 34(5):581–589, May 1942.
- [8] E. L. Cheluguet, M. E. Weber, and J. H. Vera. Modifications of the Flory-Huggins-Goldstein model for accurate description of closed-loop phase diagrams. *Chem. Eng. Sci.*, 48(8):1415–1426, 1993.
- [9] B. Coto, C. Pando, and J. A. R. Renuncio. Correlation and prediction of phase equilibria in liquid mixtures by means of the UNIQUAC equation. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(2):141–150, 1993.
- [10] K. R. Cox and C. A. Eckert. Analytical model for excess Gibbs energy of nonspecific liquid metal interactions. *AIChE J.*, 30(1):143–150, Jan. 1984.
- [11] E. L. Derr and C. H. Deal. Analytical solutions of groups: Correlation of activity coefficients through structural group parameters. *Am. Inst. Chem. Eng. Symp. Ser.*, 3(32):40–51, 1969.
- [12] Paul J. Flory. Thermodynamics of high polymer solutions. *J. Chem. Phys.*, 10:51–61, Jan. 1942.
- [13] Aage Fredenslund, Russel L. Jones, and John M. Prausnitz. Group-contribution estimation of activity coefficients in nonideal liquid mixtures. *AIChE J.*, 21(6):1086–1099, Nov. 1975.
- [14] J. Gmehling. Potential of group contribution methods for the prediction of phase equilibria and excess properties of complex mixtures. *Pure Appl. Chem.*, 75(7):875–888, 2003.
- [15] Jürgen Gmehling, Jiding Li, and Martin Schiller. A modified UNIFAC model. 2. Present parameter matrix and results for different thermodynamic properties. *Ind. Eng. Chem. Res.*, 32(1):178–193, 1993.
- [16] Robert A. Heidemann and Jamanlal M. Mandhane. Some properties of the NRTL equation in correlating liquid–liquid equilibrium data. *Chem. Eng. Sci.*, 28:1213–1221, 1973.
- [17] Robert A. Heidemann and Jamanlal M. Mandhane. Ternary liquid–liquid equilibria: The van Laar equation. *Chem. Eng. Sci.*, 30(4):425–434, 1975.
- [18] Robert A. Heidemann. Use of infinite-dilution activity coefficients for predicting azeotrope formation at constant temperature and partial miscibility in binary liquid mixtures. *Ind. Eng. Chem. Fundam.*, 14(1):72–73, 1975.
- [19] Robert A. Heidemann and Sunil L. Kokal. Combined excess free energy models and equations of state. *Fluid Phase Equilib.*, 45:17–37, 1990.
- [20] C. P. Hicks.  $n$ -fluid models and phase equilibrium. *J. Chem. Soc., Faraday Trans. 2*, 72:423–425, 1976.
- [21] J. H. Hildebrand and S. E. Wood. The derivation of equations for regular solutions. *J. Chem. Phys.*, 1(12):817–822, dec 1933.
- [22] Mats Hillert. Predictions of ternary activities from binary. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 12(3):257–259, 1988.
- [23] Mitsuyasu Hiranuma. Significance and value of the third parameter in the modified Wilson equation. *Ind. Eng. Chem. Fundam.*, 20(1):25–28, 1981.
- [24] Maurice L. Huggins. Theory of solutions of high polymers. *J. Am. Chem. Soc.*, 64:1712–1719, Jul. 1942.
- [25] H. Kehiaian. Solubility of nonelectrolytes in systems characterized by Kohler’s equation. *Bull. Acad. Pol. Sci., Ser. Sci. Chim.*, 14(2):153–155, 1966.
- [26] R. L. Klaus and H. C. van Ness. The orthogonal polynomial representation of thermodynamic excess functions. *Chem. Eng. Prog., Symp. Ser.*, 63(81):88–104, 1967.
- [27] Dana E. Knox and Hendrick C. van Ness. A model for representation of  $g^e$  and  $h^e$ . *Fluid Phase Equilib.*, 15:267–285, 1984.
- [28] J. J. van Laar. Über Dampfspannungen von binären Gemischen. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 72:723–751, 1910.
- [29] J. J. van Laar. Zur Theorie der Dampfspannungen von binären Gemischen. Erwiderung an Herrn f. Dolezalek. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 83:599–608, 1913.
- [30] J. J. van Laar. Über den Zusammenhang zwischen der Abweichung der Dampfdruckkurve von binären Gemischen Normaler Stoffe von der geraden Linie, und der Mischungswärme in der flüssigen Phase. *Z. Phys. Chem., Abt. A*, 137:421–446, 1928.
- [31] Larry Larrinaga. Graphically determining the Wilson parameters. *Chem. Eng. [Int. Ed.]*, 6:87–91, Apr. 1981.
- [32] T. W. Leland, J. S. Rowlinson, G. A. Sather, and I. D. Watson. Statistical thermodynamics of two-fluid models of mixtures. *J.C.S. Faraday 1*, 65:2034–2043, 1969.
- [33] Ján Majling, Marián Dubík, Vladimír Kovár, and Victor Jesenák. Algoritmizácia výpočtu rovnovážneho fázového zloženia mnohozložkových sústav v subsolidusovej oblasti I. *Silikaty (Prague)*, 29:343–350, Apr. 1985.
- [34] Jamanlal M. Mandhane and Robert A. Heidemann. NRTL parameters for the ternary system  $n$ -butanol,  $n$ -butylacetate and water. *Can. J. Chem. Eng.*, 51:381–385, Jun 1973. Note to the Editor.
- [35] Max Margules. Über die Zusammensetzung der gesättigten Dämpfe von Mischungen. *Monatsh. Chem. Ver. Wiss.*, 104:1243–1278, 1895.

- [36] F. Mato and F. A. Mato. Two-parameter model for correlating liquid phase activity coefficients of binary systems. *Fluid Phase Equilib.*, 20:183–188, 1985.
- [37] G. Maurer and J. M. Prausnitz. On the derivation and extension of the UNIQUAC equation. *Fluid Phase Equilib.*, 2:91–99, 1978.
- [38] Michael L. Michelsen and Robert A. Heidemann. Some properties of equation of state mixing rules derived from excess Gibbs energy expressions. *Ind. Eng. Chem. Res.*, 35:278–287, 1996.
- [39] R. V. Orye and J. M. Prausnitz. Multicomponent equilibria with the Wilson equation. *Ind. Eng. Chem.*, 57(5):18–26, May 1965.
- [40] Arthur D. Pelton and Christopher W. Bale. A modified interaction parameter formalism for non-dilute solutions. *Metall. Trans. A*, 17A:1211–1215, Jul. 1986.
- [41] P. S. Puri and J. A. Ruether. Additive excess free energy models for predicting gas solubilities in mixed solvents. *Can. J. Chem. Eng.*, 52:636–640, Oct. 1974.
- [42] Otto Redlich and A. T. Kister. Algebraic representation of thermodynamic properties and the classification of solutions. *Ind. Eng. Chem.*, 40(2):345–348, Feb. 1948.
- [43] Henri Renon and J. M. Prausnitz. Local compositions in thermodynamic excess functions for liquid mixtures. *AIChE J.*, 14(1):135–144, 1968.
- [44] George Scatchard. Equilibria in non-electrolyte solutions in relation to the vapor pressures and densities of the components. *Chem. Rev. (Washington, D. C.)*, 8(2):321–333, 1931.
- [45] George Scatchard and Walter J. Hamer. The application of equations for the chemical potentials to partially miscible solutions. *J. Am. Chem. Soc.*, 57:1805–1809, Oct. 1935.
- [46] Jacques Schwartzentruber and Henri Renon. Extension of UNIFAC to high pressures and temperatures by the use of a cubic equation of state. *Ind. Eng. Chem. Res.*, 28:1049–1055, 1989.
- [47] Paul K. Talley, James Sangster, Arthur D. Pelton, and Christopher W. Bale. Prediction of vapor–liquid equilibria and thermodynamic properties for a quinary hydrocarbon system from optimized binary data using the Kohler method. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 16(1):93–106, 1992.
- [48] Dimitrios Tassios. Limitations in correlating strongly nonideal binary systems with the NRTL and LEMF equations. *Ind. Eng. Chem. Process Des. Dev.*, 15(4):574–578, 1976.
- [49] Alessandro Vetere. Vapor–liquid equilibria with supercritical gases calculated by the excess Gibbs energy method. *Fluid Phase Equilib.*, 28:265–281, 1986.
- [50] Wenchuan Wang and Kwang-Chu Chao. The complete local concentration model activity coefficients. *Chem. Eng. Sci.*, 38(9):1483–1492, 1983.
- [51] Zhi-Chang Wang, Reinhard Lück, and Bruno Predel. New models for computing thermodynamic properties and phase diagrams of ternary systems. Part 1. Three-factor models. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(3):217–234, 1990.
- [52] Zhi-Chang Wang, Reinhard Lück, and Bruno Predel. New models for computing thermodynamic properties and phase diagrams of ternary systems Part 2. Multi-factor models. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(3):235–256, 1990.
- [53] Grant M. Wilson. Vapor–liquid equilibrium. XI. A new expression for the excess free energy of mixing. *J. Am. Chem. Soc.*, 86:127–130, Jan 1964.
- [54] Grant M. Wilson. Areas of research on activity coefficients from group contributions at the thermochemical institute. *Phase Equilib. Fluid Prop. Chem. Ind., Proc., Int. Conf., 2nd, 1980*, pages 429–444, 1980.
- [55] Jaime Wisniak and Abraham Tamir. *Mixing and Excess Thermodynamic Properties*. Physical Sciences Data 1. Elsevier, Amsterdam, Holland, 1978.
- [56] Jaime Wisniak and Alexander Apelblat. Evaluation of Wohl parameters in binary systems. *Fluid Phase Equilib.*, 90:1–13, 1993.
- [57] Jaime Wisniak. Simple graphical method for calculating the constants of the Redlich–Kister equation. *Ind. Eng. Chem. Process Des. Dev.*, 32:240, 1993.
- [58] Kurt Wohl. Thermodynamic evaluation of binary and ternary liquid systems. *Trans. Am. Inst. Chem. Eng.*, 42:215–249, 1946.
- [59] Wood, Robert H., Terence H. Lilley, and Peter T. Thompson. Rapidly converging activity expansions for representing the thermodynamic properties of fluid systems: Gases, non-electrolyte solutions, weak and strong electrolyte solutions. *J. Chem. Soc., Faraday Trans. 1*, 74:1301–1323, 1978.
- [60] Liu Xingjun and Hao Shiming. An analysis on interaction parameters of binary solid solutions. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(1):67–78, 1993.

## CHAPTER 5

# Applied Sciences

### International Union of Pure and Applied Chemistry (IUPAC)

Operative definitions and standards in chemistry and thermodynamics.

**Related keywords:** iupac | terminology | definitions | conventions.

#### Bibliography.

- [1] [Robert A. Alberty](#). Use of Legendre transforms in chemical thermodynamics. *Pure Appl. Chem.*, 73(8):1349–1380, 2001. IUPAC Technical Report.
- [2] [Robert A. Alberty](#). Use of Legendre transforms in chemical thermodynamics: International union of pure and applied chemistry, physical chemistry division, commission on thermodynamics. *J. Chem. Thermodyn.*, 34:1787–1823, 2002.
- [3] [Selby Angus](#). Guide for the preparation of thermodynamic tables and correlations of the fluid state. *CODATA Bull.*, 51:1–43, Dec 1983. IUPAC Thermodynamic Tables Project Centre, Department of Chemical Engineering and Chemical Technology, Imperial College of Science and Technology, London.
- [4] [Gerrit Ansmann](#). Natural units and the vector space of physical values. *European Journal of Physics*, 36(3):035008, 2015.
- [5] [R. E. Bedford](#), [G. Bonnier](#), [H. Maas](#), and [F. Pavese](#). Recommended values of temperature on the international temperature scale of 1990 for a selected set of secondary reference points. *Metrologia*, 33(2):133–154, 1996.
- [6] [W. Blanke](#). Änderungen durch die einfürung der neuen internationalen temperaturskala von 1990. *Brennst.-Waerme-Kraft*, 43(12):571–574, dec 1991.
- [7] [Kenneth Nolon Carter](#) and [Kenneth N. Carter, Jr.](#) Meaningful melting points. *J. Chem. Educ.*, 72(7):647–648, Jul. 1995.
- [8] [A. Carotenuto](#) and [M. Dell’Isola](#). An experimental verification of saturated salt solution-based humidity fixed points. *Int. J. Thermophys.*, 17(6):1423–1439, 1996.
- [9] [E. Richard Cohen](#) and [B. N. Taylor](#). The 1973 least-squares adjustment of the fundamental constants. *J. Phys. Chem. Ref. Data*, 2(4):663–734, 1973.
- [10] [E. Richard Cohen](#) and [B. N. Taylor](#). The 1986 CODATA recommended values of the fundamental physical constants. *J. Phys. Chem. Ref. Data*, 17(4):1795–1803, 1988.
- [11] [E. Richard Cohen](#) and [Barry N. Taylor](#). The fundamental physical constants. *Phys. Today*, 45(8):9–13, Aug. 1992.
- [12] [E. Richard Cohen](#) and [Barry N. Taylor](#). The fundamental physical constants. *Phys. Today*, 52(8):5–9, Aug. 1999.
- [13] [T. B. Coplen](#). Atomic weights of the elements 1995. *J. Phys. Chem. Ref. Data*, 26(5):1239–1253, 1997.
- [14] [J. D. Cox](#) et al. A report of IUPAC commission 1.2 on thermodynamics: Notation for states and processes significance of the word “standard” in chemical thermodynamics, and remarks on the commonly tabulated forms of the thermodynamic functions: Appendix no. IV to manual of symbols and terminology for physicochemical quantities and units. *J. Chem. Thermodyn.*, 14:805–815, 1982. Also appeared in *Pure Appl. Chem.*, 54(6) 1239-1250 (1982).
- [15] [Luigi Crovini](#). Thermal metrology after the introduction of the ITS-90. *Meas. Tech. (Engl. Transl.)*, 36(12):1393–1405, 1993. International Collaboration.
- [16] [Jitka Eysseltová](#) and [Thedford P. Dirkse](#). IUPAC–NIST solubility data series 66. Ammonium phosphates. *J. Phys. Chem. Ref. Data*, 27(6), 1998.
- [17] [Fischer, J.](#) and [J. Ullrich](#). The new system of units. *Nature Physics*, 12:4–7, jan 2016.
- [18] [Robert D. Freeman](#). Conversion of standard thermodynamic data to the new standard-state pressure. *J. Chem. Educ.*, 62(8):681–686, aug 1985.
- [19] [R. N. Goldberg](#) and [Ron D. Weir](#). Conversion of temperatures and thermodynamic properties to the basis of the international temperature scale of 1990 (technical report). *Pure Appl. Chem.*, 64(10):1545–1562, 1992.
- [20] [N. N. Greenwood](#) et al. Atomic weights of the elements. *Pure Appl. Chem.*, 21:95–108, 1970.

- [21] N. E. Holden, R. L. Martin, et al. Atomic weights of the elements 1981. *Pure Appl. Chem.*, 55(7):1545–1562, 1983.
- [22] Peter J. Mohr and Barry N. Taylor. CODATA recommended values of the fundamental physical constants: 2002. *Reviews of Modern Physics*, 77:1–107, Jan 2005.
- [23] Peter J. Mohr, Barry N. Taylor, and David B. Newell. CODATA recommended values of the fundamental physical constants: 2010. *Reviews of Modern Physics*, 84:1527–1605, Oct.–Dec. 2012.
- [24] Peter J. Mohr and Barry N. Taylor. CODATA recommended values of the fundamental physical constants: 1998. *J. Phys. Chem. Ref. Data*, 28(6):1713–1852, 1999.
- [25] G. Olofsson, S. Angus, G. T. Armstrong, and A. N. Kornilov. Assignment and presentation of uncertainties of the numerical results of thermodynamic measurements. *Pure Appl. Chem.*, 53:1805–1825, 1981.
- [26] H. Preston-Thomas. The international temperature scale of 1990 (ITS-90). *Metrologia*, 27(1):3–10, 1990.
- [27] K. J. R. Rosman and P. D. P. Taylor. Isotopic compositions of the elements 1997. *J. Phys. Chem. Ref. Data*, 27(6):1275–1287, 1998.
- [28] Ron D. Weir and Robert N. Golberg. On the conversion of thermodynamic properties to the basis of the international temperature scale of 1990. *J. Chem. Thermodyn.*, 28(3):261–276, 1996.
- [29] D. R. White and J. Fischer. The Boltzmann constant and the new Kelvin. *Metrologia*, 52(1):S213–S216, 2015.
- [30] D. H. Whiffen, editor. *Manual of Symbols and Terminology for Physicochemical Quantities and Units: 1979 Edition*. International Union of Pure and Applied Chemistry: Nomenclature & Symbols. Pergamon Press, Oxford, London, 1979.
- [31] M. E. Wieser. Atomic weights of the elements 2005 (IUPAC technical report). *Pure Appl. Chem.*, 78(11):2051–2066, 2006.
- [32] C. L. Young, R. Battino, and H. L. Clever. The solubility of gases and liquids. IUPAC Solubility Data Series.

## Physics (Phys)

Classical physics, fluid mechanics, relativity, quantum mechanics, etc.

**Related keywords: physics | (continuum | fluid) mechanics & radiation | field & theory.**

### Bibliography.

- [1] Samuel K. Allison. Enrico Fermi: 1901 – 1954. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1957.
- [2] Marcelo Alonso and Edward J. Finn. *Physics*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1981.
- [3] Isaac Asimov. Motion, sound, and heat. In *Understanding Physics. 3 Volumes in 1*, volume I. Barnes & Noble, Inc., New York, 1966.
- [4] Isaac Asimov. Light, magnetism, and electricity. In *Understanding Physics. 3 Volumes in 1*, volume II. Barnes & Noble, Inc., New York, 1966.
- [5] Isaac Asimov. The electron, proton, and neutron. In *Understanding Physics. 3 Volumes in 1*, volume III. Barnes & Noble, Inc., New York, 1966.
- [6] M. E. Baltatu, R. A. Chong, M. L. Huber, and A. Laesecke. Transport properties of petroleum fractions. *Int. J. Thermophys.*, 20(1):85–95, 1999.
- [7] F. J. Barclay, T. J. Ledwidge, and G. C. Cornfield. Some experiments on sonic velocity in two-phase one-component mixtures and some thoughts on the nature of two-phase critical flow. In *Proceedings of the Institution of Mechanical Engineers*, pages 185–194. Institution of Mechanical Engineers, Sep 1969. *Proceedings of the Institution of Mechanical Engineers, Conference Proceedings 1969* 184: 185.
- [8] W. H. Bates. *The Cure of Imperfect Sight by Treatment Without Glasses*. Press of Thos. B. Brooks, Inc., New York, 1920.
- [9] Olof Beckman. Anders Celsius, 300 år. Föredrag av Hr Olof Beckman vid Kungl. Vetenskaps-societetens högtidssammanträde den 10 november 2001”. Årsbok / Kungl. Vetenskaps-societeten i Uppsala (2002) sid. 20–31., 2001.
- [10] Arthur Beiser, editor. *The World of Physics*. McGraw-Hill Book Company, Inc., New York, 1960.
- [11] George W. Benthien. James Clerk Maxwell: His life and his faith, aug 2009.
- [12] Jacob Bigeleisen. Statistical mechanics of isotope effects on the thermodynamic properties of condensed systems. *J. Chem. Phys.*, 34(5):1485–1493, May 1961.

- [13] [Jacob Bigeleisen](#), Slobodan V. Ribnikar, Marvin J. Stern, and W. Alexander van Hook. Molecular geometry and the vapor pressures of the isotopic ethylenes: Chemical and structural isomers. *J. Chim. Phys. Phys.-Chim. Biol.*, 60:60–65, 1963.
- [14] [Jacob Bigeleisen](#), Slobodan V. Ribnikar, and W. Alexander van Hook. Molecular geometry and the vapor pressure of isotopic molecules. The equivalent isomers *cis*-, *gem*-, *trans*-dideuteroethylenes. *J. Chem. Phys.*, 38(2):489–496, Jan. 1963.
- [15] [Jacob Bigeleisen](#), Marvin Stern, and W. Alexander van Hook. Molecular geometry and the vapor pressure of isotopic molecules.  $C_2H_3D$  and  $C^{12}H_2 = C^{13}H_2$ . *J. Chem. Phys.*, 38(2):497–504, Jan. 1963.
- [16] [Jacob Bigeleisen](#). Vapor pressures of isotopic molecules. *J. Chim. Phys. Phys.-Chim. Biol.*, 60:35–43, 1963.
- [17] [J. Bigeleisen](#), C. B. Cragg, and M. Jeevenandam. Vapor pressures of isotopic methanes—Evidence for hindered rotation. *J. Chem. Phys.*, 47(11):4335–4346, Dec. 1967.
- [18] [J. Bindon](#). The secret workings of a transparent Pop-Pop engine. *Model Engineer*, 192, 2004.
- [19] [Sean Bird](#). James Clerk Maxwell: 1831–1879. Covenant Christian High School for History of Mathematics at IUPUI (Indiana University—Purdue University Indianapolis), apr 2002.
- [20] [E. R. Boyko](#) and J. F. Belliveau. Surface tension and Avogadro’s number. *J. Chem. Educ.*, 63(8):671–672, aug 1986.
- [21] [Derek Bray](#). Gas dynamics introduction. Course material E338, Dept. of Aerospace, Power & Sensors (DAPS), Cranfield University, 2005.
- [22] [Gyles Brandreth](#). *The Great Book of Optical Illusions*. Sterling Publishing Co., Inc., New York, 1985. Earlier published as *The Big Book of Optical Illusions*, 1979.
- [23] [E. Bretscher](#) and J. D. Cockcroft. Enrico Fermi: 1901–1954. *Biographical Memoirs of Fellows of the Royal Society*, 1:69–78, nov 1955.
- [24] [Y. Brechet](#) and Z. Néda. On the circular hydraulic jump. *Am. J. Phys.*, 67(8):723–731, 1999.
- [25] [P. W. Bridgman](#). *The Nature of Thermodynamics*. Harvard University Press, Cambridge, Massachusetts, 1943.
- [26] [S. Browne](#), J. Ziegler, and J. E. Shepherd. Numerical solution methods for shock and detonation jump conditions. Technical report, Aeronautics and Mechanical Engineering, California Institute of Technology, Pasadena, CA USA 91125, 2008. GALCIT Report FM2006.006 July 2004—Revised August 21, 2008.
- [27] [W. Byers Brown](#). Analytical representation of the adiabatic equation for detonation products based on statistical mechanics and intermolecular forces. *Philos. Trans. R. Soc. London, Ser. A*, 339:345–353, 1992.
- [28] [H. A. Buchdahl](#). *Seventeen Simple Lectures on General Relativity Theory*. John Wiley & Sons, New York, 1981.
- [29] [Ashley H. Carter](#). *Classical and Statistical Thermodynamics*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 2001.
- [30] [Manuel Cardona](#). Albert Einstein as the father of solid state physics, 2005.
- [31] [J. W. Cary](#). Onsager’s relation and the non-isothermal diffusion of water vapor. *J. Phys. Chem.*, 67:126–129, Jan 1963.
- [32] [M. Castagnino](#) and C. Laciaana. The global thermodynamic arrow of time. *Classical Quantum Gravity*, 19:2657–2670, apr 2002.
- [33] [H. B. G. Casimir](#). Hugo Martin Tetrode. NRC Handelsblad — Wetenschap en Onderwijs bijlage, jan 1984.
- [34] [F. Charlet](#), M.-L. Turkel, J.-F. Danel, and L. Kazandjian. Evaluation of various theoretical equations of state used in calculation of detonation properties. *J. Appl. Phys.*, 84(8):4227–4238, nov 1988.
- [35] [R. Chéret](#). The life and work of Pierre-Henri Hugoniot. *Shock Waves (Historical Archive)*, 2(1):1–4, mar 1992. Springer-Verlag.
- [36] [A. A. Chialvo](#), P. T. Cummings, and H. D. Cochran. Solvation structure, hydrogen bonding, and ion pairing in dilute supercritical aqueous NaCl mixtures. *Int. J. Thermophys.*, 17(1):147–156, 1996.
- [37] [R. D. Cowan](#) and W. Fickett. Calculation of the detonation products of solid explosives with the Kistiakowsky–Wilson equation of state. *J. Chem. Phys.*, 24(5):932–939, May 1956.
- [38] [William C. Davis](#). High explosives: The interaction of chemistry and mechanics. *Los Alamos Science*, 2:48–75, 1981.
- [39] [Jean-Philippe Dionne](#). *Chapman–Jouguet Properties of Heterogeneous Explosives*. PhD thesis, McGill University, Jul 1996.
- [40] [Freeman Dyson](#). A meeting with Enrico Fermi. *Nature*, 27:297, jan 2004.
- [41] [Ken Easterling](#). *Tomorrow’s Materials*. The Institute of Metals, London, 1988.
- [42] [A. Einstein](#). Ist die trägheit eines Körpers von seinem Energieinhalt abhängig? *Ann. Phys. (Leipzig)*, 18:639–641, 1905.

- [43] [A. Einstein](#). *Über die spezielle und die allgemeine Relativitätstheorie*. Friedr. Vieweg & Sohn, Braunschweig, achte auflage edition, 1920.
- [44] [James F. Ely](#) and H. J. M. Hanley. Prediction of transport properties. 1. Viscosity of fluids and mixtures. *Ind. Eng. Chem. Fundam.*, 20(4):323–332, 1981.
- [45] [James F. Ely](#) and H. J. M. Hanley. Prediction of transport properties. 2. Thermal conductivity of pure fluids and mixtures. *Ind. Eng. Chem. Fundam.*, 22(1):90–97, 1983.
- [46] [B. Fellmuth](#), Ch. Gaiser, and J. Fischer. Determination of the Boltzmann constant — Status and prospects. *Measurements Science and Technology*, 17:R145–R159, 2006.
- [47] [Enrico Fermi](#). Artificial radioactivity produced by neutron bombardment. Nobel Lecture, 1938.
- [48] [Richard P. Feynman](#), Robert B. Leighton, and Matthew Sands. *The Feynman Lectures on Physics. Volume 1: Mainly Mechanics, Radiation and Heat*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1963.
- [49] [Richard P. Feynman](#), Robert B. Leighton, and Matthew Sands. *The Feynman Lectures on Physics. Volume 2: Mainly Electromagnetism and Matter*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1964.
- [50] [Richard P. Feynman](#), Robert B. Leighton, and Matthew Sands. *The Feynman Lectures on Physics. Volume 3: Quantum Mechanics*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1964.
- [51] [Ira Flatow](#). *They All Laughed. From Light Bulbs to Lasers: The Fascinating Stories Behind the Great Inventions That Have Changed Our Lives*. HarperCollins Publishers, Inc., New York, 1993.
- [52] [Neville Fletcher](#), Lloyd Hollenberg, John Smith, and Joe Wolfe. The didjeridu and the vocal tract. In *International Symposium on Musical Acoustics*, volume 30, pages 87–90, 2001.
- [53] [D. F. Fletcher](#). Vapour explosions: Multiphase detonations or deflagrations? *Shock Waves*, 3:181–192, 1994.
- [54] [F. E. Fowle](#). The transparency of aqueous vapor. *Astrophys. J.*, 42:394–411, 1915.
- [55] [Sigmund Fritz](#). Solar energy on clear and cloudy days. *Sci. Mon.*, pages 55–65, Feb. 1957.
- [56] [Laurence E. Fried](#). CHEETAH 1.0 user’s manual. Technical report, Lawrence Livermore National Laboratory, jun 1994.
- [57] [Laurence E. Fried](#) and P. Clark Souers. BKWC: An empirical BKW parametrization based on cylinder test data. *Propellants, Explosives, Pyrotechniques*, 21:215–223, 1996.
- [58] [D. L. Frost](#), J. H. S. Lee, and G. Ciccarelli. The use of Hugoniot analysis for the propagation of vapor explosion waves. *Shock Waves*, 1:99–110, 1991.
- [59] [George Gamow](#). *Gravitasjon*. Cappelenens reallbøker. J. W. Cappelen Forlag, Oslo, Norge, 1966. Oversatt av J.Randers. Bokens originaltittel: *Gravity*. Utgitt av Educational Services, New York.
- [60] [Robert P. H. Gasser](#) and W. Graham Richards. *An Introduction to Statistical Thermodynamics*. World Scientific Publishing Co. Pte. Ltd., Singapore, second edition, 1995.
- [61] [Josiah Willard Gibbs](#). *Elementary Principles in Statistical Mechanics: Developed with Especial Reference to the Rational Foundation of Thermodynamics*. Charles Scribner’s sons, New York, 1902.
- [62] [Robert Gilmore](#). *Alice in Quantumland*. Springer-Verlag, New York, 1995.
- [63] [K. R. Glaesemann](#) and L. E. Fried. Recent advances in modeling Hugoniots with Cheetah. In *14th APS Topical Conference on Shock Compression of Condensed Matter Baltimore, MD, United States*, pages 1–4, Jul. 31 through Aug. 5 2005.
- [64] [Samuel Glasstone](#). *Theoretical Chemistry*. D. van Nostrand Company, Inc., New York, 1944.
- [65] [Herbert Goldstein](#), Charles Poole, and John Safko. *Classical Mechanics*. Addison-Wesley Publishing Company, Reading, Massachusetts, third edition, 2001.
- [66] [Jacob Goldberg](#). Robert Boyle: The man who changed the history of science & the history of science that changed the man. *Dartmouth Undergraduate Journal of Science*, 27(4):129–140, 2007.
- [67] [Arthur Good](#). *Tom Tits experiment. Vetenskapliga förströelser*. Teknisk Litteraturtjänst, Stockholm, Sverige, 1987. Faksimil av första svenska utgåvan, utgiven av Alf Samuelssons Förlag, Stockholm 1898.
- [68] [I. Grattan-Guinness](#). Daniel Bernoulli and the varieties of mechanics in the 18th century. *NAW*, 5/1(3):242–249, sep 2000. Johann Bernoulli lecture, Groningen, May 9, 2000.
- [69] [Ulrich Grigull](#). Fahrenheit, a pioneer of exact thermometry. *Heat Transfer*, 1:9–18, 1966. The Proceedings of the 8th International Heat Transfer Conference, San Francisco, 1966.
- [70] [Miroslav Grmela](#) and Hans Kristian Öttinger. Dynamics and thermodynamics of complex fluids. I. Development of a general formalism. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 56(6):6620–6632, dec 1997.
- [71] [S. A. Gubin](#), V. V. Odintsov, and V. I. Pepekin. Thermodynamic calculation of ideal and non-ideal detonation. *Fizika Goreniya*, 23(4):75–84, jul–aug 1987. Translated from Russian by Plenum Publishing Corporation.

- [72] J. Güémez, R. Valiente, C. Fiolhais, and M. Fiolhais. Experiments with the drinking bird. *Am. J. Phys.*, 71(12):1257–1263, 2003.
- [73] Bo Gustafsson and Anders Stigebrandt. Dynamics of the freshwater-influenced surface layers in the skagerrak. *J. Sea Res.*, 35(1–3):39–53, 1996.
- [74] Herm. Haeder. *Dampfmaschinen. I. Band: Berechnung und Details*. Herm. Haeder, Duisburg am Rhein, siebente edition, 1903.
- [75] Herm. Haeder. *Dampfmaschinen. II. Band: Zeichnungen und Bilder*. Herm. Haeder, Duisburg am Rhein, 1903.
- [76] Benjamin Hammond. The physics of dreamtime: An analysis of the acoustical properties of a didgeridoo. PHYS 225 Musical Acoustic, 2007.
- [77] Robert A. Hatch. Isaac Newton biography—Newton’s life, career, work. In *Encyclopedia Americana*, volume 20. Grolier Incorporated, Danbury, CT, 1998.
- [78] Robert A. Heidemann, Ayodeji A. Jeje, and Farhang Mohtadi. *An Introduction to the Properties of Fluids and Solids*. The University of Calgary Press, 1987.
- [79] Kåre Hellan. *Emner fra fasthetslæren*. Tapir Forlag, Trondheim, Norge, 1981.
- [80] Kåre Hellan. *Fluidmekanikk*. Tapir Forlag, Trondheim, Norge, 1982.
- [81] Per Chr. Hemmer. *Statistisk mekanikk*. Tapir Forlag, Trondheim, Norge, 1970.
- [82] Per Chr. Hemmer and Eivind Hiis Hauge. Lars Onsager—en stor fysiker og kjemiker, Apr. 1993. Fra fysikkens verden.
- [83] Terrell L. Hill. *An Introduction to Statistical Thermodynamics*. Addison-Wesley Series in Chemistry. Addison-Wesley Publishing Company, Reading, Massachusetts, 1960.
- [84] Richard L. Hills. *Power from Steam: A History of the Stationary Steam Engine*. Cambridge University Press, Cambridge, London, 1989.
- [85] Terrell L. Hill. *Thermodynamics of Small Systems. Part II*. Dover Publications, Inc., New York, 1994. Two volumes bound as one. Reprint of 1964 edition published by W. A. Benjamin, Inc.
- [86] M. L. Hobbs and M. R. Baer. Nonideal thermoequilibrium calculations using a large product species data base. *Shock Waves*, 2:177–187, 1992.
- [87] Walter Höflechner. Ludwig Boltzmann—Persönlichkeit—Karriere—Bedeutung. Vortrag gehalten in Wien am 24. November 2006 zu Boltzmanns 100. Todestag, 2006.
- [88] G. Holtsmark. *Lærebok i fysikk*. H. Aschehough & Co. (W. Nygaard), Kristiania (Oslo), Norge, 1922.
- [89] W. Alexander van Hook and James T. Phillips. Gas–liquid partition chromatography of perdeuterioethane. Isotope effects. Vaporization from solution. *J. Phys. Chem.*, 70(5):1515–1520, May 1966.
- [90] W. Alexander van Hook. Vapor pressures of the methylacetylenes, H<sub>3</sub>CCCH, H<sub>3</sub>CCCD, D<sub>3</sub>CCCH, and D<sub>3</sub>CCCD. *J. Chem. Phys.*, 46(5):1907–1918, Mar. 1967.
- [91] Don A. Howard. Albert Einstein as a philosopher of science. *Physics Today*, pages 34–40, dec 2005.
- [92] W. M. Howard, L. E. Fried, and P. C. Souers. Modeling of non-ideal aluminized explosives. Technical Report UCRL-JC-132717, Lawrence Livermore National Laboratory, Livermore, CA, 1999. Prepared for submittal to the: 11th APS Topical Conference on Shock Compression of Condensed Matter, Snowbird (UT), June 27–July 2.
- [93] John R. Hull and Kenneth L. Uherka. Magnetic heat pumps for near-room-temperature applications. *Energy (Oxford)*, 14(4):177–185, 1989.
- [94] A. T. Humphrey. Lord Rayleigh—the last of the great victorian polymaths. *GEC Review*, 7(3):167–179, 1992.
- [95] Takano Ishida and Jacob Bigeleisen. Vapor pressure of the isotopic ethylenes. IV. Liquid ethylene-d<sub>3</sub>, and -d<sub>4</sub>. *J. Chem. Phys.*, 49(12):5498–5509, Dec. 1968.
- [96] Yasuhiro Iwamura, Takehiko Itoh, Mitsuru Sakano, and Shizuma Kuribayashi. Observation of nuclear transmutation induced by deuterium permeation through Pd complex. *Mitsubishi Heavy Industries, Ltd. Technical Review*, 42(1):1–2, Feb 2005.
- [97] Michel Janssen. Albert Einstein: His biography in a nutshell. Hsci/Phys 1905 Spring 2003, Einstein for Everyone, 2003.
- [98] H. G. Jerrard and D. B. McNeill. *Dictionary of Scientific Units: Including Dimensionless Numbers and Scales*. Chapman & Hall, London, 1992.
- [99] James Prescott Joule. On the production of heat by voltaic electricity. *Philos. Trans. R. Soc. London*, 4:280–282, dec 1840.
- [100] James Prescott Joule. On the mechanical equivalent of heat. *Philos. Trans. R. Soc. London*, 5:839, jun 1849.
- [101] KanEnergi AS. Nye fornybare energikilder. Technical report, Norges forskningsråd i samarbeid med Norges vassdrags- og energidirektorat (NVE), Oslo, Norge, 2001. Revidert utgave 2001.
- [102] KanEnergi AS. Nye fornybare energikilder. Technical report, Norges forskningsråd i samarbeid med Norges vassdrags- og energiverk, Oslo, Norge, 1996.

- [103] [Edwin C. Kemble](#) and Francis Birch. Percy Williams Bridgman: 1882–1961. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1970.
- [104] [A. I. Khinchin](#). *Mathematical Foundations of Statistical Mechanics*. Dover Publications, Inc., New York, 1949. Translated from the Russian by G. Gamow.
- [105] [Susan Werner Kieffer](#). Sound speed in liquid–gas mixtures: Water–air and water–steam. *Journal of Geophysical Research*, 82(77):2895–2904, Jul 1977.
- [106] [Charles Kittel](#). *Introduction to Solid State Physics*. John Wiley & Sons, New York, 1953.
- [107] [M. J. Klein](#). The physics of J. Willard Gibbs in his time. In *Proceedings of The Gibbs Symposium*, pages 1–21. Yale University, May 15–17 1989.
- [108] [Knotts, Sandra](#), Peter J. Mohr, and William D. Phillips. An introduction to the new SI. *The Physics Teacher*, 55(1):16–21, 2017.
- [109] [Knut Jostein Knutsen](#). *Formler og data i fysikk*. Tapir Forlag, Trondheim, Norge, 7 edition, 1994.
- [110] [A. S. Kompaneys](#). *Theoretical Physics*. Dover Publications, Inc., New York, 1962. Reprint of 1961 edition published by the Foreign Languages Publishing House, Moscow. Translated from the Russian by George Yankovsky.
- [111] [R. Krishna](#) and J. A. Wesselingh. The Maxwell–Stefan approach to mass transfer. *Chem. Eng. Sci.*, 52(6):861–911, 1997.
- [112] [J. Krol](#). The automatic hydraulic ram. *Proc. - Inst. Mech. Eng.*, 165:53–73, 1951.
- [113] [L. D. Landau](#) and E. M. Lifshitz. The classical theory of fields. In *Course of Theoretical Physics*, volume 2. Pergamon Press, Oxford, London, 1951. Translated from the Russian by Morton Hamermesh. Third revised English Edition (1971).
- [114] [L. D. Landau](#) and E. M. Lifshitz. Quantum mechanics: Non-relativistic theory. In *Course of Theoretical Physics*, volume 3. Pergamon Press, Oxford, London, 1958. Translated from the Russian by J. B. Sykes and J. S. Bell. Second edition, revised and enlarged (1965).
- [115] [L. D. Landau](#) and E. M. Lifshitz. Statistical physics. In *Course of Theoretical Physics*, volume 5. Pergamon Press, Oxford, London, 1958. Translated from the Russian by J. B. Sykes and M. J. Kearsley. Second Revised and Enlarged Edition (1969).
- [116] [L. D. Landau](#) and E. M. Lifshitz. Fluid mechanics. In *Course of Theoretical Physics*, volume 6. Pergamon Press, Oxford, London, 1959. Translated from the Russian by J. B. Sykes and W. H. Reid. Third Revised English Edition (1966).
- [117] [L. D. Landau](#) and E. M. Lifshitz. Theory of elasticity. In *Course of Theoretical Physics*, volume 7. Pergamon Press, Oxford, London, 1959. Translated from the Russian by J. B. Sykes and W. H. Reid. Second Revised and Enlarged Edition (1970).
- [118] [L. D. Landau](#) and E. M. Lifshitz. Mechanics. In *Course of Theoretical Physics*, volume 1. Pergamon Press, Oxford, London, 1960. Translated from the Russian by J. B. Sykes and J. S. Bell. Second edition (1969).
- [119] [L. D. Landau](#) and E. M. Lifshitz. Electrodynamics of continuous media. In *Course of Theoretical Physics*, volume 8. Pergamon Press, Oxford, London, 1960. Translated from the Russian by J. B. Sykes and J. S. Bell. Second impression (1963).
- [120] [D. F. Lawden](#). *Principles of Thermodynamics and Statistical Mechanics*. John Wiley & Sons, New York, 1987.
- [121] [Rhonda Lee-Desautels](#). Theory of van der Waals forces as applied to particulate materials. *Educ. Reso. for Part. Techn.*, pages 1–7, 2005. Number 051Q-Lee.
- [122] [G. N. Lewis](#). The fundamental laws of matter and energy. *Science (Washington, D. C.)*, 30:84–86, 1909.
- [123] [Ju Liu](#). *Thermodynamically Consistent Modeling and Simulation of Multiphase Flows*. PhD thesis, University of Texas, Austin, Dec 2014.
- [124] [Jing Ping Lu](#). Evaluation of the thermochemical code—CHEETAH 2.0 for modelling explosives performance. Technical report, DSTO Aeronautical and Maritime Research Laboratory, 506 Lorimer St Fishermans Bend, Victoria 3207 Australia, aug 2001. DSTO-TR-1199.
- [125] [Leif Lundby](#). *Forbrenningsmotorer. I. Grunnlag*. Universitetsforlaget, Oslo, Norge, 2 edition, 1975.
- [126] [Aksel Lydersen](#). Fasthetsberegninger, mar 1978.
- [127] [James Prescott Joule \(1818–1889\)](#): Converting work into heat. The Museum of Science and Industry in Manchester, 2004.
- [128] [William Francis Magie](#). *A Source Book in Physics*. Source Books in the History of the Sciences. McGraw-Hill Book Company, Inc., New York, 1935.
- [129] [E. A. Mason](#). Estimate of molecular sizes and Avogadro’s number from surface tension. *Am. J. Phys.*, 34(12):1193, dec 1966.
- [130] [B. C. McGee](#), M. L. Hobbs, and M. R. Baer. Exponential 6 parameterization for the JCZ3-EOS. Technical report, Energetic and Multi-Phase Processes, Sandia National Laboratories, P. O. Box 5800 Albuquerque, New Mexico 87185-0834, jul 1998. SAND98-1191.

- [131] **Fundamentals of Mass Determination**. Handbook published by Mettler – Toledo AG.
- [132] **Glossary of Weighing Terms**. Handbook published by Mettler – Toledo AG.
- [133] **V. N. Mikhalkin**. Thermodynamic calculation of detonation in poorly mixed gas mixtures. *Combustion, Explosion, and Shock Waves*, 32(1):57–60, 1996. Translated from Fizika Goreniya i Vzryva, Vol. 32, No. 1, pp. 66–70, January–February 1995.
- [134] **Donald G. Miller**. Ternary isothermal diffusion and the validity of the Onsager reciprocity relations. *J. Phys. Chem.*, 63:570–578, Apr 1959.
- [135] **Parry Moon**. Proposed standard solar-radiation curves for engineering use. *J. Franklin Inst.*, 230:583–617, 1940.
- [136] **W. Muschik**, C. Papenfuss, and H. Ehrentraut. A sketch of continuum thermodynamics. *J. Non-Newtonian Fluid Mech.*, 96:255–290, 2001.
- [137] **N. V. E. Nordenmark**. Anders Celsius. In *Lychnos-bibliotek*, volume 1. Lärdomshistoriska samfundet, 1936.
- [138] **J. J. O’Connor** and E. F. Robertson. William Thomson (Lord Kelvin). MacTutor History of Mathematics archive, University of St Andrews, oct 2003.
- [139] **Øgrim**, Ormestad, and Lunde. Elektrisitet, atomfysikk. In *Rom – stoff – tid*, bind 3. J. W. Cappelens Forlag, Oslo, Norge, 1970.
- [140] **Øgrim**, Ormestad, and Lunde. Mekanikk. In *Rom – stoff – tid*, bind 1. J. W. Cappelens Forlag, Oslo, Norge, 1971.
- [141] **Øgrim**, Ormestad, and Lunde. Varme, bølger, lys. In *Rom – stoff – tid*, bind 2. J. W. Cappelens Forlag, Oslo, Norge, 1972.
- [142] **Hans Kristian Öttinger** and Miroslav Grmela. Dynamics and thermodynamics of complex fluids. II. Illustrations of a general formalism. *Phys. Rev. E: Stat. Phys., Plasmas, Fluids, Relat. Interdiscip. Top.*, 56(6):6633–6655, dec 1997.
- [143] **E. Palm**. *Hydromekanikk*. Tapir Forlag, Trondheim, Norge, 1977.
- [144] **Faidra Papanelopoulou**. Gustave-Adolphe Hirn (1815–90): Engineering thermodynamics in mid-nineteenth-century France. *British Society for the History of Science*, 39(2):231–254, jun 2006.
- [145] **Robert Pehrson**. *Boltzmann Temperature*. Kolofon AS, www.kolofon.com, 2004.
- [146] **D. J. Picard** and P. R. Bishnoi. Calculation of the thermodynamic velocity in two-phase multicomponent fluids. *Int. J. Multiphase Flow*, 13(3):295–308, 1987.
- [147] **W. O. Price** and Sheldon Krinsky. On thought experiments: Ernst Mach. *Philosophical Forum*, pages 449–457, jul 1973. Translated from “Über Gedankenexperimente”, Zeitschrift für den physikalischen und chemischen Unterricht, 10, 1–5 (1897).
- [148] **S. G. Rajeev**. A Hamilton–Jacobi formalism for thermodynamics. *Ann. Phys. (Leipzig)*, 323:2265–2285, 2008.
- [149] **F. Reif**. *Fundamentals of Statistical and Thermal Physics*. McGraw-Hill Book Company, Inc., New York, 1965.
- [150] **James G. Rizzo**. *The Stirling Engine Manual*, chapter 21, pages 153–161. Camden Miniature Steam Services, Bath, Somerset, 1995.
- [151] **R. H. Romer**. What do “voltmeters” measure?: Faraday’s law in a multiply connected region. *Am. J. Phys.*, 50(12):1089–1093, 1982.
- [152] **Rumford, Count of, Benjamin**. An inquiry concerning the source of the heat which is excited by friction. *Philos. Trans. R. Soc. London*, 88:80–102, jan 1798.
- [153] **Manuel D. Salas**. The curious events leading to the theory of shock waves. *Shock Waves*, 16:477–487, 2007.
- [154] **F. Schlögl**. Thermodynamic uncertainty relation. *J. Phys. Chem. Solids*, 49(6):679–683, 1988.
- [155] **Erwin Schrödinger**. *Statistical Thermodynamics*. Dover Publications, Inc., New York, 1989. Reprint of 1952 edition (2nd) published by Cambridge University Press with the following subtitle: *A Course of Seminar Lectures Delivered in January – March 1944, At the School of Theoretical Physics, Dublin Institute For Advanced Studies*.
- [156] **J. R. Senft**. A mathematical model for Ringbom engine operation. *J. Eng. Gas Turbines Power*, 107:590–595, 1985.
- [157] **Franz Seufert**. *Versuche an Dampfmaschinen, Dampfkesseln, Dampfturbinen und Dieselmotoren*. Julius Springer, Berlin, fünfte edition, 1919.
- [158] **Ascher H. Shapiro**. *Shape and Flow. The Fluid Dynamics of Drag*. Number 20 in The Science Study Series. Heinemann Educational Books Ltd., London, 1981.
- [159] **Morris H. Shamos**. *Great Experiments in Physics. Firsthand Accounts from Galileo to Einstein*. Dover Publications, Inc., New York, 1987. Reprint of 1959 edition published by Holt, Rinehart and Winston, New York.
- [160] **R. M. Sillitto**. Clerk Maxwell’s life and work. Talk given at JCMB (James Clerk Maxwell Building, University of Edinburgh), for a party of visiting Dutch schoolteachers, 1993.

- [161] J. R. Simões-Moreira. Oblique evaporation waves. *Shock Waves*, 10:229–234, 2000.
- [162] Nils Skogen. *Fysikk utan svarte boksar*. Fysisk institutt, Universitetet i Oslo, 1993.
- [163] Jannike Solsvik and Hugo A. Jakobsen. A review of the concepts for deriving the equations of change from the classical kinetic theory of gases: Single-component, multicomponent, and reactive gases. *European Journal of Mechanics B/Fluids*, 56:46–65, 2016.
- [164] Euan Squires. *The Mystery of the Quantum World*. Institute of Physics Publishing, Bristol, UK, second edition, 1994.
- [165] Srivastava, Y. N., A. Widom, and L. Larsen. A primer for electro-weak induced low energy nuclear reactions. *arXiv:0810.0159v1*, pages 1–8, Oct 2008.
- [166] H. Eugene Stanley. *Introduction to Phase Transitions and Critical Phenomena*. International Series on Monographs on Physics. Oxford University Press, London, 1971.
- [167] Wolfgang Steinicke. Wolfgang Pauli—leben und werk. Seminar Physik- und Chemiegeschichte IV, Universität Hamburg Fachbereich 11: Mathematik, May 2004.
- [168] Marvin J. Stern, W. Alexander van Hook, and Max Wolfberg. Isotope effects on internal frequencies in the condensed phase resulting from interactions with the hindered translations and rotations. The vapor pressures of the isotopic ethylenes. *J. Chem. Phys.*, 39(12):3179–3196, Dec. 1963.
- [169] Johan Stivimoen. Geopatogene strålingssoner: jordstrålingspåvirkninger fra vannårer og lignende, aug 1988.
- [170] George P. Sutton. *Rocket Propulsion Elements: An Introduction to the Engineering of Rockets*. John Wiley & Sons, New York, 1949.
- [171] Hitoshi Takeuchi and Hiroo Kanamori. Equations of state of matter from shock wave experiments. *Journal of Geophysical Research*, 71(16):3985–3994, Aug 1966.
- [172] J. L. Threlkeld and R. C. Jordan. Direct solar radiation available on clear days. *Trans. Am. Soc. Heat., Refrig. Air-Cond. Eng.*, 64:45–56, 1958.
- [173] R. H. Thurston. *A History of the Growth of the Steam Engine*. D. Appleton & Company, 1878.
- [174] Gaston Tissandier. *Videnskabelige adspredelser*. Alb. Cammermeyer, Kristiania (Oslo), Norge, 1882. Oversatt af kommandørkaptein H. J. Müller.
- [175] L. D. Trowbridge. An equilibrium-based model of gas reactions and detonations. Technical report, Oak Ridge National Laboratory, Oak Ridge, Tennessee 37831-6285, apr 2000. ORNL/TM-2000/123.
- [176] David J. Turner. The missing science of ball lightning. *Journal of Scientific Exploration*, 17(3):435–496, 2003.
- [177] Annanias Tveit. Vanndampdiffusjonstabell for papp- og trefiberplater. Technical Report 9, Norges Byggforskningsinstitutt, Oslo, Norge, 1954.
- [178] Annanias Tveit. Fukt og fukttransport i porøse materialer. Technical Report 39, Norges Byggforskningsinstitutt, Oslo, Norge, 1964. En litteraturstudie.
- [179] Gheorghe Vasaru, Günther Müller, Günther Reinhold, and Traian Fodor. The thermal diffusion column. Theory and practice with particular emphasis on isotope separation. In E. Krell, editor, *Physikalisch-chemische Trenn- und Messmethoden*, band 13. VEB Deutscher Verlag der Wissenschaften, Berlin, 1969.
- [180] Vladimir I. Vysotskii and Alla A. Kornilova. Transmutation of stable isotopes and deactivation of radioactive waste in growing biological systems. *Annals of Nuclear Energy*, 62:626–633, 2013.
- [181] Jearl Walker. *The Flying Circus of Physics with Answers*. John Wiley & Sons, New York, 1977.
- [182] D. Weihs. The “directional” source—A new type of singularity in potential flow. *Isr. J. Technol.*, 14:3–8, 1976.
- [183] Steven Weinberg. *De første tre minutter. Moderne forskning om universets opprinnelse*. EFI Forlag, Oslo, Norge, 1984. Oversatt av A. Quale. Bokens originale tittel: *The Thirst Three Minutes*.
- [184] Robert L. Welsh. Imploding shocks and detonations. *J. Fluid Mech.*, 29(1):61–79, 1967.
- [185] Thorstein Wereide. *Statistical Theory of Energy and Matter*. Gyldendalske Boghandel, Kristiania, 1915.
- [186] Thorstein Wereide. *Relativitetsprincippet eller tidrummets struktur*. Gyldendalske Boghandel, Kristiania, 1922.
- [187] Cheryl Weyant and Ben Valentine. The physics of the didgeridoo. UIUC Physics 199POM, 2003.
- [188] John Archibald Wheeler. Albert Einstein: 1879–1955. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1980.
- [189] Widom, A. and L. Larsen. Ultra low momentum neutron catalyzed nuclear reactions on metallic hydride surfaces. *arXiv:cond-mat/0505026v1*, pages 1–4, May 2005.
- [190] A. S. Wightman. On the prescience of J. Willard Gibbs. In *Proceedings of The Gibbs Symposium*, pages 23–38. Yale University, May 15–17 1989.
- [191] Jaime Wisniak. Pierre Maurice Marie Duhem: A polemical scientist. *Chem. Educator*, 5:156–161, 2000.

- [192] [M. M. Woolfson](#). *An Introduction to X-ray Crystallography*. Cambridge University Press, Cambridge, London, 1979.
- [193] [Mark W. Zemansky](#). *Temperatures Very Low and Very High*. Dover Publications, Inc., New York, 1981. Reprint of 1964 edition published by D. van Nostrand Company, Inc.
- [194] [G. van der Zwan](#). *Theories for the dielectric constant*, 2003.

## Chemistry (Chem)

Inorganic, organic, and physical chemistry.

**Related keywords: chemistry.**

### Bibliography.

- [1] [Jacob Aall](#). Om bjergværksvæsenet fornemmeligen med hensyn til Norges jernværker. In *Facsimilia scientia et technica Norvegica*, volume 45. NTH-trykk, Trondheim, 1981. Faksimile eksemplar 1. Opptrykk av: *Nutid og Fortid. Et Hæfteskift af Jacob Aall, Andet Hæfte, Arendal 1833*.
- [2] [Walter L. Badger](#) and [Warren L. McCabe](#). *Elements of Chemical Engineering*. Chemical Engineering Series. McGraw-Hill Book Company, Inc., New York, second edition, 1936.
- [3] [Gordon M. Barrow](#). *Physical Chemistry*. McGraw-Hill Book Company, Inc., New York, fourth edition, 1979.
- [4] [Carl Friedrich Böbert](#). Über das modumer blaifarbenwerk in norwegen. In *Facsimilia scientia et technica Norvegica*, volume 42. NTH-trykk, Trondheim, 1978. Faksimile eksemplar 279. Faksimile *Archiv für Mineralogie, Geognosie, Bergbau und Hüttenkunde, Band 21, Berlin 1847*.
- [5] [Ildfast stein](#). Håndbok utgitt av Borgestad Fabrikker A/S.
- [6] [David Eberhard Bradt](#). Kort beskrivelse over det kongelige Modumske blaafarveverk i Buskeruds amt 1781. In *Facsimilia scientia et technica Norvegica*, volume 22. NTH-trykk, Trondheim, 1966. Faksimile eksemplar 119. Opptrykk av: *Topographisk Journal for Kongeriket Norge, Hefte 29, Kiøbenhavn 1802* side 145–179.
- [7] [Georg Bredig](#). Kopp, Hermann. *Allgemeine Deutsche Biographie*, 55:820–826, 1910.
- [8] [J. N. Brønsted](#). *Lærebog i fysisk kemi*. Levin & Munksgaard, København, Danmark, 1936.
- [9] [M. Th. Brünnich](#). Forsøg til mineralogie for Norge. In *Facsimilia scientia et technica Norvegica*, volume 41. NTH-trykk, Trondheim, 1975. Faksimile eksemplar 814. Opptrykk av: *Et Pris-Skrift Belønnet formedelst Hans Kongel. Høyheds Arve-Prinsens Gavnildhed af Det Kongl. Norske Videnskabers Selskab og paa dets Bekostning udgivet til alminnelig Brug, Trondhiem, 1777*.
- [10] [A. A. Burgess](#) and [D. J. Brennan](#). Application of life cycle assessment to chemical processes. *Chem. Eng. Sci.*, 56:2589–2604, 2001.
- [11] [Ragnar Bye](#). Store nordiske kjemikere. *Kjemi*, Spesial:1–41, dec 2003.
- [12] [W. C. Chapin](#). Chemical research problems. *J. Chem. Educ.*, 2(5):391–403, May 1925.
- [13] [Clair J. Collins](#) and [Newell S. Bowman](#), editors. *Isotope Effects in Chemical Reactions*. ACS Monograph. van Nostrand Reinhold Company, New York, 1970.
- [14] [F. Albert Cotton](#) and [Geoffrey Wilkinson](#). *Advanced Inorganic Chemistry*. John Wiley & Sons, New York, fourth edition, 1980.
- [15] [Day, Jr., R. A.](#) and [A. L. Underwood](#). *Quantitative Analysis*. Prentice-Hall Chemistry Series. Prentice Hall, Inc., Englewood Cliffs, New Jersey, second edition, 1967.
- [16] [Diamondoids and Their Role in Petroleum and Natural Gas Production Fouling](#).
- [17] [Ahmet Durmayaz](#), [Oguz Salim Sogut](#), [Bahrin Sahin](#), and [Hasbi Yavuz](#). Optimization of thermal systems based on finite-time thermodynamics and thermoeconomics. *Prog. Energy Combust. Sci.*, 30:175–217, 2004.
- [18] [Dyno Industrier](#). *Sprengstoffer—Sprengningsteknikk*. Dyno Industrier A. S., Oslo, 1976.
- [19] [Sven Englundh](#). Gamla kemisk tekniska recept. del I, 1980.
- [20] [Sven Englundh](#). Kemisk tekniska recept. del II, 1980.
- [21] [W. G. Ernst](#). Earth materials. In [A. Lee McAlester](#), editor, *Foundations of Earth Science Series*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 1969.
- [22] [Michael Peterson Escholt](#). Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10A. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 78. Opptrykk av: *Geologia Norvegica. Or, a brief instructive remembrancer, Concerning that very great and spacious earthquake, Which hapned almost quite through the South parts of Norway: Upon the 24<sup>th</sup> day of April, in the year 1657. In the Danish Tongue, by Michael Peterson Escholt. And Englished by Daniel Collins*.
- [23] [Michael Peterson Escholt](#). Geologia Norvegica. In *Facsimilia scientia et technica Norvegica*, volume 10. NTH-trykk, Trondheim, 1964. Faksimile eksemplar 290. Opptrykk av: *Geologia Norvegica. Eller: En kort Undervisning om det viit-begrebne Jordskjelff Som her udi Norge skeede mesten ofuer alt Syndenfields den 24. Aprilis udi nærværende Aar 1657*.

- [24] Basil T. Fedoroff, Henry A. Aaronson, Earl F. Reese, Oliver E. Sheffield, and George D. Clift. A. In *Encyclopedia of Explosives and Related Items*, volume 1. U. S. Army Research and Development Command, New Jersey, USA, 1960.
- [25] Basil T. Fedoroff and Oliver E. Sheffield. B–c. In *Encyclopedia of Explosives and Related Items*, volume 2. U. S. Army Research and Development Command, New Jersey, USA, 1962.
- [26] Basil T. Fedoroff and Oliver E. Sheffield. C–d. In *Encyclopedia of Explosives and Related Items*, volume 3. U. S. Army Research and Development Command, New Jersey, USA, 1966.
- [27] Basil T. Fedoroff and Oliver E. Sheffield. D. In *Encyclopedia of Explosives and Related Items*, volume 4. U. S. Army Research and Development Command, New Jersey, USA, 1969.
- [28] Basil T. Fedoroff and Oliver E. Sheffield. D–e. In *Encyclopedia of Explosives and Related Items*, volume 5. U. S. Army Research and Development Command, New Jersey, USA, 1972.
- [29] Basil T. Fedoroff and Oliver E. Sheffield. E–g. In *Encyclopedia of Explosives and Related Items*, volume 6. U. S. Army Research and Development Command, New Jersey, USA, 1974.
- [30] Basil T. Fedoroff and Oliver E. Sheffield. H–l. In *Encyclopedia of Explosives and Related Items*, volume 7. U. S. Army Research and Development Command, New Jersey, USA, 1975.
- [31] Seymour M. Kaye. M–p. In *Encyclopedia of Explosives and Related Items*, volume 8. U. S. Army Research and Development Command, New Jersey, USA, 1978.
- [32] Seymour M. Kaye. Q–t. In *Encyclopedia of Explosives and Related Items*, volume 9. U. S. Army Research and Development Command, New Jersey, USA, 1980.
- [33] Seymour M. Kaye. U–z. In *Encyclopedia of Explosives and Related Items*, volume 10. U. S. Army Research and Development Command, New Jersey, USA, 1983.
- [34] Thorbjørn Faarlund, Nils Sjøberg, and Thorleif Øisang, editors. *Molekylet*. Emnebiblioteket. Gyldendal Norsk Forlag, Oslo, Norge, 1970. Norsk oversettelse av *Building the Molecule* utgitt i 1969 av Marshall Cavendish Book Ltd. London.
- [35] Hans-Joachim Flechtner. *Die Welt in der Retorte*. Deutscher Verlag, Berlin, 1938.
- [36] Christie J. Geankoplis. *Transport Processes and Unit Operations*. Allyn and Bacon, Inc., Boston, Massachusetts, 1978.
- [37] S. Golubić, W. Krumbein, and J. Schneider. The carbon cycle. In P. A. Trudinger and D. J. Swaine, editors, *Biogeochemical Cycling of Mineral-Forming Elements*, volume 3 of *Studies in Environmental Science*, chapter 2.1. Elsevier, Amsterdam, Holland, 1979.
- [38] C. M. Guldberg and P. Waage. Studier over affiniteten. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapselskapets forhandlinger for 1864*.
- [39] C. M. Guldberg and P. Waage. Om den chemiske affinitet. In *Facsimilia scientia et technica Norvegica*, volume 15. NTH-trykk, Trondheim, 1964. Opptrykk av: *Videnskapselskapets forhandlinger for 1879*, nr. 4.
- [40] Jackson E. Harrar, Lester P. Rigdon, and Steven F. Rice. Raman spectral study of solutions of  $N_2O_4$  and  $N_2O_5$  in nitric acid. *Journal of Raman Spectroscopy*, 28:891–899, 1997.
- [41] Tore Haug-Warberg. *Sprengstoffer*. Eget forlag, 1979.
- [42] Walter Heinz. Kopp, Hermann. *Neue Deutsche Biographie*, 12:567–568, 1979.
- [43] Joel H. Hildebrand. Gilbert Newton Lewis: 1875–1946. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1958.
- [44] Holmes, Jr., Edward O. and David Reinson. The system ammonium nitrate–sodium nitrate. *J. Am. Chem. Soc.*, 66(3):453–457, mar 1944.
- [45] Thomas L. Jacobs, William E. Truce, and G. Ross Robertson. *Laboratory Practice of Organic Chemistry*. The Macmillan Press Ltd., London, fifth edition, 1974.
- [46] George J. Janz, Albert W. Davison, Henry S. van Klooster, and Walter H. Bauer. *Laboratory Manual of Physical Chemistry*. John Wiley & Sons, New York, fourth edition, 1956.
- [47] T. Kimura, T. Matsushita, K. Ueda, K. Tamura, and S. Takagi. Deuterium isotope effect on excess enthalpies of methanol or ethanol and their deuterium derivatives at 298.15 K. *J. Thermal Analysis Calorimetry*, 64:231–241, 2001.
- [48] Alexander Kipnis. Sackur, Otto. *Neue Deutsche Biographie*, 22:344, 2005.
- [49] Emil Kirschbaum. *Distillation and Rectification*. Chemical Publishing Co., Inc., New York, 1948. Translated by M. Wulfinhoff.
- [50] Norbert Adolph Lange and Gordon M. Forker. *Handbook of Chemistry*. McGraw-Hill Book Company, Inc., New York, tenth edition, 1961.
- [51] Pierre Lemay and Ralph E. Oesper. Pierre Louis Dulong, his life and work. *Chymia*, 1:171–190, 1948.
- [52] Octave Levenspiel. Earth’s early atmosphere. *Chemical Innovation*, 30(5):47–51, May 2000. Extended paper published on Octave Levenspiel’s home page.
- [53] Octave Levenspiel. *Chemical Reaction Engineering*. John Wiley & Sons, New York, second edition, 1972.
- [54] T. Lewis. Kitchen improvised plastic explosives, blasting caps and fertilizer explosives, 2009.

- [55] Karl Heinrich Lieser. *Einführung in die Kernchemie*, volume 1 of *Kernchemie in Einzeldarstellungen*. Verlag Chemie GmbH, Weinheim, Germany, 1969.
- [56] Lindsay. *Lindsay's Chemical Cross Reference*. Lindsay Publications Inc., Bradley, Illinois, 1989.
- [57] James Lovelock. *Gaia: The Practical Science of Planetary Medicine*. Gaia Books Limited, London, 1991.
- [58] Belle Lowe. *Experimental Cookery. From the Chemical and Physical Standpoint*. John Wiley & Sons, New York, 1943.
- [59] Einar Wang Lund. *Kvalitativ uorganisk analyse*. Universitetsforlaget, Oslo, Norge, 1976.
- [60] Aksel Lydersen. *Kjemiteknikk*. Tapir Forlag, Trondheim, Norge, 1972.
- [61] J. J. Manley. VIII. On the apparent change in weight during chemical reaction. *Proc. R. Soc. London, Ser. A*, 87(594):227–260, 21 Aug. 1912. Communicated by Prof. J. H. Poynting.
- [62] Karl Meyer. *Meyers Vareleksikon*. Gyldendalske Boghandel, Kristiania, Norge, 2 edition, 1907.
- [63] Karl Meyer. *Meyers Vareleksikon*. Gyldendalske Boghandel, Kristiania, Norge, 3 edition, 1918.
- [64] Robert H. Perry, Don W. Green, and James O. Maloney, editors. *Perry's Chemical Engineers' Handbook*. McGraw-Hill Book Company, Inc., New York, seventh edition, 1997.
- [65] K. S. Pitzer. Origin of the acentric factor. *ACS Symp. Ser.*, 60:1–10, 1977.
- [66] L. Pogliani. Graphs and thermodynamics. *J. Math. Chem.*, 46:15–23, 2009.
- [67] J. E. Prue. Ionic equilibria. In R. A. Robinson, editor, *Topic 15. Equilibrium Properties of Electrolyte Solutions*, volume 3 of *The International Encyclopedia of Physical Chemistry and Chemical Sciences*. Pergamon Press, Oxford, London, 1966.
- [68] John D. Roberts and Marjorie C. Caserio. *Basic Principles of Organic Chemistry*. W. A. Benjamin, Inc., Menlo Park, California, second edition, 1977.
- [69] A. B. Ronov and A. A. Yaroshevsky. Chemical composition of the earth's crust. In Pembroke J. Hart, editor, *The Earth's Crust and Upper Mantle*, number 13 in Geophysical Monograph. American Geophysical Union, Washington, D. C., 1969.
- [70] P. M. Røwde. *Nye råstoffer: et streiftog til verdenshusholdningens nye eventyrkilder*. P. M. Bye & Co. AS, Oslo, Norge, 1940.
- [71] Aamund Salvesson. *Lærebok i kjemi*. Yrkesopplæringsrådet for håndverk og industri, Oslo, Norge, 1956.
- [72] Heinrich Schmidhuber. Bericht über das Kobaltwerk Snarum in Norwegen nach denen in monat juni 1846 daselbst gesammelten unterlagen bearbeitet von Heinrich Schmidhuber, k. s. berggeschwornem. In *Facsimilia scientia et technica Norvegica*, volume 38. NTH-trykk, Trondheim, 1974. Faksimile eksemplar 508. Publisert: Leipzig, 1847.
- [73] Joseph A. Shaw and Elton Fisher. A new acetylene silver nitrate complex. *J. Am. Chem. Soc.*, 68(12):2745, dec 1946.
- [74] Michell J. Sienko and Robert A. Plane. *Chemistry*. McGraw-Hill Book Company, Inc., New York, fifth edition, 1976.
- [75] Lars Gunnar Sillén, Paul W. Lange, and C. Olof Gabrielson. *Fysikalisk-Kemiska Räkneuppgifter*. Almqvist & Wiksells Akademiska handböcker. Almqvist & Wiksells, Uppsala, Sverige, 1951.
- [76] Robert Stadler. Analytische und Sprengstofftechnische Untersuchungen an Azethylensilber. *Z. Gesamte Schiess- Sprengstoffwes.*, 10(33):269–338, 1938.
- [77] Roger Y. Stanier, Michael Doudoroff, and Edward A. Adelberg. *General Microbiology*. Macmillans Student Editions. The Macmillan Press Ltd., London, third edition, 1972.
- [78] Th. Tharaldsen. *Ekspløstoffer*. Dreyers Forlag, Oslo, Norge, 1950.
- [79] E. V. Veitsman. On the rate of relativistic surface chemical reactions. *Journal of Colloid and Interface Science*, 275:555–559, 2004.
- [80] Ragnar Vestin. Silver and copper compounds of acetylene. *Sven. Kem. Tidskr.*, 66(3):65–95, 1954.
- [81] Klaus Volkamer, Christoph Streicher, Kenneth Walton, John Fagan, Hartmut Schenkluhn, and Harry Marlot. Experimental re-examination of the law of conservation of mass in chemical reactions. *Journal of Scientific Exploration*, 8(2):217–250, Summer 1994.
- [82] Robert Weast, editor. *CRC Handbook of Chemistry and Physics*. CRC Press, Inc., Boca Raton, Florida, 58th edition, 1978.
- [83] Ralph R. Wenner. *Thermochemical Calculations*. McGraw-Hill Book Company, Inc., New York, 1941.
- [84] Richard Williams. François-Marie Raoult and Raoult's law: May 23, 1887. *APS News*, 20(5):2, May 2015. Column *This Month in Physics History*.
- [85] Jaime Wisniak. William Henry: His achievements and his law. *Chem. Educator*, 6:62–68, 2001.
- [86] J. H. Wolfenden. *Numerical Problems in Advanced Physical Chemistry*. Oxford University Press, London, 1938.
- [87] C. T. Zahn. The significance of chemical bond energies. *J. Chem. Phys.*, 2(10):671–680, oct 1934.
- [88] Philipp Zilles. Protokoll zum experimentalvortrag: Pyrotechnik. Technical report, Philipps-Universität Marburg, Fachbereich Chemie, Philipps-Universität Marburg, Germany, 2002.

- [89] [Steven S. Zumdahl](#). *Chemical Principles*. D. C. Heath and Company, Lexington, Massachusetts, second edition, 1995.

### Chemical Engineering (ChemEng)

Traditional chemical engineering disciplines: two-film theory, distillation, heat exchangers, membranes, etc.

**Related keywords: chemical & engineering | unit & operations | twofilm (very specific).**

#### Bibliography.

- [1] [Ville Alopaeus](#) and Harry V. Nordén. A calculation method for multicomponent mass transfer coefficient. *Comput. Chem. Eng.*, 23:1177–1182, 1999.
- [2] [S. P. S. Andrew](#) and D. Hanson. D1. The dynamics of nitrous gas absorption. *Chem. Eng. Sci.*, 14(1):105–113, 1961.
- [3] [Dick Bedeaux](#) and Signe Kjelstrup. Irreversible thermodynamics—A tool to describe phase transitions far from global equilibrium. *Chem. Eng. Sci.*, 59:109–118, 2004.
- [4] [Erich Brünnler](#). Reaktionsgeschwindigkeit in heterogenen systemen. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 47:56–102, 1904.
- [5] [Henry Jermain Maude Creighton](#). How the nitrogen problem has been solved. *J. Franklin Inst.*, 187(6):705–735, 1919.
- [6] [F. G. Donnan](#) and Irvine Masson. Theory of gas scrubbing towers with internal packing. *J. Soc. Chem. Ind. (London)*, 39(14):236–241, jul 1920.
- [7] [W. K. Lewis](#). The principles of counter-current extraction. *The Journal of Industrial and Engineering Chemistry*, 8(9):825–833, sep 1916. Laboratory and plant.
- [8] [W. K. Lewis](#) and W. G. Whitman. Principles of absorption. *Industrial and Engineering Chemistry*, 16(12):1215–1220, dec 1924. Absorption symposium.
- [9] [Yunda Liu](#), David Bluck, and Francisco Brana-Mulero. Static and dynamic simulation of NO<sub>x</sub> absorption tower based on a hybrid kinetic-equilibrium reaction model. In Mario R. Eden, John D. Sirola, and Gavin P. Towler, editors, *Proceedings of the 8th International Conference on Foundations of Computer-Aided Process Design*, volume 34 of *Computer Aided Chemical Engineering*. Elsevier, Amsterdam, Holland, 2014.
- [10] [Susumu Miyamoto](#). A theory of the rate of solution of gas into liquid. *Bull. Chem. Soc. Jpn.*, 7(1):8–17, 1932.
- [11] [W. Nernst](#). Theorie der reaktionsgeschwindigkeit in heterogenen systemen. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 47:52–55, 1904.
- [12] [Arthur A. Noyes](#) and Willis R. Whitney. The rate of solution of solid substances in their own solutions. *J. Am. Chem. Soc.*, 19:930–934, oct 1897.
- [13] [Pratibha Pandey](#) and R. S. Chauhan. Membranes for gas separation. *Prog. Polym. Sci.*, 26:853–893, 2001.
- [14] [J. Y. Park](#) and D. R. Paul. Correlation and prediction of gas permeability in glassy polymer membrane materials via a modified free volume based group contribution method. *J. Membr. Sci.*, 125:23–39, 1997.
- [15] [J. M. Prausnitz](#) and F. W. Tavares. Thermodynamics of fluid-phase equilibria standard chemical engineering operations. *AIChE J.*, 50(4):739–761, 2004.
- [16] [M. P. Pradhan](#), N. J. Suchak, P. R. Walse, and J. B. Joshi. Multicomponent gas absorption with multiple reactions: Modelling and simulation of NO<sub>x</sub> absorption in nitric acid manufacture. *Chem. Eng. Sci.*, 52(24):4569–4591, 1997. R. A. Mashelkar (editor), Festschrift for Professor M. M. Sharma, 52(24), pp. 4437–4698 (1997).
- [17] [Olaf Strelow](#). A general calculation method for plate heat exchangers. *Int. J. Therm. Sci.*, 39:645–658, 2000.
- [18] [R. Taylor](#), K. Achuthan, and A. Lucia. Complex domain distillation calculations. *Comput. Chem. Eng.*, 20(1):93–111, 1996.
- [19] [Stefan Tessendorf](#), Rafiqul Gani, and Michael L. Michelsen. Modeling, simulation and optimization of membrane-based gas separation systems. *Chem. Eng. Sci.*, 54:943–955, 1999.
- [20] [W. G. Whitman](#) and J. L. Keats. Rates of absorption and heat transfer between gases and liquids. *J. Ind. & Eng. Chem.*, 14(3):186–190, 1922.
- [21] [Walter G. Whitman](#). A preliminary experimental confirmation of the two-film theory of gas absorption. *Chemical and Metallurgical Engineering*, 29(4):146–148, 1923.

## Metallurgy & metallic materials (Metal)

Extractive and physical metallurgy, iron and steels and metal based materials.

**Related keywords:** metals | metallurgy.

### Bibliography.

- [1] **Jacob Aall**. Om jernmalmeier og jerntilvirkning i Norge. In *Facsimilia scientia et technica Norvegica*, volume 12. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 97. Opptrykk av: *Om Jernmalmeier og Jerntilvirkning i Norge, Et fragmentarisk Forsøg*. Det Skandinaviske Litteraturselskaps Skrifter.
- [2] **A. Almar-Næss**. *Metalliske Materialer: Struktur og egenskaper*. Tapir Forlag, Trondheim, Norge, 1981.
- [3] **Anthon Beuther**. Bergkordnung des löblichen neuen bergkwerchs auff dem Golmsbergk im königreich Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 13. NTH-trykk, Trondheim, 1963. Faksimile eksemplar 207. Publisert: *Zwickau, 1540*.
- [4] **Henrich Frantzen Blichfeld**. Kort efterretning om bergverket i Sundhordlehn udi Bergens Stift i Norge, til oplysning om sandhed og befording af rigets nytte, samt for at forebygge saavel de nærværende som tilkommende verkets eieres skade. In *Facsimilia scientia et technica Norvegica*, volume 34. NTH-trykk, Trondheim, 1970. Faksimile eksemplar 216. Publisert: *Kjøbenhavn, 1771*.
- [5] **D. I. Bloemacher**. Carbonyl iron powders: Its production and new developments. *Metal Powder Report*, 45(2):117–119, 1990.
- [6] **Sven Brenner**. *Materiallära*. Number 5 in Karlebo-serien. Maskinaktiebolaget Karlebo, Stockholm, Sweden, eighth edition, 1973.
- [7] **Georg Brochmann**. *Jern*. J. W. Cappelens Forlag, Oslo, Norge, 1939.
- [8] **K. R. Cox** and **C. A. Eckert**. Analytical model for excess Gibbs energy of nonspecific liquid metal interactions. *AIChE J.*, 30(1):143–150, Jan. 1984.
- [9] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh. 1790* side 387–449 og *Indhold, pl. I–II* side 450.
- [10] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. I* side 450.
- [11] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Facsimilia scientia et technica Norvegica*, volume 1. NTH-trykk, Trondheim, 1960. Opptrykk av: *TAB. II* side 450.
- [12] **Ole Evenstad**. Afhandling om jern-malm, som findes i myrer og moradser i Norge, og omgangsmaaden med at forvandle den til jern og staal. In *Det Konglige Danske Landhuusholdnings-Selskabs Skrifter, D.3. Kbh.*, volume D3, pages 387–449, Indhold, pl. I–II, 1790. Et priisskrift, som vandt det Kongelige Landhuusholdnings-Selskabs 2den Guldmedalje i Aaret 1782.
- [13] **Jens Gram**. Tanker om aarsagen til de norske jern-verkers nærværende tilstand og nogle herunder blandede anmerkninger samt forslag af hielpe-midler til deres opkomst. In *Facsimilia scientia et technica Norvegica*, volume 25. NTH-trykk, Trondheim, 1967. Faksimile eksemplar 59. Publisert: *Christiania, 1774*.
- [14] **Gustav Hallin**. *Handbok för stålbehandlare*. Josef Bergendahls Boktryckeri, Göteborg, 1937.
- [15] **von Haxthausen, Ferdinand Christian**. Kort efterretning og beregning om Sølv-Verkets og grubernes drift, som ogsaa bergstaden Kongsberg i Norge. In *Facsimilia scientia et technica Norvegica*, volume 30. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 275. Publisert: *Kjøbenhavn, 1776*.
- [16] **Gregorius Henckel**. Et par ord til mine landsmænd om og fra den gamle hyttemand Gregorius Henckel, pensioneret hytteskriver ved Kongsberg Sølvværk. In *Facsimilia scientia et technica Norvegica*, volume 5. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 049. Publisert: *Drammen, 1830*.
- [17] **Peder Hiort**. Historisk beretning om Røraas Kobberværk, fra dets første udfindelse og anlæg 1646 intil aar 1679. In *Facsimilia scientia et technica Norvegica*, volume 29. NTH-trykk, Trondheim, 1968. Faksimile eksemplar 40. Opptrykk av: *Journal for Politik, Natur- og Menneske-Kundskab, for Julii 1819*.
- [18] **[Christian Ditlev] Reventlow**, **Moltke**, **[Morten] Wormskiold**, **[Mads] Fridsch**, and **M. von Essen**. Berganordning for kongeriget Norge. In *Facsimilia scientia et technica Norvegica*, volume 19. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 213. Publisert: *Fredriksberg, den 7 de September 1812*.
- [19] **E. C. Rollason**. *Metallurgy for Engineers*. Edward Arnold (Publishers) Ltd., London, second edition, 1955.
- [20] **Terkel Rosenqvist**. *Principles of Extractive Metallurgy*. McGraw-Hill Series in Materials Science and Engineering. McGraw-Hill Book Company, Inc., New York, 1974.
- [21] **Nils Ryum**. *Metallene: Deres oppdagelse, egenskaper og anvendelser*. Eget forlag, 1985.

- [22] **Metaller**. Katalog No. I utgitt av A/S Schreiner & Co.
- [23] **Spesialstål og slitogods**. Katalog No. IV utgitt av A/S Schreiner & Co.
- [24] **Alf Søreide**. *Metallene i handverk og industri: Arbeidskunnskap for elever i framhaldsskolen, realskolen, yrkes- og bedriftsskoler*. J. W. Cappelen's Forlag, Oslo, Norge, 1949.
- [25] **Marianne Videm**. *Cyclic Deformation of Pure Aluminum*. Dr.ing. thesis 1992:36, Norwegian Institute of Technology, 1992.
- [26] **A. Visnapuu** and L. C. George. Recovery of critical metals by carbonyl processing. Technical Report RI-9087, U.S. Bureau of Mines, 1987. Bureau of Mines Report of Investigations.
- [27] **Johan Friederich Christian Voss**. Det frivillige slaverie, eller kort begrep om Sølvverket Kongsberg nærværende tilstand. In *Facsimilia scientia et technica Norvegica*, volume 33. NTH-trykk, Trondheim, 1969. Faksimile eksemplar 330. Publisert: *Kjøbenhavn, 1771*.
- [28] **Söfren Wesing**. Kurtze, doch warhafftige beschreibung der silber und anderer bergwercke in Norwegen. In *Facsimilia scientia et technica Norvegica*, volume 21. NTH-trykk, Trondheim, 1965. Faksimile eksemplar 98. Publisert: 2. utgave fra 1685.

### Computer Programming (Prog)

Programming languages and issues, and tools for dealing with such things.

**Related keywords:** programming | paradigms | software | implementation | algorithm | computer & language.

#### Bibliography.

- [1] **Krister Åhlander**, Magne Haveraaen, and Hans Z. Munthe-Kaas. On the role of mathematical abstractions for scientific computing. In *Proceedings of the IFIP TC2/WG2.5 Working Conference on the Architecture of Scientific Software*, volume 188, pages 145–158, 2000.
- [2] **K. Åhlander**. Supporting tensor symmetries in EinSum. *Computers & Mathematics with Applications*, 45:789–803, 2003.
- [3] **H. Abelson**, N. I. Adams, IV, et al. Revised<sup>4</sup> report on the algorithmic language Scheme. In William Clinger and Jonathan Rees, editors, *Special Interest Group on Programming Languages (SIGPLAN): Lisp Pointers*. Association for Computing Machinery (ACM), New York, 1991.
- [4] **Paul W. Abrahams**, Kathryn A. Hargreaves, and Karl Berry. *TeX for the Impatient*. Addison-Wesley Publishing Company, Reading, Massachusetts, 2003.
- [5] **Berend de Boer**.  $\LaTeX$  in proper Con $\TeX$ t, Feb. 2000.
- [6] **John Boyer**. Canonical XML version 1.0. Technical report, W3C, Mar. 2001. W3C Recommendation 15 March 2001.
- [7] **Structure and Interpretation of Computer Programs**. Building a system for symbolic differentiation, 2005. Computer Science 21B (Autumn Term 2005).
- [8] **David A. Carlson**. Time-space tradeoffs for tree search and traversal. In *Proceedings of 1986 ACM Fall Joint Computer Conference*, pages 585–594, 1986.
- [9] **Jeffrey A. Chard** and Vadim Shapiro. A multivector data structure for differential forms and equations. *Math. Comput. Simul.*, 54:33–64, 2000.
- [10] **Badford L. Chamberlain**, E. Christopher Lewis, and Calvin Lin. Regions: An abstraction for expressing array computation. In *ACM SIGAPL/SIGPLAN Conference APL'99*, pages 41–49, 1999.
- [11] **B. F. Clavinness**. On canonical forms and simplification. *Journal of Association for Computing Machinery*, 17:385–396, 1970.
- [12] **Alistar Cockburn** and Laurie Williams. The cost and benefits of pair programming. In Giancarlo Succi and Michele Marchesi, editors, *Extreme Programming Examined*, chapter 14. Pearson Education, 2001.
- [13] **Allin Cottrell**. Word processors: Stupid and inefficient, Apr. 2001.
- [14] **Harvey Davies**. Infinity arithmetic, comparisons and J. *APL Quote Quad*, 25(4):28–34, 1995.
- [15] **C. H. Q. Ding**. An optimal index reshuffle algorithm for multidimensional arrays and its applications for parallel architectures. *IEEE Transactions on Parallel and Distributed Systems*, 12(3):306–315, Mar. 2001.
- [16] **Michael Doob**. A gentle introduction to  $\TeX$ : A manual for self-study. Department of Mathematics, University of Manitoba, Canada, 2007.
- [17] **Günther Dotzel**. Oberon-2, a hi-performance alternative to C++. Technical report, ModulaWare, Jan. 1997.
- [18] **Eiffel Syntax Diagrams**, 1995.
- [19] **Victor Eijkhout**. *TeX by Topic, a T $\TeX$ nician's Reference*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1991.
- [20] **H. T. Esendal**. The selection of first programming language, 1993.

- [21] D. Eyheramendy and Th. Zimmermann. Object-oriented finite elements IV. Symbolic derivations and automatic programming of nonlinear formulations. *Computer Methods in Applied Mechanics and Engineering*, 190(22–23):2729–2751, 2001.
- [22] D. Eyheramendy and Th. Zimmermann. Object-oriented finite elements II. A symbolic environment for automatic programming. *Computer Methods in Applied Mechanics and Engineering*, 132(3–4):277–304, 1996.
- [23] D. Eyheramendy and Th. Zimmermann. Object-oriented finite elements III. Theory and application of automatic programming. *Computer Methods in Applied Mechanics and Engineering*, 154(1–2):41–68, 1998.
- [24] Robin Fairbairns. The UK TeX FAQ: Your 396 questions answered, 2005. Version 3.15a, date 2005/11/29.
- [25] Emden Gansner, Eleftherios Koutsofios, and Stephen North. Drawing graphs with *dot*, jan 2006.
- [26] Robert L. Glass. Editor’s corner. Some thoughts on automatic code generation. *Journal of systems software*, 37:1–3, 1997.
- [27] Sanjay Goil and Alok Choudhary. Parsimony: An infrastructure for parallel multidimensional analysis and data mining. *Journal of Parallel and Distributed Computing*, 61(3):285–321, 2001.
- [28] John Graham-Cumming. *The GNU make Book*. No starch press, San Francisco, 2015.
- [29] Perry Greenfield, Todd Miller, Rick White, J. C. Hsu, Paul Barrett, and Jochen Kupper. Numarray. An open source project. Release 0.8, Jan. 2004.
- [30] Stephen G. Hartke. A survey of free math fonts for T<sub>E</sub>X and L<sup>A</sup>T<sub>E</sub>X, feb 2006.
- [31] A. Harmens. Wilhoit’s formulae for ideal gas state thermodynamic properties. In *Fluid and Fluid Mixtures*, pages 112–117, Teddington Middlesex, UK, 1979. IPC Science and Technology Press. Proceedings of the NPL Conference: Chemical Thermodynamic Data on Fluid and Fluid Mixtures: Their Estimation, Correlation and Use, 11–12 September 1978.
- [32] O’Haskell Rationale, 2005.
- [33] A Survey of O’Haskell, 2005.
- [34] Tore Haug-Warberg. Computer science buzz-words, Sep. 2004.
- [35] Jim Hefferon, A. Goreham, and Michael J. Downes. Managing a one-person T<sub>E</sub>X system, 2001.
- [36] Amy Hendrickson. Getting T<sub>E</sub>Xnical: Insight into T<sub>E</sub>X macro writing techniques. *MAPS 5*, pages 55–66, 1990.
- [37] Paul Hudak. Conceptions, evolution, and application of functional programming languages. *ACM Computing Surveys*, 21(3):359–411, sep 1989.
- [38] Paul Hudak and Jonathan Young. Collecting interpretations of expressions. *ACM Transactions on Programming Languages and Systems*, 13(2):269–290, Apr. 1991.
- [39] Paul Hudak and Mark P. Jones. Haskell vs. Ada vs. C++ vs. Awk vs. . . . An experiment in software prototyping productivity, Jul. 1994.
- [40] Roger K. W. Hui. Rank and uniformity. *APL Quote Quad*, 25(4):83–90, 1995.
- [41] Roberto Ierusalimsky, Luiz Henrique de Figureiredo, and Waldemar Celes. Lua 5.0 reference manual. Technical report, PUC-Rio, Computer Science Department, Nov. 2003.
- [42] Kenneth E. Iverson. Computers and mathematical notation, 2000.
- [43] K. E. Iverson. Programming notation in systems design. *IBM Syst. J.*, 2:117–128, 1963.
- [44] Kenneth E. Iverson. Notation as a tool of thought. *Commun. ACM*, 23(8):444–465, 1980. 1979 ACM Turing Award Lecture.
- [45] K. E. Iverson. A personal view on APL. *IBM Syst. J.*, 30(4):582–593, 1991.
- [46] K. E. Iverson. Programming notation in systems design. *IBM Syst. J.*, 38(2/3):139–150, 1999.
- [47] Hrvoje Jasak. Openfoam: Open source CFD in research and industry. *International Journal of Naval Architecture and Ocean Engineering*, 1(2):89–94, 2009.
- [48] Paul Johnson. The Eiffel contract for C++ programmers, 1998.
- [49] Paul Johnson. Why Eiffel is better than C++ (for big projects), Mar. 1998.
- [50] Jerzy Karczmarczuk. Adjoint codes in functional framework, 2000.
- [51] Jerzy Karczmarczuk. Functional programming and mathematical objects. In Pieter H. Hartel and Rinus Plasmeijer, editors, *Functional Programming Languages in Education*, pages 1–19, Dec. 1995. First International Symposium, FPLE’95, Nijmegen, The Netherlands.
- [52] Jerzy Karczmarczuk. Generating power of lazy semantics. *Theoretical Computer Science*, 187:203–219, 1997.
- [53] Jerzy Karczmarczuk. Functional coding of differential forms, 1999.
- [54] Jerzy Karczmarczuk. Scientific computation and functional programming. *Computing in Science and Engineering*, 1(3):64–72, 1999.
- [55] D. Kaustub and David Grimes. A comparison of object oriented scripting languages: Python and Ruby, Dec. 2001.

- [56] Naoki Kawanabe, Takeshi Hourinouchi, Masato Shiotana, and Toshihiro Sakakima. Development for scientific analysis and visualization with the object-oriented language Ruby, 2004.
- [57] Brian W. Kernighan and Dennis M. Ritchie. *The C Programming Language*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 1988.
- [58] Ki-Joo Kim. *L<sup>A</sup>T<sub>E</sub>X fonts*, 2005.
- [59] Donald E. Knuth. *The T<sub>E</sub>Xbook*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1996.
- [60] Helmut Kopka and Patrick W. Daly. *A Guide to L<sup>A</sup>T<sub>E</sub>X<sub>2 $\epsilon$</sub> : Document Preparation for Beginners and Advanced Users*. Addison-Wesley Publishing Company, Reading, Massachusetts, second edition, 1995.
- [61] J. Korelc. Multi-language and multi-environment generation of nonlinear finite element codes. *Engineering with Computers*, 18(4):312–327, 2002.
- [62] Jože Korelc. Automatic generation of finite-element code by simultaneous optimization of expressions. *Theoretical Computer Science*, 187:231–248, 1997.
- [63] E. Krishnan. *L<sup>A</sup>T<sub>E</sub>X tutorials*, 2003.
- [64] *L<sup>A</sup>T<sub>E</sub>X and Fonts*, 2000.
- [65] Gary T. Leavens and Todd D. Millstein. Multiple dispatch as dispatch on tuples, 1998.
- [66] Éric Lévénéz. *Languages*, Apr. 5 2006.
- [67] Chun-Yuan Lin, Jen-Shiuh Liu, and Yeh-Ching Chung. Efficient representation scheme for multidimensional array operations. *IEEE Transactions on Computers*, 51(3):327–345, Mar. 2002.
- [68] João Madureira, Marco Silva, Rui Meireles, and Tiago Pereira. The Ruby programming language: A pragmatic analysis, 2005.
- [69] L. R. U. Manssur and R. Portugal. The canon package: A fast kernel for tensor manipulators. *Computer Physics Communications*, 157(2):173–180, Feb. 2004.
- [70] John McCarthy. Recursive functions of symbolic expressions and their computation by machine, part i. *Communications of the ACM archive*, 3(4):1829–1839, apr 1960.
- [71] D. B. McIntyre. Language as an intellectual tool: From hieroglyphics to APL. *IBM Syst. J.*, 30(4):554–581, 1991.
- [72] D. B. McIntyre. Mastering J. In *Proceedings of the International Conference on APL'91*, pages 264–273, Palo Alto, California, United States, 1991. ACM NY, USA.
- [73] McKim, Jr., James C. Programming by contract: Designing for correctness, 1999.
- [74] Bertrand Meyer. Eiffel vs C++: One language designer's view, Jun. 1989.
- [75] Cleve Moler. Floating points. IEEE standard unifies arithmetic model, 1996.
- [76] Bernd Mösl. A comparison of C++, FORTRAN 90 and Oberon-2 for scientific programming. Technical report, Arithmetica, May 1995.
- [77] Michael Neumann. A comparison between BETA, C++, Eiffel, Java, Object Pascal, Ruby and Smalltalk, Jul. 2000.
- [78] Michael Neumann. Comparing and introducing Ruby, Feb. 2000.
- [79] K. W. Ng and C. K. Luk. A survey of languages integrating functional, object-oriented and logic programming. *Microprocessing and Microprogramming*, 41:5–36, 1995.
- [80] Stephen Omohundro. The Sather language: Efficient, interactive, object-oriented programming, 1991.
- [81] Scott Pakin. The comprehensive L<sup>A</sup>T<sub>E</sub>X symbol list, sep 2005.
- [82] Scott Pakin. The comprehensive L<sup>A</sup>T<sub>E</sub>X symbol list, Jan 2017.
- [83] Todd Plessel. Design by contract: A missing link in the quest for quality software, Aug. 1998.
- [84] Robert F. Pointon, Steffen Priebe, Hans-Wolfgang Loidl, Rita Loogen, and Phil W. Trinder. Functional vs object-oriented distributed languages. In *Lecture Notes in Computer Science*, volume 2178. Springer-Verlag, New York, 2001.
- [85] Madhusudana Rao, R., R. Rengaswamy, A. K. Suresh, and K. S. Balaraman. Industrial experience with object-oriented modelling: FCC case study. *Chemical Engineering Research and Design: Transactions of the Institution of Chemical Engineers Part A*, 82(4):527–552, apr 2004.
- [86] Keith Reckdahl. Using EPS graphics in L<sup>A</sup>T<sub>E</sub>X<sub>2 $\epsilon$</sub>  documents, 1997.
- [87] Glenn C. Reid. *Thinking in Postscript*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1990.
- [88] Garrett Rooney. Extending Ruby with C, Nov. 2004.
- [89] Peter Van Roy, Per Brand, Denys Duchier, Seif Haridi, Martin Henz, and Christian Schulte. Logic programming in the context of multiparadigm programming: The Oz experience. *Theory and Practice of Logic Programming*, 7:717–763, 2003.
- [90] Ruby QuickRef, 2004.
- [91] Ruby Syntax, 2004.
- [92] Ruby User's Guide, 2004.

- [93] [Heinz W. Schmidt](#) and Stephen M. Omohundro. CLOS, Eiffel, and Sather: A comparison. Technical Report TR-91-047, International Computer Science Institute, Berkeley, California, Sep. 1991.
- [94] [K. E. Seamons](#) and M. Winslett. An efficient abstract interface for multidimensional array i/o. In *Supercomputing'94. Proceedings*, pages 650–659, 1994.
- [95] *L<sup>A</sup>T<sub>E</sub>X Tutorials*: A primer. Indian T<sub>E</sub>X Users Group, sep 2003.
- [96] *L<sup>A</sup>T<sub>E</sub>X2<sub>ε</sub> for Class and Package Writers*. The L<sup>A</sup>T<sub>E</sub>X3 Project, dec 2003.
- [97] [Tadao Takaoka](#). O(1) time algorithms for combinatorial generation by tree traversal. *The Computer Journal*, 42(5):400–408, 1999.
- [98] [Ross Taylor](#). Thermodynamics with Maple. IV—The properties of steam. *MAPLETECH*, 3(2):61–68, 1996.
- [99] [Ross Taylor](#). Thermodynamics with Maple. I—Symbolic computation. *Mathematics and Computers in Simulation*, 45:101–119, 1998.
- [100] [Ross Taylor](#). Thermodynamics with Maple. II—Numerical and graphical applications. *Mathematics and Computers in Simulation*, 45:121–146, 1998.
- [101] [Josef Templ](#). Oberon-2 vs. C++, Oct. 1994.
- [102] [Hàn Thế Thành](#). Micro-typographic extensions to the T<sub>E</sub>X typesetting system. Master's thesis, Masaryk University Brno, Faculty of Informatics, Oct. 2000.
- [103] [T. Tsuji](#), [A. Isshiki](#), [T. Hochin](#), and [K. Higuchi](#). An implementation scheme of multidimensional arrays for molap. In *13th International Workshop on Database and Expert Systems Applications*, pages 773–778, 2002.
- [104] [Ken Turner](#). BibT<sub>E</sub>X style examples. Department of Science and Mathematics, University of Stirling, 2005.
- [105] [J. Vignes](#). New methods for evaluating the validity of the result of mathematical computations. *Math. Comput. Simul.*, 20:227–249, 1978.
- [106] [Knut L. Vik](#). *APL—A Programming Language*. Regnesenteret ved Universitetet i Trondheim, Trondheim, Norge, 1981.
- [107] [Peter J. Weingartner](#). A first guide to Postscript, Apr. 1997.
- [108] [Jim Weirich](#). Design by contract for Java, 1999.
- [109] [Darrell Whitley](#). A genetic algorithm tutorial. *Statistics and Computing*, 4:65–85, 1994.
- [110] [Niklaus Wirth](#). What can we do about the unnecessary diversity of notation for syntactic definitions? *Commun. ACM*, 20(11):822–823, Nov 1977.
- [111] [Yaml Cookbook at the YamlForRuby Site](#), 2004.
- [112] [Gerhard Zimmermann](#). Efficient creation of building performance simulation using automatic code generation. *Energy and Buildings*, 34:973–983, 2002.
- [113] [Th. Zimmermann](#) and [D. Eyheramendy](#). Object-oriented finite elements I. Principles of symbolic derivations and automatic programming. *Computer Methods in Applied Mechanics and Engineering*, 132(3–4):259–276, 1996.

## Mathematics (Math)

Applied and engineering mathematics, optimization, algebra, calculus, etc.

**Related keywords:** **mathematics.**

### Bibliography.

- [1] [N. H. Abel](#). Démonstration de l'impossibilité de la résolution générale des équations du cinquième degré. In *Facsimilia scientia et technica Norvegica*, volume 40. NTH-trykk, Trondheim, 1976. Faksimile eksemplar 302. Opptrykk av: *Mémoire sur les équations algébriques où on démontre l'impossibilité de la résolution générale des équations du cinquième degré par N. H. Abel, Christiania 1824*.
- [2] [Knut Alfsen](#) and [Erik Alfsen](#). *Matematikk for gymnaset: funksjonslære 1*. H. Aschehough & Co. (W. Nygaard), Kristiania (Oslo), Norge, 1974.
- [3] [Fr. Chr. Holb. Arentz](#). Undersøgning om hvorledes man paa korteste maade kan opløse saadanne æqvationer, som indeholde flere eller mange ubekjendte størrelser tillige. In *Facsimilia scientia et technica Norvegica*, volume 2. NTH-trykk, Trondheim, 1961. Faksimile eksemplar 334. Opptrykk av: *Nye Samling af det Kongelige Norske Videnskabers Selskabs Skrifter, Andet bind, Kiøbenhavn 1788* side 251–286.
- [4] [V. I. Arnold](#). Contact geometry: The geometrical method of Gibbs's thermodynamics. In *Proceedings of The Gibbs Symposium*, pages 163–179. Yale University, May 15–17 1989.
- [5] [Karl Egil Aubert](#). Niels Henrik Abel. *Normat*, 27(4):129–140, 1979.
- [6] [Adam Berger](#). Convexity, maximum likelihood and all that. School of Computer Science, Carnegie Mellon University, 2002.

- [7] [John D. Blanton](#). *Foundations of Differential Calculus: Euler*. Springer-Verlag, New York, 2000. Translated from the Latin *Institutiones Calculi Diffefferentialis*, Chapters 1 to 9, by Leonhard Euler, 1755.
- [8] [Jacqueline Boniface](#). Leopold Kronecker’s conception of the foundations of mathematics. *Philosophia Scientiæ*, CS 5:143–156, 2005.
- [9] [Louis de Branges de Bourcia](#). Apology for the proof of the Riemann hypothesis, Jan. 2005.
- [10] [C. Brezinski](#). The life and work of André Cholesky. *Numer. Algor.*, 43:279–288, 2006.
- [11] [I. N. Bronshtein](#) and [K. A. Semendyayev](#). *Handbook of Mathematics*. van Nostrand Reinhold Company, New York, 1985.
- [12] [Viggo Brun](#). Det gjenfunne manuskript til Abels parisavhandling. *Normat*, 7:91–97, 1953.
- [13] [Roland Bulirsch](#) and [Josef Stoer](#). Numerical treatment of ordinary differential equations by extrapolation methods. *Numer. Math.*, 8:1–13, 1966.
- [14] [J. C. Butcher](#). *Numerical Methods for Ordinary Differential Equations*. John Wiley & Sons, New York, second edition, 2008.
- [15] [J. C. Butcher](#) and [G. Wanner](#). Runge-kutta methods: Some historical notes. *Applied Numerical Mathematics*, 22:113–151, 1996.
- [16] [S. H. Choi](#), [D. A. Harney](#), and [N. L. Book](#). A robust path tracking algorithm for homotopy continuation. *Comput. Chem. Eng.*, 20(6/7):647–655, 1996.
- [17] [Alan Cooper](#), [Joel Feldman](#), and [Lon Rosen](#). Higher Legendre transforms and their relationship to Bethe–Salpeter kernels and R-field projectors. *J. Math. Phys. (Woodbury, N. Y.)*, 23(5):845–868, May 1982.
- [18] [J. Čretnik](#). A simple method for computation of a minimal representation of a system in state space. *Electronic Letters*, 24(12):752–753, jun 1988.
- [19] [P. Deuffhard](#). Order and stepsize in extrapolation methods. *Numer. Math.*, 41:399–422, 1983.
- [20] [M. Duží](#). Kurt Gödel. Metamathematical results on formally undecidable propositions: Completeness vs. incompleteness. *Organon F*, XII(4):447–474, 2005.
- [21] [W. Dyck](#), [S. Gundelfinger](#), [J. Lüroth](#), and [M. Noether](#). *Ludwig Otto Hesse’s gesammelte Werke*, pages 711–721. Der Mathematisch-Physikalische Classe der Königlich Bayerische Akademie der Wissenschaften, München, 1897.
- [22] [C. W. Gear](#) and [L. R. Petzold](#). ODE methods for the solution of differential/algebraic systems. *SIAM J. Numer. Anal.*, 21(4):716–728, 1984.
- [23] [C. W. Gear](#). Differential algebraic equations, indices, and integral algebraic equations. *SIAM J. Numer. Anal.*, 27(6):1527–1534, dec 1990.
- [24] [R. P. Gillespie](#). Partial differentiation. In [Alexander C. Aitken](#) and [Daniel E. Rutherford](#), editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, 1951.
- [25] [Michael D. Greenberg](#). *Foundations of Applied Mathematics*. Prentice Hall, Inc., Englewood Cliffs, New Jersey, 1978.
- [26] [K. A. Green](#), [S. Zhou](#), and [K. D. Luks](#). The fractal response of robust solution techniques to the stationary point problem. *Fluid Phase Equilib.*, 84:49–78, 1993.
- [27] [Nora Hartsfield](#) and [Gerhard Ringel](#). *Pearls in Graph Theory: A Comprehensive Introduction*. Dover Publications, Inc., New York, 2003. First Published by Academic Press in 1990.
- [28] [Stephen G. Hartke](#). A survey of free math fonts for  $\text{T}_{\text{E}}\text{X}$  and  $\text{L}_{\text{A}}\text{T}_{\text{E}}\text{X}$ , feb 2006.
- [29] [André Hauser](#). Auf den spuren Leonhard Eulers. *Geschichte der Mathematik*, Pädagogische Hochschule Aargau, Institut Sekundarstufe 1, 2008.
- [30] [Arnljot Høyland](#). Sannsynlighetslære. In *Sannsynlighetsregning og statistisk metodelære*, del I. Tapir Forlag, Trondheim, Norge, second edition, 1976.
- [31] [Arnljot Høyland](#). Statistisk metodelære. In *Sannsynlighetsregning og statistisk metodelære*, del II. Tapir Forlag, Trondheim, Norge, second edition, 1977.
- [32] [James M. Hyslop](#). Infinite series. In [Alexander C. Aitken](#) and [Daniel E. Rutherford](#), editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, fifth edition, 1959.
- [33] [D. C. Joyce](#). Survey of extrapolation processes in numerical analysis. *SIAM Review*, 13(4):435–490, 1971.
- [34] [Thomas W. Judson](#). *Abstract Algebra*. The Prindle, Weber & Schmidt Series in Advanced Mathematics. PWS Publishing Company, Boston, Massachusetts, 1994.
- [35] [Achim Jung](#). A short introduction to the lambda calculus, mar 2004.
- [36] [Stephen C. Kleene](#). Kurt Gödel: 1906–1978. In *Biographical Memoir*. National Academy of Sciences, Washington, D. C., 1987.
- [37] [Jerzy Kocik](#). On geometry of phenomenological thermodynamics. In *Symmetries in Science II*, ([B. Gruber](#), ed.), *Plenum*, New York, pages 279–287, 1986.
- [38] [Erwin Kreyszig](#). *Advanced Engineering Mathematics*. John Wiley & Sons, New York, fourth edition, 1979.

- [39] [Morten Rode Kristensen](#). Parameter estimation in nonlinear dynamical systems. Master's thesis, Department of Chemical Engineering, Technical University of Denmark, jun 2004.
- [40] [A. V. Levy](#) and A. Montalvo. The tunneling algorithm for the global minimization of functions. *SIAM J. Sci. Stat. Comput.*, 6(1):15–29, Jan 1985.
- [41] [Lixiang Li](#), Yixian Yang, and Haipeng Peng. Computation of multiple global optima through chaotic ant swarm. *Chaos, Solitons and Fractals*, 40:1399–1407, 2009.
- [42] [S. Liu](#) and G. Trenkler. Hadamard, Khatri–Rao, Kronecker and other matrix products. *Int. J. Inf. Sys. Sci.*, 4(1):160–177, 2008.
- [43] [J. David Logan](#). *Applied Mathematics: A Contemporary Approach*. John Wiley & Sons, New York, 1987.
- [44] [Tambs Lyche, R.](#) Funksjoner av én fri variabel. In *Lærebok i Matematisk analyse*, del I. Gyldendal Norsk Forlag, Oslo, Norge, 1940.
- [45] [Tambs Lyche, R.](#) Funksjoner av flere fri variable. In *Lærebok i Matematisk analyse*, del II. Gyldendal Norsk Forlag, Oslo, Norge, 1941.
- [46] [Tambs Lyche, R.](#) Differensiallikninger. In *Lærebok i Matematisk analyse*, del III. Gyldendal Norsk Forlag, Oslo, Norge, 1941.
- [47] [Simon J. A. Malham](#). An introduction to Lagrangian and Hamiltonian mechanics, Mar 2014.
- [48] [Cesar Martínez-Garza](#). Generalized quasilinearization versus Newton's method. *Journal of Mathematics Research*, 2(3), Aug 2010.
- [49] [Henry Margenau](#) and George Moseley Murphy. *The Mathematics of Chemistry and Physics*. D. van Nostrand Company, Inc., New York, 1951.
- [50] [Jon Mathews](#) and R. L. Walker. *Mathematical Methods of Physics*. The Benjamin/Cummings Publishing Company, Menlo Park, California, second edition, 1970.
- [51] [Melissa Moore](#). Pierre-Simon Laplace—biography paper. MATH 4010, University of Colorado, Denver, 2008.
- [52] [Joel Moses](#). Algebraic simplification: A guide for the perplexed. *Commun. ACM*, 14(8):527–537, 1971.
- [53] [Roman F. Nalewajski](#) and Robert G. Parr. Legendre transforms and Maxwell relations in density functional theory. *J. Chem. Phys.*, 77(1):399–407, Jul. 1982.
- [54] [Chi-Kong Ng](#), Duan Li, and Lian-Sheng Zhang. Global descent method for global optimization. *SIAM J. Optim.*, 80(6):3161–3184, 2010.
- [55] [J. J. O'Connor](#) and E. F. Robertson. Carl Gustav Jacob Jacobi. MacTutor History of Mathematics archive, University of St Andrews, jan 2000.
- [56] [J. J. O'Connor](#) and E. F. Robertson. Pierre-Simon Laplace. MacTutor History of Mathematics archive, University of St Andrews, jan 1999.
- [57] [J. J. O'Connor](#) and E. F. Robertson. Adrien-Marie Legendre. MacTutor History of Mathematics archive, University of St Andrews, jan 1999.
- [58] [L. R. Petzold](#). Order results for implicit Runge–Kutta methods applied to differential/algebraic systems. *SIAM J. Numer. Anal.*, 23(4):837–852, 1986.
- [59] [E. G. Phillips](#). Functions of a complex variable. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, eighth edition, 1957.
- [60] [William H. Press](#), Saul A. Teukolsky, William T. Vetterling, and Brian P. Flannery. *Numerical Recipes in FORTRAN 77: The Art of Scientific Computing*. Cambridge University Press, Cambridge, London, second edition, 2001.
- [61] [Lennart Råde](#) and Bertil Westergren. *BETA Mathematics Handbook*. Studentlitteratur, Lund, Sweden, second edition, 1990.
- [62] [Renuka Ravindran](#), C. R. Pranesachar, and D. P. Patil. Joseph Louis Lagrange (1736–1813). *Resonance*, pages 2–4, apr 2006.
- [63] [Patrick Reany](#). Structured differentiation, 1987.
- [64] [Patrick Reany](#). A structured differentiation for physicists (revised). *Arizona Journal of Natural Philosophy (AJNP)*, 4:4–20, jan 1992.
- [65] [Patrick Reany](#). More structured differentiation. *Arizona Journal of Natural Philosophy (AJNP)*, 8:16–32, apr 1996.
- [66] [Werner C. Rheinboldt](#). Differential–algebraic systems as differential equations on manifolds. *Math. of Comput.*, 43(168):473–482, Oct. 1984.
- [67] [Werner W. Rogosinski](#). Volume and integral. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, 1952.
- [68] [Walter Rudin](#). *Principles of Mathematical Analysis*. McGraw-Hill Kogakusha, Ltd., Tokyo, 1976.
- [69] [Nobuku Sagara](#) and Masao Fukushima. An efficient predictor-corrector method for solving nonlinear equations. *Journal of Computational and Applied Mathematics*, 19:343–349, 1987.

- [70] R. S. Schechter. *The Variational Method in Engineering*. McGraw-Hill Chemical Engineering Series. McGraw-Hill Book Company, Inc., New York, 1967.
- [71] M. J. Sewell. On Legendre transformations and elementary catastrophes. *Math. Proc. Camb. Phil. Soc.*, 82:147–163, 1977.
- [72] M. J. Sewell. On Legendre transformations and umbilic catastrophes. *Math. Proc. Camb. Phil. Soc.*, 83:273–288, 1978.
- [73] Janak Raj Sharma and Himani Arora. An efficient derivative free iterative method for solving systems of nonlinear equations. *Applicable Analysis and Discrete Mathematics*, 7(2):390–403, 2013.
- [74] Janak Raj Sharma and Himani Arora. On efficient weighted-Newton methods for solving systems of nonlinear equations. *Applied Mathematics and Computation*, 222:497–506, 2013.
- [75] Janak Raj Sharma and Puneet Gupta. An efficient fifth order method for solving systems of nonlinear equations. *Computers and Mathematics with Applications*, 67:591–601, 2014.
- [76] George F. Simmons. Introduction to topology and modern analysis. In William Ted Martin and E. H. Spanier, editors, *International Series in Pure and Applied Mathematics*. McGraw-Hill Book Company, Inc., New York, 1963.
- [77] M. A. Soliman. On the convergence acceleration of iterative processes. *Comput. Chem. Eng.*, 9(1):97–98, 1985.
- [78] Michael Spivak. *Calculus on Manifolds: A Modern Approach to Classical Theorems of Advanced Calculus*. Addison-Wesley Publishing Company, Reading, Massachusetts, 1965.
- [79] Knut Sydsæter, Atle Seierstad, and Arne Strøm. *Matematisk analyse. Bind II*. Universitetsforlaget, Oslo, Norge, third edition, 1996.
- [80] Thomas, Jr., George B. *Calculus and Analytical Geometry*. Addison-Wesley Series in Mathematics. Addison-Wesley Publishing Company, Reading, Massachusetts, 1972.
- [81] David J. Thomas and Judith M. Smith. Joseph Raphson, F. R. S. *Notes Rec. R. Soc. Lond.*, 44(5):151–167, 1990.
- [82] Hrishikesh D. Vinod. Matrix algebra, class notes (part 1). Fordham University, New York, 1998.
- [83] Hrishikesh D. Vinod. Matrix algebra, class notes (part 2). Fordham University, New York, 1998.
- [84] Herbert S. Wilf. *Generatingfunctionology*. Academic Press, Inc., Orlando, Florida, 1994.
- [85] D.-H. Xu, A. Danesh, and A. C. Todd. An accelerated successive substitution method for calculation of saturation pressure of multicomponent fluids. *Fluid Phase Equilib.*, 72:15–24, 1992.
- [86] Daniel Zwillinger. *Handbook of Differential Equations*. Academic Press, Inc., Orlando, Florida, second edition, 1992.

## Optimization (OptMet)

Theory and applications of optimization, and (non)linear programming.

**Related keywords:** optimization.

### Bibliography.

- [1] Masanao Aoki. Introduction to optimization techniques: Fundamentals and applications of nonlinear programming. In Gerald Estrin, editor, *Macmillan Series in Applied Computer Sciences*. The Macmillan Company, New York, 1971.
- [2] Christopher W. Bale and A. D. Pelton. Optimization of binary thermodynamic and phase diagram data. *Metall. Trans. B*, 14B:77–83, Mar. 1983.
- [3] Yonathan Bard. *Nonlinear Parameter Estimation*, appendix A, pages 287–309. Academic Press, Inc., Orlando, Florida, 1974.
- [4] James V. Beck and Kenneth J. Arnold. *Parameter Estimation in Engineering and Science*, chapter 7, pages 334–379. Wiley series in probability and mathematical statistics. John Wiley & Sons, New York, 1977.
- [5] James V. Beck and Kenneth J. Arnold. *Parameter Estimation in Engineering and Science*. Wiley series in probability and mathematical statistics. John Wiley & Sons, New York, 1977.
- [6] Lorenz T. Biegler and Ignacio E. Grossmann. Retrospective on optimisation. *Comput. Chem. Eng.*, 28:1169–1192, 2004.
- [7] Branin, Jr., Franklin H. and Stanley K. Hoo. A method for finding multiple extrema of a function of  $n$  variables. In F. A. Lootsma, editor, *Numerical Methods for Non-Linear Optimization*. Academic Press, Inc., Orlando, Florida, 1972. Papers presented at the Conference on Numerical Methods for Non-Linear Optimization, held at the University of Dundee (Scotland) 28th June to 1st July 1971.
- [8] Branin, Jr., F. H. Widely convergent method for finding multiple solutions of simultaneous nonlinear equations. *IBM J. Res. Dev.*, pages 504–522, Sep. 1972.

- [9] Branin, Jr., Franklin H. and Stanley K. Hoo. A method for finding multiple extrema of a function of  $n$  variables. In F. A. Lootsma, editor, *Numerical Methods for Non-Linear Optimization*, pages 231–237, Orlando, Florida, 1972. Academic Press, Inc.
- [10] Theodore W. C. Chen and Vassilios S. Vassiliadis. Solution of nonlinear optimization problems using the penalty/modified barrier method with the use of exact Hessians. *Comput. Chem. Eng.*, 27:501–525, 2003.
- [11] Hern-Shann Chen and Mark A. Stadtherr. A modification of Powell’s dogleg method for solving systems of nonlinear equations. *Comput. Chem. Eng.*, 5(3):143–150, 1981.
- [12] C. R. Corles. The use of regions of attraction to identify global minima. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 55–95, Amsterdam, Oct. 1974. North-Holland Publishing Company.
- [13] L. C. Dixon, J. Gomulka, and G. P. Szegö. Towards global optimization. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 29–54, Amsterdam, Oct. 1974. North-Holland Publishing Company.
- [14] J. C. Dunn. Newton’s method and the Goldstein step-length rule for constrained minimization problems. *SIAM J. Control and Optimiz.*, 18(6):659–674, 1980.
- [15] Everett, III, Hugh. Generalized Lagrange multiplier method for solving problems of optimum allocation of resources. *Operations Research*, 11(3):399–417, 1963.
- [16] R. Fletcher. *Practical Methods of Optimization*. Wiley Interscience. John Wiley & Sons, New York, second edition, 1978.
- [17] Saul B. Gelfand and Sanjoy K. Mitter. Metropolis-type annealing algorithms for global optimization in  $\mathbb{R}^d$ . *SIAM J. Control and Optimiz.*, 31(1):111–131, Jan. 1993.
- [18] Philip E. Gill and Walter Murray. Newton-type methods for unconstrained and linearly constrained optimization. *Math. Prog.*, 7:311–350, 1974.
- [19] Philip E. Gill, Walther Murray, and Margaret H. Wright. *Practical Optimization*. Academic Press, Inc., Orlando, Florida, 1981.
- [20] A. A. Goldstein. On steepest descent. *J. Soc. Ind. Appl. Math., Ser. A*, 3:147–151, 1965.
- [21] A. A. Goldstein. On Newton’s method. *Numer. Math.*, 7:391–393, 1965.
- [22] A. A. Goldstein and J. F. Price. On descent local minima. *Math. Comput.*, 25:569–574, 1971.
- [23] J. Gomulka. Remarks on Branin’s method for solving nonlinear equations. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 96–106, Amsterdam, Oct. 1974. North-Holland Publishing Company.
- [24] J. Gomulka. Two implementations of Branin’s method: Numerical experience. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization 2*, pages 151–163, Amsterdam, 1978. North-Holland Publishing Company.
- [25] John Greenstadt. On the relative efficiencies of gradient methods. *Math. Comput.*, 21:360–367, 1967.
- [26] Ignacio E. Grossmann and Lorenz T. Biegler. Part II. Future perspective on optimization. *Comput. Chem. Eng.*, 28:1193–1218, 2004.
- [27] J. W. Hardy. An implemented extension of Branin’s method. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 117–142, Amsterdam, Oct. 1974. North-Holland Publishing Company.
- [28] C. den Heijer and W. C. Rheinboldt. On steplength algorithms for a class of continuation methods. *SIAM J. Numer. Anal.*, 18(5):925–948, Oct. 1981.
- [29] K. Ichida and Y. Fujii. Multicriterion optimization using interval analysis. *Computing*, 44:47–57, 1990.
- [30] C. M. Keller, A. V. Cabot, and B. G. Flury. Two algorithms for global optimization. *Mathl. Comput. Modelling*, 21(12):47–59, 1995.
- [31] Hervé Lamure and Dominique Michelucci. Solving geometric constraints by homotopy. *IEEE Trans. Visual. Computer Graphics*, 2(1):28–34, Mar. 1996.
- [32] Victor J. Law and Robert H. Fariss. Sufficiency conditions for constrained extrema. *AIChE J.*, 17(2):425–427, mar 1971.
- [33] V. J. Law and R. H. Fariss. Transformational discrimination for unconstrained optimization. *Ind. Eng. Chem. Fundam.*, 11(2):154–161, 1972.
- [34] Angelo Lucia and Yang Feng. Multivariate terrain methods. *AIChE J.*, 49(10):2553–2563, oct 2003.
- [35] David G. Luenberger. *Linear and Nonlinear Programming*. Addison-Wesley Publishing Company, Reading, Massachusetts, second edition, 1989.
- [36] Mitsunori Makino. A priori estimation of Newton type homotopy method for calculating an optimal solution of convex optimization problem. *IEICE Trans. Fundam. Electronics Commun. Comput.*, E78-A(10):1339–1344, Oct. 1995.
- [37] M. Minoux. *Mathematical Programming. Theory and Algorithms*. John Wiley & Sons, New York, 1986. Reprint of 1983 edition published by Bordas Dunod Gauthier, Villars, Paris 1983. Original title *Programmation Mathématique: Théorie et Algorithmes*. Translated from the French by Steven Vajda.

- [38] Jorge J. Moré and Danny C. Sorensen. On the use of directions of negative curvature in a modified Newton method. *Math. Prog.*, 16:1–20, 1979.
- [39] W. Murray. Second derivative methods. In W. Murray, editor, *Numerical Methods for Unconstrained Optimization*, chapter 4. Academic Press, Inc., Orlando, Florida, 1972.
- [40] J. L. Nazareth and Liqun Qi. Globalization of Newton’s method for solving non-linear equations. *Num. Lin. Alg. Appl.*, 3(3):239–249, 1996.
- [41] F. Michael Rabinowitz. Algorithm 744: A stochastic algorithm for global optimization with constraints. *ACM Trans. Math. Software*, 21(2):194–213, Jun. 1995.
- [42] S. S. Rao. *Optimization: Theory and Applications*. Wiley Eastern Limited, New Delhi, 1978.
- [43] Werner C. Rheinboldt. Solution fields of nonlinear equations and continuation methods. *SIAM J. Numer. Anal.*, 17(2):221–237, Apr. 1980.
- [44] Werner C. Rheinboldt. Numerical analysis of continuation methods for nonlinear structural problems. *Comput. & Struct.*, 13:103–113, 1981.
- [45] M. J. Sewell. *Maximum and Minimum Principles. A Unified Approach with Applications*. Cambridge University Press, Cambridge, London, 1987.
- [46] Harold W. Sorenson. Parameter estimation: Principles and problems. In Jerry M. Mendell, editor, *Control and Systems Theory*. MarcelDekker, Inc., New York, 1980.
- [47] Danny C. Sorensen. Newton’s method with a model trust region modification. *SIAM J. Numer. Anal.*, 19(2):409–426, Apr. 1982.
- [48] Peter Spellucci. *Numerische Verfahren der nichtlinearen Optimierung*. Internationale Schriftenreihe zur numerischen Mathematik: Lehrbuch. Birkhäuser Verlag, Basel, 1993.
- [49] C. P. Stephens and W. Baritomba. Global optimization requires global information. *J. Optimization Theory Appl.*, 96(3):575–588, Mar. 1998.
- [50] G. Treccani. On the convergence of Branin’s method: A counter example. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 107–116, Amsterdam, Oct. 1974. North-Holland Publishing Company.
- [51] G. Treccani. A new strategy for global minimization. In L. C. W. Dixon and G. P. Szegö, editors, *Towards Global Optimization*, pages 143–148, Amsterdam, Oct. 1974. North-Holland Publishing Company.

## Linear Algebra (LinAlg)

Numerical linear algebra and numerical tensor & vector calculus.

**Related keywords:** linear & algebra | matrix | tensor | vector & space.

### Bibliography.

- [1] K. Åhlander. Einstein summation for multidimensional arrays. *Computers and Mathematics with Applications*, 44(8):1007–1017, 2002.
- [2] Daniel G. Antzoulatos and Alexander A. Sawchuk. Hypermatrix algebra: Theory. *CVGIP: Image Understanding*, 57(1):24–41, Jan. 1993.
- [3] Rutherford Aris. *Vectors, Tensors and the Basic Equations of Fluid Mechanics*. Dover Publications, Inc., New York, 1989. Reprint of 1962 edition published by Prentice-Hall, Inc., Englewood Cliffs, New Jersey.
- [4] Alan H. Barr. The Einstein summation notation: Introduction to cartesian tensors and extensions to the notation, 1989. In SIGGRAPH 89 Course notes #30 on Topics in Physically Based Modeling, pages J1–J12. Lecture notes from Californial Institute of Technology.
- [5] Stephen Barnett. *Matrices: Methods and Applications*. Oxford Applied Mathematics and Computing Science Series. Clarendon Press, Oxford, 1990.
- [6] Gerhard Berge. *Vektor og tensoranalyse*. Matematisk institutt, Universitetet i Bergen, 1988.
- [7] R. A. Brualdi and J. Csima. Small matrices of large dimension. *Linear Algebra and its Applications*, 150:227–241, 1991.
- [8] Miroslav Fiedler. Matrix inequalities. *Numer. Math.*, 9:109–119, 1966.
- [9] V. N. Gorbuzov. Autonomy of a system of equations in total differentials. *Differ. Equations*, 34(2):153–160, 1998.
- [10] Alexander Graham. *Kronecker Products and Matrix Calculus: with Applications*. Ellis Horwood Limited, Chichester, England, 1981.
- [11] Harald Hanche-Olsen. Vektorrom og lineæravbildninger, 1995/96.
- [12] Harald Hanche-Olsen. The derivative of a determinant, 1997.
- [13] Pascal Joly. Solution of systems of linear algebraic equations. In *Encyclopedia of Life Support Systems (EOLSS)*, theme Computational Methods and Algorithms. UNESCO, Eolss Publishers, Oxford ,UK, 2000. Honorary Theme Editors: Vladimir V. Shaidurov and Olivier Pironneau.

- [14] Dong Kyue Kim, Yoo Ah Kim, and Kunsoo Park. Generalizations of suffix arrays to multi-dimensional matrices. *Theoretical Computer Science*, 302:223–238, 2003.
- [15] Serge Lang. *Linear Algebra*. Springer-Verlag, New York, third edition, 1987.
- [16] Terry Lawson. *Linear Algebra*. John Wiley & Sons, New York, 1996.
- [17] Shuangzhe Liu and Heinz Neudecker. Several matrix Kantorovich-type inequalities. *J. Math. Anal. Appl.*, 197:23–26, 1996.
- [18] Shuangzhe Liu, W. Polasek, and H. Neudecker. Equality conditions for matrix Kantorovich-type inequalities. *J. Math. Anal. Appl.*, 212:517–528, 1997.
- [19] Shuangzhe Liu and Heinz Neudecker. A survey of Cauchy–Schwarz and Kantorovich-type matrix inequalities. *Stat. Pap.*, 40:55–73, 1999.
- [20] Marvin Marcus and Henryk Minc. *A Survey of Matrix Theory and Matrix Inequalities*. Allyn and Bacon, Inc., Boston, Massachusetts, 1964.
- [21] Vlastimil Pták. The Kantorovich inequality. *Am. Math. Monthly*, 102:820–821, Nov. 1995.
- [22] D. E. Rutherford. Vector methods. In Alexander C. Aitken and Daniel E. Rutherford, editors, *University Mathematical Texts*. Oliver and Boyd Ltd., Edinburgh, Scotland, eighth edition, 1954.
- [23] A. V. Smirnov. Introduction to tensor calculus, 2003.
- [24] Gilbert Strang. *Linear Algebra and Its Applications*. Academic Press, Inc., Orlando, Florida, second edition, 1980.
- [25] B. Wissman. Introduction to tensors, covariant differentiation and applications, oct 2003.

## CHAPTER 6

# Numerical Calculations

## Numerical Methods (NumMet)

Applied & Engineering numerical methods.

**Related keywords:** numerical & (method | mathematics | algorithm).

### Bibliography.

- [1] A. B. Coon and M. A. Stadtherr. Generalized block-tridiagonal matrix orderings for parallel computation in process flowsheeting. *Comput. Chem. Eng.*, 19(6/7):787–805, 1995.
- [2] Iain S. Duff and Jennifer A. Scott. A parallel direct solver for large sparse highly unsymmetric linear systems. *ACM Transactions on Mathematical Software*, 30(2):95–117, jun 2004.
- [3] I. S. Duff and J. K. Reid. The design of MA48: A code for the direct solution of sparse unsymmetric linear systems of equations. *ACM Transactions on Mathematical Software*, 22(2):187–226, 1996.
- [4] Alan George and Esmond Ng. An implementation of Gaussian elimination with partial pivoting for sparse systems. *SIAM J. Sci. Stat. Comput.*, 6(2):390–409, apr 1985.
- [5] Alan George and Esmond Ng. Symbolic factorization for sparse Gaussian elimination with partial pivoting. *SIAM J. Sci. Stat. Comput.*, 8(6):877–898, nov 1987.
- [6] Alan George, Joseph Liu, and Esmond Ng. A data structure for sparse *qr* and *lu* factorizations. *SIAM J. Sci. Stat. Comput.*, 9(1):100–121, jan 1988.
- [7] Claes Johnson. *Numerical Solution of Partial Differential Equations by the Finite Element Method*. Studentlitteratur, Lund, Sweden, 1987.
- [8] X. Joulia, B. Koehret, and M. Enjalbert. Simulateur modulaire sequentiel a convergence simultane. *Chem. Eng. J. (Lausanne)*, 30:113–127, 1985.
- [9] K. Kubota. Matrix inversion algorithms by means of automatic differentiation. *Appl. Math. Lett.*, 7(4):19–22, 1994.
- [10] Harry M. Markowitz. The elimination form of the inverse and its application to linear programming. *Management Science*, 3(3):255–269, apr 1957.
- [11] William H. Press, Saul A. Teukolsky, William T. Vetterling, and Brian P. Flannery. *Numerical Recipes in C. The Art of Scientific Computing*. Cambridge University Press, Cambridge, London, second edition, 1992.
- [12] Robert Schreiber. A new implementation of sparse Gaussian elimination. *ACM Transactions on Mathematical Software*, 8(3):256–276, sep 1982.
- [13] Zahari Zlatev. On some pivotal strategies in Gaussian elimination by sparse technique. *SIAM J. Numer. Anal.*, 17(1):18–30, feb 1980.

## Calculating Phase Diagrams (CALPHAD)

So-called CALPHAD assessments.

**Related keywords:** (assessed | assessment) & (phase & diagrams) & (solid | liquid | vapor | vapour) | (metal | ceramic | salt).

### Bibliography.

- [1] R. G. J. Ball, M. A. Mignanelli, T. I. Barry, and J. A. Gisby. The calculation of phase equilibria of oxide core-concrete systems. *J. Nucl. Mater.*, 201:238–249, 1993.
- [2] T. I. Barry and A. T. Dinsdale. Thermodynamics of metal – gas – liquid reactions. *Mater. Sci. Technol.*, 3:501–512, Jul 1987.
- [3] Thomas I. Barry, Alan T. Dinsdale, and John A. Gisby. Predictive thermochemistry and phase equilibria of slags. *JOM*, pages 32–38, Apr 1993.
- [4] Patrice Chartrand and Arthur D. Pelton. Thermodynamic evaluation and optimization of the LiCl–NaCl–KCl–RbCl–CsCl–MgCl<sub>2</sub>–CaCl<sub>2</sub> system using the modified quasi-chemical model. *Metall. Mater. Trans. A*, 32A:1361–1383, Jun. 2001.

- [5] Patrice Chartrand and Arthur D. Pelton. Thermodynamic evaluation and optimization of the LiF–NaF–KF–MgF<sub>2</sub>–CaF<sub>2</sub> system using the modified quasi-chemical model. *Metall. Mater. Trans. A*, 32A:1385–1396, Jun. 2001.
- [6] Patrice Chartrand and Arthur D. Pelton. Thermodynamic evaluation and optimization of the Li,Na,K,Mg,Ca // F,Cl reciprocal system using the modified quasi-chemical model. *Metall. Mater. Trans. A*, 32A:1417–1430, Jun. 2001.
- [7] Tim Chart, Frances Putland, and Alan Dinsdale. Calculated phase equilibria for the Cr–Fe–Ni–Si system—I ternary equilibria. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 4(1):27–46, 1980.
- [8] Hun-Jae Chung, Jae-Hyoek Shim, and Dong Nyung Lee. Thermodynamic evaluation and calculation of phase equilibria of the Ti–Mo–C–N quaternary system. *J. Alloys Compd.*, 282:142–148, 1999.
- [9] Yuwen Cui, Xiaogang Lu, and Zhanpeng Jin. Experimental study and thermodynamic assessment of the Ni–Mo–Ta ternary system. *Metall. Mater. Trans. A*, 30(11):2735–2744, Nov. 1999.
- [10] R. H. Davies, A. T. Dinsdale, T. G. Chart, T. I. Barry, and M. H. Rand. Application of MTDATA to the modeling of multicomponent equilibria. *High Temp. Sci.*, 26:251–262, 1990.
- [11] A. T. Dinsdale, J. A. Gisby, T. I. Barry, and A. L. Davies. Predictive thermochemistry and phase equilibria for slags and other oxide systems. *International Minerals and Metals Technology*, pages 150–155, 1995.
- [12] P. Dörner, L. J. Gauckler, H. Krieg, H. L. Lukas, G. Petzow, and J. Weiss. On the calculation and representation of multicomponent systems. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 3(4):241–257, 1979.
- [13] P. Fellner, I. Horsák, I. Košťenská, A. Žúžiová, and A. Michalíová. Thermodynamic analysis and calculation of the ternary phase diagram of the system water–urea–ammonium nitrate. *Chem. Papers*, 42(6):721–730, 1988.
- [14] Frederic H. Hayes, Hans Leo Lukas, Günter Effenberg, and Günter Petzow. Thermodynamic calculation of the Al-rich corner of the Al–Ti–B system. *Z. Metallkd.*, 80(5):361–365, 1989.
- [15] Mats Hillert and L.-I. Staffansson. An analysis of the phase equilibria in the Fe–FeS system. *Metall. Trans. B*, 6:37–41, 1975.
- [16] Weiming Huang, Mats Hillert, and Xizhen Wang. Thermodynamic assessment of the CaO–MgO–SiO<sub>2</sub> system. *Metall. Mater. Trans. A*, 26A:2293–2310, Sep. 1995.
- [17] Wang Jixin. Thermodynamic optimization for Al–La system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 18(3):269–272, 1994.
- [18] Larry Kaufman. Calculation of multicomponent ceramic phase diagrams. *Physica B (Amsterdam)*, 150:99–114, 1988.
- [19] P. Liang, H. J. Seifert, H. L. Lukas, G. Ghosh, G. Effenberg, and F. Aldinger. Thermodynamic modelling of the Cu–Mg–Zn ternary system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 22(4):527–544, 1998.
- [20] H. Liang and Y. A. Chang. Thermodynamic modeling of the Nb–Si–Ti ternary system. *Intermetallics*, 7:561–570, 1999.
- [21] V. A. Lysenko. Thermodynamic calculation of phase equilibria in the Ba–Y–O system. *Inorg. Mater. (Transl. of Neorg. Mater.)*, 35(11):1163–1168, 1999.
- [22] Henrik Nicolaisen, Peter Rasmussen, and Jens M. Sørensen. Correlation and prediction of mineral solubilities in the reciprocal salt system (Na<sup>+</sup>, K<sup>+</sup>)(Cl<sup>-</sup>, SO<sub>4</sub><sup>2-</sup>)–H<sub>2</sub>O at 0–100°C. *Chem. Eng. Sci.*, 48(18):3149–3158, sep 1993.
- [23] Chang-Seok Oh and Dong Nyung Lee. Thermodynamic assessment of the Ga–Te system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 16(3):317–330, 1992.
- [24] H. Okamoto, D. J. Chakrabarti, D. E. Laughlin, and T. B. Massalski. The Au–Cu (gold–copper) system. *Bull. Alloy Phase Diagrams*, 8(5):454–470, 1987.
- [25] A. D. Pelton and C. W. Bale. Computational techniques for the treatment of thermodynamic data in multicomponent systems and the calculation of phase equilibria. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 1(3):253–273, 1977.
- [26] Frances Putland, Tim Chart, and Alan Dinsdale. Thermodynamically calculated phase diagrams for the Co–Cr–Ta and Co–Cr–Nb systems. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 4(2):133–141, 1980.
- [27] S. Ravindra Reddy and J. P. Hajra. Thermodynamics and phase equilibria in the In–Sb system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 18(1):39–45, 1994.
- [28] N. Saunders and A. P. Miodownik. *CALPHAD. Calculation of Phase Diagrams: A Comprehensive Guide*. Pergamon Materials Series. Pergamon Press, Oxford, London, 1998.
- [29] Jain-Yun Shen, Christian Chatillon, Ibrahim Ansara, Andy Watson, Bridget Rugg, and Tim Chart. Optimisation of the thermodynamic and phase diagram data in the ternary As–Ga–In system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 19(2):215–226, 1995.

- [30] A. K. Singh and Romesh C. Sharma. Phase equilibria calculation of Zn–Cd–Te system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 16(2):161–172, 1992.
- [31] S. Srikanth and A. Petric. Optimization and calculation of the Fe–Ta phase diagram. *J. Alloys Compd.*, 203:281–288, Jul. 1994.
- [32] J. R. Taylor and A. T. Dinsdale. Thermodynamic and phase diagram data for the CaO–SiO<sub>2</sub> system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 14(1):71–88, 1990.
- [33] Jeffrey R. Taylor and Alan T. Dinsdale. A thermodynamic assessment of the Ni–O, Cr–O and Cr–Ni–O systems using the ionic liquid and compound energy models. *Z. Metallkd.*, 81(5):354–366, 1990.
- [34] J. R. Taylor, A. T. Dinsdale, M. Hillert, and M. Selleby. A critical assessment of thermodynamics and phase diagram data for the Al–O system. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 16(2):173–179, 1992.
- [35] Jeffrey R. Taylor and Alan T. Dinsdale. A thermodynamic assessment of the Cr–Fe–O system. *Z. Metallkd.*, 84(5):335–345, 1993.
- [36] Zhang Zhaochun, Su Hang, Wu Zhu, Chen Nianyi, and Peng Ruiwu. Calculation of thermodynamic properties from the miscibility gap in the phase diagram of Zn–Pb system by means of NRTL equation. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 22(3):313–322, 1998.
- [37] Y. Zuo and Y. A. Chang. Thermodynamic calculation of the Al–Mg phase diagram. *CALPHAD: Comput. Coupling Phase Diagrams Thermochem.*, 17(2):161–174, 1993.

### Flowsheet calculations (Flowsheet)

Chemical process calculations.

**Related keywords:** flowsheets | networks | solvers | chemical & processes.

#### Bibliography.

- [1] A. R. J. Arendsen and G. F. Versteeg. Dynamic thermodynamics with internal energy, volume, and amount of moles as states: Application to liquified gas tank. *Ind. Eng. Chem. Res.*, 48:3167–3176, 2009.
- [2] A. Barrett and J. J. Walsh. Improved chemical process simulation using local thermodynamic approximations. *Comput. Chem. Eng.*, 3:397–402, 1979.
- [3] F. Berger and F. A. Perris. Flowpack II—A new generation of system for steady state process flowsheeting. *Comput. Chem. Eng.*, 3:309–317, 1979.
- [4] L. T. Biegler. On the simultaneous solution and optimization of large scale engineering systems. *Comput. Chem. Eng.*, 12(5):357–369, 1988.
- [5] R. Bogusch, B. Lohmann, and W. Marquardt. Computer-aided process modeling with MODKIT. *Comput. Chem. Eng.*, 25:963–995, 2001.
- [6] I. D. L. Bogle and J. D. Perkins. Sparse Newton-like methods in equation oriented flowsheeting. *Comput. Chem. Eng.*, 12(8):791–805, 1988.
- [7] R. Bogusch and W. Marquardt. A formal representation of process model equations. *Comput. Chem. Eng.*, 21(10):1105–1115, 1997.
- [8] Stefan Brüggermann, Jan Oldenburg, Ping Zhang, and Wolfgang Marquardt. Robust dynamic simulation of three-phase reactive batch distillation columns. *Ind. Eng. Chem. Res.*, 43:3672–3684, 2004.
- [9] Marcelo Castier. Dynamic simulation of fluids in vessels via entropy maximization. *J. Ind. & Eng. Chem.*, 16:122–129, 2010.
- [10] S. Charton, V. Blet, and J. P. Corriou. A simplified model for real gas expansion between two reservoirs connected by a thin tube. *Chem. Eng. Sci.*, 51(2):295–308, 1996.
- [11] H. N. Cofer and M. A. Stadtherr. Reliability of iterative linear equation solvers in chemical process simulation. *Comput. Chem. Eng.*, 20(9):1123–1132, 1996.
- [12] Jan Dudczak. Optimal structuring of modular computations of chemical engineering systems—The general strategy. *Comput. Chem. Eng.*, 10(1):7–18, 1986.
- [13] L. B. Evans, J. F. Boston, H. I. Britt, P. W. Gallier, P. K. Gupta, B. Joseph, V. Mahalic, E. Ng, W. D. Seider, and H. Yagi. Aspen: An advanced system for process engineering. *Comput. Chem. Eng.*, 3:319–327, 1979.
- [14] J. C. Fagley and B. Carnahan. The sequential-clustered method for dynamic chemical plant simulation. *Comput. Chem. Eng.*, 14(2):161–177, 1990.
- [15] F. M. Gonçalves, M. Castier, and O. Q. F. Araújo. Dynamic simulation of flash drums using rigorous physical property calculations. *Braz. J. Chem. Eng.*, 24(2):277–286, Apr–Jun 2007.
- [16] E. W. Gorczynski, H. P. Hutchison, and A. R. M. Wajih. Development of a modularly organised equation-oriented process simulator. *Comput. Chem. Eng.*, 3:353–356, 1979. Paper 6A.5.
- [17] K. M. Hangos and I. T. Cameron. A formal representation of assumptions in process modelling. *Comput. Chem. Eng.*, 25:237–255, 2001.

- [18] R. Hernandez and R. W. H. Sargent. A new algorithm for process flowsheeting. *Comput. Chem. Eng.*, 3:363–371, 1979.
- [19] M. Hillestad and T. Hertzberg. Dynamic simulation of chemical engineering systems by the sequential modular approach. *Comput. Chem. Eng.*, 10(4):377–388, 1986.
- [20] M. Hillestad and T. Hertzberg. Convergence and stability of the sequential modular approach to dynamic process simulation. *Comput. Chem. Eng.*, 12(5):407–414, 1988.
- [21] H. P. Hutchison, D. J. Jackson, and W. Morton. The development of an equation-oriented flowsheet simulation and optimization package—I. The QUASILIN program. *Comput. Chem. Eng.*, 10(1):19–29, 1986.
- [22] H. P. Hutchison, D. J. Jackson, and W. Morton. The development of an equation-oriented flowsheet simulation and optimization package—II. Examples and results. *Comput. Chem. Eng.*, 10(1):31–47, 1986.
- [23] S. Kajaluoto, P. Neittaanmäki, and J. Ruhtila. Comparison of different solution algorithms for sparse linear equations arising from flowsheeting problems. *Comput. Chem. Eng.*, 13(4/5):433–439, 1989.
- [24] Kang Lee and En Sup Yoon. The flexible modular approach on dynamic process simulation. *Comput. Chem. Eng.*, 18:761–765, 1994.
- [25] Kang Wook Lee, Kang Ju Lee, Soo Hyoung Choi, and En Sup Yoon. Stochastic dynamic simulation of chemical processes with changing uncertainties. *Comput. Chem. Eng.*, 20:557–562, 1996.
- [26] Xiang Li, Zhijiang Shao, and Jixin Qian. Module-oriented automatic differentiation in chemical process systems optimization. *Comput. Chem. Eng.*, 28:1551–1561, 2004.
- [27] Eduardo R. A. Lima, Marcelo Castier, and Evaristo C. Biscaia, Jr. Differential-algebraic approach to dynamic simulations of flash drums with rigorous evaluation of physical properties. *Oil & Gas Science and Technology — Rev. IFP*, 63(5):677–686, 2008.
- [28] Angelo Lucia and Sandro Macchietto. New approach to approximation of quantities involving physical properties derivatives in equation-oriented process design. *AIChE J.*, 29(5):705–712, 1983.
- [29] W. Marquardt. Trends in computer-aided process modeling. *Comput. Chem. Eng.*, 20(6/7):591–609, 1996.
- [30] B. A. Murtagh. On the solution and optimization of large-scale engineering systems. *Comput. Chem. Eng.*, 6(1):1–5, 1982.
- [31] M. Oh and C. C. Pantelides. A modelling and simulation language for combined lumped and distributed parameter systems. *Comput. Chem. Eng.*, 20(6/7):611–633, 1996.
- [32] G. M. Ostrovsky, A. S. Ocheretovy, and T. A. Berezinsky. Approach to steady-state calculations of chemical plants. *Comput. Chem. Eng.*, 16(9):893–899, 1992.
- [33] Derya B. Özyurt and Ralph W. Pike. Theory and practice of simultaneous data reconciliation and gross error detection for chemical processes. *Comput. Chem. Eng.*, 28:381–402, 2004.
- [34] G. Pagani, A. D’Arminio Monforte, and A. de Mitri. Improving an equation-oriented package. *Comput. Chem. Eng.*, 13(8):931–945, 1989.
- [35] C. C. Pantelides, D. Gritsis, K. R. Morison, and R. W. H. Sargent. The mathematical modelling of transient systems using differential-algebraic equations. *Comput. Chem. Eng.*, 12(5):449–454, 1988.
- [36] C. C. Pantelides and P. I. Barton. Equation-oriented dynamic simulation. Current status and future perspectives. In *European Symposium on Computer Aided Process Engineering-2*, 1993.
- [37] J. D. Perkins. Efficient solution of design problems using a sequential-modular flowsheeting programme. *Comput. Chem. Eng.*, 3:375–381, 1979.
- [38] J. Perregaard. Model simplification and reduction for simulation and optimization of chemical processes. *Comput. Chem. Eng.*, 17(5/6):465–483, 1993.
- [39] P. C. Piela, T. G. Epperly, K. M. Westerberg, and A. W. Westerberg. ASCEND: An object-oriented computer environment for modeling and analysis: The modeling language. *Comput. Chem. Eng.*, 15(1):53–72, 1991.
- [40] Michel Pons and Michael Halloran. Thermodynamic and physical properties v1.1. Technical report, CAPE-OPEN, CO-LaN consortium, Apr 2005. CO Thermo 1.1 Specification, version 2.11.
- [41] R. W. H. Sargent. Introduction: 25 years of progress in process systems engineering. *Comput. Chem. Eng.*, 28:437–439, 2004.
- [42] M. Schacham, S. Macchietto, L. F. Stutzman, and P. Babcock. Equation oriented approach to process flowsheeting. *Comput. Chem. Eng.*, 6(2):79–85, 1982.
- [43] Warren D. Seider, David D. Brengel, and Soemantri Widagdo. Nonlinear analysis in process design. *AIChE J.*, 37(1):1–38, Jan. 1991.
- [44] M. Shacham, S. Macchietto, L. F. Stutzman, and P. Babcock. Review. Equation oriented approach to process flowsheeting. *Comput. Chem. Eng.*, 6(2):79–95, 1982.
- [45] Volker Siepmann. *Process Modelling on a Canonical Basis*. Dr.ing. thesis 2006:71, Norwegian Institute of Technology, Jul 2006.

- [46] Jeffrey J. Siirola. Strategic process synthesis: Advances in the hierarchical approach. *Comput. Chem. Eng.*, 20:S1637–S1643, 1996.
- [47] Mark A. Stadtherr and E. Stephen Wood. Sparse matrix methods for equation-based chemical process flowsheeting—I. *Comput. Chem. Eng.*, 8(1):9–18, 1984.
- [48] E. T. Sun and M. A. Stadtherr. Issues in nonlinear equation solving in chemical engineering. *Comput. Chem. Eng.*, 12(11):1129–1139, 1988.
- [49] J. Tolsma and P. I. Barton. Daepack: An open modeling environment for legacy models. *Ind. Eng. Chem. Res.*, 39(6):1826–1839, 2000.
- [50] John E. Tolsma, Jerry A. Clabaugh, and Paul I. Barton. Symbolic incorporation of external procedures into process modeling environments. *Ind. Eng. Chem. Res.*, 41:3867–3876, 2002.
- [51] Victor R. Vasquez and Wallace B. Whiting. Uncertainty of predicted process performance due to variations in thermodynamics model parameter estimation from different experimental data sets. *Fluid Phase Equilib.*, 142:115–130, 1998.
- [52] M. Weiss and H. A. Preisig. Structural analysis in the dynamical modelling of chemical engineering systems. *Mathematical and Computer Modelling of Dynamical Systems*, 6(4):325–364, 2000.
- [53] M. Weiss and H. A. Preisig. Simplifying hypotheses in computer-aided modelling: A singular perturbation approach. *Comput. Chem. Eng.*, 21:721–726, 1997.
- [54] Kelly R. Westman, Angelo Lucia, and David C. Miller. Flash and distillation calculations by a Newton-like method. *Comput. Chem. Eng.*, 8(3/4):219–228, 1984.
- [55] Arthur W. Westerberg and Dean R. Benjamin. Thoughts on a future equation-oriented flowsheeting system. *Comput. Chem. Eng.*, 9(5):517–526, 1985.
- [56] C. E. Wilhelm and R. E. Swaney. Robust solution of algebraic process modelling equations. *Comput. Chem. Eng.*, 18(6):511–531, 1994.
- [57] D. Wolbert, X. Joulia, B. Koehret, and L. T. Biegler. Flowsheet optimization and optimal sensitivity analysis using analytical derivatives. *Comput. Chem. Eng.*, 18(11/12):1083–1095, 1994.
- [58] A. Yang, M. Schlüter, B. Bayer, J. Krüger, E. Haberstroh, and W. Marquardt. A concise conceptual model for material data and its applications in process engineering. *Comput. Chem. Eng.*, 27:595–609, 2003.
- [59] S. E. Zitney and M. A. Stadtherr. Computational experiments in equation-based chemical process flowsheeting. *Comput. Chem. Eng.*, 12(12):1171–1186, 1988.
- [60] Stephen E. Zitney and Mark A. Stadtherr. Supercomputing strategies for the design and analysis of complex separation systems. *Ind. Eng. Chem. Res.*, 32(4):604–612, apr 1993.

### Automatic Differentiation (Diff)

Automatic (computerized) differentiation.

**Related keywords: automatic & differentiation.**

#### Bibliography.

- [1] Claus Bendtsen and Ole Stauning. TADIFF, a flexible C++ package for automatic differentiation using Taylor series expansion. Technical report, Department of mathematical modelling, Technical university of Denmark, 1997.
- [2] C. Bischof, P. Hovland, and B. Norris. Implementation of automatic differentiation tools. In *Proceedings of the 2002 ACM SIGPLAN Workshop on Partial Evaluation and Semantics-based Program Manipulation (PEPM'02)*, pages 98–107, New York, NY, USA, 2002. ACM Press.
- [3] Christian Bischof, Andreas Griewank, and David Juedes. Exploiting parallelism in automatic differentiation. In *ICS '91: Proceedings of the 5th International Conference on Supercomputing*, International Conference of Supercomputing, pages 146–153, Jun. 1991.
- [4] Christian Bischof, Alan Carle, George Corliss, Andreas Griewank, and Paul Hovland. ADIFOR—Generating derivative codes from Fortran programs. *Scientific Programming*, 1(1):1–29, 1992. ADIFOR Working Note # 1. Also appeared as Preprint MCS-P263-0991, Mathematics and Computer Science Division, Argonne National Laboratory, Argonne, Ill., September 1991, and as Technical Report 91185, Center for Research in Parallel Computation, Rice University, Houston, Tex., 1991.
- [5] Christian Bischof, G. Corliss, L. Green, Andreas Griewank, K. Haigler, and P. Newman. Automatic differentiation of advanced CFD codes for multidisciplinary design. *Computing Systems in Engineering*, 3(6):625–637, 1992.
- [6] Christian Bischof, Alan Carle, Peyvand M. Khademi, and Gordon Pusch. Automatic differentiation: Obtaining fast and reliable derivatives—Fast. In *Proceedings of the SIAM Symposium on Control Problems in Industry*, pages 1–17, San Diego, Jul. 1994.

- [7] [Christian Bischof](#) and Andreas Griewank. Computational differentiation and multidisciplinary design. In H. Engl and J. McLaughlin, editors, *Inverse Problems and Optimal Designs in Industry*, pages 187–211, Stuttgart, 1994. Teubner Verlag.
- [8] [R. Boudjemaa](#), M. G. Cox, A. B. Forbes, and P. M. Harris. Automatic differentiation techniques and their application in metrology. Technical report, Centre for Mathematics and Scientific Computing, Jun. 2003.
- [9] [Thomas F. Coleman](#) and Arun Verma. Admat-1: Automatic differentiation and matlab interface toolbox. *ACM Trans. Math. Software*, 26(1):150–176, Mar. 2000.
- [10] [Thomas F. Coleman](#) and Arun Verma. ADMAT: An automatic differentiation toolbox for MATLAB. Technical report, Computer Science Department and Center for Applied Mathematics, Cornell University, Ithaca, NY, May 1998.
- [11] [George F. Corliss](#). Automatic differentiation bibliography. Department of Mathematics, Statistics, and Computer Science, Marquette University, Milwaukee, USA, 1991.
- [12] [George F. Corliss](#) and Andreas Griewank. Operator overloading as an enabling technology for automatic differentiation. Technical report, Argonne National Laboratory, Argonne, Illinois, May 1993.
- [13] [P. Cusdin](#) and J.-D. Müller. Automatic differentiation: Learning to speak AD. Technical Report QUB-SAE-03-05, QUB School of Aeronautical Engineering, 2003.
- [14] [L. C. W. Dixon](#), Z. Maany, and M. Mohseninia. Automatic differentiation of large sparse systems. *Journal of Economic Dynamics and Control*, 14(2):299–311, 1990.
- [15] [D. Elizondo](#), B. Cappelaere, and Ch. Faure. Automatic versus manual model differentiation to compute sensitivities and solve non-linear inverse problems. *Comput. Geosci.*, 28:309–326, 2002.
- [16] [A. Griewank](#). A mathematical view of automatic differentiation. *Acta Numerica*, 12:321–398, 2003.
- [17] [Andreas Griewank](#). On automatic differentiation. Technical report, Argonne National Laboratory, Mathematics and Computer Science Division, Argonne, Illinois, Nov. 1988.
- [18] [Andreas Griewank](#). On automatic differentiation. In Masao Iri and Kunio Tanabe, editors, *Mathematical Programming: Recent Developments and Applications*. Kluwer Academic Publishers, Dordrecht, 1989.
- [19] [Andreas Griewank](#) and Shawn Reese. On the calculation of Jacobian matrices by the Markowitz rule for vertex elimination. In Andreas Griewank and George F. Corliss, editors, *Automatic Differentiation of Algorithms: Theory, Implementation, and Application*, pages 126–135, Philadelphia, 1991. SIAM.
- [20] [Andreas Griewank](#). Achieving logarithmic growth of temporal and spatial complexity in reverse automatic differentiation. *Optimiz. Methods & Software*, 1:35–54, 1992.
- [21] [Andreas Griewank](#), David Juedes, and Jean Utke. ADOL-C: A package for the automatic differentiation of algorithms written in C/C++, version 1.5, Dec. 1993.
- [22] [Andreas Griewank](#), Christian Bischof, George Corliss, Alan Carle, and Karen Williamson. Derivative convergence for iterative equation solvers. *Optimiz. Methods & Software*, 2:321–355, 1993.
- [23] [Andreas Griewank](#). Some bounds on the complexity of gradients, jacobians, and hessians. In P. M. Pardalos, editor, *Complexity in Numerical Optimization*. World Scientific Publishing Co. Pte. Ltd., Singapore, 1993.
- [24] [Andreas Griewank](#), David Juedes, and Jean Utke. ADOL-C, a package for the automatic differentiation of algorithms written in C/C++. *ACM Trans. Math. Software*, 22(2):131–167, 1996.
- [25] [John D. Hobby](#). An application for semi-automatic differentiation. In H. M. Bücker, G. Corliss, P. Hovland, U. Naumann, and B. Norris, editors, *Automatic Differentiation*, Lecture Notes in Computational Science and Engineering. Springer-Verlag, New York, 2005.
- [26] [Paul Hovland](#), Christian Bischof, Donna Spiegelman, and Mario Casella. Efficient derivative codes through automatic differentiation and interface contraction: An application in biostatistics. *SIAM Journal on Scientific Computing*, 18(4):1056–1066, 1997.
- [27] [Max E. Jerrell](#). Function minimization and automatic differentiation using C++. In Norman Meyrowitz, editor, *OOPSLA'89: Object-Oriented Programming: Systems, Languages and Applications*, pages 169–173, New Orleans, Louisiana, Oct. 1989. Association for Computing Machinery.
- [28] [Jeremy R. Johnson](#) and Anthony F. Breitzman. Automatic derivation and implementation of fast convolution algorithms. *Journal of Symbolic Computation*, 37(2):261–293, 2004.
- [29] [David W. Juedes](#). A taxonomy of automatic differentiation tools. In Andreas Griewank and George F. Corliss, editors, *Automatic Differentiation of Algorithms: Theory, Implementation, and Application*, pages 315–329, Philadelphia, PA, 1991. SIAM. Also appeared as Preprint MCS-P265-0991, Mathematics and Computer Science Division, Argonne National Laboratory, Argonne, Ill., September 1991.
- [30] [Dan Kalman](#). Doubly recursive multivariant automatic differentiation. *Mathematics Magazine*, 75(3):187–202, Jun. 2002.
- [31] [R. Kalaba](#) and L. Tesfatsion. Automatic differentiation of functions of derivatives. *Computers & Mathematics with Applications*, 12(11):1091–1103, 1986.

- [32] [Thomas Kaminski](#), Ralf Giering, Marko Scholze, Peter Rayner, and Wolfgang Knorr. An example of an automatic differentiation-based modelling system. In *Lecture Notes in Computer Science*, volume 2668. Springer-Verlag, New York, 2003.
- [33] [Jerzy Karczmarczuk](#). Functional differentiation of computer programs. *Journal of Higher-Order and Symbolic Computation*, 14(1):35–57, Mar. 2001.
- [34] [Jerzy Karczmarczuk](#). Functional differentiation of computer programs. *ACM SIGPLAN Notices*, 34(1):195–203, 1999.
- [35] [Gershon Kedem](#). Automatic differentiation of computer-programs. *ACM Trans. Math. Software*, 6(2):150–165, 1980.
- [36] [Seppo Linnainmaa](#). Taylor expansion of the accumulated rounding error. *J-BIT*, 16:146–160, 1976.
- [37] [Andrew Mauer](#). A collection of tools in support of automatic differentiation. Technical Report 185, National Laboratory, Mathematics and Computer Science Division, Argonne, Illinois, Feb. 1994.
- [38] [Leo Michelotti](#). MXYZPTLK: A practical, user-friendly C++ implementation of differential algebra: User’s guide. Technical Report FN-535, Fermi National Accelerator Laboratory, Batavia, Ill., Jan. 1990.
- [39] [Leo Michelotti](#). Automatic differentiation of limit functions. In *Particle Accelerator Conference (PAC 93)*, Washington D. C., 17-20 May 1993, pages 495–497. IEEE PAC, 1993.
- [40] [C. Mischler](#), X. Joulia, E. Hassold, A. Galligo, and R. Esposito. Automatic differentiation applications to computer aided process engineering. *Comput. Chem. Eng.*, 19:779–784, 1995.
- [41] [Richard D. Neidinger](#). Computing multivariable Taylor series to arbitrary order. *APL Quote Quad*, 25(4):134–144, 1995.
- [42] [L. B. Rall](#). The arithmetic of differentiation. *Math. Mag.*, 59(5):275–282, Dec. 1986.
- [43] [H. Q. Tan](#), S. M. Arnold, and X. Dong. Computer simulation of the mathematical modeling involved in constitutive equation development: Via symbolic computations. *Mathl. Comput. Modelling*, 14:927–932, 1990.
- [44] [Ross Taylor](#). Automatic derivation of thermodynamic property functions using computer algebra. *Fluid Phase Equilib.*, 129:37–47, 1997.
- [45] [E. Tijsskens](#), D. Roose, H. Ramon, and J. De Baerdemaeker. Fastder++, efficient automatic differentiation for non-linear pde solvers. *Math. Comput. Simul.*, 65(1–2):177–190, 2004.
- [46] [John E. Tolsma](#) and Paul I. Barton. On computational differentiation. *Comput. Chem. Eng.*, 22(4/5):475–490, 1998.
- [47] [I. Tsukanov](#) and M. Hall. Data structure and algorithms for fast automatic differentiation. *Int. J. Numer. Meth. Engng.*, 56:1949–1972, 2003.
- [48] [J. Utke](#). Openad: Algorithm implementation user guide. Technical report, Mathematics and Computer Science Division, Apr. 2004.
- [49] [Arun Verma](#). An introduction to automatic differentiation. *Curr. Sci.*, 78(7):804–807, Apr. 2000.
- [50] [D. Villard](#) and B. Arnaldi. Symbolic differentiation library for simulation of multibody rigid systems. *Math. Comput. Simul.*, 42:659–673, 1996.
- [51] [Allan D. Wittkopf](#) and Gregory J. Reid. Fast differential elimination in c: The cdiffelim environment. *Computer Physics Communications*, 139:192–217, 2001.
- [52] [Li Xiang](#), Zhong Weitao, Shao Zhijiang, and Qian Jixin. Applying extended automatic differentiation technique to process system optimization problems. In *Proceedings of the American Control Conference*, volume 5, pages 4079–4084, 2001.

## CHAPTER 7

# Experimental Data

### Standard State Properties (DatPCD)

Standard state properties for pure components.

**Related keywords:** `barin` | `janaf` | `dippr` | `etc.`

#### Bibliography.

- [1] [Robert A. Alberty](#) and Catherine A. Gehrig. Standard chemical thermodynamic properties of alkane isomer groups. *J. Phys. Chem. Ref. Data*, 13(4):1173–1197, 1984.
- [2] [I. Ansara](#). Génération et application des bases de données thermochimiques. *Entropie*, 27(161):74–79, 1991.
- [3] [L. Ya. Aranovich](#) and R. G. Berman. Optimized standard state and solution properties of minerals. II. Comparisons, predictions, and applications. *Contrib. Mineral. Petrol.*, 126:25–37, 1996.
- [4] [G. H. Aylward](#) and T. J. V. Findlay. *SI Chemical Data*. John Wiley & Sons, New York, second edition, 1974.
- [5] [Christopher W. Bale](#) and Gunnar Erikson. Metallurgical thermochemical databases—A review. *Can. Metall. Q.*, 29(2):105–132, 1990.
- [6] [Ihsan Barin](#) and Gregor Platzki. Ag–Kr. In *Thermochemical Data of Pure Substances*, volume I. VCH Verlagsgesellschaft, Weinheim, Germany, third edition, 1995.
- [7] [Ihsan Barin](#) and Gregor Platzki. La–Zr. In *Thermochemical Data of Pure Substances*, volume II. VCH Verlagsgesellschaft, Weinheim, Germany, third edition, 1995.
- [8] [R. G. Berman](#) and L. Ya. Aranovich. Optimized standard state and solution properties of minerals. I. Model calibration for olivine, orthopyroxene, cordierite, garnet, and ilmenite in the system FeO–MgO–CaO–Al<sub>2</sub>O<sub>3</sub>–TiO<sub>2</sub>–SiO<sub>2</sub>. *Contrib. Mineral. Petrol.*, 126:1–24, 1996.
- [9] [A. Burcat](#) and Branko Ruscic. Third millennium ideal gas and condensed phase thermochemical database for combustion with updates from active thermochemical tables, 2005.
- [10] [M. W. Chase](#), C. A. Davies, J. R. Downey, D. J. Frurip, R. A. McDonald, and A. N. Syverud. JANAF thermochemical tables. Third edition. Part I, Al–Co. *J. Phys. Chem. Ref. Data, Suppl.*, 14(1):1–926, 1985.
- [11] [M. W. Chase](#), C. A. Davies, J. R. Downey, D. J. Frurip, R. A. McDonald, and A. N. Syverud. JANAF thermochemical tables. Third edition. Part II, Cr–Zr. *J. Phys. Chem. Ref. Data, Suppl.*, 14(1):927–1856, 1985.
- [12] [Jing Chao](#), Kenneth R. Hall, Kenneth N. Marsh, and Randolph C. Wilhoit. Thermodynamic properties of key organic oxygen compounds in the carbon range C<sub>1</sub> to C<sub>4</sub>. Part 2. Ideal gas properties. *J. Phys. Chem. Ref. Data*, 15(4):1369–1436, 1986.
- [13] [Chase, Malcolm W., Jr.](#) NIST–JANAF thermochemical tables. Fourth edition. Part I, Al–Co. *J. Phys. Chem. Ref. Data, Suppl.*, pages 1–957, 1998. Monograph No. 9.
- [14] [Chase, Malcolm W., Jr.](#) NIST–JANAF thermochemical tables. Fourth edition. Part II, Cr–Zr. *J. Phys. Chem. Ref. Data, Suppl.*, pages 958–1951, 1998. Monograph No. 9.
- [15] [F. J. Cook](#). Analysis of specific heat data in the critical region of magnetic solids. *J. Phys. Chem. Ref. Data*, 2(1):11–24, 1973.
- [16] [J. D. Cox](#), D. D. Wagman, and V. A. Medvedev. *CODATA Key Values for Thermodynamics*. CODATA Series on Thermodynamic Properties. Hemisphere Publishing Corporation, New York, 1989. Final Report of the CODATA Task Group on Key Values for Thermodynamics.
- [17] [Eugene S. Domalski](#). Selected values of heats of combustion and heats of formation of organic compounds containing the elements C, H, N, O, P and S. *J. Phys. Chem. Ref. Data*, 1(2):221–278, 1972.
- [18] [E. S. Domalski](#), W. H. Evans, and E. D. Hearing. Heat capacities and entropies of organic compounds in the condensed phase. *J. Phys. Chem. Ref. Data*, 13(1), 1984.
- [19] [Eugene S. Domalski](#) and Elizabeth D. Hearing. Heat capacities and entropies of organic compounds in the condensed phase. Volume II. *J. Phys. Chem. Ref. Data*, 19(4):881–1048, 1990.

- [20] Eugene S. Domalski and Elizabeth D. Hearing. Heat capacities and entropies of organic oxygen compounds in the condensed phase. Volume III. *J. Phys. Chem. Ref. Data*, 25(1):1–523, 1996.
- [21] G. T. Dyos and T. Farrell. *Electrical Resistivity Handbook*. Peter Peregrinus Ltd., London, 1992.
- [22] D. A. Everest. Physical and thermodynamic properties of organic compounds. Technical Report 41, National Physical Laboratory: Division of Chemical Standards, Teddington, London, Nov. 1975.
- [23] Yingwei Fei and Surendra K. Saxena. A thermochemical data base for phase equilibria in the system Fe–Mg–Si–O at high pressure and temperature. *Phys. Chem. Miner.*, 13:311–324, 1986.
- [24] L. P. Filippov. Equation of state for water in the metastable region. *Therm. Eng. (Transl. of Teploenergetika Moscow)*, 31(3):175–177, 1984.
- [25] Armando Fernandez Guillermet and Per Gustafson. An assessment of the thermodynamic properties and the  $p, t$  phase diagram of iron. *High Temp. - High Pressures*, 16:591–610, 1985.
- [26] Haas, Jr., John L., Gilpin R. Robinson, Jr., and Bruce S. Hemingway. Thermodynamic tabulations for selected phases in the system CaO–Al<sub>2</sub>O<sub>3</sub>–SiO<sub>2</sub>–H<sub>2</sub>O at 101.325 kPa(1 atm) between 273.15 and 1800 K. *J. Phys. Chem. Ref. Data*, 10(3):575–669, 1981.
- [27] Reed A. Howald, Richard D. Jones, Ahmed B. Rezvani, Bimalendu N. Roy, Matthew J. Scanlon, and Timothy M. Swager. Temperature dependence of the bulk modulus of solids in the MgO–Al<sub>10.5</sub>–SiO<sub>2</sub> system. *J. Phys. Chem.*, 89(12):2682–2687, 1985.
- [28] Taylor Lyman. Metallography, structures and phase diagrams. In *Metals Handbook*, volume 8. American Society for Metals, Metals Park, Ohio, eighth edition, 1961.
- [29] Bonnie J. McBride, Michael J. Zehe, and Sanford Gordon. NASA Glenn coefficients for calculating thermodynamic properties of individual species. Technical report, Glenn Research Center, Cleveland, Ohio, sep 2002. NASA/TP—2002-211556.
- [30] Eric H. Oelkers, Harold C. Helgeson, Everett L. Shock, Dimitri A. Sverjensky, James W. Johnson, and Vitalii Pokrovskii. Summary of the apparent standard partial molal Gibbs free energies of formation of aqueous species, minerals, and gases at pressures 1 to 5000 bars and temperatures 25 to 1000 °C. *J. Phys. Chem. Ref. Data*, 24(4):1401–1560, 1995.
- [31] Růžička, Jr., Vlastimil, Milan Zábanský, and Vladimír Majer. Heat capacities of organic compounds in liquid state. II. C<sub>1</sub> to C<sub>18</sub>  $n$ -alkanes. *J. Phys. Chem. Ref. Data*, 20(2):405–444, 1991.
- [32] G. V. Samsonov. *Handbook of the Physicochemical Properties of the Elements*. Plenum Press, New York, 1968. Translated from Russian.
- [33] H. M. Spiers, editor. *Technical Data of Fuel*. British National Committee World Power Conference, London, fifth edition, 1952.
- [34] G. Tammann. Die Abhängigkeit des Schmelzpunktes beim Glaubersalz vom Druck. *Z. Phys. Chem., Abt. A*, 46:818–826, 1903.
- [35] Y. S. Touloukian, R. W. Powell, C. Y. Ho, and P. G. Klemens. Thermal conductivity: Metallic elements and alloys. In *Thermophysical Properties of Matter*, volume I. IFI/Plenum, New York, 1970.
- [36] H. U. von Vogel. *Chemiker-Kalender*. Springer-Verlag, New York, 1974.
- [37] Donald D. Wagman, William H. Evans, Vivian B. Parker, Richard H. Schumm, Iva Halow, Sylvia M. Balley, Kenneth L. Churney, and Ralph L. Nuttall. Selected values for inorganic and C<sub>1</sub> and C<sub>2</sub> organic substances in SI units. *J. Phys. Chem. Ref. Data, Suppl.*, 11(2):1–380, 1982.
- [38] G. K. White and M. L. Mingos. Thermophysical properties of some key solids: An update. *Int. J. Thermophys.*, 18(5):1269–1327, 1997.
- [39] M. E. Wieser. Atomic weights of the elements 2005 (IUPAC technical report). *Pure Appl. Chem.*, 78(11):2051–2066, 2006.
- [40] R. C. Wilhoit. The design of a multipurpose file of thermodynamic data. *J. Chem. Inf. Comput. Sci.*, 20:138–143, 1980.
- [41] Randolph C. Wilhoit, Jing Chao, and Kenneth R. Hall. Thermodynamic properties of key organic oxygen compounds in the carbon range C<sub>1</sub> to C<sub>4</sub>. Part 1. Properties of condensed phases. *J. Phys. Chem. Ref. Data*, 14(1):1–175, 1985.
- [42] Milan Zábanský, Vlastimil Růžička, Jr., and Vladimír Majer. Heat capacities of organic compounds in the liquid state. I. C<sub>1</sub> to C<sub>18</sub> 1-alkanols. *J. Phys. Chem. Ref. Data*, 19(3):719–762, 1990.
- [43] A. I. Zaitsev, A. D. Litvina, N. P. Lyakishev, and B. M. Mogutnov. Thermodynamic properties and phase equilibria in the CaF<sub>2</sub>–SiO<sub>2</sub>–Al<sub>2</sub>O<sub>3</sub>–CaO system: IV. Calculation and investigation of phase equilibria in the CaO–SiO<sub>2</sub>–Al<sub>2</sub>O<sub>3</sub> system. *Inorg. Mater. (Transl. of Neorg. Mater.)*, 34(3):268–276, 1998.
- [44] Michael J. Zehe, Sanford Gordon, and Bonnie J. McBride. CAP: A computer code for generating tabular thermodynamic functions from NASA Lewis coefficients. Technical report, Glenn Research Center, Cleveland, Ohio, feb 2002. NASA/TP—2001-210959/REV1.

## Mineral Acids (DatAci)

Experimental data for aqueous solutions of HCl, H<sub>2</sub>SO<sub>4</sub>, HNO<sub>3</sub>, etc.

**Related keywords: (experimental | data) & acids.**

### Bibliography.

- [1] [Aubrey P. Altshuller](#). Thermodynamic considerations in the interactions of nitrogen oxides and oxyacids in the atmosphere. *Journal of the Air Pollution Control Association*, 6(2):97–100, 1956.
- [2] [Carrière et Arnaud, E.](#) Détermination des courbes d'ébullition et de rosée des mélanges d'acide chlorhydrique et d'eau sous la pression de 760<sup>mm</sup>. *Acad. Sci., Paris, C. R.*, 179:1265–1267, Juillet–Dec. 1924.
- [3] [T. V. Atamanenko](#) and L. A. Bernshtein. Liquid–vapor equilibrium in nitric acid–water–sulfuric acid ternary systems. *Zh. Prikl. Khim. (S.-Peterburg)*, 60(2):241–245, 1987.
- [4] [Remy Audinos](#). Étude de l'équilibre liquide-vapeur des mélanges de peroxyde d'azote, d'acide nitrique et d'eau. *Acad. Sci., Paris, C. R.*, C(266):1117–1120, Apr. 1968.
- [5] [Rémy Audinos](#). Étude des tensions de vapeur des mélanges d'acide nitrique, d'eau et de peroxyde d'azote. *J. Chim. Phys. Phys.-Chim. Biol.*, 66(3):489–495, 1969. English translation of title: Vapor Pressure of Mixtures of Nitric Acid, Water and Nitrogen Peroxide.
- [6] [G. Aunis](#). Mesure de tensions de vapeur partielles des mélanges NO<sup>3</sup>H–H<sup>2</sup>O à 20° C et vérification de l'équation de Margules–Duhem. 2<sup>e</sup> partie.—verification de l'équation de Margules–Duhem pour les mélanges NO<sub>3</sub>H–H<sub>2</sub>O à 20° C. *J. Chim. Phys. Phys.-Chim. Biol.*, 49:103–108, see also pp. 99–102, 1952. English translation by K. S. Bevis in AEC-tr-2914 entitled: Part II—Verification of the Margules–Duhem Equation for HNO<sub>3</sub>–H<sub>2</sub>O Mixtures at 20 C.
- [7] [Axtmann, Robert C.](#), Woodfin E. Shuler, and Bruce B. Murray. Proton resonance shifts in nitric acid solutions of aluminum nitrate. *J. Phys. Chem.*, 64(1):57–61, 1960.
- [8] [Rolf Werner Bach](#), Hans Adolf Friedrichs, and Hans Rau. *p–v–t* relations for HCl–H<sub>2</sub>O mixtures up to 500° C and 1500 bars. *High Temp. - High Pressures*, 9:305–312, 1977.
- [9] [Fanny Balbaud](#), Gérard Sanchez, Gérard Santarini, and Gérard Picard. Equilibria between gas and liquid phases for concentrated aqueous solutions of nitric acid. *European Journal of Inorganic Chemistry*, 1999(2):277–285, feb 1999.
- [10] [Bassett, Henry, Jr.](#) and Hugh Stott Taylor. LXVI.—Calcium nitrate. Part I. The two-component system: Calcium nitrate–water. Part II. The three-component system: Calcium nitrate–nitric acid–water at 25° C. *J. Chem. Soc., Trans.*, 101:576–585, 1912.
- [11] [Ernst Berl](#) and H. H. Saenger. Über das system N<sub>2</sub>O<sub>5</sub>–HNO<sub>3</sub>. *Monatsh. Chem. Ver. Wiss.*, 53(1):1036–1056, Dec 1929.
- [12] [E. Berl](#) and H. Staudinger. Über die Bestimmung der Siedepunkts- und Destillationskurve von Salzsäure–Wasser–Gemischen. *Z. Angew. Chem.*, 43:1019–1022, 1930.
- [13] [Keith D. Beyer](#) and Anne R. Hansen. Phase diagram of the nitric acid/water system: Implications for polar stratospheric clouds. *J. Phys. Chem. A*, 106(43):10275–10284, oct 2002.
- [14] [Keith D. Beyer](#), Anne R. Hansen, and Nick Raddatz. Experimental determination of the H<sub>2</sub>SO<sub>4</sub>/HNO<sub>3</sub>/H<sub>2</sub>O phase diagram in regions of stratospheric importance. *J. Phys. Chem. A*, 108(5):770–780, feb 2004.
- [15] [Walter D. Bonner](#) and Roland E. Wallace. The boiling points of constant boiling hydrochloric acids. *J. Am. Chem. Soc.*, 52:1747–1750, May 1930.
- [16] [A. Bosen](#) and H. Engels. Description of the phase equilibrium of sulfuric acid with the NRTL equation and a solvation model in a wide concentration and temperature range. *Fluid Phase Equilib.*, 43:213–230, 1988.
- [17] [William Robert Bousfield](#). CLV.—The study of the density and viscosity of aqueous solutions, with special reference to nitric acid. Part I. Densities. *J. Chem. Soc., Trans.*, 107:1405–1426, 1915.
- [18] [William Robert Bousfield](#). III—Mixtures of nitrogen peroxide and nitric acid. *J. Chem. Soc., Trans.*, 115:45–55, 1919.
- [19] [T. Boublik](#) and K. Kuchynka. Gleichgewicht flüssigkeit–dampf XXII. Abhängigkeit der zusammensetzung des azeotropischen gemisches des systems salpetersäure–wasser vom druck. *Collection of Czechoslovak chemical communications*, 25(2):579–582, 1960.
- [20] [Stefano Brandani](#) and Vincenzo Brandani. Vapor–liquid equilibrium calculation of the system water–nitric acid over the entire concentration range. *Fluid Phase Equilib.*, 114:37–45, 1996.
- [21] [Peter Brimblecombe](#) and Simon L. Clegg. Equilibrium partial pressures of strong acids over concentrated solutions—III. The temperature variation of HNO<sub>3</sub> solubility. *Atmospheric Environment. Part A. General Topics*, 24(7):1945–1955, 1990.

- [22] R. R. Brown, G. E. Daut, R. V. Mrazek, and N. A. Gokcen. Solubility and activity of aluminium chloride in aqueous hydrochloric acid solutions. Technical Report 8379, U.S. Bureau of Mines, 1979. Bureau of Mines Report of Investigations.
- [23] Charles L. Burdick and E. Stanley Freed. The equilibrium between nitric oxide, nitrogen peroxide and aqueous solution of nitric acid. *J. Am. Chem. Soc.*, 43:518–530, 1921.
- [24] Clifford D. Carpenter and Joseph Babor. Concentrating dilute nitric acid solutions: A study of rate of concentration—Application of results to fractionation—Determination of composition of vapor above boiling solution of known composition. *Transactions of the American Institute of Chemical Engineers*, XIV:67–79, 1922. The years 1921–22 are collated in one volume.
- [25] Clifford D. Carpenter and Joseph A. Babor. Concentrating dilute nitric acid. *Chem. Metall. Eng.*, 31(7):260–262, Aug. 1924.
- [26] Carpenter, C. D. and J. A. Babor. Concentrating dilute nitric acid—A study of the relation between the vapor and liquid phases in the distillation of nitric acid–water solutions, and of nitric acid–sulphuric acid–water solutions. *Trans. Amer. Inst. Chem. Engrs.*, 16:111–148, 1924.
- [27] Giorgio Carta and Robert L. Pigford. Absorption of nitric oxide in nitric acid and water. *Ind. Eng. Chem. Fundam.*, 22:329–335, 1983.
- [28] Kenneth S. Carslaw, Simon L. Clegg, and Peter Brimblecombe. A thermodynamic model of the system  $\text{HCl}-\text{HNO}_3-\text{H}_2\text{SO}_4-\text{H}_2\text{O}$ , including solubilities of  $\text{HBr}$  from  $< 200$  to  $328\text{ K}$ . *J. Phys. Chem.*, 99(29):11557–11574, 1995.
- [29] Alessandro Cerquetti, Paolo Longhi, and Torquato Mussini. Thermodynamics of aqueous hydrochloric acid from E. M. F.'s of hydrogen-chlorine cells. *J. Chem. Eng. Data*, 13(4):458–461, Oct. 1968.
- [30] E. C. Chen and George McGuire. Vapor–liquid equilibria of the hydrochloric acid–water system. *AIChE J.*, 16(4):686–687, Jul. 1970.
- [31] V. S. Chupalov, N. K. Muzychenko, and Yu. B. Stepanov. Study of equilibrium between nitrogen oxides and 60–80% nitric acid. *Zh. Prikl. Khim. (S.-Peterburg)*, 60(10):2179–2181, 1988.
- [32] Jean-Luc Clavelin and Philippe Mirabel. Détermination des pressions partielles du mélange eau–acide nitrique. *J. Chim. Phys. Phys.-Chim. Biol.*, 76(83):533–537, 1979.
- [33] Simon L. Clegg and Peter Brimblecombe. Equilibrium partial pressures and mean activity and osmotic coefficients of 0–100% nitric acid as a function of temperature. *J. Phys. Chem.*, 94(13):5369–5380, 1990.
- [34] S. L. Clegg, K. S. Carslaw, and P. Brimblecombe. Comment on “vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination” by D.-E. Taleb, J.-L. Ponche, and P. Mirabel: Part 1. *Journal of Geophysical Research: Atmospheres*, 102(D13):10867–10869, May 1997.
- [35] S. L. Clegg, K. S. Carslaw, and P. Brimblecombe. Comment on “vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination” by D.-E. Taleb, J.-L. Ponche, and P. Mirabel: Part 2. *Journal of Geophysical Research: Atmospheres*, 103(D13):16291–16294, Jul 1998.
- [36] A. K. Covington and J. E. Prue. 301. Activity coefficients of nitric and perchloric acids in dilute aqueous solution from e.m.f. and transport-number measurements. *J. Chem. Soc.*, pages 1567–1572, 1957.
- [37] Henry Jermain Maude Creighton and John Horace Githens. On the boiling-point of aqueous solutions of nitric acid at different pressures. Part I. *J. Franklin Inst.*, 179(2):161–169, Feb 1915.
- [38] Henry Jermain Maude Creighton and Herschel Gaston Smith. On the boiling-point of aqueous solutions of nitric acid at different pressures. Part II.—The influence of water-retaining agents on the composition of the mixture of maximum boiling-point. *J. Franklin Inst.*, 180(6):703–709, Dec 1915.
- [39] Henry Jermain Maude Creighton. Electrolytic concentration of aqueous solutions of nitric acid. I. *J. Franklin Inst.*, 193(1):89–95, Jan 1922.
- [40] R. C. Crooks, R. Q. Wilson, A. E. Bearnse, and Jr. R. B. Filbert. Composition of vapors from boiling nitric acid solutions. Technical Report BMI-978, Battelle Memorial Institute, Columbus, Ohio, Feb 1955.
- [41] René Dalmon and René Freymann. Sur les spectres d'absorption dans le proche infrarouge de l'acide nitrique a 100% et en solution aqueuse. *Acad. Sci., Paris, C. R.*, 211(18):472–474, Nov. 1940.
- [42] W. Davis and H. J. De Bruin. New activity coefficients of 0–100 per cent aqueous nitric acid. *J. Inorg. Nucl. Chem.*, 26(6):1069–1083, 1964.
- [43] J. G. Dawber. Evidence for stability of the  $\text{H}_3\text{O}^+(\text{H}_2\text{O})_2$  and  $\text{NO}_3^- - (\text{HNO}_3)_2$  from vapour pressure data in the region of nitric acid monohydrate. *J. Inorg. Nucl. Chem.*, 37:1043–1044, 1975.
- [44] Edoardo Decanini, Giuliano Nardini, and Alessandro Paglianti. Absorption of nitrogen oxides in columns equipped with low-pressure drops structured packings. *Ind. Eng. Chem. Res.*, 39:5003–5011, 2000.

- [45] N. C. Deno, Henry J. Peterson, and Edward Sacher. Nitric acid equilibria in water–sulfuric acid. *J. Phys. Chem.*, 65(2):199–201, 1961.
- [46] J. A. Duisman and S. A. Stern. Vapor pressure and boiling point of pure nitric acid. *J. Chem. Eng. Data*, 14(4):457–459, oct 1969.
- [47] Dunning, W. J. and C. W. Nutt. Dissociation and equilibria of pure liquid nitric acid. *Trans. Faraday Soc.*, 47:15–25, 1951.
- [48] Egan, Jr., Edward P. Vapor pressure of liquid nitric acid. *Ind. Eng. Chem.*, 37(3):303–304, Mar. 1945.
- [49] Egan, Jr., E. P. and B. B. Luff. Density of aqueous solutions of phosphoric acid. Measurements at 15° to 80°. *Ind. Eng. Chem.*, 47(6):1280–1281, Jun. 1955.
- [50] Egan, Jr., Edward P., Basil B. Luff, and Zachary T. Wakefield. Heat capacity of phosphoric acid solutions, 15° to 80°. *J. Phys. Chem.*, 62(9):1091–1095, sep 1958.
- [51] Egan, Jr., Edward P. and Basil B. Luff. Heat of solution of orthophosphoric acid. *J. Phys. Chem.*, 65(3):523–526, Mar. 1961.
- [52] Egan, Jr., Edward P. and Basil B. Luff. Heat capacities of the liquid phase in the system CaO–P<sub>2</sub>O<sub>5</sub>–H<sub>2</sub>O. *J. Chem. Eng. Data*, 11(4):509–519, oct 1966.
- [53] Egan, Jr., Edward P. and Basil B. Luff. Heats of solution at 25 degrees in the system CaO–P<sub>2</sub>O<sub>5</sub>–H<sub>2</sub>O. *J. Chem. Eng. Data*, 11(4):520–532, oct 1966.
- [54] Ellis, S. R. M. and J. M. Thwaites. Vapour–liquid equilibria of nitric acid–water–sulphuric acid mixtures. *J. appl. Chem.*, 7(4):152–160, Apr 1957.
- [55] Elverum, Jr., Gerard W. and David M. Mason. Melting point measurement of the system HNO<sub>3</sub>–N<sub>2</sub>O<sub>4</sub>–H<sub>2</sub>O. *J. Phys. Chem.*, 60:104–108, Jan. 1956.
- [56] H. Engels and A. Bosen. Description of the system HCl/H<sub>2</sub>O with a “local composition” equation for the activity coefficient and a suitable dissociation model. *Fluid Phase Equilib.*, 28:171–181, 1986.
- [57] Erlebach, Woodland Eustace. *The Boiling Points, Compositions, and Densities of the Azeotropes of Deuteriochloric Acid*. PhD thesis, University of British Columbia, 1953.
- [58] Giuseppe Faita and Torquato Mussini. Thermodynamic functions of aqueous hydrochloric acid at various concentrations and temperatures. *J. Chem. Eng. Data*, 9(3):332–335, Jul. 1964.
- [59] Thad D. Farr and Kelly L. Elmore. System CaO–P<sub>2</sub>O<sub>5</sub>–HF–H<sub>2</sub>O: Thermodynamic properties. *J. Phys. Chem.*, 66(2):315–318, 1962.
- [60] Finlayson-Pitts, B. J., L. M. Wingen, A. L. Sumner, D. Syomin, and K. A. Ramazan. The heterogeneous hydrolysis of NO<sub>2</sub> in laboratory systems and in outdoor and indoor atmospheres: An integrated mechanism. *Phys. Chem. Chem. Phys.*, 5(2):223–242, 2003.
- [61] R. Flatt and F. Benguerel. Sur l'équilibre liquide-vapeur du système binaire HNO<sub>3</sub>–H<sub>2</sub>O à 25°C. *Helv. Chim. Acta*, 45(6):1765–1772, 1962.
- [62] W. R. Forsythe and W. F. Giauque. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 1(64):48–61, 1942.
- [63] W. R. Forsythe and W. F. Giauque. Additions and corrections. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 12(64):3069, 1942.
- [64] W. R. Forsythe and W. F. Giauque. Additions and corrections. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 12(65):2479, 1943.
- [65] J. J. Fritz and C. R. Fuget. Vapor pressure of aqueous hydrogen chloride solutions, 0° to 50°C. *Ind. Eng. Chem. Chem. Eng. Data Series*, 1(1):10–12, 1956.
- [66] R. M. Gibbons and A. P. Laughton. An equation of state for hydrochloric acid solutions. *Fluid Phase Equilib.*, 18:61–68, 1984.
- [67] R. J. Gillespie, E. D. Hughes, and C. K. Ingold. 504. Cryoscopic measurements in nitric acid. Part I. The solutes dinitrogen pentoxide and water. The self-dissociation of nitric acid. *J. Chem. Soc.*, pages 2552–2558, 1950.
- [68] D. R. Goddard, E. D. Hughes, and C. K. Ingold. 505. Chemistry of nitronium salts. Part I. Isolation of some nitronium salts. *J. Chem. Soc.*, pages 2559–2575, 1950.
- [69] N. A. Gokcen. Partial pressures of gaseous HCl and H<sub>2</sub>O over aqueous solutions of HCl, AlCl<sub>3</sub> and FeCl<sub>3</sub>. Technical Report 8456, U.S. Bureau of Mines, 1980. Bureau of Mines Report of Investigations.
- [70] Stuard R. Gunn and Leroy G. Green. Heat of solution of hydrogen chloride. *J. Chem. Eng. Data*, 8:180, 1963.

- [71] Gutierrez-Cañas, C., P. L. Arias, and J. A. Legarreta. Industrial nitrogen oxides absorption simulation. *Comput. Chem. Eng.*, 13(9):985–1002, 1989.
- [72] R. Haase, H. Naas, and H. Thumm. Experimentelle Untersuchungen über das thermodynamische Verhalten konzentrierter Halogenwasserstoffsäuren. *Z. Phys. Chem., Neue Folge*, 37:210–229, 1963.
- [73] R. Haase, K.-H. Dücker, and H. A. Küppers. Aktivitätskoeffizienten und dissoziationskonstanten wäßriger salpetersäure und überchlorsäure. *Ber. Bunsen-Ges.*, 69(2):97–109, 1965.
- [74] David R. Hanson and Konrad Mauersberger. Vapor pressures of  $\text{HNO}_3/\text{H}_2\text{O}$  solutions at low temperatures. *J. Phys. Chem.*, 92(21):6167–6170, Oct 1988.
- [75] Józef Hawliczek and Jerzy Synowiec. Badania nad desorpcja chlorowodoru z roztworów wodnych cz. I. Równowaga ciecz-pary układu  $\text{H}_2\text{O}-\text{HCl}$ . *Chem. Stosow. (1957-63)*, 3:369–387, 1962.
- [76] Walter Hayduk, Haruki Asatani, and Benjamin C.-Y. Lu. Solubility of sulfur dioxide in aqueous sulfuric acid solutions. *J. Chem. Eng. Data*, 33(4):506–509, 1988.
- [77] Hedlund, Frank Huess, Jerome Frutiger, and Gürkan Sin. Svovlsyres fortyndingsvarme. *Dansk Kemi*, 99(3):16–20, 2018.
- [78] H. F. Holmes, R. H. Busey, J. M. Simonson, R. E. Mesmer, D. G. Archer, and R. H. Wood. The enthalpy of dilution of  $\text{HCl}(aq)$  to 648 K and 40 MPa: Thermodynamic properties. *J. Chem. Thermodyn.*, 19:863–890, 1987.
- [79] G. D. Honti. Nitric acid. In G. D. Honti, editor, *The Nitrogen Industry*, chapter 2.2 in volume 2 of part 1/2 of the collection. Akadémiai Kiadó, Budapest, 1976.
- [80] Hood, G. C. and C. A. Reilly. Ionization of strong electrolytes. VIII. Temperature coefficient of dissociation of strong acids by proton magnetic resonance. *J. Chem. Phys.*, 32:127–130, jan 1960.
- [81] Elmar-M. Horn, Helmuth Keiser, and Klaus Schoeller. Zur thermischen zersetzung der hochkonzentrierten salpetersäure. *Monatsh. Chem.—Chem. Monthly*, 118(11):1205–1218, Nov 1987.
- [82] Jaecker-Voirol, A., J. L. Ponche, and P. Mirabel. Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperatures. *Journal of Geophysical Research: Atmospheres*, 95(D8):11857–11863, 1990.
- [83] Jaecker-Voirol, A., J. L. Ponche, and P. Mirabel. Correction to “Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperatures” by A. Jaecker-Voirol et al. *Journal of Geophysical Research: Atmospheres*, 95(D13):22565, Dec 1990.
- [84] James T. F. Kao. Vapor–liquid equilibrium of water–hydrogen chloride system. *J. Chem. Eng. Data*, 15(3):362–367, 1970.
- [85] C. M. Karni, W. O. Rains, R. M. Counce, J. S. Watson, B. B. Spencer, and G. D. Del Cul. Water and mean ionic activities of aqueous  $\text{HNO}_3$  solutions calculated from an extension of the Brunauer–Emmett–Teller (BET) model. *Sep. Sci. Technol.*, 45(12–13):1894–1900, 2010.
- [86] S. G. Katal'nikov, A. V. Khoroshilov, and M. M. Chelyak. Isotope separation by the "nitric acid" method. Phase and isotope equilibrium in a system consisting of nitrogen oxides and nitric acid. *Soviet Atomic Energy*, 60(2):132–138, 1986. Translated from *Atomnaya Énergiya*, Vol. 60, No. 2, pp. 109–113, February, 1986.
- [87] Kay, Webster B. and S. Alexander Stern. Phase relations of nitric acid at physicochemical equilibrium. *Ind. Eng. Chem.*, 47(7):1463–1469, 1955.
- [88] Kay, W. B., S. A. Stern, and M. D. Sanghvi. Phase relations of nitric acid–nitrogen dioxide and nitric acid–water mixtures at physicochemical equilibrium. *Ind. Eng. Chem. Chem. Eng. Data Series*, 2(1):21–28, 1956.
- [89] Khoroshilov, A. V., T. V. Guseva, and S. G. Katal'nikov. Analysis of the chemical composition of the gas phase at equilibrium with nitric acid solutions. *Industrial Laboratory*, 48(10):947–951, 1982.
- [90] Wolfgang Kindler, Günter Wüster, and Hans Rau. Equation of state for the vapour of concentrated and diluted hydrochloric acid. *Ber. Bunsen-Ges.*, 82:543–545, 1978.
- [91] Alfons Klemenc and Erich Hayek. Zur kenntnis der salpetersäure. V. Über die einwirkung des stickoxydes auf salpetersäure bis zur erreichung des gleichgewichtes. (a) Das system ansteigend bis zu 1 n-salpetersäuregleichgewichtskonzentration in wäßriger lösung.  $3\text{HNO}_2 \leftrightarrow \text{HNO}_3 + 2\text{NO} + \text{H}_2\text{O}$ . *Z. Anorg. Allg. Chem.*, 186(1):181–224, 1929.
- [92] A. Klemenc and Th. Spiess. Zur kenntnis der salpetersäure. VIII. Über die lösungskurve des systems  $\text{HNO}_3-\text{NO}_2$  und  $\text{HNO}_3-\text{H}_2\text{O}-\text{NO}_2$ . *Monatsh. Chem. Ver. Wiss.*, 77(1):216–223, Dec 1947.
- [93] A. Klemenc. Zur kenntnis der salpetersäure. X. Methoden zur gasanalyse im system  $\text{N}_2-\text{NO}-\text{NO}_2-\text{N}_2\text{O}_4-\text{N}_2\text{O}_3-\text{HNO}_2-\text{HNO}_3$ . *Monatsh. Chem. Ver. Wiss.*, 83(2):334–345, Mar 1952. English translation by K. S. Bevis in AEC-tr-2450 entitled: Information on Nitric Acid. X. Gas Analysis Methods in the System  $\text{N}_2-\text{NO}-\text{NO}_2-\text{N}_2\text{O}_4-\text{N}_2\text{O}_3-\text{HNO}_2-\text{HNO}_3$ .
- [94] I. A. Leenson. Sulfuric acid and water: Paradoxes of dilution. *J. Chem. Educ.*, 81(7):991–994, Jul. 2004.
- [95] Y.-N. Lee and S. E. Schwartz. Reaction kinetics of nitrogen dioxide with liquid water at low partial pressures. *J. Phys. Chem.*, 85:840–848, 1981.

- [96] Jan Bernard Lefers. *Absorption of Nitrogen Oxides into Diluted and Concentrated Nitric Acid*. Delft University Press, 1980.
- [97] J. B. Lefers and P. J. van den Berg. Absorption of  $\text{NO}_2/\text{N}_2\text{O}_4$  into diluted and concentrated nitric acid. *Chem. Eng. J. (Lausanne)*, 23:211–221, 1982.
- [98] Robert J. Lemire, Colin P. Brown, and Allan B. Campbell. Vapor pressure of nitric acid and water in the systems  $\text{HNO}_3-\text{H}_2\text{O}$  and  $\text{HNO}_3-\text{Th}(\text{NO}_3)_4-\text{H}_2\text{O}$  at  $50^\circ\text{C}$ . *J. Chem. Eng. Data*, 30(4):421–424, 1985.
- [99] Gilbert N. Lewis and Arthur Edgar. The equilibrium between nitric acid, nitrous acid and nitric oxide. *J. Am. Chem. Soc.*, 33(3):292–299, 1911.
- [100] L. Lloyd and P. A. H. Wyatt. The vapour pressure of nitric acid solutions. Part I. New azeotropes in the water–dinitrogen pentoxide system. *J. Chem. Soc.*, pages 2248–2252, 1955.
- [101] Basil B. Luff. Heat capacity and enthalpy of phosphoric acid. *J. Chem. Eng. Data*, 1(26):70–74, 1981.
- [102] N. V. Lutugina and L. I. Kokovkina. Liquid–vapor equilibrium in water–hydrogen chloride, water–hydrogen iodide, and iodide–hydrogen chloride systems. *Zh. Prikl. Khim. (S.-Peterburg)*, 38(7):1487–1494, Jul. 1965.
- [103] Arturo Maimoni. Vapor–liquid equilibria for nitric acid–water and plutonium nitrate–nitric acid–water solutions. Technical Report UCID-18148, Lawrence Livermore Laboratory (California Univ.), Jan 1980.
- [104] Mario Massucci, Simon L. Clegg, and Peter Brimblecombe. Equilibrium vapor pressure of  $\text{H}_2\text{O}$  above aqueous  $\text{H}_2\text{SO}_4$  at low temperature. *J. Chem. Eng. Data*, 41(4):765–778, 1996.
- [105] Mario Massucci, Simon L. Clegg, and Peter Brimblecombe. Equilibrium partial pressure, thermodynamic properties of aqueous and solid phases and  $\text{Cl}_2$  production from aqueous  $\text{HCl}$  and  $\text{HNO}_3$  and their mixtures. *J. Phys. Chem. A*, 103:4209–4226, 1999.
- [106] A. H. McDaniel, J. A. Davidson, C. A. Cantrell, R. E. Shetter, and J. G. Calvert. Enthalpies of formation of dinitrogen pentoxide and the nitrate free radical. *J. Phys. Chem.*, 92:4172–4175, 1988.
- [107] H. McKenzie, J. MacDonald-Taylor, F. McLachlan, R. Orr, and D. Woodhead. Modelling of nitric and nitrous acid chemistry for solvent extraction purposes. *Procedia Chemistry*, 21:481–486, 2016. ATALANTE 2016 International Conference on Nuclear Chemistry for Sustainable Fuel Cycles.
- [108] Anderson B. McKeown and Frank E. Belles. Vapor pressures of concentrated nitric acid solutions in the composition range 83 to 97 percent nitric acid, 0 to 6 percent nitrogen dioxide, 0 to 15 percent water, and in the temperature range  $20^\circ$  to  $80^\circ\text{C}$ . Technical Report RM-E53G08, National Advisory Committee for Aeronautics (NACA), Washington, Sep 1953.
- [109] Anderson B. McKeown and Frank E. Belles. Vapor pressures and calculated heats of vaporization of concentrated nitric acid solutions in the composition range 71 to 89 percent nitric acid, 7 to 20 percent nitrogen dioxide, 1 to 10 percent water, and in the temperature range  $10^\circ$  to  $60^\circ\text{C}$ . Technical Report RM-E53L14, National Advisory Committee for Aeronautics (NACA), Washington, Feb 1954.
- [110] Anderson B. McKeown and Frank E. Belles. Nitric acid–nitrogen dioxide–water system—Vapor pressures and related properties. *Ind. Eng. Chem.*, 47(12):2540–2543, 1955.
- [111] I. R. McKinnon, J. G. Mathieson, and I. R. Wilson. Gas phase reaction of nitric oxide with nitric acid. *J. Phys. Chem.*, 83(7):779–780, 1979.
- [112] Frank Douglas Miles. *Nitric Acid: Manufacture and Uses*, pages 9–19. Oxford University Press, London, 1961.
- [113] D. N. Miller. Mass transfer in nitric acid absorption. *AIChE J.*, 33(8):1351–1358, 1987.
- [114] Claude C. van Nuys. Enthalpy and heats of dilution of the system  $\text{HCl}-\text{H}_2\text{O}$ . *Am. Inst. Chem. Eng. Symp. Ser.*, pages 663–678, 1943.
- [115] L. I. Olefir, M. I. Knyshenko, and L. K. Priima. The shift of the azeotropic point in the system  $\text{HNO}_3-\text{H}_2\text{O}$ . *The Soviet chemical industry*, 10(2):127–128, 1978. The printed article says pp.127–128 but the receipt coming back from the library says pp.123–124.
- [116] J. L. Oscarson, S. E. Gillespie, R. M. Izatt, X. Chen, and C. Pando. Thermodynamic quantities for the ionization of nitric acid in aqueous solution from 270 to  $319^\circ\text{C}$ . *J. Solution Chem.*, 21(8):789–801, 1992.
- [117] D. F. Othmer. Composition of vapors from boiling binary solutions. *Ind. Eng. Chem.*, 20(7):743–746, Jul. 1928.
- [118] Jaakko Ilmari Partanen, Pekka Mikael Juusola, and Pentti Olavi Minkkinen. Density of aqueous nitric acid solutions in the molality range 0–3.5  $\text{mol kg}^{-1}$  at 293.15, 298.15, 303.15 and 308.15 K. *Acta Chem. Scand.*, 46:338–342, 1992.
- [119] B. A. Patterson and E. M. Woolley. Thermodynamics of ionization of water at temperatures  $278.15 < T/\text{K} < 393.15$  and at the pressure  $p = 0.35\text{MPa}$ : Apparent molar volumes and apparent molar heat capacities of aqueous solutions of potassium and sodium nitrates and nitric acid. *J. Chem. Thermodyn.*, 34(4):535–556, 2002.

- [120] **Kenneth S. Pitzer**, Rabindra N. Roy, and Leonard F. Silvester. Thermodynamics of electrolytes. 7. Sulfuric acid. *J. Am. Chem. Soc.*, 99(15):4930–4936, Jul. 1977.
- [121] **Antoine Potier**. Préparation de l'acide nitrique au laboratoire. Densité à 15°C. *C. R. Hebd. Seances Acad. Sci.*, 233(19):1113–1115, 1951. Présentée par Paul Pascal.
- [122] **Jacqueline Potier** and Antoine Potier. Ébulliométrie de l'acide nitrique. *Bull. Soc. Chim. Fr.*, 19(3–4):250–251, 1952.
- [123] **A. Potier**. Préparation du peroxyde d'azote pur au laboratoire. *Bull. Soc. Chim. Fr., Mem.*, 20(139):709–711, 1953.
- [124] **A. Potier**. Propriétés ionisantes et ionisation de l'acide nitrique anhydre. *Bull. Soc. Chim. Fr.*, 21(2):166, Jan 1954. Section d'Alger.
- [125] **A. Potier**. Calcul des pressions partielles isothermes des systèmes binaires liquides volatils à partir des courbes isothermes pressions totales. Composition. *Bull. Soc. Chim. Fr.*, 22(3):316, 1955. Section d'Alger.
- [126] **Antoine Potier**. I. Tensions de vapeur totales et partielles des solutions acides nitrique–eau entre 0° et 20°C. Densités des solutions concentrées à 15°C. *Bull. Soc. Chim. Fr., Mem.*, 23(5):47–49, 1956.
- [127] **Antoine Potier**. II. Tensions de vapeur totales et partielles des solutions acides nitrique–peroxyde d'azote entre –10° et +10°C, densités à 15°C. *Bull. Soc. Chim. Fr., Mem.*, 23(6):50–53, 1956.
- [128] **J. Potier** and A. Potier. Sur le système acide nitrique–eau. *Bull. Soc. Chim. Fr.*, 21(2):1119–1120, 1956. Section d'Alger.
- [129] **A. Potier**. Étude des propriétés thermodynamiques des systèmes. Acide nitrique–Eau et peroxyde d'azote. Acide nitrique. *Annales de la Faculté des sciences de Toulouse: Mathématiques*, 4<sup>e</sup> série, 20:1–98, 1956. Published by: Gauthier-Villars, Libraire-Éditeur, Paris.
- [130] **Jacqueline Potier**. Étude ébulliométrique du système  $\text{NO}_3\text{H}-\text{OH}_2$ . *Alger. Sciences Physiques.*, IV(1):61–86, 1957.
- [131] **Antoine Potier** and Jacqueline Potier. Les binaires acide nitrique–nitrates «alcalins». II. Équilibres liquide–solide dans le système  $\text{HNO}_3-\text{NH}_4\text{NO}_3$ . *Bull. Soc. Chim. Fr., Mem.*, 26(51):311–314, 1959.
- [132] **K. J. Prösek**. Gleichgewicht zwischen der flüssigen und der gasförmigen phase in salpetersäure enthaltenden systemen. *Coll. of Czechoslovak Chem. Comm.*, 32(7):2397–2404, 1967.
- [133] **Hans Rau** and T. R. Narayanan Kuttu. PVT-Messungen an konzentrierter Salzsäure bis 500°C und 1500 bar. *Ber. Bunsen-Ges.*, 76(7):645–651, 1972.
- [134] **O. Redlich** and J. Bigeleisen. The ionization of strong electrolytes. I. General remarks, nitric acid. *J. Am. Chem. Soc.*, 65:1883–1887, 1943.
- [135] **O. Redlich** and G. C. Hood. Ionic interaction, dissociation and molecular structure. *Discuss. Faraday Soc.*, 24:87–93, 1957.
- [136] **O. Redlich**, W. E. Gargrave, and W. D. Krostek. Thermodynamics of solutions. consistency of sparse data. Activities of nitric and perchloric acids. *Ind. Eng. Chem. Fundam.*, 7(2):211–214, May 1968.
- [137] **R. C. Reid**, A. B. Reynolds, D. T. Morgan, G. W. Bond, J. E. Savolainen, and M. L. Hyman. Vapor-liquid equilibria. Dilute nitric acid, hydrochloric acid, and hydrochloric acid–nitric acid–water solutions. *Ind. Eng. Chem.*, 49(8):1307–1310, 1957.
- [138] **Theodore W. Richards** and Allan Winter Rowe. The heats of dilution and the specific heats of dilute solutions of nitric acid and of hydroxides and chlorides and nitrates of lithium, sodium, potassium, and cesium. *J. Am. Chem. Soc.*, 43(4):770–796, 1921.
- [139] **Eric K. Rideal**. On the absorption of oxides of nitrogen by nitric acid. *J. Ind. & Eng. Chem.*, 12(6):531–538, 1920.
- [140] **Fernand Rivenq**. Étude ébulliométrique des solutions aqueuses d'acide nitrique sous pression réduite. *Bull. Soc. Chim. Fr.*, 26(6):822–826, 1959. English translation of title: Boiling point diagrams of aqueous nitric acid mixtures under reduced pressure.
- [141] **Ruas, Alexandre**, Patrick Pochon, Jean-Pierre Simonin, and Philippe Moisy. Nitric acid: Modeling osmotic coefficients and acid–base dissociation using the BIMSA theory. *Dalton Trans.*, 39(42):10148–10153, 2010.
- [142] **Takeshi Sako**, Toshikatsu Hakuta, and Hiroshi Yoshitome. Salt effects on vapor–liquid equilibria for volatile strong electrolyte-water systems. *J. Chem. Eng. Jpn.*, 17(4):381–388, 1984.
- [143] **Takeshi Sako**, Toshikatsu Hakuta, and Hiroshi Yoshitome. Vapor pressures of binary ( $\text{H}_2\text{O}-\text{HCl}$ ,  $-\text{MgCl}_2$ , and  $-\text{CaCl}_2$ ) and ternary ( $\text{H}_2\text{O}-\text{MgCl}_2-\text{CaCl}_2$ ) aqueous solutions. *J. Chem. Eng. Data*, 30(2):224–228, 1985.
- [144] **Bo Sander**, Peter Rasmussen, and Aage Fredenslund. Calculation of vapour–liquid equilibria in nitric acid–water–nitrate salt systems using an extended UNIQUAC equation. *Chem. Eng. Sci.*, 41(5):1185–1195, 1986.
- [145] **Akira Sasahira**, Tadahiro Hoshikawa, and Fumio Kawamura. Hydration model for evaluation of thermodynamic activities in nitric acid–nitrate–water system. *J. Nucl. Sci. Technol.*, 29(6):566–575, Jun. 1992.

- [146] Alfred Schmidt. Über das System Chlorwasserstoff/Wasser. *Chem.-Ing.-Tech.*, 25(8/9):455–466, 1953.
- [147] James A. Snyder, Denise Hanway, Julio Mendez, Alan J. Jamka, and Fu-Ming Tao. A density functional theory study of the gas-phase hydrolysis of dinitrogen pentoxide. *J. Phys. Chem.*, 103:9355–9358, 1999.
- [148] N. N. Solov'ev, V. P. Panov, and L. Ya. Tereshchenko. Calculating the equilibrium composition for the liquid-phase  $\text{HNO}_x - \text{NO}_x - \text{H}_2\text{O}$ . *Zh. Prikl. Khim. (Leningrad)*, 57(10):2029–2031, oct 1984.
- [149] William C. Sproesser and Guy B. Taylor. Vapor pressures of aqueous solutions of nitric acid. *J. Am. Chem. Soc.*, 43(8):1782–1787, 1921.
- [150] B. G. Staples and J. M. Procopio, Jr. Vapor-pressure data for common acids at high temperature. *Chem. Eng. [Int. Ed.]*, 77(25):113–115, Nov. 1970.
- [151] Stern, S. Alexander, J. T. Mullhaupt, and Webster B. Kay. The physicochemical properties of pure nitric acid. *Chem. Rev. (Washington, D. C.)*, 60(2):185–207, 1960.
- [152] H. I. Stonehill. 170. Transport numbers of nitric acid in water at 25° from e.m.f. measurements. *J. Chem. Soc.*, pages 647–651, 1943.
- [153] R. Stryjek and J. H. Vera. Vapor–liquid equilibrium of hydrochloric acid solutions with the PRSV equation of state. *Fluid Phase Equilib.*, 25:279–290, 1986.
- [154] Taleb, Djamel-Eddine, Jean-Luc Ponche, and Philippe Mirabel. Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination. *Journal of Geophysical Research: Atmospheres*, 101(D20):25967–25977, Nov 1996.
- [155] Ignatius N. Tang. On the equilibrium partial pressures of nitric acid and ammonia in the atmosphere. *Atmospheric Environment (1967)*, 14(7):819–828, 1980.
- [156] I. N. Tang, H. R. Munkelwitz, and J. H. Lee. Vapor–liquid equilibrium measurements for dilute nitric acid solutions. *Atmospheric Environment (1967)*, 22(11):2579–2585, 1988.
- [157] Guy B. Taylor. Vapor pressure of aqueous solutions of nitric acid. *Ind. Eng. Chem.*, 17(6):633–635, 1925.
- [158] L. Ya. Tereshchenko, M. I. Kucha, V. P. Panoc, and V. V. Zubov. Equilibrium of nitrogen oxides with nitric acid solutions. *Zh. Prikl. Khim. (S.-Peterburg)*, 52(8):1657–1661, 1980.
- [159] H. Theobald. Messungen zum Gleichgewicht Salpetersäure/nitrose Gase. *Chem.-Ing.-Tech.*, 40(15):763–765, 1968.
- [160] Emmanuel Thibert and Florent Domine. Thermodynamics and kinetics of the solid solution of  $\text{HNO}_3$  in ice. *J. Phys. Chem. B*, 102(22):4432–4439, 1998.
- [161] N. N. Treushchenko, M. E. Pozin, G. V. Bel'chenko, and E. Ya. Ivanova. Vapor pressure of the system  $\text{HNO}_3 - \text{NH}_4\text{NO}_3 - \text{H}_2\text{O}$ . *J. Appl. Chem. USSR*, 44(8):1775–1780, 1997. Translated from Zhurnal Prikladnoi Khimii, 44(8), August 1971, pp. 1753–1760.
- [162] V. G. Tsvetkov, V. F. Sopin, V. N. Marshева, L. Ya. Tsvetkova, E. M. Belova, and G. N. Marchenko. Enthalpy of water mixturing with sulfuric and nitric acid mixtures. *Zh. Obshch. Khim.*, 57(5):997–1000, 1987.
- [163] V. G. Tsvetkov, V. F. Sopin, V. N. Marshева, L. Ya. Tsvetkova, E. M. Belova, and G. N. Marchenko. Enthalpy of water mixturing with sulfuric and nitric acid mixtures. *J. Gen. Chem. USSR (Transl. of Zh. Obshch. Khim.)*, 57(5):888–991, 1987. Translated from Zhurnal Obshchei Khimii, 57(5), May 1987, pp. 997–1000.
- [164] A. G. Udovenko, V. I. Konvisar, A. V. Shapka, and A. I. Mishchenko. Equilibrium of nitrogen oxides with nitric acid solutions under pressure. *Zh. Prikl. Khim. (S.-Peterburg)*, 49(6):1274–1276, 1976.
- [165] Robert Vandoni and Michele Laudy. Détermination des tensions de vapeurs partielles des mélanges  $\text{NO}^3\text{H} - \text{N}^2\text{O}^4 - \text{NO}^2$ . *Mem. services chim. etat (Paris)*, 36:261–268, 1951.
- [166] R. Vandoni and M. Laudy. Mesure de tensions de vapeur partielles des mélanges  $\text{NO}^3\text{H} - \text{H}^2\text{O}$  à 20° C et vérification de l'équation de Margules–Duhem. 1<sup>re</sup> partie.—Mesure des tensions de vapeur partielles des mélanges  $\text{NO}^3\text{H} - \text{H}^2\text{O}$  à 20° C. *J. Chim. Phys. Phys.-Chim. Biol.*, 49:99–102, see also pp. 103–108, 1952. English translation by K. S. Bevis in AEC-tr-2913 entitled: Measurement of Partial Pressures of  $\text{HNO}_3 - \text{H}_2\text{O}$  mixtures at 20 C and Verification of the Margules-Duhem Equation.
- [167] Frank H. Verhoek and Farrington Daniels. The dissociation constants of nitrogen tetroxide and of nitrogen trioxide. *J. Am. Chem. Soc.*, 53(4):1250–1263, Apr. 1931.
- [168] Klaus Vetter. Das gleichgewicht zwischen stickstoffdioxyd und salpetriger säure in salpetersäure. *Z. Anorg. Chem.*, 260(4–5):242–248, 1949.
- [169] Wm. H. Waggaman. *Phosphoric Acid, Phosphates and Phosphatic Fertilizers*, chapter 17, pages 308–344. American Chemical Society Monograph Series. Reinhold, New York, second edition, 1952.
- [170] Zachary T. Wakefield, Basil B. Luff, and Robert B. Reed. Heat capacity and enthalpy of phosphoric acid. *J. Chem. Eng. Data*, 4(17):420–423, 1972.
- [171] D. M. Waldorf and A. L. Babb. Vapor-phase equilibrium of  $\text{NO}$ ,  $\text{NO}_2$ ,  $\text{H}_2\text{O}$ , and  $\text{HNO}_2$ . *J. Chem. Phys.*, 39(2):432–435, Jul. 1963.

- [172] [Werner Weisweiler](#) and Karl-Heinz Deiß. Influence of electrolytes on the absorption of nitrogen oxide components  $\text{N}_2\text{O}_4$  and  $\text{N}_2\text{O}_3$  in aqueous absorbents. *Chem. Eng. Technol.*, 10:131–142, 1987.
- [173] [Werner Weisweiler](#), Klaus Eidam, Michael Thiemann, Erich Scheibler, and Karl Wilhelm Wiegand. Absorption of  $\text{NO}_2/\text{N}_2\text{O}_4$  in nitric acid. *Chem. Eng. Technol.*, 13:97–101, 1990.
- [174] [G. L. Wilson](#) and F. D. Miles. The partial pressures of nitric acid–water mixtures from  $0^\circ$ – $20^\circ\text{C}$ . *Trans. Faraday Soc.*, 35:356–363, 1940.
- [175] [M. Wrewsky](#), N. Sawaritzky, and L. Scharloff. Die Bestimmung des Dampfdruckes und der Zusammensetzung des Dampfes wässriger Lösungen von Chlorwasserstoff und Bromwasserstoff bei verschiedenen Temperaturen. *Z. Phys. Chem., Stoichiom. Verwandtschaftsl.*, 112:97–109, 1924.
- [176] [I. I. Zakharov](#). Quantum-chemical modeling of the mechanism for formation of  $\text{HNO}_3$  from  $\text{NO}_2$  and water. *Theoretical and Experimental Chemistry*, 48(4):233–239, Sep 2012. Translated from *Teoreticheskaya i Éksperimental'naya Khimiya*, 48(4), pp. 217–223, July–August, 2012.
- [177] [Frank J. Zeleznik](#). Thermodynamic properties of the aqueous sulfuric acid system to 350 K. *J. Phys. Chem. Ref. Data*, 20(6):1157–1200, 1991.
- [178] [Renyi Zhang](#), Paul J. Wooldridge, and Mario J. Molina. Vapor pressure measurements for the  $\text{H}_2\text{SO}_4/\text{HNO}_3/\text{H}_2\text{O}$  and  $\text{H}_2\text{SO}_4/\text{HCl}/\text{H}_2\text{O}$  systems: Incorporation of stratospheric acids into background sulfate aerosols. *J. Phys. Chem.*, 97:8541–8548, 1993.
- [179] [B. Ya. Zilberman](#), D. V. Ryabkov, E. A. Puzikov, E. V. Andreeva, and N. E. Mishina. Influence of pressure (temperature) on the nitric acid distribution between the liquid and vapor in the course of evaporation of nitric acid radioactive waste. *Radiochemistry*, 58(3):237–242, Sep 2016. Translated from *Radiokhimiya*, 58(3), pp. 206–210, 2016.

### Nitric acid (hno3)

Papers related to the thermophysical properties of pure  $\text{HNO}_3$  and its aqueous solutions (nitric acid). Including thermodynamic models and other theoretical aspects as well as experimental data. Beyond also the simple binary  $\text{HNO}_3$ – $\text{H}_2\text{O}$ .

**Related keywords:** no2 | n2o4 | n2o5 | hno2 | nitric\ acid | (nitrogen | nitrous | nitric)\ (per)?oxide(s)?

#### Bibliography.

- [1] [Aubrey P. Altshuller](#). Thermodynamic considerations in the interactions of nitrogen oxides and oxyacids in the atmosphere. *Journal of the Air Pollution Control Association*, 6(2):97–100, 1956.
- [2] [S. P. S. Andrew](#) and D. Hanson. D1. The dynamics of nitrous gas absorption. *Chem. Eng. Sci.*, 14(1):105–113, 1961.
- [3] [T. V. Atamanenko](#) and L. A. Bernshtein. Liquid–vapor equilibrium in nitric acid–water–sulfuric acid ternary systems. *Zh. Prikl. Khim. (S.-Peterburg)*, 60(2):241–245, 1987.
- [4] [Remy Audinos](#). Étude de l'équilibre liquide-vapeur des mélanges de peroxyde d'azote, d'acide nitrique et d'eau. *Acad. Sci., Paris, C. R.*, C(266):1117–1120, Apr. 1968.
- [5] [Rémy Audinos](#). Étude des tensions de vapeur des mélanges d'acide nitrique, d'eau et de peroxyde d'azote. *J. Chim. Phys. Phys.-Chim. Biol.*, 66(3):489–495, 1969. English translation of title: Vapor Pressure of Mixtures of Nitric Acid, Water and Nitrogen Peroxide.
- [6] [G. Aunis](#). Mesure de tensions de vapeur partielles des mélanges  $\text{NO}^3\text{H}$ – $\text{H}^2\text{O}$  à  $20^\circ\text{C}$  et vérification de l'équation de Margules–Duhem. 2<sup>e</sup> partie.—verification de l'équation de Margules–Duhem pour les mélanges  $\text{NO}_3\text{H}$ – $\text{H}_2\text{O}$  à  $20^\circ\text{C}$ . *J. Chim. Phys. Phys.-Chim. Biol.*, 49:103–108, see also pp. 99–102, 1952. English translation by K. S. Bevis in AEC-tr-2914 entitled: Part II—Verification of the Margules–Duhem Equation for  $\text{HNO}_3$ – $\text{H}_2\text{O}$  Mixtures at  $20^\circ\text{C}$ .
- [7] [Axtmann, Robert C.](#), Woodfin E. Shuler, and Bruce B. Murray. Proton resonance shifts in nitric acid solutions of aluminum nitrate. *J. Phys. Chem.*, 64(1):57–61, 1960.
- [8] [Fanny Balbaud](#), Gérard Sanchez, Gérard Santarini, and Gérard Picard. Equilibria between gas and liquid phases for concentrated aqueous solutions of nitric acid. *European Journal of Inorganic Chemistry*, 1999(2):277–285, feb 1999.
- [9] [Bassett, Henry, Jr.](#) and Hugh Stott Taylor. LXVI.—Calcium nitrate. Part I. The two-component system: Calcium nitrate–water. Part II. The three-component system: Calcium nitrate–nitric acid–water at  $25^\circ\text{C}$ . *J. Chem. Soc., Trans.*, 101:576–585, 1912.
- [10] [Ernst Berl](#) and H. H. Saenger. Über das system  $\text{N}_2\text{O}_5$ – $\text{HNO}_3$ . *Monatsh. Chem. Ver. Wiss.*, 53(1):1036–1056, Dec 1929.
- [11] [Keith D. Beyer](#) and Anne R. Hansen. Phase diagram of the nitric acid/water system: Implications for polar stratospheric clouds. *J. Phys. Chem. A*, 106(43):10275–10284, oct 2002.

- [12] Keith D. Beyer, Anne R. Hansen, and Nick Raddatz. Experimental determination of the  $\text{H}_2\text{SO}_4/\text{HNO}_3/\text{H}_2\text{O}$  phase diagram in regions of stratospheric importance. *J. Phys. Chem. A*, 108(5):770–780, feb 2004.
- [13] William Robert Bousfield. CLV.—The study of the density and viscosity of aqueous solutions, with special reference to nitric acid. Part I. Densities. *J. Chem. Soc., Trans.*, 107:1405–1426, 1915.
- [14] William Robert Bousfield. III—Mixtures of nitrogen peroxide and nitric acid. *J. Chem. Soc., Trans.*, 115:45–55, 1919.
- [15] W. R. Bousfield. The determination of the ionisation of an aqueous solution. *Trans. Faraday Soc.*, 15:47–73, 1919.
- [16] T. Boublik and K. Kuchynka. Gleichgewicht flüssigkeit–dampf XXII. Abhängigkeit der zusammensetzung des azeotropischen gemisches des systems salpetersäure–wasser vom druck. *Collection of Czechoslovak chemical communications*, 25(2):579–582, 1960.
- [17] Stefano Brandani and Vincenzo Brandani. Vapor–liquid equilibrium calculation of the system water–nitric acid over the entire concentration range. *Fluid Phase Equilib.*, 114:37–45, 1996.
- [18] Peter Brimblecombe and Simon L. Clegg. Equilibrium partial pressures of strong acids over concentrated solutions—III. The temperature variation of  $\text{HNO}_3$  solubility. *Atmospheric Environment. Part A. General Topics*, 24(7):1945–1955, 1990.
- [19] Charles L. Burdick and E. Stanley Freed. The equilibrium between nitric oxide, nitrogen peroxide and aqueous solution of nitric acid. *J. Am. Chem. Soc.*, 43:518–530, 1921.
- [20] Clifford D. Carpenter and Joseph Babor. Concentrating dilute nitric acid solutions: A study of rate of concentration—Application of results to fractionation—Determination of composition of vapor above boiling solution of known composition. *Transactions of the American Institute of Chemical Engineers*, XIV:67–79, 1922. The years 1921–22 are collated in one volume.
- [21] Clifford D. Carpenter and Joseph A. Babor. Concentrating dilute nitric acid. *Chem. Metall. Eng.*, 31(7):260–262, Aug. 1924.
- [22] Carpenter, C. D. and J. A. Babor. Concentrating dilute nitric acid—A study of the relation between the vapor and liquid phases in the distillation of nitric acid–water solutions, and of nitric acid–sulphuric acid–water solutions. *Trans. Amer. Inst. Chem. Engrs.*, 16:111–148, 1924.
- [23] Giorgio Carta and Robert L. Pigford. Absorption of nitric oxide in nitric acid and water. *Ind. Eng. Chem. Fundam.*, 22:329–335, 1983.
- [24] Kenneth S. Carslaw, Simon L. Clegg, and Peter Brimblecombe. A thermodynamic model of the system  $\text{HCl}-\text{HNO}_3-\text{H}_2\text{SO}_4-\text{H}_2\text{O}$ , including solubilities of  $\text{HBr}$  from  $< 200$  to  $328$  K. *J. Phys. Chem.*, 99(29):11557–11574, 1995.
- [25] V. S. Chupalov, N. K. Muzychenko, and Yu. B. Stepanov. Study of equilibrium between nitrogen oxides and 60–80% nitric acid. *Zh. Prikl. Khim. (S.-Peterburg)*, 60(10):2179–2181, 1988.
- [26] Jean-Luc Clavelin and Philippe Mirabel. Détermination des pressions partielles du mélange eau–acide nitrique. *J. Chim. Phys. Phys.-Chim. Biol.*, 76(83):533–537, 1979.
- [27] Simon L. Clegg and Peter Brimblecombe. Equilibrium partial pressures and mean activity and osmotic coefficients of 0–100% nitric acid as a function of temperature. *J. Phys. Chem.*, 94(13):5369–5380, 1990.
- [28] S. L. Clegg, K. S. Carslaw, and P. Brimblecombe. Comment on “vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination” by D.-E. Taleb, J.-L. Ponche, and P. Mirabel: Part 1. *Journal of Geophysical Research: Atmospheres*, 102(D13):10867–10869, May 1997.
- [29] Simon L. Clegg, Peter Brimblecombe, and Anthony S. Wexler. Thermodynamic model of the system  $\text{H}^+ - \text{NH}_4^+ - \text{Na}^+ - \text{SO}_4^{2-} - \text{NO}_3^- - \text{Cl}^- - \text{H}_2\text{O}$  at 298.15 K. *J. Phys. Chem.*, 102(12):2155–2171, 1998.
- [30] Simon L. Clegg, Peter Brimblecombe, and Anthony S. Wexler. Thermodynamic model of the system  $\text{H}^+ - \text{NH}_4^+ - \text{SO}_4^{2-} - \text{NO}_3^- - \text{H}_2\text{O}$  at tropospheric temperatures. *J. Phys. Chem.*, 102(12):2137–2154, 1998.
- [31] S. L. Clegg, K. S. Carslaw, and P. Brimblecombe. Comment on “vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination” by D.-E. Taleb, J.-L. Ponche, and P. Mirabel: Part 2. *Journal of Geophysical Research: Atmospheres*, 103(D13):16291–16294, Jul 1998.
- [32] A. K. Covington and J. E. Prue. 301. Activity coefficients of nitric and perchloric acids in dilute aqueous solution from e.m.f. and transport-number measurements. *J. Chem. Soc.*, pages 1567–1572, 1957.
- [33] Henry Jermain Maude Creighton and John Horace Githens. On the boiling-point of aqueous solutions of nitric acid at different pressures. Part I. *J. Franklin Inst.*, 179(2):161–169, Feb 1915.
- [34] Henry Jermain Maude Creighton and Herschel Gaston Smith. On the boiling-point of aqueous solutions of nitric acid at different pressures. Part II.—The influence of water-retaining agents on the composition of the mixture of maximum boiling-point. *J. Franklin Inst.*, 180(6):703–709, Dec 1915.

- [35] Henry Jermain Maude Creighton. How the nitrogen problem has been solved. *J. Franklin Inst.*, 187(6):705–735, 1919.
- [36] Henry Jermain Maude Creighton. Electrolytic concentration of aqueous solutions of nitric acid. I. *J. Franklin Inst.*, 193(1):89–95, Jan 1922.
- [37] R. C. Crooks, R. Q. Wilson, A. E. Bearnse, and Jr. R. B. Filbert. Composition of vapors from boiling nitric acid solutions. Technical Report BMI-978, Battelle Memorial Institute, Columbus, Ohio, Feb 1955.
- [38] René Dalmon and René Freymann. Sur les spectres d'absorption dans le proche infrarouge de l'acide nitrique a 100% et en solution aqueuse. *Acad. Sci., Paris, C. R.*, 211(18):472–474, Nov. 1940.
- [39] W. Davis and H. J. De Bruin. New activity coefficients of 0–100 per cent aqueous nitric acid. *J. Inorg. Nucl. Chem.*, 26(6):1069–1083, 1964.
- [40] J. G. Dawber. Evidence for stability of the  $\text{H}_3\text{O}^+(\text{H}_2\text{O})_2$  and  $\text{NO}_3^- - (\text{HNO}_3)_2$  from vapour pressure data in the region of nitric acid monohydrate. *J. Inorg. Nucl. Chem.*, 37:1043–1044, 1975.
- [41] Edoardo Decanini, Giuliano Nardini, and Alessandro Paglianti. Absorption of nitrogen oxides in columns equipped with low-pressure drops structured packings. *Ind. Eng. Chem. Res.*, 39:5003–5011, 2000.
- [42] N. C. Deno, Henry J. Peterson, and Edward Sacher. Nitric acid equilibria in water–sulfuric acid. *J. Phys. Chem.*, 65(2):199–201, 1961.
- [43] J. A. Duisman and S. A. Stern. Vapor pressure and boiling point of pure nitric acid. *J. Chem. Eng. Data*, 14(4):457–459, oct 1969.
- [44] Dunning, W. J. and C. W. Nutt. Dissociation and equilibria of pure liquid nitric acid. *Trans. Faraday Soc.*, 47:15–25, 1951.
- [45] Egan, Jr., Edward P. Vapor pressure of liquid nitric acid. *Ind. Eng. Chem.*, 37(3):303–304, Mar. 1945.
- [46] Ellis, S. R. M. and J. M. Thwaites. Vapour–liquid equilibria of nitric acid–water–sulphuric acid mixtures. *J. appl. Chem.*, 7(4):152–160, Apr 1957.
- [47] Elverum, Jr., Gerard W. and David M. Mason. Melting point measurement of the system  $\text{HNO}_3 - \text{N}_2\text{O}_4 - \text{H}_2\text{O}$ . *J. Phys. Chem.*, 60:104–108, Jan. 1956.
- [48] Finlayson-Pitts, B. J., L. M. Wingen, A. L. Sumner, D. Syomin, and K. A. Ramazan. The heterogeneous hydrolysis of  $\text{NO}_2$  in laboratory systems and in outdoor and indoor atmospheres: An integrated mechanism. *Phys. Chem. Chem. Phys.*, 5(2):223–242, 2003.
- [49] R. Flatt and F. Benguerel. Sur l'équilibre liquide-vapeur du système binaire  $\text{HNO}_3 - \text{H}_2\text{O}$  á 25°C. *Helv. Chim. Acta*, 45(6):1765–1772, 1962.
- [50] R. Flatt and F. Benguerel. Sur l'équilibre liquide-vapeur á 25°C de systèmes ternaires composés d'un nitrate, d'acide nitrique et d'eau. *Helv. Chim. Acta*, 45(6):1772–1776, 1962.
- [51] W. R. Forsythe and W. F. Giauque. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 1(64):48–61, 1942.
- [52] W. R. Forsythe and W. F. Giauque. Additions and corrections. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 12(64):3069, 1942.
- [53] W. R. Forsythe and W. F. Giauque. Additions and corrections. The entropies of nitric acid and its mono- and tri-hydrates. Their heat capacities from 15 to 300°K. The heats of dilution at 298.1°K. The internal rotation and free energy of nitric acid gas. The partial pressures over its aqueous solutions. *J. Am. Chem. Soc.*, 12(65):2479, 1943.
- [54] R. J. Gillespie, E. D. Hughes, and C. K. Ingold. 504. Cryoscopic measurements in nitric acid. Part I. The solutes dinitrogen pentoxide and water. The self-dissociation of nitric acid. *J. Chem. Soc.*, pages 2552–2558, 1950.
- [55] D. R. Goddard, E. D. Hughes, and C. K. Ingold. 505. Chemistry of nitronium salts. Part I. Isolation of some nitronium salts. *J. Chem. Soc.*, pages 2559–2575, 1950.
- [56] Gutierrez-Cañas, C., P. L. Arias, and J. A. Legarreta. Industrial nitrogen oxides absorption simulation. *Comput. Chem. Eng.*, 13(9):985–1002, 1989.
- [57] R. Haase, K.-H. Dücker, and H. A. Küppers. Aktivitätskoeffizienten und dissoziationskonstanten wäßriger salpetersäure und überchlorsäure. *Ber. Bunsen-Ges.*, 69(2):97–109, 1965.
- [58] David R. Hanson and Konrad Mauersberger. Vapor pressures of  $\text{HNO}_3/\text{H}_2\text{O}$  solutions at low temperatures. *J. Phys. Chem.*, 92(21):6167–6170, Oct 1988.
- [59] Jackson E. Harrar, Lester P. Rigdon, and Steven F. Rice. Raman spectral study of solutions of  $\text{N}_2\text{O}_4$  and  $\text{N}_2\text{O}_5$  in nitric acid. *Journal of Raman Spectroscopy*, 28:891–899, 1997.

- [60] G. D. Honti. Nitric acid. In G. D. Honti, editor, *The Nitrogen Industry*, chapter 2.2 in volume 2 of part 1/2 of the collection. Akadémiai Kiadó, Budapest, 1976.
- [61] Hood, G. C. and C. A. Reilly. Ionization of strong electrolytes. VIII. Temperature coefficient of dissociation of strong acids by proton magnetic resonance. *J. Chem. Phys.*, 32:127–130, jan 1960.
- [62] Elmar-M. Horn, Helmuth Keiser, and Klaus Schoeller. Zur thermischen zersetzung der hochkonzentrierten salpetersäure. *Monatsh. Chem.—Chem. Monthly*, 118(11):1205–1218, Nov 1987.
- [63] Jaecker-Voirol, A., J. L. Ponche, and P. Mirabel. Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperatures. *Journal of Geophysical Research: Atmospheres*, 95(D8):11857–11863, 1990.
- [64] Jaecker-Voirol, A., J. L. Ponche, and P. Mirabel. Correction to “Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperatures” by A. Jaecker-Voirol et al. *Journal of Geophysical Research: Atmospheres*, 95(D13):22565, Dec 1990.
- [65] A. M. Kalinkin. Calculation of phase equilibria in the  $\text{NaNO}_3\text{--KNO}_3\text{--HNO}_3\text{--H}_2\text{O}$  system at 25 degrees C over the range from 0 to 20 m  $\text{HNO}_3$ . *Russian J. Appl. Chem.*, 69(5):664–668, May 1996. Translated from *Zhurnal Prikladnoi Khimii*, 69(5), 1996, pp. 743–748.
- [66] C. M. Karni, W. O. Rains, R. M. Counce, J. S. Watson, B. B. Spencer, and G. D. Del Cul. Water and mean ionic activities of aqueous  $\text{HNO}_3$  solutions calculated from an extension of the Brunauer–Emmett–Teller (BET) model. *Sep. Sci. Technol.*, 45(12–13):1894–1900, 2010.
- [67] S. G. Katal'nikov, A. V. Khoroshilov, and M. M. Chelyak. Isotope separation by the "nitric acid" method. Phase and isotope equilibrium in a system consisting of nitrogen oxides and nitric acid. *Soviet Atomic Energy*, 60(2):132–138, 1986. Translated from *Atomnaya Énergiya*, Vol. 60, No. 2, pp. 109–113, February, 1986.
- [68] Kay, Webster B. and S. Alexander Stern. Phase relations of nitric acid at physicochemical equilibrium. *Ind. Eng. Chem.*, 47(7):1463–1469, 1955.
- [69] Kay, W. B., S. A. Stern, and M. D. Sanghvi. Phase relations of nitric acid–nitrogen dioxide and nitric acid–water mixtures at physicochemical equilibrium. *Ind. Eng. Chem. Chem. Eng. Data Series*, 2(1):21–28, 1956.
- [70] Khoroshilov, A. V., T. V. Guseva, and S. G. Katal'nikov. Analysis of the chemical composition of the gas phase at equilibrium with nitric acid solutions. *Industrial Laboratory*, 48(10):947–951, 1982.
- [71] Alfons Klemenc and Erich Hayek. Zur kenntnis der salpetersäure. V. Über die einwirkung des stickoxydes auf salpetersäure bis zur erreichung des gleichgewichtes. (a) Das system ansteigend bis zu 1 n-salpetersäuregleichgewichtskonzentration in wäßriger lösung.  $3\text{HNO}_2 \leftrightarrow \text{HNO}_3 + 2\text{NO} + \text{H}_2\text{O}$ . *Z. Anorg. Allg. Chem.*, 186(1):181–224, 1929.
- [72] A. Klemenc and Th. Spiess. Zur kenntnis der salpetersäure. VIII. Über die lösungskurve des systems  $\text{HNO}_3\text{--NO}_2$  und  $\text{HNO}_3\text{--H}_2\text{O--NO}_2$ . *Monatsh. Chem. Ver. Wiss.*, 77(1):216–223, Dec 1947.
- [73] A. Klemenc. Zur kenntnis der salpetersäure. X. Methoden zur gasanalyse im system  $\text{N}_2\text{--NO--NO}_2\text{--N}_2\text{O}_4\text{--N}_2\text{O}_3\text{--HNO}_2\text{--HNO}_3$ . *Monatsh. Chem. Ver. Wiss.*, 83(2):334–345, Mar 1952. English translation by K. S. Bevis in AEC-tr-2450 entitled: Information on Nitric Acid. X. Gas Analysis Methods in the System  $\text{N}_2\text{--NO--NO}_2\text{--N}_2\text{O}_4\text{--N}_2\text{O}_3\text{--HNO}_2\text{--HNO}_3$ .
- [74] Alexander Kolker and Juan J. de Pablo. Thermodynamic modeling of concentrated aqueous electrolyte and nonelectrolyte solutions. *AIChE J.*, 41(6):1563–1571, 1995.
- [75] Y.-N. Lee and S. E. Schwartz. Reaction kinetics of nitrogen dioxide with liquid water at low partial pressures. *J. Phys. Chem.*, 85:840–848, 1981.
- [76] Jan Bernard Lefers. *Absorption of Nitrogen Oxides into Diluted and Concentrated Nitric Acid*. Delft University Press, 1980.
- [77] J. B. Lefers and P. J. van den Berg. Absorption of  $\text{NO}_2/\text{N}_2\text{O}_4$  into diluted and concentrated nitric acid. *Chem. Eng. J. (Lausanne)*, 23:211–221, 1982.
- [78] Robert J. Lemire, Colin P. Brown, and Allan B. Campbell. Vapor pressure of nitric acid and water in the systems  $\text{HNO}_3\text{--H}_2\text{O}$  and  $\text{HNO}_3\text{--Th(NO}_3)_4\text{--H}_2\text{O}$  at 50°C. *J. Chem. Eng. Data*, 30(4):421–424, 1985.
- [79] Gilbert N. Lewis and Arthur Edgar. The equilibrium between nitric acid, nitrous acid and nitric oxide. *J. Am. Chem. Soc.*, 33(3):292–299, 1911.
- [80] Yunda Liu, David Bluck, and Francisco Brana-Mulero. Static and dynamic simulation of  $\text{NO}_x$  absorption tower based on a hybrid kinetic-equilibrium reaction model. In Mario R. Eden, John D. Siirola, and Gavin P. Towler, editors, *Proceedings of the 8th International Conference on Foundations of Computer-Aided Process Design*, volume 34 of *Computer Aided Chemical Engineering*. Elsevier, Amsterdam, Holland, 2014.
- [81] L. Lloyd and P. A. H. Wyatt. The vapour pressure of nitric acid solutions. Part I. New azeotropes in the water–dinitrogen pentoxide system. *J. Chem. Soc.*, pages 2248–2252, 1955.

- [82] [Arturo Maimoni](#). Vapor–liquid equilibria for nitric acid–water and plutonium nitrate–nitric acid–water solutions. Technical Report UCID-18148, Lawrence Livermore Laboratory (California Univ.), Jan 1980.
- [83] [Mario Massucci](#), Simon L. Clegg, and Peter Brimblecombe. Equilibrium partial pressure, thermodynamic properties of aqueous and solid phases and  $\text{Cl}_2$  production from aqueous  $\text{HCl}$  and  $\text{HNO}_3$  and their mixtures. *J. Phys. Chem. A*, 103:4209–4226, 1999.
- [84] [H. McKenzie](#), J. MacDonald-Taylor, F. McLachlan, R. Orr, and D. Woodhead. Modelling of nitric and nitrous acid chemistry for solvent extraction purposes. *Procedia Chemistry*, 21:481–486, 2016. ATALANTE 2016 International Conference on Nuclear Chemistry for Sustainable Fuel Cycles.
- [85] [Anderson B. McKeown](#) and Frank E. Belles. Vapor pressures of concentrated nitric acid solutions in the composition range 83 to 97 percent nitric acid, 0 to 6 percent nitrogen dioxide, 0 to 15 percent water, and in the temperature range  $20^\circ$  to  $80^\circ\text{C}$ . Technical Report RM-E53G08, National Advisory Committee for Aeronautics (NACA), Washington, Sep 1953.
- [86] [Anderson B. McKeown](#) and Frank E. Belles. Vapor pressures and calculated heats of vaporization of concentrated nitric acid solutions in the composition range 71 to 89 percent nitric acid, 7 to 20 percent nitrogen dioxide, 1 to 10 percent water, and in the temperature range  $10^\circ$  to  $60^\circ\text{C}$ . Technical Report RM-E53L14, National Advisory Committee for Aeronautics (NACA), Washington, Feb 1954.
- [87] [Anderson B. McKeown](#) and Frank E. Belles. Nitric acid–nitrogen dioxide–water system—Vapor pressures and related properties. *Ind. Eng. Chem.*, 47(12):2540–2543, 1955.
- [88] [I. R. McKinnon](#), J. G. Mathieson, and I. R. Wilson. Gas phase reaction of nitric oxide with nitric acid. *J. Phys. Chem.*, 83(7):779–780, 1979.
- [89] [Frank Douglas Miles](#). *Nitric Acid: Manufacture and Uses*, pages 9–19. Oxford University Press, London, 1961.
- [90] [D. N. Miller](#). Mass transfer in nitric acid absorption. *AIChE J.*, 33(8):1351–1358, 1987.
- [91] [L. I. Olefir](#), M. I. Knyshenko, and L. K. Priima. The shift of the azeotropic point in the system  $\text{HNO}_3\text{--H}_2\text{O}$ . *The Soviet chemical industry*, 10(2):127–128, 1978. The printed article says pp.127–128 but the receipt coming back from the library says pp.123–124.
- [92] [J. L. Oscarson](#), S. E. Gillespie, R. M. Izatt, X. Chen, and C. Pando. Thermodynamic quantities for the ionization of nitric acid in aqueous solution from 270 to  $319^\circ\text{C}$ . *J. Solution Chem.*, 21(8):789–801, 1992.
- [93] [Jaakko Ilmari Partanen](#), Pekka Mikael Juusola, and Pentti Olavi Minkkinen. Density of aqueous nitric acid solutions in the molality range 0– $3.5\text{ mol kg}^{-1}$  at 293.15, 298.15, 303.15 and  $308.15\text{ K}$ . *Acta Chem. Scand.*, 46:338–342, 1992.
- [94] [J. P. Passarello](#) and W. Fürst. Representation of the equilibrium properties of the  $\text{H}_2\text{O--HNO}_3\text{--N}_2\text{O}_5$  systems using the MSA electrolyte model. *Fluid Phase Equilib.*, 116:177–184, 1996.
- [95] [B. A. Patterson](#) and E. M. Woolley. Thermodynamics of ionization of water at temperatures  $278.15 < T/\text{K} < 393.15$  and at the pressure  $p = 0.35\text{ MPa}$ : Apparent molar volumes and apparent molar heat capacities of aqueous solutions of potassium and sodium nitrates and nitric acid. *J. Chem. Thermodyn.*, 34(4):535–556, 2002.
- [96] [F. P. Pike](#). Vapor–liquid equilibrium in water–nitric acid systems at 760mmHg. Technical Report CF-50-6-143, Oak Ridge National Laboratory (US), 1950.
- [97] [M. H. R. J. Plusjé](#). Theory and practice in the treatment of phosphate rock with nitric acid. In *Nitrophosphates*, number 13 in Proceedings (Technical papers), pages 3–31. International Fertilizer Society, 1951.
- [98] [Antoine Potier](#). Préparation de l’acide nitrique au laboratoire. Densité à  $15^\circ\text{C}$ . *C. R. Hebd. Seances Acad. Sci.*, 233(19):1113–1115, 1951. Présentée par Paul Pascal.
- [99] [Jacqueline Potier](#) and Antoine Potier. Ébulliométrie de l’acide nitrique. *Bull. Soc. Chim. Fr.*, 19(3–4):250–251, 1952.
- [100] [A. Potier](#). Préparation du peroxyde d’azote pur au laboratoire. *Bull. Soc. Chim. Fr., Mem.*, 20(139):709–711, 1953.
- [101] [A. Potier](#). Propriétés ionisantes et ionisation de l’acide nitrique anhydre. *Bull. Soc. Chim. Fr.*, 21(2):166, Jan 1954. Section d’Alger.
- [102] [A. Potier](#). Calcul des pressions partielles isothermes des systèmes binaires liquides volatils à partir des courbes isothermes pressions totales. Composition. *Bull. Soc. Chim. Fr.*, 22(3):316, 1955. Section d’Alger.
- [103] [Antoine Potier](#). I. Tensions de vapeur totales et partielles des solutions acides nitrique–eau entre  $0^\circ$  et  $20^\circ\text{C}$ . Densités des solutions concentrées à  $15^\circ\text{C}$ . *Bull. Soc. Chim. Fr., Mem.*, 23(5):47–49, 1956.
- [104] [Antoine Potier](#). II. Tensions de vapeur totales et partielles des solutions acides nitrique–peroxyde d’azote entre  $-10^\circ$  et  $+10^\circ\text{C}$ , densités à  $15^\circ\text{C}$ . *Bull. Soc. Chim. Fr., Mem.*, 23(6):50–53, 1956.
- [105] [J. Potier](#) and A. Potier. Sur le système acide nitrique–eau. *Bull. Soc. Chim. Fr.*, 21(2):1119–1120, 1956. Section d’Alger.

- [106] **A. Potier**. Étude des propriétés thermodynamiques des systèmes. Acide nitrique–Eau et peroxyde d’azote. Acide nitrique. *Annales de la Faculté des sciences de Toulouse: Mathématiques*, 4<sup>e</sup> série, 20:1–98, 1956. Published by: Gauthier-Villars, Libraire-Éditeur, Paris.
- [107] **Jacqueline Potier**. Étude ébulliométrique du système  $\text{NO}_3\text{H}-\text{OH}_2$ . *Alger. Sciences Physiques.*, IV(1):61–86, 1957.
- [108] **Antoine Potier** and Jacqueline Potier. Les binaires acide nitrique–nitrates «alcalins». II. Équilibres liquide–solide dans le système  $\text{HNO}_3-\text{NH}_4\text{NO}_3$ . *Bull. Soc. Chim. Fr., Mem.*, 26(51):311–314, 1959.
- [109] **M. P. Pradhan**, N. J. Suchak, P. R. Walse, and J. B. Joshi. Multicomponent gas absorption with multiple reactions: Modelling and simulation of  $\text{NO}_x$  absorption in nitric acid manufacture. *Chem. Eng. Sci.*, 52(24):4569–4591, 1997. R. A. Mashelkar (editor), Festschrift for Professor M. M. Sharma, 52(24), pp. 4437–4698 (1997).
- [110] **K. J. Prösek**. Gleichgewicht zwischen der flüssigen und der gasförmigen phase in salpetersäure enthaltenden systemen. *Coll. of Czechoslovak Chem. Comm.*, 32(7):2397–2404, 1967.
- [111] **O. Redlich** and J. Bigeleisen. The ionization of strong electrolytes. I. General remarks, nitric acid. *J. Am. Chem. Soc.*, 65:1883–1887, 1943.
- [112] **O. Redlich** and G. C. Hood. Ionic interaction, dissociation and molecular structure. *Discuss. Faraday Soc.*, 24:87–93, 1957.
- [113] **O. Redlich**, W. E. Gargrave, and W. D. Krostek. Thermodynamics of solutions. consistency of sparse data. Activities of nitric and perchloric acids. *Ind. Eng. Chem. Fundam.*, 7(2):211–214, May 1968.
- [114] **R. C. Reid**, A. B. Reynolds, D. T. Morgan, G. W. Bond, J. E. Savolainen, and M. L. Hyman. Vapor-liquid equilibria. Dilute nitric acid, hydrochloric acid, and hydrochloric acid–nitric acid–water solutions. *Ind. Eng. Chem.*, 49(8):1307–1310, 1957.
- [115] **Robert T. Rewick** and B. J. Gikis. The heat of solution of ammonium nitrate in nitric acid. *J. Chem. Eng. Data*, 25(2):127–131, 1980.
- [116] **Theodore W. Richards** and Allan Winter Rowe. The heats of dilution and the specific heats of dilute solutions of nitric acid and of hydroxides and chlorides and nitrates of lithium, sodium, potassium, and cesium. *J. Am. Chem. Soc.*, 43(4):770–796, 1921.
- [117] **Eric K. Rideal**. On the absorption of oxides of nitrogen by nitric acid. *J. Ind. & Eng. Chem.*, 12(6):531–538, 1920.
- [118] **Fernand Rivenq**. Étude ébulliométrique des solutions aqueuses d’acide nitrique sous pression réduite. *Bull. Soc. Chim. Fr.*, 26(6):822–826, 1959. English translation of title: Boiling point diagrams of aqueous nitric acid mixtures under reduced pressure.
- [119] **Ruas, Alexandre**, Patrick Pochon, Jean-Pierre Simonin, and Philippe Moisy. Nitric acid: Modeling osmotic coefficients and acid–base dissociation using the BIMSA theory. *Dalton Trans.*, 39(42):10148–10153, 2010.
- [120] **Bo Sander**, Peter Rasmussen, and Aage Fredenslund. Calculation of vapour–liquid equilibria in nitric acid–water–nitrate salt systems using an extended UNIQUAC equation. *Chem. Eng. Sci.*, 41(5):1185–1195, 1986.
- [121] **Akira Sasahira**, Tadahiro Hoshikawa, and Fumio Kawamura. Hydration model for evaluation of thermodynamic activities in nitric acid–nitrate–water system. *J. Nucl. Sci. Technol.*, 29(6):566–575, Jun. 1992.
- [122] **James A. Snyder**, Denise Hanway, Julio Mendez, Alan J. Jamka, and Fu-Ming Tao. A density functional theory study of the gas-phase hydrolysis of dinitrogen pentoxide. *J. Phys. Chem.*, 103:9355–9358, 1999.
- [123] **N. N. Solov’ev**, V. P. Panov, and L. Ya. Tereshchenko. Calculating the equilibrium composition for the liquid-phase  $\text{HNO}_x-\text{NO}_x-\text{H}_2\text{O}$ . *Zh. Prikl. Khim. (Leningrad)*, 57(10):2029–2031, oct 1984.
- [124] **William C. Sproesser** and Guy B. Taylor. Vapor pressures of aqueous solutions of nitric acid. *J. Am. Chem. Soc.*, 43(8):1782–1787, 1921.
- [125] **B. G. Staples** and J. M. Procopio, Jr. Vapor-pressure data for common acids at high temperature. *Chem. Eng. [Int. Ed.]*, 77(25):113–115, Nov. 1970.
- [126] **Stern, S. Alexander**, J. T. Mullhaupt, and Webster B. Kay. The physicochemical properties of pure nitric acid. *Chem. Rev. (Washington, D. C.)*, 60(2):185–207, 1960.
- [127] **H. I. Stonehill**. 170. Transport numbers of nitric acid in water at 25° from e.m.f. measurements. *J. Chem. Soc.*, pages 647–651, 1943.
- [128] **Taleb, Djamel-Eddine**, Jean-Luc Ponche, and Philippe Mirabel. Vapor pressures in the ternary system water–nitric acid–sulfuric acid at low temperature: A reexamination. *Journal of Geophysical Research: Atmospheres*, 101(D20):25967–25977, Nov 1996.
- [129] **Ignatius N. Tang**. On the equilibrium partial pressures of nitric acid and ammonia in the atmosphere. *Atmospheric Environment (1967)*, 14(7):819–828, 1980.
- [130] **I. N. Tang**, H. R. Munkelwitz, and J. H. Lee. Vapor–liquid equilibrium measurements for dilute nitric acid solutions. *Atmospheric Environment (1967)*, 22(11):2579–2585, 1988.

- [131] [Guy B. Taylor](#). Vapor pressure of aqueous solutions of nitric acid. *Ind. Eng. Chem.*, 17(6):633–635, 1925.
- [132] [L. Ya. Tereshchenko](#), M. I. Kucha, V. P. Panoc, and V. V. Zubov. Equilibrium of nitrogen oxides with nitric acid solutions. *Zh. Prikl. Khim. (S.-Peterburg)*, 52(8):1657–1661, 1980.
- [133] [H. Theobald](#). Messungen zum Gleichgewicht Salpetersäure/nitrose Gase. *Chem.-Ing.-Tech.*, 40(15):763–765, 1968.
- [134] [Emmanuel Thibert](#) and Florent Domine. Thermodynamics and kinetics of the solid solution of HNO<sub>3</sub> in ice. *J. Phys. Chem. B*, 102(22):4432–4439, 1998.
- [135] [N. N. Treushchenko](#), M. E. Pozin, G. V. Bel'chenko, and E. Ya. Ivanova. Vapor pressure of the system HNO<sub>3</sub>–NH<sub>4</sub>NO<sub>3</sub>–H<sub>2</sub>O. *J. Appl. Chem. USSR*, 44(8):1775–1780, 1997. Translated from Zhurnal Prikladnoi Khimii, 44(8), August 1971, pp. 1753–1760.
- [136] [V. G. Tsvetkov](#), V. F. Sopin, V. N. Marshева, L. Ya. Tsvetkova, E. M. Belova, and G. N. Marchenko. Enthalpy of water mixturing with sulfuric and nitric acid mixtures. *Zh. Obshch. Khim.*, 57(5):997–1000, 1987.
- [137] [V. G. Tsvetkov](#), V. F. Sopin, V. N. Marshева, L. Ya. Tsvetkova, E. M. Belova, and G. N. Marchenko. Enthalpy of water mixturing with sulfuric and nitric acid mixtures. *J. Gen. Chem. USSR (Transl. of Zh. Obshch. Khim.)*, 57(5):888–991, 1987. Translated from Zhurnal Obshchei Khimii, 57(5), May 1987, pp. 997–1000.
- [138] [A. G. Udovenko](#), V. I. Konvisar, A. V. Shapka, and A. I. Mishchenko. Equilibrium of nitrogen oxides with nitric acid solutions under pressure. *Zh. Prikl. Khim. (S.-Peterburg)*, 49(6):1274–1276, 1976.
- [139] [Robert Vandoni](#) and Michele Laudy. Détermination des tensions de vapeurs partielles des mélanges NO<sup>3</sup>H–N<sup>2</sup>O<sup>4</sup>–NO<sup>2</sup>. *Mem. services chim. etat (Paris)*, 36:261–268, 1951.
- [140] [R. Vandoni](#) and M. Laudy. Mesure de tensions de vapeur partielles des mélanges NO<sup>3</sup>H–H<sup>2</sup>O à 20° C et vérification de l'équation de Margules–Duhem. 1<sup>re</sup> partie.—Mesure des tensions de vapeur partielles des mélanges NO<sup>3</sup>H–H<sup>2</sup>O à 20° C. *J. Chim. Phys. Phys.-Chim. Biol.*, 49:99–102, see also pp. 103–108, 1952. English translation by K. S. Bevis in AEC-tr-2913 entitled: Measurement of Partial Pressures of HNO<sub>3</sub>–H<sub>2</sub>O mixtures at 20 C and Verification of the Margules-Duhem Equation.
- [141] [Frank H. Verhoek](#) and Farrington Daniels. The dissociation constants of nitrogen tetroxide and of nitrogen trioxide. *J. Am. Chem. Soc.*, 53(4):1250–1263, Apr. 1931.
- [142] [Klaus Vetter](#). Das gleichgewicht zwischen stickstoffdioxyd und salpetriger säure in salpetersäure. *Z. Anorg. Chem.*, 260(4–5):242–248, 1949.
- [143] [D. M. Waldorf](#) and A. L. Babb. Vapor-phase equilibrium of NO, NO<sub>2</sub>, H<sub>2</sub>O, and HNO<sub>2</sub>. *J. Chem. Phys.*, 39(2):432–435, Jul. 1963.
- [144] [Meng Wang](#), Maximilian B. Gorensek, and Chau-Chyun Chen. Thermodynamic representation of aqueous sodium nitrate and nitric acid solution with electrolyte NRTL model. *Fluid Phase Equilib.*, 407:105–116, 2016.
- [145] [Meng Wang](#), Harnoor Kaur, and Chau-Chyun Chen. Thermodynamic modeling of HNO<sub>3</sub>–H<sub>2</sub>SO<sub>4</sub>–H<sub>2</sub>O ternary system with symmetric electrolyte NRTL model. *AIChE J.*, 63(7):3110–3117, 2017.
- [146] [Werner Weisweiler](#) and Karl-Heinz Deiß. Influence of electrolytes on the absorption of nitrogen oxide components N<sub>2</sub>O<sub>4</sub> and N<sub>2</sub>O<sub>3</sub> in aqueous absorbents. *Chem. Eng. Technol.*, 10:131–142, 1987.
- [147] [Werner Weisweiler](#), Klaus Eidam, Michael Thiemann, Erich Scheibler, and Karl Wilhelm Wiegand. Absorption of NO<sub>2</sub>/N<sub>2</sub>O<sub>4</sub> in nitric acid. *Chem. Eng. Technol.*, 13:97–101, 1990.
- [148] [G. L. Wilson](#) and F. D. Miles. The partial pressures of nitric acid–water mixtures from 0°–20°C. *Trans. Faraday Soc.*, 35:356–363, 1940.
- [149] [P. A. H. Wyatt](#). Vapour phase reactions and the Duhem–Margules equation (with some reference to nitric acid). *Trans. Faraday Soc.*, 50:352–357, 1954.
- [150] [I. I. Zakharov](#). Quantum-chemical modeling of the mechanism for formation of HNO<sub>3</sub> from NO<sub>2</sub> and water. *Theoretical and Experimental Chemistry*, 48(4):233–239, Sep 2012. Translated from Teoreticheskaya i Éksperimental'naya Khimiya, 48(4), pp. 217–223, July–August, 2012.
- [151] [Renyi Zhang](#), Paul J. Wooldridge, and Mario J. Molina. Vapor pressure measurements for the H<sub>2</sub>SO<sub>4</sub>/HNO<sub>3</sub>/H<sub>2</sub>O and H<sub>2</sub>SO<sub>4</sub>/HCl/H<sub>2</sub>O systems: Incorporation of stratospheric acids into background sulfate aerosols. *J. Phys. Chem.*, 97:8541–8548, 1993.
- [152] [B. Ya. Zilberman](#), D. V. Ryabkov, E. A. Puzikov, E. V. Andreeva, and N. E. Mishina. Influence of pressure (temperature) on the nitric acid distribution between the liquid and vapor in the course of evaporation of nitric acid radioactive waste. *Radiochemistry*, 58(3):237–242, Sep 2016. Translated from Radiokhimiya, 58(3), pp. 206–210, 2016.

## Formaldehyde (DatCHOH)

Experimental data for formaldehyde.

Related keywords: (experimental | data) & hcoh | (ch3coh & h2o).

### Bibliography.

- [1] M. Albert, B. C. Garcia, C. Kuhnert, R. Peschla, and G. Maurer. Vapor–liquid equilibrium of aqueous solutions of formaldehyde and methanol. *AIChE J.*, 46(8):1676–1687, 2000.
- [2] M. Albert, I. Hahnenstein, H. Hasse, and G. Maurer. Vapor–liquid and liquid–liquid equilibria in binary and ternary mixtures of water, methanol, and methylal. *J. Chem. Eng. Data*, 46(4):897–903, 2001.
- [3] M. Albert, I. Hahnenstein, H. Hasse, and G. Maurer. Vapor–liquid equilibrium of formaldehyde mixtures: New data and model revision. *AIChE J.*, 42(6):1741–1752, 1996.
- [4] M. Albert, B. C. Garcia, C. Kreiter, and G. Maurer. Vapor–liquid and chemical equilibria of formaldehyde–water mixtures. *AIChE J.*, 45(9):2024–2033, 1999.
- [5] P. Alessi, R. Bugarin, M. Fermeglia, and I. Kikic. Thermodynamics of water formaldehyde systems. *World Congress III of Chemical Engineering, Tokyo*, 7a(307):164–167, 1986.
- [6] V. Brandani, G. Digiaco, and V. Mucciante. A test for the thermodynamic consistency of VLE data for the systems water formaldehyde and methanol formaldehyde. *Ind. Eng. Chem. Res.*, 26(6):1162–1167, 1987.
- [7] S. Brandani, V. Brandani, and G. Digiaco. A physical theory superimposed onto the chemical theory for describing vapor–liquid-equilibria of binary-systems of formaldehyde in active solvents. *Ind. Eng. Chem. Res.*, 30(2):414–420, 1991.
- [8] S. Brandani, V. Brandani, and G. Digiaco. Vapor–liquid-equilibrium of water trioxane formaldehyde mixture. *Fluid Phase Equilib.*, 63:27–41, 1991.
- [9] S. Brandani, V. Brandani, and G. Digiaco. The system formaldehyde water–methanol—, thermodynamics of solvated and associated solutions. *Ind. Eng. Chem. Res.*, 31(7):1792–1798, 1992.
- [10] Stefano Brandani, Vincenzo Brandani, and Domenica Flammini. Isothermal vapor–liquid-equilibria for the water–1,3,5-trioxane system. *J. Chem. Eng. Data*, 39(1):184–185, 1994.
- [11] Stefano Brandani and Vincenzo Brandani. Isothermal vapor–liquid-equilibria and solubility in the system methanol plus 1,3,5-trioxane. *J. Chem. Eng. Data*, 39(2):203–204, 1994.
- [12] Stefano Brandani, Vincenzo Brandani, and Domenica Flammini. Solubility of trioxane in water. *J. Chem. Eng. Data*, 39(2):201–202, 1994.
- [13] S. Brandani, V. Brandani, and I. Tarquini. Vapor–liquid equilibrium of formaldehyde mixtures containing methanol. *Ind. Eng. Chem. Res.*, 37(8):3485–3489, 1998.
- [14] Stanley J. Green and Raymond E. Vener. Vapor–liquid equilibria of formaldehyde–methanol–water. *Ind. Eng. Chem.*, 47(1):103–109, jan 1955.
- [15] H. Hasse, I. Hahnenstein, and G. Maurer. Revised vapor–liquid-equilibrium model for multicomponent formaldehyde mixtures. *AIChE J.*, 36(12):1807–1814, 1990.
- [16] H. Hasse and G. Maurer. Kinetics of the poly(oxyethylene) glycol formation in aqueous formaldehyde solutions. *Ind. Eng. Chem. Res.*, 30(9):2195–2200, 1991.
- [17] H. Hasse and G. Maurer. Vapor–liquid-equilibrium of formaldehyde-containing mixtures at temperatures below 320 K. *Fluid Phase Equilib.*, 64:185–199, 1991.
- [18] H. Hasse and G. Maurer. Heat of dilution in aqueous and methanolic formaldehyde solutions. *Ber. Bunsen-Ges.*, 96(1):83–96, 1992.
- [19] Robert A. Kumpf and James R. DameWood, Jr. Interaction of formaldehyde with water. *J. Phys. Chem.*, 93:4478–4486, 1989.
- [20] Y. Q. Liu, H. Hasse, and G. Maurer. Enthalpy change on vaporization of aqueous and methanolic formaldehyde solutions. *AIChE J.*, 38(11):1693–1702, 1992.
- [21] J. Masamoto and K. Matsuzaki. A study of vapor–liquid-equilibrium of formaldehyde–water mixtures using chemical theory. *J. Chem. Eng. Jpn.*, 27(1):6–11, 1994.
- [22] J. Masamoto, K. Matsuzaki, and H. Morishita. Purification of formaldehyde by distillation. *J. Chem. Eng. Jpn.*, 27(1):12–16, 1994.
- [23] G. Maurer. Vapor–liquid equilibrium of formaldehyde- and water-containing multicomponent mixtures. *AIChE J.*, 32(6):932–948, 1986.
- [24] Edgar L. Piret and M. W. Hall. Distillation principles of formaldehyde solutions. *Ind. Eng. Chem.*, 40(4):661–672, apr 1948.
- [25] Z. J. Shan, Y. R. Wang, S. C. Qiu, C. Y. Zheng, and J. Shi. Vapor–liquid-equilibria for the quaternary system of formaldehyde(1)–methanol(2)–methylal(3)–water(4). *Fluid Phase Equilib.*, 111(1):113–126, 1995.
- [26] O. D. Soboleva, S. V. Kazakova, Yu. M. Blazhin, and S. K. Ogorodnikov. Liquid–vapor equilibrium in the formaldehyde–water system. *Zh. Prikl. Khim. (S.-Peterburg)*, 57(4):860–865, 1984.

## Cryolite (DatCry)

Experimental data for  $\text{NaF} - \text{AlF}_3 - \text{Al}_2\text{O}_3$  plus  $\text{CaF}_2$ ,  $\text{LiF}$ , etc.

**Related keywords: (experimental | data) & cryolite | naf | alf3 | na3alf6 | al2o3.**

### Bibliography.

- [1] Maggie Aldén. On the homogeneity range of  $\beta''$ -alumina in the  $\text{Na}_2\text{O} - \text{MgO} - \text{Al}_2\text{O}_3$  system. *Solid State Ionics*, 20(1):17–23, 1986.
- [2] Mohammad Alivaliollahi, Johannes Gerlach, and Ursula Henning. Der Einfluß von Badzusätzen auf die Dampfdrücke  $\text{NaF} - \text{AlF}_3$ -haltiger Schmelzen. *Erzmetall*, 31(5):220–224, 1978.
- [3] M. E. Baud. Chaleurs spécifiques de l'aluminium et de ses sels solides. *J. Phys. Radium*, 2:569–573, 1903.
- [4] J. Brynestad, K. Grjotheim, and S. Urnes. Structural interpretation of molten cryolite and of cryolite melts containing aluminium oxide. *Metall. Ital.*, 8:495–502, 1960.
- [5] Gary L. Bullard and Dennis D. Przybycien. DTA determination of bath liquidus temperatures: Effect of  $\text{LiF}$ . In R. E. Miller, editor, *Light Metals 1986*, number 115 in Annual Meeting, pages 437–444, Warrendale, Pennsylvania, 2<sup>nd</sup>–6<sup>th</sup> Mar. 1986. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [6] Stanley Cantor and Wilfred T. Ward. Freezing point depressions in sodium fluoride. IV. Effects on trivalent fluorides. *J. Phys. Chem.*, 67:2766–2769, Dec. 1963.
- [7] Douglas F. Craig and Jesse J. Brown, Jr. Phase equilibria in the system  $\text{CaF}_2 - \text{AlF}_3 - \text{Na}_3\text{AlF}_6$  and part of the system  $\text{CaF}_2 - \text{AlF}_3 - \text{Na}_3\text{AlF}_6 - \text{Al}_2\text{O}_3$ . *J. Am. Ceram. Soc.*, 63(5/6):254–261, May/Jun. 1980.
- [8] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part I: The equilibrium  $\text{NaF}(s) + \text{Al} = \text{Na}_3\text{AlF}_6(s) + 3\text{Na}$ . *Metall. Trans.*, 1:1691–1694, Jun. 1970.
- [9] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part II: The free energies of formation of cryolite ( $\text{Na}_3\text{AlF}_6$ ) and chiolite ( $\text{Na}_5\text{Al}_3\text{F}_{14}$ ). *Metall. Trans.*, 1:2211–2215, Aug. 1970.
- [10] Ernest W. Dewing. Liquidus curves for aluminium cell electrolyte V. Representation by regression equations. *J. Electrochem. Soc.*, 117(6):780–781, Jun. 1970.
- [11] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part III: Activities in liquid mixtures. *Metall. Trans.*, 3:495–501, Feb. 1972.
- [12] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part IV: The cryolite liquidus curve. *Metall. Trans.*, 3:2699–2702, Oct. 1972.
- [13] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part V: Solid solution in cryolite. *Metall. Trans. B*, 9B:687–690, Dec. 1978.
- [14] Ernest W. Dewing. Thermodynamic functions for  $\text{LiF} - \text{AlF}_3$  mixtures at 1293 K. *Metall. Trans. B*, 11B:245–249, Jun. 1980.
- [15] Ernest W. Dewing. The solubility of titanium diboride in aluminium. *Metall. Trans. A*, 20A:2185–2187, Oct. 1989.
- [16] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ . Part VI: Revision. *Metall. Trans. B*, 21B:285–294, Apr. 1990.
- [17] Ernest W. Dewing. Thermodynamics of the system  $\text{NaF} - \text{AlF}_3$ : Part VII. Non-stoichiometric solid cryolite. *Metall. Mater. Trans. B*, 28B:1095–1097, Dec 1997.
- [18] E. W. Dewing, S. Rolseth, L. Støen, and J. Thonstad. The solubility of  $\text{ZnO}$  and  $\text{ZnAl}_2\text{O}_4$  in cryolite melts. *Metall. Mater. Trans. B*, 28B:1099–1101, Dec 1997.
- [19] Ernest W. Dewing and Jomar Thonstad. Activities in the system cryolite–alumina. *Metall. Mater. Trans. B*, 28:1089–1093, Dec 1997.
- [20] Thomas B. Douglas and David A. Ditmars. Measured relative enthalpy and derived thermodynamic properties of anhydrous crystalline aluminium trifluoride,  $\text{AlF}_3$ , from 273 to 1173°K. Technical Report 9389, National Bureau of Standards, 1966.
- [21] Thomas B. Douglas and David A. Ditmars. Measured relative enthalpy of anhydrous crystalline aluminium trifluoride,  $\text{AlF}_3$ , from 273 to 1173°K and derived thermodynamic properties from 273 to 1600°K. *J. Res. Natl. Bur. Stand., Sect. A*, 71A(3):185–193, May/Jun. 1967.
- [22] Anne Fenerty and E. A. Hollingshead. Liquidus curves for aluminium cell electrolyte III. Systems cryolite and cryolite–alumina with aluminium fluoride and calcium fluoride. *J. Electrochem. Soc.*, 107(12):993–997, Dec. 1960.
- [23] Foster, Jr., Perry A. Phase diagram of a portion of the system  $\text{Na}_3\text{AlF}_6 - \text{AlF}_3 - \text{Al}_2\text{O}_3$ . *J. Am. Ceram. Soc.*, 58(7/8):288–291, Jul./Aug. 1975.
- [24] W. B. Frank and L. M. Foster. The constitution of cryolite and  $\text{NaF} - \text{AlF}_3$  melts. *J. Phys. Chem.*, 64:95–98, Jan. 1960.

- [25] W. B. Frank. Thermodynamic considerations in the aluminium-producing electrolyte. *J. Phys. Chem.*, 65:2081–2087, Nov 1961.
- [26] Johannes Gerlach, Ursula Henning, and Manfred Mücke. Dampfdruckmessungen an NaF–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>–Al-Schmelzen. *Erzmetall*, 26(10):496–500, 1973.
- [27] J. Gerlach, U. Hennig, and R. Roedel. The solubility of aluminium in NaF–AlF<sub>3</sub> melts. *Metall (Berlin)*, 29(3):267–269, 1975.
- [28] H. Ginsberg and K. Resch. Beiträge zur Kenntnis des Zustandes der Salzschnmelzen der Systeme NaF–AlF<sub>3</sub> und NaF–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>. *Erzmetall*, 8(11):523–533, 1960.
- [29] Tor Grande and Jørn Rutlin. Termisk analyse. Sluttappoort, Institutt for uorganisk kjemi, Norges Tekniske Høgskole, Universitetet i Trondheim, Aug 1995.
- [30] Kai Grjotheim. Contribution to the theory of the aluminium electrolysis. In *Det Kgl. Norske Videnskabers Selskabs Skrifter*, volume 5, pages 16–18. Det Kongelige Norske Videnskabers Selskab, 1956.
- [31] K. Grjotheim, H. Kvande, and K. Motzfeldt. Vapor–liquid equilibria in the system NaF–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>. In *Light Metals 1975*, Annual Meeting, pages 125–137, Warrendale, Pennsylvania, 1975. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [32] Kai Grjotheim and Halvor Kvande, editors. *Introduction to Aluminium Electrolysis. Understanding the Hall–Héroult Process*, pages 53–54. Aluminium-Verlag, Düsseldorf, second edition, 1993.
- [33] Juan Guzman, Kai Grjotheim, and Terje Østvold. The influence of different fluoride additions on the vapor pressure of molten cryolite. In *Light Metals 1986*, Annual Meeting, pages 425–429, Warrendale, Pennsylvania, 1986. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [34] J. P. Harrison, G. Lombardo, and P. P. Peressini. Specific heat of NaF. *J. Phys. Chem. Solids*, 29:557–559, 1968.
- [35] Birgit Jenssen Holm and Fredrik Grønvold. Enthalpies of fusion of the alkali cryolites determined by drop calorimetry. *Acta Chem. Scand. (1947-1973)*, 27(6):2043–2050, 1973.
- [36] Holm, Birgit Jenssen and Jan Lützow Holm. Differential thermal analysis equipment for the study of molten fluoride equilibria. *Thermochim. Acta*, 5:273–283, 1973.
- [37] Jan Lützow Holm and Birgit Jenssen Holm. Phase relations and thermodynamic properties in the ternary reciprocal system LiF–NaF–Na<sub>3</sub>AlF<sub>6</sub>–Li<sub>3</sub>AlF<sub>6</sub>. *Thermochim. Acta*, 6:375–398, 1973.
- [38] Jan Lützow Holm. Thermochemistry of molten cryolite mixtures I. Sodium fluoride–aluminium fluoride mixtures. *High Temp. Sci.*, 6(1):16–36, Mar. 1974.
- [39] Birgit Jenssen Holm and Jan Lützow Holm. The enthalpy of mixing of liquid NaF and NaMgF<sub>3</sub> from drop calorimetry. *Thermochim. Acta*, 10:393–397, 1974.
- [40] K. C. Hong and O. J. Kleppa. Enthalpy of fusion of AlF<sub>3</sub> and enthalpies of mixing of AlF<sub>3</sub> with Li<sub>3</sub>AlF<sub>6</sub> and Na<sub>3</sub>AlF<sub>6</sub>. *High Temp. Sci.*, 8:299–307, 1976.
- [41] K. C. Hong and O. J. Kleppa. Thermochemistry of the liquid mixtures of aluminium fluoride with alkali fluorides and with zinc fluoride. *J. Phys. Chem.*, 82(2):176–182, 1978.
- [42] Juan Guzman Jaimes. *Vaporization of Molten Fluorides. The Systems Sodium Fluoride–Aluminium Fluoride Cryolite–Alkali Fluoride and Cryolite–Alkaline Earth Fluoride*. Dr.ing. thesis, Norwegian Institute of Technology, Feb 1986.
- [43] G. J. Janz, R. P. T. Tomkins, C. B. Allen, J. R. Downey, Jr., G. L. Gardner, U. Krabs, and S. K. Singer. Molten salts: Volume 4, part 2, chlorides and mixtures. Electrical conductance, density, viscosity and surface tension data. *J. Phys. Chem. Ref. Data*, 4(4):871–1178, 1975.
- [44] G. J. Janz, R. P. T. Tomkins, C. B. Allen, J. R. Downey, Jr., and S. K. Singer. Molten salts: Volume 4, part 3, bromides and mixtures; iodides and mixtures. Electrical conductance, density, viscosity and surface tension data. *J. Phys. Chem. Ref. Data*, 6(2):409–596, 1977.
- [45] Grini Johansen, Helge, Åsmund Sterten, and Jomar Thonstad. The phase diagrams of the systems Na<sub>3</sub>AlF<sub>6</sub>–Fe<sub>0.947</sub>O and Na<sub>3</sub>AlF<sub>6</sub>–FeF<sub>2</sub> and related activities of FeF<sub>2</sub> from emf measurements. *Acta Chem. Scand.*, 43:417–420, 1989.
- [46] J. Joly. On the specific heats of minerals. *Proc. R. Soc. London*, 41:250–274, 1886.
- [47] Stein Julsrud. A thermodynamic model for the system NaF–AlF<sub>3</sub> based on a Temkin model with complex formation. Institute of Inorganic Chemistry, NTH, and Norsk Hydro Research Centre.
- [48] E. G. King. Low temperature heat capacities and entropies at 298.15°K of cryolite, anhydrous aluminium fluoride and sodium fluoride. *J. Am. Chem. Soc.*, 79:2056–2057, May 1957.
- [49] A. J. Kirkham and B. Yates. The low-temperature specific heats and related thermodynamic properties of sodium fluoride and caesium bromide. *J. Phys. C: Solid State Phys.*, 1(2):1162–1170, 1968.
- [50] J. Kleicker, O. Knacke, and D. Lücke. Thermodynamics of Na–Al-fluorides. In *Conference on User Applications of Alloy Phase Diagrams*, pages 177–184, Lake Buena Vista, Florida, USA, 1986. ASM’s Materials Week’86.
- [51] E. N. Kolosov, V. B. Shol’ts, and L. N. Sidorov. Mass-spectrometric study of the thermodynamic properties of the system sodium fluoride–aluminium fluoride. III. Investigation of molecular velocities

- with the aid of a selector. Enthalpies of dissociation of complex molecules. *Russ. J. Phys. Chem. (Transl. of Zh. Fiz. Khim.)*, 48(9):1303–1305, 1974.
- [52] Ulrich Kuxmann and Ulrich Tillessen. Dampfdruckmessungen in den Systemen NaF–AlF<sub>3</sub> und LiF–NaF–AlF<sub>3</sub>. *Z. Erzbergbau Metallhuettenwes.*, 20(4):147–194, Apr. 1967.
- [53] Halvor Kvande. Determination of the average molar mass of the vapour above molten sodium chloride by the transpiration method. *Acta Chem. Scand., Ser. A*, 33(6):407–412, 1979.
- [54] Halvor Kvande. *Thermodynamics of the System NaF–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>–Al Studied by Vapour Pressure Measurements*. Dr.ing. thesis, Norwegian Institute of Technology, 1979.
- [55] Halvor Kvande. Vaporization behaviour of the system AlF<sub>3</sub>(s)–Al(l). *High Temp. - High Pressures*, 14:245–252, 1982.
- [56] Halvor Kvande. Vapor-phase studies of NaF–AlF<sub>3</sub> melts. 2. The NaF-rich part (cryolite). *High Temp. - High Pressures*, 15:63–71, 1983.
- [57] Halvor Kvande. Vapor-phase studies of NaF–AlF<sub>3</sub> melts. 1. The AlF<sub>3</sub>-rich part. *High Temp. - High Pressures*, 15:51–62, 1983.
- [58] Halvor Kvande. Vapor-phase studies of NaF–AlF<sub>3</sub> melts. 3. Thermodynamic data. *High Temp. - High Pressures*, 15:677–683, 1983.
- [59] Halvor Kvande. Vapor-phase studies of NaF–AlF<sub>3</sub> melts. 4. Activity data. *High Temp. - High Pressures*, 15:685–691, 1983.
- [60] G. J. Landon and A. R. Ubbelohde. Melting and crystal structure of cryolite (3NaF, AlF<sub>3</sub>). *Proc. R. Soc. London, Ser. A*, 240A:160–172, 1957.
- [61] Steven S. Lee, Kuan-Shaur Lei, Paul Xu, and Jesse J. Brown, Jr. Determination of melting temperatures and Al<sub>2</sub>O<sub>3</sub> solubilities for Hall cell electrolyte compositions. In J. P. McGeer, editor, *Light Metals 1984*, number 113 in Annual Meeting, pages 841–855, Warrendale, Pennsylvania, 27<sup>th</sup> Feb.–1<sup>st</sup> Mar. 1984. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [62] P. L. Lin, A. D. Pelton, and M.-L. Saboungi. Computer analysis of phase diagrams and thermodynamic properties of cryolite based systems: Part II. The AlF<sub>3</sub>–CaF<sub>2</sub>–LiF, AlF<sub>3</sub>–CaF<sub>2</sub>–NaF, and CaF<sub>2</sub>–LiF–NaF systems. *Metall. Trans. B*, 13B:61–69, Mar 1982.
- [63] V. S. Lyashenko. Heat capacity and heat of fusion for some chloride and fluoride containing melts. *Metallurg (Leningrad)*, 10(11):85–98, 1935.
- [64] G. Mesrobian, M. Rolin, and H. Pham. Étude sous pression des mélanges fluorure de sodium-fluorure d'aluminium riches en fluorure d'aluminium. *Rev. Int. Hautes Temp. Refract.*, 9(1):139–146, 1972.
- [65] J.-P. Millet, H. Pham, and M. Rolin. Les diagrammes de phase des systèmes binaires NaCl–AlF<sub>3</sub> et CaF<sub>2</sub>–AlF<sub>3</sub>: Extrapolation du point triple du fluorure d'aluminium. *Rev. Int. Hautes Temp. Refract.*, 2(4):277–284, 1974.
- [66] G. D. Milova and A. I. Gusev. A method of calculating the solubility in cryolite. *Raspilavy*, pages 398–401, 1991. English translation of *Raspilavy*, No. 5, pp. 79–83 (1991).
- [67] J. L. Murray. The Al–Na (aluminium–sodium) system. *Bull. Alloy Phase Diagrams*, 4(4):407–410, 1983.
- [68] C. J. O'Brien and K. K. Kelley. High temperature heat contents of cryolite, anhydrous aluminium fluoride and sodium fluoride. *J. Am. Chem. Soc.*, 79(21):5616–5618, 1957.
- [69] R. Ødegård, Å. Sterten, and J. Thonstad. The solubility of aluminium in cryolite melts. In R. D. Zabreznik, editor, *Light Metals 1987*, number 116 in Annual Meeting, pages 389–398, Warrendale, Pennsylvania, 24<sup>th</sup>–26<sup>th</sup> Feb. 1987. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [70] T. G. Pearson and J. Waddington. Electrode reactions in the aluminium reduction cell. *Discuss. Faraday Soc.*, 1:307–320, 1947.
- [71] A. D. Pelton, P. K. Talley, and R. A. Sharma. Thermodynamic evaluation of phase equilibria in the CaCl<sub>2</sub>–MgCl<sub>2</sub>–CaF<sub>2</sub>–MgF<sub>2</sub> system. *J. Phase Equilib.*, 13(4):384–390, 1992.
- [72] Ramon Fernandez Perez. *Physico-Chemical Properties of Cryolite Based Melts. Effect of Additives on the Temperature of Primary Crystallization, Density and Surface Tension*. Dr.ing. thesis, Norwegian Institute of Technology, Oct 1986.
- [73] Ray D. Peterson and Alton T. Tabereaux. Liquidus curves for the cryolite–AlF<sub>3</sub>–CaF<sub>2</sub>–Al<sub>2</sub>O<sub>3</sub> system in aluminium cell electrolytes. In R. D. Zabreznik, editor, *Light Metals 1987*, number 116 in Annual Meeting, pages 383–388, Warrendale, Pennsylvania, 24<sup>th</sup>–26<sup>th</sup> Feb. 1987. TMS Light Metals Committee, The Metallurgical Society of AIME.
- [74] N. W. F. Phillips, R. H. Singleton, and E. A. Hollingshead. Liquidus curves for aluminium cell electrolyte II. Ternary systems of cryolite–alumina with sodium fluoride, sodium chloride, and aluminium fluoride. *J. Electrochem. Soc.*, 102(12):690–692, Dec. 1955.
- [75] N. W. F. Phillips, R. H. Singleton, and E. A. Hollingshead. Liquidus curves for aluminium cell electrolyte I. Cryolite–alumina. *J. Electrochem. Soc.*, 102(11):648–649, Nov. 1955.

- [76] D. V. Pruttskov, P. I. Titaev, A. A. Andriiko, V. M. Mozhaev, and P. V. Polyakov. Solubility of an aluminium-silicon alloy in molten cryolite. *Rasplavy*, pages 389–393, 1991. English translation of *Rasplavy*, No. 5, pp. 69–73 (1991).
- [77] Zhuxian Qiu, Jinseng Zhang, Kai Grjotheim, and Halvor Kvande. Phase equilibrium studies on the cryolite–aluminium fluoride system. In Elwin L. Rooy, editor, *Light Metals 1991*, number 120 in Annual Meeting, pages 315–320, Warrendale, Pennsylvania, 17<sup>th</sup>–21<sup>st</sup> Feb. 1991. TMS Light Metals Committee, The Minerals, Metals & Materials Society.
- [78] C. W. Ransley and H. Neufeld. The solubility relationship in the aluminium–sodium and aluminium–silicon–sodium systems. *J. Inst. Met.*, 78:25–46, 1950.
- [79] E. Robert, J. E. Olsen, V. Danek, E. Tixhon, T. Østvold, and B. Gilbert. Structure and thermodynamics of alkali fluoride–aluminium fluoride–alumina melts. Vapor pressure, solubility, and Raman spectroscopic studies. *J. Phys. Chem. B*, 101(46):9447–9457, 1997.
- [80] Maurice Rolin. Sur la structure ionique de la cryolithe pure fundue. II—Détermination expérimentale du diagramme binaire cryolithe–AlF<sub>3</sub>. *Bull. Soc. Chim. Fr.*, 27(4):671–677, 1960.
- [81] Maurice Rolin and Marc Bernard. Sur la structure ionique de la cryolithe pure fundue. VI—Reconnaissance du schéma de dissociation des ions AlF<sub>6</sub><sup>3-</sup> et calcul de la constante d'équilibre. *Bull. Soc. Chim. Fr.*, 29(3):429–433, 1962.
- [82] Maurice Rolin and Marcel Rey. Étude thermodynamique des solutions dans la cryolithe fondue. I—Le système cryolithe-fluorure de sodium. *Bull. Soc. Chim. Fr.*, 33(9):2785–2790, 1966.
- [83] A. Røstum, A. Solheim, and A. Sterten. Phase diagram data in the system Na<sub>3</sub>AlF<sub>6</sub>–Li<sub>3</sub>AlF<sub>6</sub>–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>. Part I: Liquidus temperatures for primary cryolite crystallization. In Christian M. Bickert, editor, *Light Metals 1990*, number 119 in Annual Meeting, pages 311–316, Warrendale, Pennsylvania, 18<sup>th</sup>–22<sup>nd</sup> Feb. 1990. TMS Light Metals Committee, The Minerals, Metals & Materials Society.
- [84] M. L. Saboungi, P. L. Lin, P. Cerisier, and A. D. Pelton. Computer analysis of phase diagrams and thermodynamic properties of cryolite based systems: I. The AlF<sub>3</sub>–LiF–NaF system. *Metall. Trans. B*, 11B:493–501, Sep. 1980.
- [85] L. N. Sidorov and E. N. Kolosov. Mass-spectrometric study of the thermodynamic properties of the system sodium fluoride–aluminium fluoride. I. Calculations of the enthalpy and free energy of formation of sodium tetrafluoroaluminate (NaAlF<sub>4</sub>) in the gas phase. *Russ. J. Phys. Chem. (Transl. of Zh. Fiz. Khim.)*, 42(10):1382–1384, 1968.
- [86] L. N. Sidorov, E. N. Kolosov, and V. B. Shol'ts. Mass-spectrometric study of the thermodynamic properties of the system sodium fluoride–aluminium fluoride. II. Thermodynamic properties of the condensed phase in the system sodium fluoride–aluminium fluoride. *Russ. J. Phys. Chem. (Transl. of Zh. Fiz. Khim.)*, 42(10):1384–1386, 1968.
- [87] E. Skybakmoen, A. Solheim, and A. Sterten. Phase diagram data in the system Na<sub>3</sub>AlF<sub>6</sub>–Li<sub>3</sub>AlF<sub>6</sub>–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub>. Part II: Alumina solubility. In Christian M. Bickert, editor, *Light Metals 1990*, number 119 in Annual Meeting, pages 317–323, Warrendale, Pennsylvania, 18<sup>th</sup>–22<sup>nd</sup> Feb. 1990. TMS Light Metals Committee, The Minerals, Metals & Materials Society.
- [88] A. Solheim, S. Rolseth, E. Skybakmoen, and L. Støen. Liquidus temperatures and alumina solubility in the system Na<sub>3</sub>AlF<sub>6</sub>–AlF<sub>3</sub>–LiF–CaF<sub>2</sub>–MgF<sub>2</sub>. In James W. Evans, editor, *Light Metals 1995*, number 124 in Annual Meeting, pages 451–460, Warrendale, Pennsylvania, 12<sup>th</sup>–16<sup>th</sup> Feb. 1995. TMS Light Metals Committee, The Minerals, Metals & Materials Society.
- [89] Åsmund Sterten. Structural entities in NaF–AlF<sub>3</sub> melts containing alumina. *Electrochim. Acta*, 25:1673–1677, 1980.
- [90] Åsmund Sterten, Kjell Hamberg, and Inge Mæland. Activities and phase diagram data of NaF–AlF<sub>3</sub>–Al<sub>2</sub>O<sub>3</sub> mixtures derived from electromotive force and cryoscopic measurements. Standard thermodynamic data of β-Al<sub>2</sub>O<sub>3</sub>(s), Na<sub>3</sub>AlF<sub>6</sub>(s), Na<sub>5</sub>Al<sub>3</sub>F<sub>14</sub>(s) and NaAlF<sub>4</sub>(l). *Acta Chem. Scand., Ser. A*, 36:329–344, 1982.
- [91] Åsmund Sterten and Inge Mæland. Thermodynamics of molten mixtures of Na<sub>3</sub>AlF<sub>6</sub>–Al<sub>2</sub>O<sub>3</sub> and NaF–AlF<sub>3</sub>. *Acta Chem. Scand., Ser. A*, 39:241–257, 1985.
- [92] Åsmund Sterten and O. Skar. Some binary Na<sub>3</sub>AlF<sub>6</sub>–M<sub>x</sub>O<sub>y</sub> phase diagrams. *Aluminium (Isernhagen, Ger.)*, 64(10):1051–1054, 1988.
- [93] J. M. Stuve and M. J. Ferrante. Low-temperature heat capacities and high-temperature enthalpies of chiolite (Na<sub>5</sub>Al<sub>3</sub>F<sub>14</sub>). Technical Report 8442, U.S. Bureau of Mines, 1980. Bureau of Mines Report of Investigations.
- [94] A. Vajna and R. Bacchiega. Tensione divapore dei bagni criolitici struttura dei bagni criolitici allo stato solido liquido de vapore. *Metall. Ital.*, 8:481–494, 1960.
- [95] Haiming Xiao, Jomar Thonstad, and Sverre Rolseth. A study on the solubility of SnO<sub>2</sub> in cryolite-based melts. Technical Report STF34/F93043, SINTEF Metallurgy, Trondheim, Norway, Feb 1993.
- [96] K. Yoshida and Ernest W. Dewing. Activities in NaF–AlF<sub>3</sub> melts saturated with Al<sub>2</sub>O<sub>3</sub>. *Metall. Trans.*, 3:683–686, Mar. 1972.

- [97] He-Hua Zhou, Oddvin Herstad, and Terje Østvold. Vapour pressure studies of and complex formation in NaF–AlF<sub>3</sub> and Na<sub>3</sub>AlF<sub>6</sub>–MgF<sub>2</sub> melts. In Euel Ray Cutshall, editor, *Light Metals 1992*, number 121 in Annual Meeting, pages 511–520, Warrendale, Pennsylvania, 1<sup>st</sup>–5<sup>th</sup> Mar. 1992. TMS Light Metals Committee, The Minerals, Metals & Materials Society.

### Electrolytes (DatEle)

Experimental data for aqueous salt solutions and electrolytes in general.

**Related keywords: (experimental | data) & electrolytes | aqueous & solutions.**

#### Bibliography.

- [1] A. G. Bergman and L. V. Oprelenskova. Solubility polytherms of the calcium nitrate–potassium nitrate–water and calcium nitrate–potassium chloride–water ternary systems. *Russ. J. Inorg. Chem. (Transl. of Zh. Neorg. Khim.)*, 14(8):1144–1146, 1969.
- [2] V. P. Bykov, L. S. Skum, T. V. Kozenkova, L. D. Polevaya, and L. V. Konvisar. Physicochemical properties of a solution of ammonium and calcium nitrates. *Khim. Prom-st. (Moscow)*, 23(12):730–732, 1991.
- [3] Dale D. Ensor and Henry L. Anderson. Heats of dilution of NaCl: Temperature dependence. *J. Chem. Eng. Data*, 18(2):205–212, 1973.
- [4] Warren W. Ewing. Calcium nitrate. II. The vapor pressure–temperature relations of the binary system calcium nitrate–water. *J. Am. Chem. Soc.*, 49:1963–1973, Aug. 1927.
- [5] Warren W. Ewing, Alfred N. Rogers, John Z. Miller, and Edward McGovern. Calcium nitrate. III. Heats of hydration and of solution of the binary system calcium nitrate–water. *J. Am. Chem. Soc.*, 54:1335–1343, Apr. 1932.
- [6] Warren W. Ewing and Alfred N. Rogers. Calcium nitrate. IV. Heats of dilution of solutions calcium nitrate in water. *J. Am. Chem. Soc.*, 55:3603–3609, Sep. 1933.
- [7] Rainer Feistel. Equilibrium thermodynamics of seawater revisited. *Progress in Oceanography*, 31(2):101–179, 1993.
- [8] R. Flatt and P. Bocherens. Sur le système ternaire Ca<sup>2+</sup>–K<sup>+</sup>–NO<sub>3</sub><sup>–</sup>–H<sub>2</sub>O. *Helv. Chim. Acta*, 45(1):187–195, 1962.
- [9] R. Flatt and F. Benguerel. Sur l'équilibre liquide-vapeur à 25°C de systèmes ternaires composés d'un nitrate, d'acide nitrique et d'eau. *Helv. Chim. Acta*, 45(6):1772–1776, 1962.
- [10] Jeffrey A. Gates, David M. Tillett, Dorothy E. White, and Robert H. Wood. Apparent molar heat capacities of aqueous NaCl solutions from 0.05 to 3.0 mol kg<sup>–1</sup>, 350 to 600 K, and 2 to 18 MPa. *J. Chem. Thermodyn.*, 19:131–146, 1987.
- [11] Gibbard, Jr., H. Frank, George Scatchard, Raymond A. Rousseau, and Jefferson L. Creek. Liquid–vapor equilibrium of aqueous sodium chloride, from 298 to 373 K and from 1 to 6 mol kg<sup>–1</sup>, and related properties. *J. Chem. Eng. Data*, 19(3):281–288, 1974.
- [12] Robert D. Howell, Krishnam Raju, and Gordon Atkinson. Thermodynamics of “scale” mineral solubilities. 4. Experimental measurement of SrSO<sub>4</sub>(s) in H<sub>2</sub>O and aqueous NaCl from 25 to 250°C and from 1 to 500 bar. *J. Chem. Eng. Data*, 37(4):464–469, 1992.
- [13] Ito. 20°C (NH<sub>4</sub>SO<sub>3</sub>)<sub>2</sub>NH–NH<sub>4</sub>SO<sub>3</sub>NH<sub>2</sub>–(NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>–H<sub>2</sub>O. *Kogyo Kagaku Zasshi*, 63:1909–1912, 1960.
- [14] Ito. 25°C (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>–NH<sub>4</sub>NO<sub>3</sub>–NH<sub>4</sub>SO<sub>3</sub>NH<sub>2</sub>–H<sub>2</sub>O. *Kogyo Kagaku Zasshi*, 63:1913–1916, 1960.
- [15] V. I. Kosterina, A. A. Ponzivskii, E. A. Konstantinova, V. T. Orlova, and I. N. Lepeshkov. The KNO<sub>3</sub>–Ca(NO<sub>3</sub>)<sub>2</sub>–H<sub>2</sub>O systems at 55°C. *Russ. J. Inorg. Chem. (Transl. of Zh. Neorg. Khim.)*, 30(1):129–131, 1985.
- [16] Josette Lamberger and René Paris. É du système eau-nitrate de calcium-nitrate d'ammonium. *Bull. Soc. Chim. Fr.*, 17(5–6):546–552, Mar. 1950.
- [17] E. Lange and H. Streeck. Verdünnungswärmen einiger 2–1-wertiger bzw. 1–2-wertiger Salze in grosser Verdünnung bei 25°C. II. Mg(NO<sub>3</sub>)<sub>2</sub>, Ca(NO<sub>3</sub>)<sub>2</sub>, Sr(NO<sub>3</sub>)<sub>2</sub>, Ba(NO<sub>3</sub>)<sub>2</sub>, Li<sub>2</sub>SO<sub>4</sub>, Na<sub>2</sub>SO<sub>4</sub>, K<sub>2</sub>SO<sub>4</sub>, Rb<sub>2</sub>SO<sub>4</sub>, Cs<sub>2</sub>SO<sub>4</sub>, einschliesslich einiger neuerer Messungen am MgSO<sub>4</sub> und CaSO<sub>4</sub>. *Z. Phys. Chem., Abt. A*, 157:1–31, 1931.
- [18] E. E. Messikomer and R. H. Wood. The enthalpy of dilution of aqueous sodium chloride at 298.15 to 373.15 K, measured with a flow calorimeter. *J. Chem. Thermodyn.*, 7:119–130, 1975.
- [19] Frank J. Millero and Wing H. Leung. The thermodynamics of seawater at one atmosphere. *Am. J. Sci.*, 276:1035–1077, Nov. 1976.
- [20] Frank J. Millero. The thermodynamics of seawater. Part II. Thermochemical properties. *Ocean Sci. Eng.*, 8(1):1–40, 1983.
- [21] Arne Ölander and Halvard Liander. The phase diagram of sodium chloride and steam above the critical point. *Acta Chem. Scand. (1947-1973)*, 4(9):1437–1445, 1950.

- [22] Roberto T. Pabalan and Kenneth S. Pitzer. Apparent molar heat capacity and other thermodynamic properties of aqueous KCl solutions to high temperatures and pressures. *J. Chem. Eng. Data*, 33(3):354–362, 1988.
- [23] J. N. Pearce and L. E. Blackman. The vapor pressures and activity coefficients of aqueous solutions of calcium and aluminium nitrates at 25°. *J. Am. Chem. Soc.*, 57:24–27, Jan. 1935.
- [24] Kenneth S. Pitzer and Yi gui Li. Thermodynamics of aqueous sodium chloride to 823 K and 1 kilobar (100 MPa). *Proc. Natl. Acad. Sci. U. S. A.*, 80:7689–7693, 1983.
- [25] Kenneth S. Pitzer. Thermodynamics of sodium chloride solutions in steam. *J. Phys. Chem.*, 87(7):1120–1125, 1983.
- [26] Kenneth S. Pitzer and Yi gui Li. Critical phenomena and thermodynamics of dilute aqueous sodium chloride to 823 K. *Proc. Natl. Acad. Sci. U. S. A.*, 81:1268–1271, 1984.
- [27] M. H. R. J. Plusjé. Theory and practice in the treatment of phosphate rock with nitric acid. In *Nitrophosphates*, number 13 in Proceedings (Technical papers), pages 3–31. International Fertilizer Society, 1951.
- [28] M. E. Pozin, B. A. Kopylev, and Li shen Van. Investigation of metastable equilibria in the system  $\text{Ca}(\text{NO}_3)_2 - \text{H}_2\text{O}$ . *J. Appl. Chem. (Leningrad)*, 33:2642–2649, 1960. Translated from Zh. Prikl. Khim. (Leningrad), Vol. 33, No. 12, pp. 2675–2684, December, 1960.
- [29] R. A. Robinson. The activity coefficient of calcium nitrate in aqueous solution at 25°C from isopiestic vapor pressure measurements. *J. Am. Chem. Soc.*, 62:3130–3131, Nov. 1940.
- [30] I. M. Rodnyanskii, V. I. Korobkov, and I. S. Galinker. Specific volumes of aqueous electrolyte solutions at high temperatures. *Russ. J. Inorg. Chem. (Transl. of Zh. Neorg. Khim.)*, 36(10):1192–1194, Oct. 1962.
- [31] G. Scatchard. Thermal expansion and the Debye–hückel heat of dilution. *J. Am. Chem. Soc.*, 53(6):2037–2039, 1931.
- [32] A. Sieverts and W. Petzold. Binäre Systeme: Nitrate von Metallen der Zweiten Gruppe des Periodischen Systems und Wasser. III  $\text{Ca}(\text{NO}_3)_2 - \text{H}_2\text{O}$ ,  $\text{Sr}(\text{NO}_3)_2 - \text{H}_2\text{O}$  und  $\text{Ba}(\text{NO}_3)_2 - \text{H}_2\text{O}$ . *Z. Anorg. Allg. Chem.*, 212:233–241, 1933.
- [33] S. Sourirajan and G. C. Kennedy. The system  $\text{H}_2\text{O} - \text{NaCl}$  at elevated temperatures and pressures. *Am. J. Sci.*, 260:115–141, 1962.
- [34] Hugh Stott Taylor and William Noland Henderson. The solubility curves of salt hydrates: Calcium nitrate. *J. Am. Chem. Soc.*, 37:1688–1694, 1915.
- [35] P. Wang and A. Anderko. Computation of dielectric constants of solvent mixtures and electrolyte solutions. *Fluid Phase Equilib.*, 186:103–122, 2001.

### Gases and Fluids (DatGas)

Experimental  $p, V, T$  and thermal data for gases and dense fluids.

**Related keywords: (experimental | data) & (gases | dense & fluids) | pvt | enthalpy | heat & capacity.**

#### Bibliography.

- [1] Graydon K. Anderson. Solubility of carbon dioxide in water under incipient clathrate formation conditions. *J. Chem. Eng. Data*, 47:219–222, 2002.
- [2] G. J. Ashton and G. G. Haselden. Measurements of enthalpy and phase equilibrium for simulated natural gas mixtures and correlation of the results by a modified starling equation. *Cryogenics*, pages 41–47, 1980.
- [3] Edward P. Bartlett. The concentration of water vapor in compressed hydrogen, nitrogen and a mixture of these gases in the presence of condensed water. *J. Am. Chem. Soc.*, 27:65–78, jan 1927.
- [4] V. I. Baranenko and V. S. Kirov. Solubility of hydrogen in water in a broad temperature and pressure range. *Atomnaya Énergiya*, 66(1):24–28, jan 1989. Translated from Russian by Plenum Publishing Corporation.
- [5] W. P. Baskett and G. P. Matthews. An experiment to measure gas imperfection. *J. Chem. Educ.*, 62(4):353–355, 1985.
- [6] Rubin Battino and H. Lawrence Clever. The solubility of gases in liquids. *Chem. Rev. (Washington, D. C.)*, 66:395–463, 1966.
- [7] Manson Benedict. Pressure, volume, temperature properties of nitrogen at high density. I. Results obtained with a weight piezometer. *J. Am. Chem. Soc.*, 59(11):2224–2242, Nov. 1937.
- [8] K. Bier, J. Kunze, and G. Maurer. Non-ideal mixing behaviour of  $(0.5\text{C}_2\text{H}_6 + 0.5\text{C}_3\text{H}_8)$ . II. Excess properties and their representation by equations of state. *J. Chem. Thermodyn.*, 12:151–164, 1980.
- [9] T. J. Bruno, G. L. Hume, and J. F. Ely. Hydrogen component fugacities in binary mixtures with methane and propane. *Int. J. Thermophys.*, 7(5):1033–1051, 1986.

- [10] T. J. Bruno and G. L. Hume. Hydrogen component fugacities in binary mixtures with carbon dioxide. *Int. J. Thermophys.*, 7(5):1053–1063, 1986.
- [11] T. J. Bruno. Hydrogen component fugacities in binary mixtures with carbon dioxide: Pressure dependence. *Int. J. Thermophys.*, 8(2):205–216, 1987.
- [12] H. B. Brugge, C.-A. Hwang, W. J. Rogers, J. C. Holste, and K. R. Hall. Experimental cross virial coefficients for binary mixtures of carbon dioxide with nitrogen, methane and ethane at 300 K and 320 K. *Physica A (Amsterdam)*, 156:382–416, 1989.
- [13] John J. Carroll, John D. Slupsky, and Alan E. Mather. The solubility of carbon dioxide in water at low pressure. *J. Phys. Chem. Ref. Data*, 20(6):1201–1209, 1991.
- [14] John J. Carroll and Alan E. Mather. The system carbon dioxide–water and the Krichevsky–Kasarnovsky equation. *J. Solution Chem.*, 21(7):607–621, 1992.
- [15] A. Chapoy, A. H. Mohammadi, A. Chareton, B. Tohidi, and D. Richon. Measurement and modeling of gas solubility and literature review of the properties for the carbon dioxide–water system. *Ind. Eng. Chem. Res.*, 43:1794–1802, 2004.
- [16] Rosa Crovetto. Evaluation of solution data of the system CO<sub>2</sub>–H<sub>2</sub>O from 273 K to the critical point of water. *J. Phys. Chem. Ref. Data*, 20(3):575–589, 1991.
- [17] E. M. Dantzler, C. M. Knobler, and M. L. Windsor. Interaction virial coefficients in hydrocarbon mixtures. *J. Phys. Chem.*, 72(2):676–684, feb 1968.
- [18] Walter Dannhauser. Determination of compressibility factor. *J. Chem. Educ.*, 49(8):563–564, 1972.
- [19] R. D. Deshmukh and A. E. Mather. A mathematical model for equilibrium solubility of hydrogen sulfide and carbon dioxide in aqueous alkanolamine solutions. *Chem. Eng. Sci.*, 36:355–362, 1980.
- [20] Larry W. Diamond and Nikolay N. Akinfiyev. Solubility of CO<sub>2</sub> in water from –1.5 to 100°C and from 0.1 to 100 MPa: Evaluation of literature data and thermodynamic modelling. *Fluid Phase Equilib.*, 208:265–290, 2003.
- [21] Richard D. Dick. Shock wave compression of benzene, carbon disulfide, carbon tetrachloride, and liquid nitrogen. *J. Chem. Phys.*, 52:6021–6032, 1970.
- [22] H. A. Dirksen and C. H. Riesz. Equilibrium in the steam reforming of natural gas. *Ind. Eng. Chem.*, 45(7):1562–1565, Jul. 1953.
- [23] Donald R. Douslin, Roland H. Harrison, Richard T. Moore, and John P. McCullough.  $p-v-t$  relations for methane. *J. Chem. Eng. Data*, 9(3):358–363, Jul 1964.
- [24] J. H. Dymond, K. N. Marsh, R. C. Wilhoit, and K. C. Wong. The virial coefficients of pure gases and mixtures. In *Group IV: Physical Chemistry*, volume 21A. Landolt–Börnstein, 2002.
- [25] J. H. Dymond, K. N. Marsh, R. C. Wilhoit, and K. C. Wong. The virial coefficients of pure gases and mixtures. In *Group IV: Physical Chemistry*, volume 21B. Landolt–Börnstein, 2002.
- [26] Tim Eggeman and Steve Chafin. Pitfalls of CO<sub>2</sub> freezing prediction. In *Gas Processors Association 82nd Annual Convention, San Antonio, Texas*, pages 1–19, mar 2003.
- [27] P. T. Eubank, J. J. Scheloske, and L. L. Joffrion. Dew points of wet gases: New data and a new correlation using interaction third virial coefficients. *Fluid Phase Equilib.*, 30:255–264, 1986.
- [28] Katherine A. Evelein, R. Gordon Moore, and Robert A. Heidemann. Correlation of the phase behavior in the systems hydrogen sulphide–water and carbon dioxide–water. *Ind. Eng. Chem. Process Des. Dev.*, 15(3):423–428, 1976.
- [29] E. R. Grilly and R. L. Mills. Volume change on melting of N<sub>2</sub> up to 3500 kg/cm<sup>2</sup>. *Phys. Rev.*, 105(4):1140–1145, Feb. 1957.
- [30] R. D. Gunn, P. L. Chueh, and J. M. Prausnitz. Inversion temperatures and pressures for cryogenic gases and their mixtures. *Cryogenics*, 6:324–329, Dec 1966.
- [31] Raja F. Hajjar, Webster B. Kay, and Glen F. Leverett. Second virial coefficients of eight compounds in the range 40° to 200°C. *J. Chem. Eng. Data*, 14(3):377–380, jul 1969.
- [32] Robert A. Heidemann and John M. Prausnitz. Equilibrium data for wet-air oxidation. Water content and thermodynamic properties of saturated combustion gases. *Ind. Eng. Chem. Process Des. Dev.*, 16(3):375–381, 1977.
- [33] Matthew P. Hodges, Richard J. Wheatley, and Gregory K. Schenter. Intermolecular potential and second virial coefficient of the water–hydrogen complex. *J. Chem. Phys.*, 120(2):710–720, jan 2004.
- [34] Von L. Holborn and M. Jakob. Die spezifische Wärme  $c_p$  der Luft bei 60°C und 1 bis 200 at.. *Z. Ver. Dtsch. Ing.*, 40(58):1429–1436, 1914.
- [35] L. G. Hoxton. The Joule–Thomson effect for air at moderate temperature and pressures. *Phys. Rev.*, 13:438–479, 1919.
- [36] Richard T. Jacobsen and Richard B. Steward. Thermodynamic properties of nitrogen including liquid and vapor phases from 63 K to 2000 K with pressures to 10,000 bar. *J. Phys. Chem. Ref. Data*, 2(4):757–922, 1973.
- [37] K. Kerl and H. Häusler. Mean polarizabilities and second virial coefficients of the gases Ar, CH<sub>4</sub>, C<sub>2</sub>H<sub>6</sub>, C<sub>3</sub>H<sub>8</sub> and C(CH<sub>3</sub>)<sub>4</sub>. *Ber. Bunsen-Ges.*, 88:992–997, 1984.

- [38] R. Kleinrahm, W. Duschek, and W. Wagner. (Pressure, density, temperature) measurements in the critical region of methane. *J. Chem. Thermodyn.*, 18:1103–1114, 1986.
- [39] N. M. Lancaster and C. J. Wormald. The excess molar enthalpies of  $(0.5\text{H}_2\text{O} + 0.5\text{C}_2\text{H}_4)(g)$  and  $(0.5\text{H}_2\text{O} + 0.5\text{C}_2\text{H}_6)(g)$  at high temperatures and pressures. *J. Chem. Thermodyn.*, 19:89–97, 1987.
- [40] Eric W. Lemmon, Richard T. Jacobsen, Steven G. Penoncello, and Daniel G. Fried. Thermodynamic properties of air and mixtures of nitrogen, argon, and oxygen from 60 to 2000 K at pressures to 2000 MPa. *J. Phys. Chem. Ref. Data*, 29(3):331–362, 2000.
- [41] Taras Y. Makogon and Dendy E. Sloan, Jr. Phase equilibrium for methane hydrate from 190 to 262 K. *J. Chem. Eng. Data*, 39(2):351–353, 1994.
- [42] Nikos J. P. Maris and Leonard I. Stiel. Interaction second virial coefficients of polar mixtures from parameters of the components. *Ind. Eng. Chem. Process Des. Dev.*, 24(1):183–187, 1985.
- [43] A. Michels, Hub. Wijker, and Hk Wijker. Isotherms of argon between  $0^\circ\text{C}$  and  $150^\circ\text{C}$  and pressures up to 2900 atmospheres. *Physica (Amsterdam)*, 15(7):627–633, Aug. 1949.
- [44] A. Michels, J. M. Levelt, and W. de Graaff. Compressibility isotherms of argon at temperatures between  $-25^\circ\text{C}$  and  $-155^\circ\text{C}$ , and densities up to 640 amagat (pressures up to 1050 atmospheres). *Physica (Amsterdam)*, 24:659–671, 1958.
- [45] Donald G. Miller. Joule–Thomson inversion curve, corresponding states, and simpler equations of state. *Ind. Eng. Chem. Fundam.*, 9(4):585–589, 1970.
- [46] Charles A. Passut and Ronald P. Danner. Correlation of ideal gas enthalpy, heat capacity, and entropy. *Ind. Eng. Chem. Process Des. Dev.*, 11(4):543–546, 1972.
- [47] Mukund R. Patel, L. Lane Joffrion, and Philip T. Eubank. A simple procedure for estimating virial coefficients from Burnett *pvt* data. *AIChE J.*, 34(7):1229–1232, 1988.
- [48] Andrey V. Plyasunov and Everett L. Shock. Second cross virial coefficients for interactions involving water. Critical data compilation. *J. Chem. Eng. Data*, 48:808–821, 2003.
- [49] Sandrine Portier and Christopher Rochelle. Modelling  $\text{CO}_2$  solubility in pure water and NaCl-type waters from 0 to  $300^\circ\text{C}$  and from 1 to 300 bar. Application to the Utsira formation at Sleipner. *Chemical Geology*, 217:187–189, 2005.
- [50] V. A. Rabinovich and V. G. Beketov. *Moist Gases: Thermodynamic Properties*. Begell House Inc., New York, 1995.
- [51] R. E. Randelman and L. A. Wenzel. Joule–Thomson coefficients of hydrogen and methane mixtures. *J. Chem. Eng. Data*, 33:293–299, 1988.
- [52] P. Richards and C. J. Wormald. The enthalpy of mixing of (water + argon) vapour. *Z. Phys. Chem., Neue Folge*, 128:35–42, 1981.
- [53] J. R. Roebuck and H. Osterberg. The Joule–Thomson effect in nitrogen. *Phys. Rev.*, 48:450–457, 1935.
- [54] Marvin Ross. The dissociation of dense liquid nitrogen. *J. Chem. Phys.*, 86(12):7110–7118, Jun. 1987.
- [55] Werner Schulze. Vapor–liquid equilibria and saturation densities in the ternary system helium–nitrogen–methane. *Fluid Phase Equilib.*, 134:213–224, 1997.
- [56] E. S. Sellers and D. R. Augood. The distillation characteristics of liquid hydrogen. *Trans. Inst. Chem. Eng.*, 34:53–88, 1956.
- [57] H. Silberberg, K. A. Kobe, and J. J. McKetta. Gas compressibilities with the Burnett apparatus. *J. Chem. Eng. Data*, 4(4):314–323, 1959.
- [58] G. Smith, A. Sellars, T. K. Yerlett, and C. J. Wormald. The excess enthalpy of (water + hydrogen) vapour and (water + methane) vapour. *J. Chem. Thermodyn.*, 15:29–35, 1983.
- [59] G. R. Smith, M. J. Fahy, and C. J. Wormald. The excess molar enthalpy of  $\{x\text{H}_2\text{O} + (1-x)\text{C}_n\text{H}_{2n+2}\}(g)$  for  $n = 5, 6, 7$  and  $8$ . *J. Chem. Thermodyn.*, 16:825–831, 1984.
- [60] G. R. Smith and C. J. Wormald. The excess molar enthalpies of  $\{x\text{H}_2\text{O} + (1-x)\text{CO}\}(g)$  and  $\{x\text{H}_2\text{O} + (1-x)\text{CO}_2\}(g)$ . *J. Chem. Thermodyn.*, 16:543–550, 1984.
- [61] R. Span and W. Wagner. A new equation of state for carbon dioxide covering the fluid region from the triple-point temperature to 1100 K at pressures up to 800 MPa. *J. Phys. Chem. Ref. Data*, 25(6):1509–1596, 1996.
- [62] Hugh M. Spencer and John L. Justice. Empirical heat capacity equations for simple gases. *J. Am. Chem. Soc.*, 56:2311–2312, nov 1934.
- [63] Hugh M. Spencer. Empirical heat capacity equations of various gases. *J. Am. Chem. Soc.*, 67:1859–1860, oct 1945.
- [64] N. G. Sretenskaja, Richard J. Sadus, and E. Ulrich Franck. High-pressure phase equilibria and critical curve of the water + helium system to 200 MPa and 723 K. *J. Phys. Chem.*, 99:4273–4277, 1995.
- [65] Paul B. Stewart and Prem Munjal. Solubility of carbon dioxide in pure water, synthetic sea water, and synthetic sea water concentrates at  $-5^\circ$  to  $25^\circ\text{C}$  and 10 to 45 atm pressure. *J. Chem. Eng. Data*, 15(1):67–71, 1970.

- [66] W. B. Streett and J. C. G. Calado. Liquid–vapor equilibrium for hydrogen + nitrogen at temperatures from 63 K to 110 K and pressures to 57 MPa. *J. Chem. Thermodyn.*, 10:1089–1100, 1978.
- [67] H. Teng, A. Yamasaki, M.-K. Chun, and H. Lee. Solubility of liquid CO<sub>2</sub> in water at temperatures from 278 K to 293 K and pressures from 6.44 MPa to 29.49 MPa and densities of the corresponding aqueous solutions. *J. Chem. Thermodyn.*, 29:1301–1310, 1997.
- [68] Constantine Tsonopoulos. Second virial coefficients of water pollutants. *AIChE J.*, 24(6):1112–1115, Nov. 1978.
- [69] G. A. Vliegthart and H. N. W. Lekkerkerker. Predicting the gas–liquid critical point from the second virial coefficient. *J. Chem. Phys.*, 112(12):5364–5369, 2000.
- [70] I. Wängberg, E. Ljungström, B. E. R. Olsson, and J. Davidsson. Temperature dependence of the reaction NO<sub>3</sub> + NO<sub>2</sub> → NO + NO<sub>2</sub> + O<sub>2</sub> in the range from 296 to 332 K. *J. Phys. Chem.*, 96:7640–7645, 1992.
- [71] Emmerich Wilhelm, Rubin Battino, and Robert J. Wilcock. Low-pressure solubility of gases in liquid water. *Chem. Rev. (Washington, D. C.)*, 77(2):219–262, 1977.
- [72] Richard A. Wilsak and George Thodos. Coexistence behavior of the vapor–liquid–solid equilibrium states for argon. *J. Chem. Eng. Data*, 29(3):255–262, 1984.
- [73] C. J. Wormald and B. Wurzbeger. Second virial cross coefficients for (ammonia + water) derived from gas phase excess enthalpy measurements. *J. Chem. Thermodyn.*, 33:1193–1210, 2001.
- [74] C. J. Wormald and B. Wurzbeger. Ammonia–carbon dioxide association. Second virial cross coefficients for (ammonia–carbon dioxide) derived from gas phase excess enthalpy measurements. *J. Chem. Thermodyn.*, 34:1567–1573, 2002.
- [75] C. J. Wormald. Water–hydrogen chloride association. Second virial cross coefficients for water–hydrogen chloride from gas phase excess enthalpy measurements. *J. Chem. Thermodyn.*, 35:417–431, 2003.
- [76] Christopher J. Wormald, Keith L. Lewis, and Stephen Mosedale. The excess enthalpies of hydrogen + methane, hydrogen + nitrogen, methane + nitrogen, methane + argon, and nitrogen + argon at 298 and 201 K at pressures up to 10.2 MPa. *J. Chem. Thermodyn.*, 9:27–42, 1977.
- [77] Christopher J. Wormald. A differential-flow mixing calorimeter. The excess enthalpy of methane + benzene, methane + cyclohexane, and benzene + cyclohexane. *J. Chem. Thermodyn.*, 9:901–910, 1977.
- [78] C. J. Wormald and C. N. Colling. Excess enthalpies of some mixtures containing steam. *Int. Conf. Prop. Steam*, 9th, 1979, pages 655–663, 1980.
- [79] C. J. Wormald and C. N. Colling. Excess enthalpies for (water + methane) vapour up to 698.2 K and 12.6 MPa. *AIChE J.*, 30(3):386–393, May 1984.
- [80] C-s. Zha, R. Boehler, D. A. Young, and M. Ross. The argon melting curve to very high pressures. *J. Chem. Phys.*, 85(2):1034–1036, Jul. 1986.
- [81] Wei Zhang and Jan A. Schouten. *pvt*–*x* behavior of He–N<sub>2</sub> mixtures from 270 to 353 K and up to 280 bar. *J. Chem. Eng. Data*, 37:114–119, 1992.

## Water and Heavy Water (DatH2O)

Experimental data for water and heavy water.

**Related keywords: (experimental | data) & (water | h2o | heavy | d2o).**

### Bibliography.

- [1] J. Ananthaswamy and Gordon Atkinson. Thermodynamics of concentrated electrolyte mixtures. 4. Pitzer–Debye–Hückel limiting slopes for water from 0 to 100°C and from 1 atm to 1 kbar. *J. Chem. Eng. Data*, 29(1):81–87, 1984.
- [2] Donald G. Archer and Peiming Wang. The dielectric constant of water and Debye–Hückel limiting law slopes. *J. Phys. Chem. Ref. Data*, 19(2):371–411, 1990.
- [3] Francis Birch, J. F. Schairer, and H. Cecil Spicer, editors. *Handbook of Physical Constants*, pages 204–212. Number 15 in Geological Society of America Special Papers. The Geological Society of America, New York, Jan 1942.
- [4] Diego P. Fernández, Y. Mulev, A. R. H. Goodwin, and J. M. H. Levelt Sengers. A database for the static dielectric constant of water and steam. *J. Phys. Chem. Ref. Data*, 24(1):33–69, 1995.
- [5] D. P. Fernández, A. R. H. Goodwin, E. W. Lemmon, J. M. H. Levelt Sengers, and R. C. Williams. A formulations for the static permittivity of water and steam at temperatures from 238 K to 873 K at pressures up to 1200 MPa, including derivatives and Debye–Hückel coefficients. *J. Phys. Chem. Ref. Data*, 26(4):1125–1166, 1997.

- [6] Wely Brasil Floriano and Marco Antonio Chaer Nascimento. Dielectric constant and density of water as a function of pressure at constant temperature. *Brazilian Journal of Physics*, 34(1):38–41, mar 2004.
- [7] E. U. Franck, S. Rosenzweig, and M. Christoforakos. Calculation of the dielectric-constant of water to 1000 degrees C and very high-pressures. *Ber. Bunsen-Ges.*, 94(2):199–203, feb 1990.
- [8] W. F. Giaouque and J. W. Stout. The entropy of water and the third law of thermodynamics. The heat capacity of ice from 15 to 273 K. *J. Am. Chem. Soc.*, 58:1144–1150, 1936.
- [9] Saul Goldman, Chris Joslin, and Evgeny A. Wasserman. Theoretical calculation of the static dielectric-constant of water at high-temperatures and pressures. *J. Phys. Chem.*, 98(24):6231–6233, jun 1994.
- [10] V. A. Gorodysky. A new semiempirical method of calculation of the static dielectric-constant of water. *Dokl. Akad. Nauk*, 325(2):311–313, 1992.
- [11] D. E. Hare and C. M. Sorensen. The density of supercooled water. II. Bulk samples cooled to the homogeneous nucleation limit. *J. Chem. Phys.*, 87(8):4840–4845, Oct 1987.
- [12] P. G. Hill and D. Chris MacMillan. Saturation states of heavy water. *J. Phys. Chem. Ref. Data*, 9(3):735–749, 1980.
- [13] Philip G. Hill. A unified fundamental equation for the thermodynamic properties of H<sub>2</sub>O. *J. Phys. Chem. Ref. Data*, 19(5):1233–1274, 1990.
- [14] W. Holzapfel and E. U. Franck. Leitfähigkeit und Ionendissoziation des Wassers bis 1000° und 100 kbar. *Ber. Bunsen-Ges.*, 70(9/10):1105–1112, 1966.
- [15] W. Alexander van Hook. Vapor pressures of the isotopic waters and ices. *J. Phys. Chem.*, 72(4):1234–1244, Apr 1968.
- [16] W. Alexander van Hook. Condensed phase isotope effects. *Isotopenpraxis*, 4(5):161–169, 1968.
- [17] John R. Huffman and H. C. Urey. Separation of oxygen isotopes by a fractionating column. *Ind. Eng. Chem.*, 29(5):531–535, May 1937.
- [18] Richard T. Jacobsen, Richard B. Steward, and Majid Jahangiri. Thermodynamic properties of nitrogen from the freezing line to 2000 K at pressures to 10000 MPa. *J. Phys. Chem. Ref. Data*, 15(2):735–908, 1986.
- [19] Gabor Jancso, Jovan Pupezin, and W. Alexander van Hook. The vapor pressure of ice between +10<sup>-2</sup> and -10<sup>+2</sup> °C. *J. Phys. Chem.*, 74(15):2984–2989, 1970.
- [20] George S. Kell. Effects of isotopic composition, temperature, pressure and dissolved gases on the density of liquid water. *J. Phys. Chem. Ref. Data*, 6(4):11091131, 1977.
- [21] J. Kestin, J. V. Sengers, B. Kamgar-Parsi, and J. M. H. Levelt Sengers. Thermophysical properties of fluid D<sub>2</sub>O. *J. Phys. Chem. Ref. Data*, 13(2):601–609, 1984.
- [22] J. Kestin, J. V. Sengers, B. Kamgar-Parsi, and J. M. H. Levelt Sengers. Thermophysical properties of fluid H<sub>2</sub>O. *J. Phys. Chem. Ref. Data*, 13(1):175–183, 1984.
- [23] J. Kestin and J. V. Sengers. New international formulations for the thermodynamic properties of light and heavy water. *J. Phys. Chem. Ref. Data*, 15(1):305–320, 1986.
- [24] M. L. S. Matos Lopes and C. A. Nieto de Castro. Liquid mutual diffusivities of the H<sub>2</sub>O / D<sub>2</sub>O system. *Int. J. Thermophys.*, 7(3):699–707, 1986.
- [25] William L. Marshall and E. U. Franck. Ion product of water substance, 0–1000° C, 1–10,000 bars. New international formulation and its background. *J. Phys. Chem. Ref. Data*, 10(2):295–304, 1981.
- [26] Yu. V. Mulev and S. N. Smirnov. Calculation of the dielectric permittivity of water up to 923 K and 1 GPA. *High Temp. (Transl. of Teplofiz. Vys. Temp.)*, 30(1):51–54, jan/feb 1992.
- [27] Tariq Muneer and Simon Michael Scott. Density and temperature explicit equations of state for steam. *Energy Convers. Manage.*, 31(4):315–325, 1991.
- [28] N. S. Osborne, H. F. Stimson, E. F. Fiock, and D. C. Ginnings. The pressure of saturated water vapor in the range 100° to 374°C. *Bur. Stand. J. Res.*, 10:155–188, 1933.
- [29] Nathan S. Osborne. Heat of fusion of ice. A revision. *J. Res. Natl. Bur. Stand. (U. S.)*, 23:643–646, Dec 1939.
- [30] Nathan S. Osborne, Harold F. Stimson, and Defoe C. Ginnings. Thermal properties of saturated water and steam. *J. Res. Natl. Bur. Stand. (U. S.)*, 23:261–270, Aug 1939.
- [31] Kenneth S. Pitzer. Dielectric constant of water at very high temperature and pressure. *Proc. Natl. Acad. Sci. U. S. A.*, 80:4575–4576, jul 1983.
- [32] H. Sato, M. Uematsu, K. Watanabe, A. Saul, and W. Wagner. New international skeleton tables for the thermodynamic properties of ordinary water substance. *J. Phys. Chem. Ref. Data*, 17(4):1439–1540, 1988.
- [33] H. Sato, K. Watanabe, J. M. H. Levelt Sengers, J. S. Gallagher, P. G. Hill, J. Straub, and W. Wagner. Sixteen thousand evaluated experimental thermodynamic property data for water and steam. *J. Phys. Chem. Ref. Data*, 20(5):1023–1044, 1991.
- [34] A. Saul and W. Wagner. International equations for the saturation properties of ordinary water substance. *J. Phys. Chem. Ref. Data*, 16(4):893–901, 1987.

- [35] A. Saul and W. Wagner. A fundamental equation for water covering the range from the melting line to 1273 K at pressures up to 25,000 MPa. *J. Phys. Chem. Ref. Data*, 18(4):1537–1564, 1989.
- [36] P. F. Scholander. Tensile water. *Am. Sci.*, 60:584–590, Sep/Oct 1972.
- [37] Ernst Schmidt. Properties of water and steam in SI-units. Second, revised and updated, printing. In Ulrich Grigull, editor, *Properties of Water and Steam in SI-Units*. Springer-Verlag, New York, 1979.
- [38] Levelt Sengers, J. M. H., J. Straub, K. Watanabe, and P. G. Hill. Assessment of critical parameter values for H<sub>2</sub>O and D<sub>2</sub>O. *J. Phys. Chem. Ref. Data*, 14(1):193–207, 1985.
- [39] R. Tufeu, P. Bury, Y. Garrabos, and B. Le Neindre. Comparison of the thermal conductivities of water and heavy water. *Int. J. Thermophys.*, 7(3):663–673, 1986.
- [40] M. Uematsu and E. U. Franck. Static dielectric constant of water and steam. *J. Phys. Chem. Ref. Data*, 9(4):1291–1306, 1980.
- [41] W. Wagner and A. Pruß. The IAPWS formulation 1995 for the thermodynamic properties of ordinary water substance for general and scientific use. *J. Phys. Chem. Ref. Data*, 31(2):387–535, 2002.
- [42] W. Wagner and A. Pruß. International equations for the saturation properties of ordinary water substance. Revised according to the international temperature scale of 1990. Addendum to *J. Phys. Chem. Ref. Data* 16, 893 (1987). *J. Phys. Chem. Ref. Data*, 22(3):783–787, 1993.
- [43] W. Wagner, A. Saul, and A. Pruß. International equations for the pressure along the melting and along the sublimation curve of ordinary water substance. *J. Phys. Chem. Ref. Data*, 23(3):515–525, 1994.
- [44] Evgeny Wasserman, Bernard Wood, and John Brodholt. The static dielectric constant of water at pressures up to 20 kbar and 1273 K: Experiment, simulations, and empirical equations. *Geochim. Cosmochim. Acta*, 59(1):1–6, 1995.
- [45] Ho-Ming Yeh and Shyh-Ching Yang. The enrichment of heavy water in a batch-type thermal diffusion column. *Chem. Eng. Sci.*, 39(7/8):1277–1282, 1984.

### Hydrocarbons (DatHyC)

Experimental data for alkanes, alkenes, alkynes, aromates, etc.

**Related keywords: (experimental | data) & hydrocarbons.**

#### Bibliography.

- [1] Julio C. Acosta, Edison Hevia, and Stuart Leipziger. Dew and bubble point measurements for carbon dioxide–propane mixtures. *J. Chem. Eng. Data*, 29:304–309, 1984.
- [2] Douglas Ambrose and Constantine Tsonopoulos. Vapor–liquid critical properties of elements and compounds. 2. Normal alkanes. *J. Chem. Eng. Data*, 40:531–546, 1995.
- [3] T. S. Brown, A. J. Kidnay, and E. D. Sloan. Vapor–liquid equilibria in the carbon dioxide–ethane system. *Fluid Phase Equilib.*, 40:169–184, 1988.
- [4] T. S. Brown, V. G. Niesen, E. D. Sloan, and A. J. Kidnay. Vapor–liquid equilibria for the binary systems of nitrogen, carbon dioxide, and *n*-butane at temperatures from 220 to 344 K. *Fluid Phase Equilib.*, 53:7–14, 1989.
- [5] E. Brunner. Fluid mixtures at high pressures. *J. Chem. Thermodyn.*, 22:335–351, 1990.
- [6] J. F. Connolly. Solubility of hydrocarbons in water near the critical solution temperatures. *J. Chem. Eng. Data*, 11(1):13–16, 1966.
- [7] O. L. Culberson, A. B. Horn, and J. J. McKetta, Jr. Phase equilibria in hydrocarbon–water systems—The solubility of ethane in water at pressures to 120 pounds per square inch. *Pet. Trans.*, 189:1–6, 1950.
- [8] O. L. Culberson and J. J. McKetta, Jr. Phase equilibria in hydrocarbon–water systems II—The solubility of ethane in water at pressures to 10,000 psi. *Pet. Trans.*, 189:319–322, 1950.
- [9] O. L. Culberson and J. J. McKetta, Jr. Phase equilibria in hydrocarbon–water systems III—The solubility of methane in water at pressures to 10,000 psia. *Pet. Trans.*, 192:223, 1951.
- [10] Harold G. Donnelly and Donald L. Katz. Phase equilibria in the carbon dioxide–methane system. *Ind. Eng. Chem.*, 46(3):511–517, Mar 1954.
- [11] Åge Fredenslund and Jørgen Møllerup. Measurement and prediction of equilibrium ratios for the C<sub>2</sub>H<sub>6</sub> + CO<sub>2</sub> system. *J. Chem. Soc., Faraday Trans. 1*, 70:1653–1660, 1974.
- [12] Luis A. Galicia. Phase equilibria and critical points in *n*-alkanes + CO<sub>2</sub> + N<sub>2</sub> systems. *Fluid Phase Equilib.*, 27:331–339, 1986.
- [13] Salah E. M. Hamam and Benjamin C.-Y. Lu. Isothermal vapor–liquid equilibria in binary system propane–carbon dioxide. *J. Chem. Eng. Data*, 21(2):200–204, 1976.
- [14] J. L. Heidman, C. Tsonopoulos, C. J. Brady, and G. M. Wilson. High-temperature mutual solubilities of hydrocarbons and water. Part II: Ethylbenzene, ethylcyclohexane, and *n*-octane. *AIChE J.*, 31(3):376–384, 1985.

- [15] Jack J.-C. Hsu, N. Nagarajan, and R. L. Robinson, Jr. Equilibrium phase compositions, phase densities, and interfacial tensions for CO<sub>2</sub> + hydrocarbon systems. 1. CO<sub>2</sub> + *n*-butane. *J. Chem. Eng. Data*, 30:485–491, 1985.
- [16] H. Inomata, K. Arai, and S. Saito. Vapor–liquid equilibria for CO<sub>2</sub> / hydrocarbon mixtures at elevated temperatures and pressures. *Fluid Phase Equilib.*, 36:107–119, 1987.
- [17] L. L. Larson, M. K. Silva, M. A. Taylor, and F. M. Orr, Jr. Temperature dependence of L<sub>1</sub>/L<sub>2</sub>/V behavior in CO<sub>2</sub> / hydrocarbon systems. *SPE Reservoir Eng.*, pages 105–114, Feb 1989.
- [18] D. Legret, D. Richon, and H. Renon. Critical evaluation of methane–hydrocarbon high-pressure experimental vapour–liquid equilibrium data using equations of state. *Fluid Phase Equilib.*, 17:323–350, 1984.
- [19] E. W. Lemmon and A. R. H. Goodwin. Critical properties and vapor pressure equation for alkanes C<sub>*n*</sub>H<sub>2*n*+2</sub>: Normal alkanes with *n* ≤ 36 and isomers for *n* = 4 through *n* = 9. *J. Phys. Chem. Ref. Data*, 29(1):1–39, 2000.
- [20] Ho-Mu Lin. Peng–Robinson equation of state for vapor–liquid equilibrium calculations for the carbon dioxide + hydrocarbon mixtures. *Fluid Phase Equilib.*, 16:151–169, 1984.
- [21] Feliciano M. Llave, Kraemer D. Luks, and James P. Kohn. Three-phase liquid–liquid–vapor equilibria in the methane + ethane + *n*-hexane and methane + ethane + *n*-heptane systems. *J. Chem. Eng. Data*, 31:418–421, 1986.
- [22] Feliciano M. Llave and Ting Horng Chung. Vapor–liquid equilibria of nitrogen–hydrocarbon systems at elevated pressures. *J. Chem. Eng. Data*, 33:123–128, 1988.
- [23] John J. McKetta and Donald L. Katz. Methane–*n*-butane–water system in two- and three-phase regions. *Ind. Eng. Chem.*, 40(5):853–863, 1948.
- [24] J. M. Moysan and H. Paradowski. Prediction of phase behaviour of gas-containing systems with cubic equations of state. *Chem. Eng. Sci.*, 41(8):2069–2074, 1986.
- [25] Kunio Nagahama, Hitoshi Konishi, Daisuke Hoshino, and Mitsuho Hirata. Binary vapor–liquid equilibria of carbon dioxide–light hydrocarbons at low temperature. *J. Chem. Eng. Jpn.*, 7(5):291–298, 1974.
- [26] Kazunari Ohgaki and Takashi Katayama. Isothermal vapor–liquid equilibrium data for the ethane–carbon dioxide system at high pressures. *Fluid Phase Equilib.*, 1:27–32, 1977.
- [27] R. H. Olds, H. H. Reamer, B. H. Sage, and W. N. Lacey. Phase equilibria in hydrocarbon systems: The *n*-butane–carbon dioxide system. *Ind. Eng. Chem.*, 41(3):475–482, 1949.
- [28] H. H. Reamer, B. H. Sage, and W. N. Lacey. Phase equilibria in hydrocarbon systems. volumetric and phase behavior of the propane–carbon dioxide system. *Ind. Eng. Chem.*, 43(11):2515–2520, 1951.
- [29] J. G. Roof. Three-phase critical point in hydrocarbon–water system. *J. Chem. Eng. Data*, 15(2):301–303, 1970.
- [30] G. F. Russell, R. Thompson, F. P. Vance, and R. L. Huntington. Experimental determinations of water vapour content of a natural gas up to 2000 pounds pressure. *Pet. Trans.*, 160:150–156, 1945.
- [31] Steven K. Shibata and Stanley I. Sandler. High-pressure vapor–liquid equilibria involving mixtures of nitrogen, carbon dioxide, and *n*-butane. *J. Chem. Eng. Data*, 34:291–298, 1989.
- [32] W. V. Steele and R. D. Chirico. Thermodynamic properties of alkenes (mono-olefins larger than C<sub>4</sub>). *J. Phys. Chem. Ref. Data*, 22(2):377–430, 1993.
- [33] Mark A. Trebble, Paul H. Salim, and Phillip M. Sigmund. A generalized approach to prediction of two and three phase CO<sub>2</sub>–hydrocarbon equilibria. *Fluid Phase Equilib.*, 82:111–118, 1993.
- [34] Michael S.-W. Wei, Trent S. Brown, Arthur J. Kidnay, and E. Dendy Sloan. Vapor + liquid equilibria for the ternary system methane + ethane + carbon dioxide at 230 K and its constituent binaries at temperatures from 207 to 270 K. *J. Chem. Eng. Data*, 40:726–731, 1995.
- [35] David Zudkevitch and Joseph Joffe. Correlation and prediction of vapor–liquid equilibria with the Redlich–Kwong equation of state. *AIChE J.*, 16(1):112–119, Jan. 1970.

### Liquids (DatLiq)

Experimental data for liquids (different from gases and fluids).

**Related keywords: (experimental | data) & liquids.**

#### Bibliography.

- [1] Scott P. Christensen and James D. Olson. Phase equilibria and multiple azeotropy of the acetic acid–isobutyl acetate system. *Fluid Phase Equilib.*, 79:187–199, 1992.
- [2] Paul Ehrlich and Robert H. Fariss. Negative partial molal volumes in the critical region. Mixtures of ethylene and vinyl chloride. *Comput. Chem. Eng.*, 73(4):1164–1167, 1969.
- [3] Alfred W. Francis. Ternary systems with three separate binodal curves. *J. Phys. Chem.*, 60:20–27, Jan. 1956.

- [4] Alfred W. Francis. Liquid–liquid equilibrium. In Herbert M. Schoen and John J. Mcketta, Jr., editors, *Interscience Library of Chemical Engineering and Processing*, volume 3. John Wiley & Sons, New York, 1963.
- [5] Joel H. Hildebrand. Seven liquid phases in equilibrium. *J. Phys. Colloid Chem.*, 53:944–947, 1949.
- [6] Horsley, L. H. Table of azeotropes and nonazeotropes. *Anal. Chem.*, 19(8):508–600, 1947.
- [7] Horsley, Lee Herbert. *Azeotropic Data—II*. American Chemical Society, Washington, 1962.
- [8] Scott L. Kittsley and Herbert A. Goeden. Eight liquid phases in stable equilibrium. *J. Am. Chem. Soc.*, 72:4841–4842, 1950.
- [9] C. D. Maranas, C. M. McDonald, S. T. Harding, and C. A. Floudas. Locating all azeotropes in homogenous azeotropic systems. *Comput. Chem. Eng.*, 20:413–418, 1996.
- [10] Hiroyuki Matsuda, Wenbin Dai, Kiyofumi Kurihara, Kenji Ochi, and Kazuo Kojima. The calculation of excess molar enthalpies of liquid mixtures at high pressures and temperatures from measurements at ambient conditions. *Fluid Phase Equilib.*, 215:45–53, 2004.
- [11] José M. Resa, José M. Goenaga, Rafael Gonzalez-Olmos, and Miguel Iglesias. Measurement and modeling of phase equilibria for ethanol + water + methanol at isobaric condition. *J. Chem. Eng. Data*, 51:2114–2120, 2006.
- [12] J. S. Rowlinson. *Liquids and Liquid Mixtures*. Butterworths Scientific Publications. Butterworths Publications Ltd., London, 1959.
- [13] Segura, Hugo, Jaime Wisniak, Antonio Aucejo, Juan B. Montón, and Rosa Muñoz. Polyazeotropy in binary systems. 2. Association effects. *Ind. Eng. Chem. Res.*, 35(11):4194–4202, 1996.
- [14] Viktor Tekáč, Ivan Cibulka, and Robert Holub. *pvt* properties of liquids and liquid mixtures: A review of the experimental methods and the literature data. *Fluid Phase Equilib.*, 19:33–149, 1985.
- [15] Wisniak, Jaime, Hugo Segura, and Ricardo Reich. Polyazeotropy in binary systems. *Ind. Eng. Chem. Res.*, 35(10):3742–3758, 1996.

### Minerals (DatMin)

Experimental data for (clay) minerals.

**Related keywords: (experimental | data) & minerals.**

#### Bibliography.

- [1] R. G. Berman. Internally-consistent thermodynamic data for minerals in the system  $\text{Na}_2\text{O}-\text{K}_2\text{O}-\text{CaO}-\text{MgO}-\text{FeO}-\text{Fe}_2\text{O}_3-\text{Al}_2\text{O}_3-\text{SiO}_2-\text{TiO}_2-\text{H}_2\text{O}-\text{CO}_2$ . *J. Petrol.*, 29(2):445–522, 1988.
- [2] Keith Frye. The encyclopedia of mineralogy. In *Encyclopedia of Earth Sciences*, volume IVB. Hutchinson Ross Publishing Company, Stroudsburg, Pennsylvania, 1981.
- [3] Harold C. Helgeson, Joan M. Delany, H. Wayne Nesbitt, and Dennis K. Bird. Summary and critique of the thermodynamic properties of rock-forming minerals. *Am. J. Sci.*, 278A:1–229, 1978.
- [4] James Leckie and Werner Stumm. Phosphate precipitation. In Earnest F. Gloyna and W. Wesley Eckenfelder, editors, *Water Quality Improvement by Physical and Chemical Processes*, pages 237–249, Austin, Texas, 1970. Center for Research in Water Resources, University of Texas Press. Water Resources Symposium (3).
- [5] Stanley J. Lefond, editor. *Industrial Minerals and Rocks: Nonmetallics Other Than Fuels*. American Institute of Mining, Metallurgical, and Petroleum Engineers, Inc., New York, fourth edition, 1975.
- [6] Craig E. Manning. The solubility of quartz in  $\text{H}_2\text{O}$  in the lower crust and upper mantle. *Geochim. Cosmochim. Acta*, 58(22):4831–4839, 1994.
- [7] Richard A. Robie, Bruce S. Hemingway, and James R. Fisher. *Thermodynamic Properties of Minerals and Related Substances at 298.15 K and 1 bar ( $10^5$  Pascals) Pressure and at Higher Temperatures*. Number 1452 in Geological Survey Bulletin. United States Printing Office, Washington, 1978.
- [8] Philippe Vieillard and Yves Tardy. Thermochemical properties of phosphates. In Jerome O. Nriagu and P. B. Moore, editors, *Phosphate Minerals*, chapter 4. Springer-Verlag, New York, 1984.

### Ammonia (DatNH3)

Experimental data for ammonia, water, nitrogen and hydrogen needed for ammonia synthesis calculations.

**Related keywords: (experimental | data) & (ammonia | nh3).**

**Bibliography.**

- [1] Carlo G. Alesandrini, Scott Lynn, and John M. Prausnitz. Calculation of vapor–liquid equilibria for the system  $\text{NH}_3\text{--N}_2\text{--H}_2\text{--Ar--CH}_4$ . *Ind. Eng. Chem. Process Des. Dev.*, 11(2):253–259, 1972.
- [2] Fernando J. Antezana and Huk Y. Cheh. Component fugacities in hydrogen–ammonia–propane mixtures. I. The fugacity of hydrogen. *Ind. Eng. Chem. Fundam.*, 14(3):224–232, 1975.
- [3] T. J. Edwards, J. Newman, and J. M. Prausnitz. Thermodynamics of vapor–liquid equilibria for the ammonia–water system. *Ind. Eng. Chem. Fundam.*, 17(4):264–268, 1978.
- [4] Paul C. Gillespie and Grant M. Wilson. Total pressure and vapor–liquid equilibrium measurements on the ammonia/water system. Technical report, Wiltec Research Company, Inc., Provo, Utah, 1982. DIPPR project 805(A)/81,82.
- [5] Paul C. Gillespie, W. Vincent Wilding, and Grant M. Wilson. Vapor–liquid equilibrium measurements on the ammonia–water system from 313 K to 589 K. Technical report, Wiltec Research Co., Inc., Provo, Utah, 1985. DIPPR project 758–81.
- [6] Ulrich Göppert and Gerd Maurer. Vapor–liquid equilibria in aqueous solutions of ammonia and carbon dioxide at temperatures between 333 and 393 K and pressures up to 7 MPa. *Fluid Phase Equilib.*, 41:153–185, 1988.
- [7] Gianfranco Guerreri. High-pressure vapor–liquid equilibria in ammonia–water–nitrogen–hydrogen system. *AIChE J.*, 13(5):877–883, Sep. 1967.
- [8] Jean-Luc Guillevic, Dominique Richon, and Henri Renon. Vapor–liquid equilibrium data for the binary system water–ammonia at 403.1, 453.1 and 503.1 K up to 7.0 MPa. *J. Chem. Eng. Data*, 30(3):332–335, 1985.
- [9] Lester Haar and John S. Gallagher. Thermodynamic properties of ammonia. *J. Phys. Chem. Ref. Data*, 7(3):635–792, 1978.
- [10] Von Friedel Heise. Flüssigkeits-Dampfgleichgewichte zwischen Ammoniak und verschiedenen Gasen. *Ber. Bunsen-Ges.*, 76(9):938–943, 1972.
- [11] R. A. Heidemann and S. S. H. Rizvi. Correlation of ammonia–water equilibrium data with various modified Peng–Robinson equations of state. *Fluid Phase Equilib.*, 29:439–446, 1986.
- [12] F. Kurz, B. Rumpf, and G. Maurer. Vapor–liquid–solid equilibria in the system  $\text{NH}_3\text{--CO}_2\text{--H}_2\text{O}$  from around 310 to 470 K: New experimental data and modeling. *Fluid Phase Equilib.*, 104:261–275, 1995.
- [13] Miguel A. Leiva and Victor Vivanco. Vapor–liquid equilibria of aqueous solutions containing volatile weak electrolytes by using the free energy minimization method. *Fluid Phase Equilib.*, 27:483–490, 1986.
- [14] G. Maurer. Phase equilibria in chemically reactive systems. *Fluid Phase Equilib.*, 30:337–352, 1986.
- [15] E. Naumowicz-Węglińska and W. Wóycicki. Excess enthalpies of gaseous mixtures containing ammonia 2. Ammonia + nitrogen and ammonia + methane. *J. Chem. Thermodyn.*, 18:1047–1052, 1986.
- [16] Ellen M. Pawlikowski, John Newman, and John M. Prausnitz. Phase equilibria for aqueous solutions of ammonia and carbon dioxide. *Ind. Eng. Chem. Process Des. Dev.*, 21:764–770, 1982.
- [17] K. V. Reddy and A. Husain. Vapor–liquid equilibrium relationship for ammonia in presence of other gases. *Ind. Eng. Chem. Process Des. Dev.*, 19(4):580–586, 1980.
- [18] Syed S. H. Rizvi and Robert A. Heidemann. Vapor–liquid equilibria in the ammonia–water system. *J. Chem. Eng. Data*, 32(2):183–191, 1987.
- [19] Bruce E. Roberts and Peter R. Tremaine. Vapour liquid equilibrium calculations for dilute aqueous solutions of  $\text{CO}_2$ ,  $\text{H}_2\text{S}$ ,  $\text{NH}_3$ , and  $\text{NaOH}$  to 300°C. *Can. J. Chem. Eng.*, 63:294–300, 1985.
- [20] Bernd Rumpf, Frank Weyrich, and Gerd Maurer. Enthalpy of dilution in aqueous systems of single solutes ammonia, sodium sulfate and ammonium sulfate: Experimental results and modeling. *Thermochim. Acta*, 303:77–91, 1997.
- [21] Thomas M. Smolen, David B. Manley, and Bruce E. Poling. Vapor–liquid equilibrium data for the  $\text{NH}_3\text{--H}_2\text{O}$  system and its description with a modified cubic equation of state. *J. Chem. Eng. Data*, 36(2):202–208, 1991.
- [22] Reiner Tillner-Roth and Daniel G. Friend. Survey and assessment of available measurements of thermodynamic properties of the mixture {Water + Ammonia}. *J. Phys. Chem. Ref. Data*, 27(1):45–61, 1998.
- [23] Reiner Tillner-Roth and Daniel G. Friend. A Helmholtz free energy formulation of the thermodynamic properties of the mixture {Water + Ammonia}. *J. Phys. Chem. Ref. Data*, 27(1):63–96, 1998.
- [24] Hannelore Zeininger. Flüssigkeits/Dampf-Gleichgewicht des Systems  $\text{N}_2/\text{H}_2/\text{CH}_4/\text{NH}_3$  bei 25°C und Drücken bis 500 bar. *Chem.-Ing.-Tech.*, 45(17):1067–1070, 1973.

**Molten Salts (DatSal)**

Experimental data for simple molten salts of sodium, potassium, ammonium, etc.

**Related keywords: (experimental | data) & (salt | melts).**

**Bibliography.**

- [1] L. H. Adams and R. E. Gibson. Equilibrium in binary systems under pressure. III. The influence of pressure on the solubility of ammonium nitrate in water at 25°. *J. Am. Chem. Soc.*, 54:4520–4537, 1932.
- [2] C. A. Angell and J. C. Tucker. Heat capacities and fusion entropies of the tetrahydrates of calcium nitrate, cadmium nitrate, and magnesium acetate. Concordance of calorimetric and relaxational “ideal” glass transition temperatures. *J. Phys. Chem.*, 78(3):278–281, 1974.
- [3] Donald G. Archer. Thermodynamic properties of the NaCl + H<sub>2</sub>O system. I. Thermodynamic properties of NaCl(cr). *J. Phys. Chem. Ref. Data*, 21(1):1–21, 1992.
- [4] Donald G. Archer. Thermodynamic properties of the NaCl + H<sub>2</sub>O system. II. Thermodynamic properties of NaCl(aq), NaCl · H<sub>2</sub>O(cr), and phase equilibria. *J. Phys. Chem. Ref. Data*, 21(4):793–829, 1992.
- [5] E. R. van Artsdalen and I. S. Yaffe. Electrical conductance and density of molten salt systems: KCl–LiCl, KCl–NaCl and KCl–KI. *J. Phys. Chem.*, 59(2):118–127, Feb. 1955.
- [6] J. L. Barton and H. Bloom. A boiling point method for determination of vapor pressures of molten salts. *J. Phys. Chem.*, 60:1413–1416, Oct. 1956.
- [7] N. L. Bowen. Properties of ammonium nitrate. *J. Phys. Chem.*, 30:736–737, 1926.
- [8] J. D. Brandner, Norman M. Junk, J. W. Lawrence, and Jack Robins. Vapor pressures of ammonium nitrate. *J. Chem. Eng. Data*, 7(2):227–228, Apr. 1962.
- [9] R. N. Brown and A. C. McLaren. On the mechanism of the thermal transformations in solid ammonium nitrate. *Proc. R. Soc., Ser. A*, 266:329–343, 1962.
- [10] M. Buback and E. U. Franck. Measurements of the vapour pressures and critical data of ammonium halides. *Ber. Bunsen-Ges.*, 76(3/4):350–354, 1972.
- [11] A. N. Campbell, J. B. Fishman, G. Rutherford, T. P. Schaefer, and L. Ross. Vapor pressures of aqueous solutions of silver nitrate, of ammonium nitrate, and of lithium nitrate. *Can. J. Chem.*, 34:151–159, 1956.
- [12] Patrice Chartrand, Christian Robelin, and Arthur Pelton. A thermodynamic model for the NH<sub>4</sub>,K,Na // NO<sub>3</sub>,Cl–H<sub>2</sub>O system. Technical report, CRCT—École Polytechnique Montréal, Canada, 447 Berwick, Town of Mount Royal, Quebec, Canada, 2002.
- [13] A. Chmarzynski. Enthalpies of crystallization of equilibrium solid phases occurring in the system. K<sub>2</sub>SO<sub>4</sub>–(NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>–H<sub>2</sub>O at 298.15 K. *J. Therm. Anal.*, 38:2027–2032, 1992.
- [14] A. Chmarzyński and H. Piekarski. Thermochemical investigations of the systems KCl–KBr–H<sub>2</sub>O, K<sub>2</sub>SO<sub>4</sub>–(NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>–H<sub>2</sub>O and KNO<sub>3</sub>–NH<sub>4</sub>NO<sub>3</sub>–H<sub>2</sub>O at 298.15 K. *J. Therm. Anal.*, 45:869–878, 1995.
- [15] A. Chmarzyński and B. Dejewska. Enthalpies of crystallization of equilibrium solid phases occurring in the system KNO<sub>3</sub>–NH<sub>4</sub>NO<sub>3</sub>–H<sub>2</sub>O at 298.15 K. *J. Therm. Anal.*, 45:859–867, 1995.
- [16] R. V. Coates and G. D. Woodard. 381. An X-ray diffractometric study of the ammonium nitrate–potassium nitrate system. *J. Chem. Soc.*, pages 2135–2140, Mar. 1965.
- [17] L. C. de Coppet. Ueber einige ältere Bestimmungen des Gefrierpunktes Gesättigter Salzlösungen. *Z. Phys. Chem., Stoechiom. Verwandtschaftsl.*, 22:240–241, 1897.
- [18] B. Dejewska. The distribution coefficient of isomorphous admixtures for KCl–KBr–H<sub>2</sub>O, K<sub>2</sub>SO<sub>4</sub>–(NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>–H<sub>2</sub>O and KNO<sub>3</sub>–NH<sub>4</sub>NO<sub>3</sub>–H<sub>2</sub>O systems at 298 K. *Cryst. Res. Technol.*, 27(2):385–394, 1992.
- [19] Ingemar Dellien. A DSC study of the phase transformations of ammonium nitrate. *Thermochim. Acta*, 55:181–191, 1982.
- [20] P. Dingemans. Die Dampfspannung von Gesättigten NH<sub>4</sub>NO<sub>3</sub>-Lösungen. *Recl. Trav. Chim. Pays-Bas*, 60:317–328, 1941.
- [21] C. A. van Driel, A. E. D. M. van der Heijden, S. de Boer, and G. M. van Rosmalen. The III–IV phase transition in ammonium nitrate: Mechanisms. *J. Cryst. Growth*, 141:404–418, 1994.
- [22] Reginald George Early and Thomas Martin Lowry. Properties of ammonium nitrate. Part I. The freezing point and transition-temperatures. *J. Chem. Soc., Trans.*, 115(131):1387–1405, 1919.
- [23] Reginald George Early and Thomas Martin Lowry. The properties of ammonium nitrate. Part III. Ammonium nitrate and sodium nitrate. *J. Chem. Soc., Trans.*, CXI:963–969, 1922.
- [24] George Feick and R. M. Hainer. On the thermal decomposition of ammonium nitrate. steady state reaction. Temperatures and reaction rate. *J. Am. Chem. Soc.*, 76:5860–5863, nov 1954.
- [25] George Feick. The dissociation pressure and free energy of formation of ammonium nitrate. *J. Am. Chem. Soc.*, 76:5858–5860, Nov. 1954.
- [26] E. E. Ferguson, D. C. Levendis, and F. R. L. Schoening. X-ray diffraction study of the orientational relation between the IV and III phases of ammonium nitrate. *Chem. Mater.*, 5(9):1293–1298, 1993.

- [27] L. Filipescu, D. Fătu, Micaela Mocioi, and E. Segal. On the chemical and thermal stabilization of  $\text{NH}_4\text{NO}_3(iv)$ . *Thermochim. Acta*, 97:229–241, 1986.
- [28] Frédéric Gemme and Patrice Chartrand. Thermodynamic optimization and calculation of the  $\text{NH}_4$ , K //  $\text{NO}_3$ ,  $\text{H}_2\text{PO}_4$ , Cl system. Technical report, CRCT—École Polytechnique Montréal, Canada, 447 Berwick, Town of Mount Royal, Quebec, Canada, 2002.
- [29] Ernst Jänecke and Erich Rahlfs. Über das System  $\text{NH}_4\text{NO}_3 - \text{H}_2\text{O}$ . *Z. Anorg. Allg. Chem.*, 192:237–244, 1930.
- [30] Ernst Jänecke. Über das ternäre System der Nitrate von Kalium, Natrium und Ammonium (K, Na,  $\text{NH}_4$ ) $\text{NO}_3$ . *Z. Anorg. Chem.*, 259:92–106, 1949.
- [31] Xiaoyan Ji, Xiaohua Lu, Luzheng Zhang, Ningzhong Bao, Yanru Wang, Jun Shi, and Benjamin C.-Y. Lu. A further study of solid–liquid equilibrium for the  $\text{NaCl} - \text{NH}_4\text{Cl} - \text{H}_2\text{O}$  system. *Chem. Eng. Sci.*, 55:4993–5001, 2000.
- [32] A. G. Keenan. The cryoscopic heat of fusion of ammonium nitrate. *J. Phys. Chem.*, 60:1356–1361, 1956.
- [33] A. D. Kirshenbaum, J. A. Cahill, P. J. McGonigal, and A. V. Grosse. The density of liquid  $\text{NaCl}$  and  $\text{KCl}$  and an estimate of their critical constants together with those of the other alkali halides. *J. Inorg. Nucl. Chem.*, 24:1287–1296, 1962.
- [34] Kjersti Kleveland and Tor Grande. DSC measurements in the  $\text{KNO}_3 - \text{NH}_4\text{NO}_3$  system. Technical report, Department of Chemistry, Norwegian University of Science and Technology, nov 2000.
- [35] R. J. M. Konings and E. H. P. Cordfunke. The vapour pressures of hydroxides. I. The alkali hydroxides  $\text{KOH}$  and  $\text{CsOH}$ . *J. Chem. Thermodyn.*, 20:103–108, 1988.
- [36] C. M. Kramer, Z. A. Munir, and K. H. Stern. Evaporation of  $\text{NaNO}_3$ ,  $\text{KNO}_3$  and  $\text{NaNO}_2$ . *High Temp. Sci.*, 16:257–267, 1983.
- [37] B. W. Lucas, M. Ahtee, and A. W. Hewat. The crystal structure of phase II ammonium nitrate. *Acta Crystallogr., Sect. A: Cryst. Phys., Diffr., Theor. Gen. Crystallogr.*, 35:1038–1041, 1979.
- [38] N. W. Luft. Sublimation pressures and latent heats of ammonium salts. *Ind. Chem.*, pages 502–504, Oct. 1955.
- [39] Vadim A. Medvedev, Michael E. Efimov, Paul J. Cerutti, Robert M. McKay, Lawrence H. Johnson, and Loren G. Hepler. Standard enthalpy of solution of ammonium nitrate in water at 298 K. *Thermochim. Acta*, 23:87–92, 1978.
- [40] Alan W. C. Menzies and N. N. Dutt. The liquidus surface of the ternary system composed of the nitrates of potassium, sodium and calcium. *J. Am. Chem. Soc.*, 33:1366–1375, 1911.
- [41] Ida L. Millican, Alfred Francis Joseph, and Thomas Martin Lowry. The properties of ammonium nitrate. Part II. Ammonium nitrate and water. *J. Chem. Soc., Trans.*, 121:959–963, 1922.
- [42] Mirosław Miller and Krzysztof Skudlarski. Thermodynamic properties of  $\{x\text{NaCl} + (1 - x)\text{NaBr}\}$  at 840 K investigated by mass spectrometry. *J. Chem. Thermodyn.*, 19:565–570, 1987.
- [43] Erwin Mirsch and Fritz Serowy. Die Lösungsgleichgewichte des salzsauren quinären Systems  $\text{MgCl}_2 - \text{KCl} - \text{NaCl} - \text{HCl} - \text{H}_2\text{O}$ . *Freiberg. Forschungsh. C*, 118:1–70, Dec 1961.
- [44] L. Misane, El Allali, S., M. Kaddami, A. Zrineh, R. Tenu, J. Berthet, and J. J. Counioux. Le système ternaire  $\text{H}_2\text{O} - \text{NH}_4\text{NO}_3 - \text{Mg}(\text{NO}_3)_2$  isothermes  $-25, -14, 0, 30$  et  $45^\circ\text{C}$ . *Thermochim. Acta*, 356(1–2):117–126, 2000.
- [45] Wolf Müller and Paul Kaufmann. Über die Löslichkeit von Ammoniumnitrat in Wasser zwischen  $12$  und  $40^\circ$ . *Z. Phys. Chem., Stoechiom. Verwandtschaftsl.*, 42:497–500, 1903.
- [46] Masanori Nagatani and Tetsuro Seiyama. Heat capacities and thermodynamic properties of ammonium nitrate crystal: Phase transitions between stable and metastable phases. *Bull. Chem. Soc. Jpn.*, 40(8):1833–1844, 1967.
- [47] I. N. Nikonova and A. G. Bergman. Nitrophoska 1 (sulphate-type) article II polytherm of a reversible system of potassium and ammonium sulphates and nitrates. *Zh. Prikl. Khim. (Leningrad)*, 15(6):438–446, 1942.
- [48] Donald F. Othmer and Gerhard J. Frohlich. Correlating vapor pressures and heats of solution for the ammonium nitrate–water system, an enthalpy–concentration diagram. *AIChE J.*, 6(2):210–214, Jun. 1960.
- [49] Edgar Philip Perman and William John Howells. The properties of ammonium nitrate. Part VI. The reciprocal salt pair ammonium nitrate and potassium sulphate. *J. Chem. Soc., Trans.*, 123:2128–2134, 1923.
- [50] G. Rasulić, S. Jovanović, and Lj. Milanović. Ammonium nitrate changes during thermal analysis. *J. Therm. Anal.*, 30:65–72, 1985.
- [51] Robert T. Rewick and B. J. Gikis. The heat of solution of ammonium nitrate in nitric acid. *J. Chem. Eng. Data*, 25(2):127–131, 1980.

- [52] G. A. Sacchetto, G. G. Bombi, and C. Macca. Vapour pressures of very concentrated electrolyte solutions. I. Measurements in  $\{(1-x)\text{H}_2\text{O} + x\text{LiNO}_3\}$  and  $\{(1-x)\text{H}_2\text{O} + x\text{NH}_4\text{NO}_3\}$  by dew-point apparatus. *J. Chem. Thermodyn.*, 13:31–40, 1981.
- [53] Giuseppe A. Sacchetto and Zdeněk Kodejš. Water solubility in molten-salt mixtures. *J. Chem. Soc., Faraday Trans. 1*, 78(12):3519–3527, 1982.
- [54] James Sangster and Arthur D. Pelton. Phase diagrams and thermodynamic properties of the 70 binary alkali halide systems having common ions. *J. Phys. Chem. Ref. Data*, 16(3):509–561, 1987.
- [55] H. Scherzberg, R. Schmitz, and W. Wöhlk. Messo pilots new potassium sulphate process. *Phosphorus & Potassium*, pages 20–27, Mar./Apr. 1992. No. 178.
- [56] K. D. Shah and A. G. Roberts. Properties of ammonium nitrate. In Cornelius Keleti, editor, *Nitric Acid and Fertilizer Nitrates*, Fertilizer Science and Technology Series, volume 4. MarcelDekker, Inc., New York, 1985.
- [57] Toshiaki Shirai and Toshio Ishibashi. Studies on fused ammonium nitrate. II. Specific volume of fused ammonium nitrate and its solutions of salts. *Sci. Pap. Coll. Gen. Educ., Univ. Tokyo*, 7:53–60, 1957.
- [58] C. Howard Shomate. High-temperature heat contents of magnesium nitrate, calcium nitrate and barium nitrate. *J. Am. Chem. Soc.*, 66:928–929, Jun. 1944.
- [59] Christer Sjölin. Mechanism of caking of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) prills. *J. Agric. Food Chem.*, 20(4):895–900, 1972.
- [60] James P. Smith, James R. Lehr, and A. William Frazier. Crystallographic properties of the ammonium nitrate-sulfates  $3\text{NH}_4\text{NO}_3 \cdot (\text{NH}_4)_2\text{SO}_4$  and  $2\text{NH}_4\text{NO}_3 \cdot (\text{NH}_4)_2\text{SO}_4$ . *J. Agric. Food Chem.*, 10(1):77–78, Jan-Feb 1962.
- [61] G. A. Sorina, M. B. Blinova, and Yu. V. Tsekhanskaya. Total vapor pressure and differential heats of vaporization of water in the system ammonium nitrate - water. *J. Appl. Chem. (Leningrad)*, 48(8):1788–1791, 1976. Translated from *Zh. Prikl. Khim. (Leningrad)*, Vol. 48, No. 8, pp. 1720–1724, August, 1975.
- [62] Luke E. Steiner and John Johnston. Development of a method of radiation calorimetry, and the heat of fusion or of transition of certain substances. *J. Phys. Chem.*, 32(6):912–939, 1928.
- [63] C. C. Stephenson, D. R. Bentz, and D. A. Stevenson. The heat capacity of ammonium nitrate nitrate from 15 to 315°K. *J. Am. Chem. Soc.*, 77:2161–2164, Apr. 1955.
- [64] I. Konkoly Thege, I. Manninger, and E. E. Zapp. Thermal analysis of the polymorphic transformations of ammonium nitrate. *J. Therm. Anal.*, pages 111–116, 1980.
- [65] C. E. Vanderzee, Dan H. Waugh, Norman C. Haas, and Dwight A. Wigg. The standard enthalpy of solution of  $\text{NH}_4\text{NO}_3(c, iv)$  in water at 298.15 K. (A search for the standard thermodynamic state.). *J. Chem. Thermodyn.*, 12:27–40, 1980.
- [66] C. E. Vanderzee, Dan H. Waugh, and Norman C. Haas. Enthalpies of dilution and relative apparent molar enthalpies of aqueous ammonium nitrate. The case of a weekly hydrolysed (dissociated) salt. *J. Chem. Thermodyn.*, 12:21–25, 1980.
- [67] B. F. Wishaw and R. H. Stokes. The osmotic and activity coefficients of aqueous solutions of ammonium chloride and ammonium nitrate at 25°. *Trans. Faraday Soc.*, 49:27–31, 1953.

### Solids (DatSol)

Experimental  $p, V, T$  data of pure solid phases.

**Related keywords: (experimental | data) & solids.**

#### Bibliography.

- [1] D. Batani, F. Strati, H. Stabile, M. Tomasini, G. Lucchini, A. Ravasio, M. Koenig, A. Benuzzi-Mounaix, H. Nishimura, Y. Ochi, J. Ullschmied, J. Skala, B. Kralikova, M. Pfeifer, Ch. Kadlec, T. Mocek, A. Präg, T. Hall, P. Milani, E. Barborini, and P. Piseri. Hugoniot data for carbon at megabar pressures. *Phys. Rev. Lett.*, 92(6):065503.1–065503.4, feb 2004.
- [2] Jitka Eysseltová and Thedford P. Dirkse. IUPAC–NIST solubility data series 66. Ammonium phosphates. *J. Phys. Chem. Ref. Data*, 27(6), 1998.
- [3] W. G. Manzhelii, A. M. Tolkachev, M. I. Bagatskii, and E. I. Voitovitch. Thermal expansion, heat capacity, and compressibility of solid  $\text{CO}_2$ . *Phys. Stat. Sol. (b)*, 44(1):39–49, 1971.

### Urea (DatUrea)

Experimental data for urea, biuret,  $\text{CO}_2$  and water at urea synthesis conditions.

**Related keywords: (experimental | data) & urea | biuret.**

**Bibliography.**

- [1] [W. Durisch](#), S. M. Lemkowitz, and P. J. van der Berg. Constituent and component measurements and calculations of the vapour/liquid equilibrium of the ternary system carbon dioxide–ammonia–water under urea synthesis conditions. *Chimia*, 34(7):314–322, jul 1980.
- [2] [Egan, Jr., Edward P.](#) and Basil B. Luff. Heat of solution, heat capacity, and density of aqueous urea solutions at 25°C. *J. Chem. Eng. Data*, 11(2):192–194, Apr. 1966.